Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Which of the following can form H-bonds	A. NH₃ B. C₂H₆ C. NaCl D. CHCl₃
2	The long chains of amino acids are coiled about one another onto a spiral by	A. Ionic bond B. Van der walls forces C. Hydrogen bonding D. Overlapping of orbitals
3	The polarizabilities of elements mostly increase down the group due to the reason that	A. The atomic numbers increase B. Number of protons increase C. Number of shells increase along, with increase of shielding effect D. The behavior of the elements remain the same
4	The repulsion of electronic clouds of the molecules are responsible for the attractive forces among the molecules. These forces are	A. Dipole-induced dipole forces B. Ion-dipole forces C. Instantaneous dipole-induced dipole forces D. Dipole-dipole forces
5	In order to mention the B.P. of water at 110°C, the external pressure should be	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 760 torr D. Any value of pressure
6	When water freezes at 0°C, its density decreases due to	A. Cubic structure of ice B. Empty spaces present in the structure of ice C. Change of bond lengths D. Change of bond angles
7	NH_3shows a maximum boiling point among the hydrides of V-A group elements due to	A. Very small size of nitrogen B. Lone pair electrons present on Nitrogen C. Enhanced electronegative character of Nitrogen D. Pyramidla structure of NH₃
8	Acetone and chloroform are soluble in each other due to	A. Intermolecualr hydrogen bonding B. Dipole-dipole interaction C. Instantaneous dipoles D. All of the above
9	London dispersion forces are the only forces present among the	A. Molecules of water in liquid state B. Atoms of helium in gaseous state at high temperature C. Molecules of solid iodine D. Molecules of hydrogen chloride gas
10	Noon has low critical temperature and pressure as compared to other gases. the most probable reason is that	A. Its octet is complete B. It is a monoatomic gas C. It has very low polarizability D. It has least forces of attraction
11	The highest temperature above which a gas cannot be liquified, no matter how much the pressure is applied is known as	A. Boiling temperature B. Consulate temperature C. Absolute zero D. Critical temperature
12	The molecules of a gas show more deviation from ideal behaviour at low temperature, because	A. Attractive force dominate at low temperature B. Kinetic energies are increased C. Collisions become less frequent D. Densities of the gases increase
13	The free expansion of the gas from high pressure towards the low pressure causes	A. Increase of temperature B. Decrease of temperature C. Grater number of collisions among the molecules D. Decrease of velocity of gas molecules

14	Rate of diffusion of CO and N_2 are same at room temperature due to the reason, that	A. Both are diatomic molecules B. Both have same multiple bond in them C. Both have lone pairs in them D. Both have same molar masses
15	In gas occupies a volume of 2 dm^3 at 27°C and I atm pressure. The expression for its volume at S.T.P. is	
16	Gas equation is derived by combining	A. Avogadro's and Charles's Law B. Boyle's and Charles's Law C. Avogadro's and Boyle's Law D. Avogadro's, Boyle's and Charles's Law
17	Normal temperature and pressure (S.T.P) of gas rafers to	A. 273 K and 76 mm Hg B. 273° C and 760 mm Hg C. 273 K and 760 mm Hg D. 273° C and 76 mm Hg