Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Those elements whose electronegativities are 1.2 and 3.2, react to form	A. Ionic bond B. Covalent bond C. Gaseous substance D. Definiting a liquid substance
2	The shielding effect of the inner electrons is responsible for	A. Increasing ionization energy values B. Decreasing ionization energy values C. Increasing electron affinity D. Increasing electonegativity
3	The octet rule in not followed in the formation of	A. NF₃ B. CF₄ C. CCl₄ D. PCl₅
4	Which of the following species has unpaired electrons in anti-bonding molecular orbitals	A. O₂⁺² B. N₂⁻² C. B₂ D. F₂
5	Which of the hydrogen halides has the highest percentage of ionic character	A. HF B. HBr C. HCl D. HI
6	Which of the following molecules has zero dipole moment	A. NH₃ B. CHCl₃ C. H₂O D. BF₃
7	Which of the following statements is not correct regarding bonding molecular orbitals	A. Bonding molecular orbitals possess less energy than atomic orbitals from which they are formed B. Bonding molecular orbitals have low electron density between the two nuclei C. Every electron in the bonding molecular orbitals contributes to the attraction between atoms D. Bonding molecular orbitals are formed when the electron waves undergo constructive interference
8	The number of bonds in nitrogen molecule is	A. One o and oneπ B. One o and twoπ C. Three sigma only D. Two o and twoπ
9	An ionic compound $A^{+}B^{-}$is most likely to be formed when	A. The ionization energy of A is high and electron affinity of B is low B. The ionization energy of A is low and electron affinity of B is high C. Both the ionization energy of A and electron affinity of B are high D. Both the ionization energy of A and electron affinity of B are low
		A. Planck quantum theory

10	Bohr's model of atom is contradicted by	B. Quantization of energy of electrons C. Heisenberg's uncertainty principle D. Quantization of angular members
11	The charge on electron was determined by millikan in his oil drop experiment and its value is	A. 6.023 x 10⁻²³C B. 1.602 x 10⁻²³C C. 1.602 x 10⁻¹⁹C D. 6.625 x 10⁻³⁴C
12	Which of the following represents electronic configuration of the most electropositive elements	A. He [2s⁻¹] B. Xe [6s¹] C. He [2s²] D. Xe [6s²]
13	Which of the following particles would on losing an electron has its outermost p-orbital as half filled	A. Nitrogen atom B. O⁺ion C. P⁻¹ion D. S⁺¹ion
14	The azimuthal quantum number / = 2, then M(Magnetic quantum number) can have values as	A. +1, -1 B. +1, 0, -1 C. +2, +1, 0,1, -1, 2 D. +3, +2, +1, 0, 1, -2, -3
15	If uncertainty position of an electron is zero, the uncertainty in its momentum would be	A. Zero B. Infinite C. Both a and b D. None of these
16	The line of the balmer series in the visible region of the spectrum, but the limiting line, in the series lies in	A. Visible region B. X-Ray region C. I.R region D. U.V. region
17	Anode rays were discovered by	A. J. Stoney B. Rutherford C. J.J. Thomson D. Goldstein