

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Le-Chatelier Braun principle is sometimes known as	A. Law of mass action B. Law of mobile equilibrium C. Law of active mass D. All of these above
2	The relationship between K_p and K_c is given by	
3	If the volumes of reactants and products are same in a gaseous phase reaction, then the equilibrium state is not affected by	A. Change of temperature B. Change of pressure C. Change of concentration D. Catalyst
4	Question Image	A. $\text{dm}^3 + 6 \text{ mole}^{-2}$ B. $\text{mole}^2 \text{ dm}^{-6}$ C. Mole dm^{-3} D. Having no units
5	At equilibrium stage of chemical reaction	A. The concentration of reaction is equal to concentration of products B. The rate constant of forward reaction is equal to rate constant of backward reaction C. The rate of forward reaction is equal rate of backward reaction D. The energy of activation of forward step is equal to energy of activation of backward step
6	An excess of aqueous silver nitrate is added to aqueous barium chloride and precipitate is removed by filtration. What are the main ions in the filtrate	
7	The solubility product of AgCl is $2.0 \times 10^{-10} \text{ mol}^2 \text{ dm}^{-6}$. The maximum concentration of Ag^+ ions in the solution is	A. $2.0 \times 10^{-10} \text{ mol dm}^{-3}$ B. $1.41 \times 10^{-5} \text{ mol dm}^{-3}$ C. $1.0 \times 10^{-10} \text{ mol dm}^{-3}$ D. $4.0 \times 10^{-20} \text{ mol dm}^{-3}$
8	The pH of $10^{-3} \text{ mol dm}^{-3}$ of an aqueous solution of H_2SO_4 is	A. 3.0 B. 2.7 C. 2.0 D. 1.5
9	Question Image	A. The value of K_p falls with a rise in temperature B. The value of K_p falls with increasing pressure C. Adding 2O_5 catalyst increase the equilibrium yield of sulphur trioxide D. The value of K_p is equal to K_c
10	For which system does the equilibrium constant K_c has units of $(\text{concentration})^{-1}$	
11	Whenever a reaction is endothermic, then it means that	A. Heat is transferred from surrounding to the system B. Heat is transferred system to the surrounding C. Heat content of the product is greater than that of reactants D. Heat content of the reactants is greater than the products
12	One of the following statements about Born-Haber cycle is correct. Which is that statement	A. Born-Haber cycle is different from Hess's law B. The energy change in a cyclic process is not zero C. The lattice energy of the crystalline

substances can be calculated easily
D. Heat of formation of the product and the lattice energy of the substance can be calculated simultaneously

13	The net change in a chemical reaction is same whether it takes place directly or indirectly is	A. Henry's law B. Charles's C. Hess's lass D. Graham's law
14	The S.I. units for the molar heat capacity are	A. Joule Cm ⁻³ deg ⁻¹ B. Joule deg ⁻¹ atm ⁻¹ C. Joule deg ⁻¹ mol ⁻¹ D. Joule deg ⁻¹ kg ⁻¹
15	<div>Question Image</div>	
16	At constant volume of a system remains constant and the heat is absorbed by the system, them amount of heat absorbed is called	A. Enthaply change of the system B. Internal energy change of the system C. Total enthalphy of the system D. Total internal energy of the system
17	What happens to the enthalpy change when the coefficients of a chemical equation are doubled	A. It doubles B. It becomes half C. It does not change D. It cannot be predicted