Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Colligative properites are the properties of	A. Dilute solutions which behave as nearly ideal solution B. Concentrated solution which behave as nearly non-ideal solution C. Both a and b D. Neither a nor b
2	The molal boiling point constant is the ration of the elevation in boiling point to	A. Molarity B. Molality C. Mole fraction of solvent D. Mole fraction of solute
3	Two solution of NaCl and KCl are prepared separately by dissolving same moles of them in the fixed amount of solvent. Which of the following statements is true for these solution	A. KCl solution will have higher boiling point than NaCl solution B. Both the solutions have different boiling point C. KCl and NaCl solution possess same vapour pressure D. KCl solution possesses lower freezing point than NaCl solution
4	Which of the following solutions has the highest boiling point	A. 5.85% solution of sodium chloride B. 18.0% solution of glucose C. 6.0% solution of urea D. All have the same boiling points
5	In azeotropic mixture showing positive deviation from Raoult's law the volume of the mixture is	A. Slightly more than the total volume of the components B. Slightly less than the total volume of the components C. Equal to the total volume of the components D. None of these
6	Azeotrpic mixture of two liquids boils at a lower temperature than either of them, when	A. It is saturated B. Is shows positive deviation from Raoult's law C. It shows negative deviation from Raoult's law D. Is is metastable
7	An aqueous solution of ethanol in water has vapour pressure	A. Equal to that water B. Equal to that of ethanol C. More than that of H_2O D. less than that of water
8	A solution of glucose is 10% to volume in which 1 g mole of it is dissolved will be	A. 1 dm^3 B. 1.8 dm^3 C. 200 cm^3 D. 900 cm^3
9	18 g of glucose is dissolved in 90 g of water. The relative lowering of vapour pressure is equal to	A. 1/5 B. 5.1 C. 1/51 D. 6
10	Molarity of pure water is	A. 1 B. 18 C. 55.5 D. 6
11	The pK_a value of CH_3COOH is 4.74 when we mix CH_3COOH and CH_3COONa in the ratio of 10:1, tehn the pH of the buffer is	A. 4.74 B. 5.74 C. 3.74 D. 7.00
12	The number of moles of acid or base required by one dm^3 of buffer to alter its pH by one unit is called	A. Buffer efficiency B. Buffer capacity C. Buffer action D. None
13	A buffer solution can be prepared by mixing	A. Weak acid and its salt with weak base B. Weak base and its salt with strong acid

		C. Strong acid and its salts with weak base
		D. Strong base and its salt with weak acid
14	The value of pH and P^{OH} of pure water at 25° C is	A. 14 B. 7 C. 1×10^{-14} D. 1×10^{14}
15	The pH of human blood is	A. 7.0 B. 7.4 C. 4.0 D. 6.5
16	K_a and K_b of a conjugate acid and are related with k_w as	A. $K_a + K_b = K_w$ B. $K_a - K_b = K_w$ C. $K_a . K_b = K_w$ D. $K_a / K_b = K_w$
17	The effect of temperature on equilibrium was studied by	A. Lewis B. Van der wall C. Arrhenius D. Vant hoff