

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	What is the mass of aluminium is 204 g of the aluminum oxide Al_2O_3 .	A. 26 g B. 27 g C. 54 g D. 108 g
2	How many molecules of CO_2 are formed by burning 12 g carbon with excess of oxygen.	A. 3.01×10^{23} B. 1×10^{23} C. 6.02×10^{23} D. 1.03×10^{23}
3	How many moles of water results by burning 4 mole of H_2 with excess of oxygen.	A. 1 mol B. 2 mol C. 3 mol D. 4 mol
4	Stoichiometric calculations cannot applied to reversible reactions because.	A. Product again changes to reactant B. Less product is formed C. Reaction goes only in one direction D. Products do not disappear.
5	The number of electrons in one mole of hydrogen gas is.	A. 6.02×10^{23} B. 12.04×10^{23} C. Only two D. Indefinite
6	The Avogadro constant is the number of.	A. Atoms in 1 g of helium ga B. Molecules in 35.5 g of chlorine gas C. Atoms in 6 h graphite D. Atoms in 24 g of magnesium
7	Which of the following contains the same number of molecules as 9 g of water.	A. 2 g of hydrogen gas B. 14 g of nitrogen gas C. 32 g of oxygen gas D. 44 g of carbon dioxide gas
8	What has a mass equal to that of one mole of water.	A. 22.4 dm ³ of water B. One mole of steam C. One molecule of water D. Two moles of hydrogen molecules and one mole of oxygen molecules.
9	12.04×10^{23} atoms of nitrogen gas is equal to.	A. 1 mol B. 2 mol C. 3 mol D. 4 mol
10	Which statement is incorrect about 64 g of SO_2 .	A. It is one mole SO_2 B. The number of SO_2 molecule are 6.02×10^{23} C. The number of oxygen atoms are 6.02×10^{23} D. The number sulphur atom are 6.02×10^{23}
11	How many atoms are present in half mol of oxygen gas. Gas exist in diatomic state.	A. 3.01×10^{23} B. 6.02×10^{23} C. 2×10^{23} D. 1.003×10^{23}
12	How many moles of CO are present having 12.04×10^{23} molecules of CO_2 .	A. 0.5 mol B. 1.0 mol C. 1.5 mol D. 2.0 mol
13	A glass is full of water and contains 6.02×10^{23} molecules of H_2O The mass of water molecules is.	A. 18 gm B. 90 g C. 120 g D. 180 g
14	The mass fo 1.505×10^{23} atoms of sulphur is.	A. 0.5 g B. 0.6 g C. 0.7 g D. 0.8 g

15	The number of H ₂ O molecules in 9 grams of ice is	A. 3.01×10^{23} B. 6.02×10^{23} C. 6.02×10^8 D. 12.04×10^{23}
16	The number of moles of hydrogen atoms in 3.2 g of methane CH ₄ .	A. 0.2 B. 0.4 C. 0.6 D. 0.8
17	One mole of Carbon -12 has mass	A. 0.012 kg B. 1 kg C. 0.022 kg D. 12 kg