Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	More ideal gas at room temperature is.	A. CO2 B. NH3 C. SO2 D. N2
2	A real gas can be liquefied if.	A. Temperature is more than critical temperature. B. Temperature is less than critical temperature C. Pressure is more than critical pressure and temperature is less than critical temperature D. Its pressure is less than critical pressure
3	An Ideal gas can not be liquefied because.	A. Its critical temperature is always above 0 ^oC B. It molecules are relatively smaller in size C. Its solidify before becoming a liquid D. Force operative between its molecules are negligible
4	Which one of the following gases cannot be liquefied by Line's method.	A. Water vapours B. NH3 C. Nitrogen D. H2
5	Gas molecules show more deviation from ideal behaviour at high pressure because.	A. Velocity of molecules increases B. Velocity of molecules decreases C. Force of attraction between molecules increases D. Force of collision per unit area increases
6	Under which conditions of temperature and pressure will a real gas behave most like an ideal gas.	A. Low temperature and how pressure B. High temperature and high pressure C. Low temperature and high pressure D. High temperature and low pressure
7	The concept of distribution of velocities among the gas molecules was given by.	A. Clausius B. Mexwell C. Boltzmann D. Vander waal
8	Which is not example of natural plasma.	A. Lightening bolt B. Aurora C. Neon sign D. Sum
9	The total kinetic energy of one mole of an ideal gas is given by	A. 3/2 RT B. 1/2 KT C. 1/2 RT D. 3/2 KT
10	Total pressure of mixture of two gases is.	A. The sum of their partial pressures. B. The difference of their partial pressures C. The product of their practical pressures D. The ratio of their partial pressures
11	Which mixture of gases is used by the deep sea divers.	A. Oxygen and nitrogen B. Oxygen and helium C. Oxygen and carbon di oxide D. Oxygen and water vapours
12	Gases exert pressure on the walls of the container becasue the gas molecules.	A. Collide with each other B. Collide with walls of container C. Have definite volume

		D. Obey the gas laws
13	The diffusion of gases at absolute zero will be	A. Unchanged B. Zero C. slightly decreases D. Slightly increases
14	The rate of diffusion of a gas of molar mass 72 as compared to H2 will be.	A. 1/6 times B. 1.4 times C. 6 times D. same
15	The rate of diffusion of a gas is	A. Directly proportional to its density B. Directly proportional to molecular mass C. Inversely proportional to its density D. Inversely proportional to square root of its molecular mass
16	Which one of the following gases diffuse more rapidly.	A. Cl2 B. CO2 C. CH4 D. N2
17	At which distance a molecule is present from its neighbor molecules of its own diameter, at room temperature.	A. 100 times B. 200 times C. 300 times D. 400 times