

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Long chains of amino acids are coiled about one another into a spiral by	A. Covalent bond B. Ionic bond C. Hydrogen bond D. Van Der Waal's forces
2	Exceptionally low acidic strength of HF is due to.	A. Strong polar bond between H and F B. Smaller size of fluorine C. Strong hydrogen bonding D. electronegativity of fluorine
3	Down the VII -A group, polarizability generally.	A. Increases B. Decreases C. Remain constant D. Negligible
4	Diamond is a bad conductor because.	A. It has tight structure. B. It has a high density C. There is no free electron present in the crystal of diamond to conduct electricity D. None of the above
5	Which of the following is psuedo solid	A. CaF ₂ B. Glass C. NaCl D. Al
6	The molecule of CO ₂ in dry ice form the.	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystals
7	Amorphous solids.	A. Have sharp melting points. B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms.
8	Ionic solids are characterized by.	A. Low melting points B. High vapour pressures C. Good conductivity in solid state D. Solubility in polar solvents
9	In order to mention the oiling point of water at 110 °C, the external pressure should be.	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 765 torr D. Any value of pressure
10	When water freezes at 0 °C, its density decrease due to.	A. Cubic structure of ice B. Empty spaces present in the structure of ice C. Change of bond lengths D. Change of bond angles
11	NH ₃ shows a maximum boiling point among the hydrides of the group V elements due to.	A. Very small size of nitrogen B. Long pair of electrons present on nitrogen C. Enhanced electronegative character of nitrogen D. Pyramidal structure of NH ₃
12	Select the correct answer out of the following alternative suggestions London dispersion forces are the only forces present among the.	A. Molecules of water in liquid state B. Atoms of helium is gaseous state at high temperature. C. Molecules of solid I ₂ D. Molecule of H-Cl gas
13	Observed pressure is less than ideal pressure for any gas due to	A. Intermolecular forces B. Size of molecules C. Boiling point of molecules D. Both a and c

14	What are the SI units of Van der Waal constant 'a'	A. atm dm ³ mol ⁻² B. atm dm ⁶ mol ⁻² C. Nm ⁴ mol ⁻² D. Nm mol ⁻¹
15	What are the SI of excluded volume 'b' in Vander waal equation.	A. dm ³ mol ⁻¹ B. m ³ mol ⁻¹ C. mol dm ⁻³ D. mol m ⁻³
16	The Van der Waals' equation explain the behaviour of.	A. Ideal gas B. Real gas C. Vapours D. Non ideal gases
17	Critical temperature of CO ₂ gas is.	A. 31.1 °C B. 13.1 K C. 13.1 °C D. 1.31 °C