

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	According to MOT, which molecular orbital has highest energy.	A. sigma 1s B. pi+ 2S C. pi 2py D. Pi+ 2px
2	Which one of the following molecules is paramagnetic.	A. H2 B. He C. N2 D. O2
3	Which of the following molecules have unpaired electrons is the bonding molecular orbitals.	A. N2 B. O2 C. B2 D. F2
4	In BeCl2, the covalent bond is formed due to overlap of	A. sp -s B. sp -p C. sp2 -p D. sp3 -p
5	The percentage of s characters in sp3 hybrid orbital is.	A. 25% B. 33.3% C. 50% D. 75%
6	As compared to pure atomic orbitals the hybrid orbitals have.	A. Low energy B. High energy C. Same energy D. None of these
7	How many sigma and pi bonds are present in C2H2.	A. one sigma and two pi B. two pi and one sigma C. Two pi and three sigma D. Three pi and two sigma
8	In which molecule all atoms are coplanar.	A. CH4 B. BF3 C. NH3 D. PH3
9	In which one of the following pairs do the molecule have similar shape.	A. BF3 and AlCl3 B. CO2 and H2O C. CH4 and PH3 D. NH3 and BCl3
10	Both CH4 and NH3 are four electron pair system the angles of CH4 and NH3 are 109.5° and 107.5 ° respectively. This deviation is due to.	A. Hydrogen bonding in ammonia B. Lone pair attraction C. Lone pair occupy more space and repel to other bond pairs D. Lone pair lone pair repulsion
11	A molecule has two lone pairs and two bond pairs around the central atom. The shape of molecule is.	A. Linear B. Pyramidal C. Angular D. Tetrahedral
12	Which one of the following molecule have angle of 120 °	A. Be Cl2 B. BF3 C. CH4 D. NH3
13	Which one of the following correctly describe the shape of NH3 molecules.	A. tetrahedral B. Pyramidal C. Angular D. Square planar
14	Ionic bond is formed by combination of groups	A. IA and VIII B. II A and VII A C. IV A and VA D. VIA and VII A
15	Which molecules is 100% covalent	A. H2 B. H2O C. HF D. HCl

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| 16 | Suppose a new element 'J' has discovered and has seven electron in the valence shell. Which statement about this element would be correct. | A. It is monatomic
B. It form covalent bond with hydrogen
C. It forms stable positive ion
D. It forms covalent bond with group IA element |
| 17 | In which of the following Pairs, do the elements form a compound by sharing electrons. | A. carbon and chlorine
B. Lithium and iodine
C. Neon and oxygen
D. Potassium and bromine |