Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Standard enthalpy of combustion of carbon is -394 kJ mol-1 than, which is the standard enthalpy of formation of CO2	A. +394 kJ B. -394 kJ C. 0 kJ D. +197 kJ
2	Which is not is state function	A. Enthalpy B. Entropy C. Pressure D. work
3	Which substance have delta E = Delta H and no pressure - volume work.	A. Liquids only B. Solids only C. Gases only D. Liquids and solids
4	The sum of all kinds of energies of atoms, ions or molecules of a system is known as.	A. Kinetic energy B. Potential energy C. Internal energy D. Solar energy
5	Which process is endothermic and spontaneous	A. Neutralization of NaOH with HCl B. Formation of NH3 from H2 and N2 C. Formation of H2O from H2 and O2 D. Evaporation of sea water
6	Spontaneous processes are mostly	A. Reversible B. Irreversible C. Not irreversible D. None of these
7	Which of the following value of heat of formation indicates that the product is least stable.	A. -94 KJ B. -231.6 KJ C. +21.4 KJ D. +70 KJ
8	the value of delta H and delta E for liquids and solids is almost same because.	A. No change in temperature B. Heat absorbed C. No change in volume D. Heat evolved
9	Which one of the following process is endothermic	A. condensation of steam B. Freezing of water C. electrolysis of water D. None of these
10	A chemical change always involve	A. Absorption of heat B. Evolution of heat C. Either absorption or evolution of heat D. The liberation of heat and light energy
11	BF3 has zero while NH3 has 1.49 D dipole moment because.	A. B is less electronegative than N B. F is more electronegative than N C. BF3 is pyramidal while NH3 is planar D. NH3 is pyramidal while BF3 is trigonal planar
12	Dipole moment is defined as.	A. Charge x distance B. Charge x Debye C. Charge x displacement D. Charge x bond energy
13	Bond energy depends upon	A. Electronegativity B. Size of atom C. Bond length D. All of these
14	Which bond has more ionic characters in it.	A. C - F B. N - F C. O - F D. F - F

15	Which of the following molecules have highest bond energy	A. F2 B. Cl2 C. Br2 D. I2
16	What is bond order.	A. Number of unpaired electrons B. Number of paired electrons C. Number of electrons present in antibonding molecular orbital D. Number of bond formed between two atoms after overlap
17	A molecular orbital can accommodate maximum electron	A. 2 B. 6 C. 8 D. 10