

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Which one of the following has highest pH	A. Distilled water B. 1 M NH_4OH C. 1 M NaOH D. Water saturated with chlorine gas
2	Which one of the following aqueous solutions has the highest pH	A. 0.1 M NaOH B. 0.1 M HCl C. 0.2 M H_2SO_4 D. 0.1 M HNO_3
3	The sum of pH and pOH is	A. 0 B. 7 C. 14 D. 10
4	When concentration of one product is removed at equilibrium stage, in which direction it moves to reestablish equilibrium.	A. Forward B. Reverse C. Neither forward nor reverse D. Equally move in both direction
5	When solid KI dissolved in water, its heat of solution is positive. What would happen to dissolution when temperature is increased.	A. Increases B. Decreases C. Remain same D. First increases then decreases
6	For the equilibrium system $\text{N}_2 + \text{O}_2 + \text{Heat} = 2\text{NO}$ the equilibrium constant decreases by	A. Decreasing the temperature B. Adding a catalyst C. Adding N_2 D. Adding NO
7	the substance which increase the rate of reaction but remains unchanged at the end of the reaction is called.	A. Indicator B. Promoter C. Catalyst D. Activated complex
8	Almost forward reaction is complete when value of K_c is	A. very high B. Very small C. Neither large nor very small D. No correlation
9	The unit of K_c for the reaction $\text{N}_2 + \text{O}_2 = 2\text{NO}$ will be	A. mol dm^{-3} B. $\text{mol}^{-1} \text{ dm}^3$ C. $\text{mol}^{-2} \text{ dm}^6$ D. No units
10	A chemical reaction $\text{A} \rightleftharpoons \text{B}$ is said to be in equilibrium when	A. Complete conversion of A to B has taken place B. Conversion of A to B is 50% complete C. Rate of transformation of A to B is equal to B to A D. 50% Reactant have been changed to B
11	The pH of 10^{-3} mole dm^{-3} of an aqueous solution of H_2SO_4 is.	A. 3.0 B. 2.7 C. 2.0 D. 1.5
12	Born Haber cycle is used to determine the	A. Lattice energy B. Enthalpy of formation C. Enthalpy of ionization D. Enthalpy of dissociation
13	Which one of the following enthalpies is always an exothermic process.	A. Enthalpy of atomization B. Enthalpy of neutralization C. Enthalpy of ionization D. Enthalpy of dissociation
14	The enthalpy of combustion is	A. Positive B. Negative C. Either positive or negative D. No correlation

15	The heat contents of the system is known as.	A. Entropy B. Enthalpy C. Work D. Free energy
16	Standard enthalpy change when one mole of compound is formed from their elements at standard state is.	A. Heat of formation B. Standard heat of formation C. Heat of combustion D. Standard Heat of neutralization
17	The standard heat of formation is measured at 1 atmosphere and	A. 0 °C B. 100 °C C. 293 °C D. 25 °C