Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	A solution of glucose is 10% The volume in which 1 g mole of it dissolved will be.	A. 1 dm3 B. 1.8 dm3 C. 900 cm3 D. 200 cm3
2	18 g glucose is dissolved in 90 g o water, The relative lowering of vapour pressure is equal to.	A. 1/5 B. 5.1 C. 1/51 D. 6
3	Which statement is correct about solubility product constant.	A. It is applicable at highly soluble substances. B. Value of Kap is independent of temperature C. It is used for homogeneous aquarium system D. It can be used to predict that precipitation will take place or not by combining two ions
4	A solution will be unsaturated if	A. Ionic product = Kap B. Ionic product < Ksp C. Ionic Product > Ksp D. both 'a' and 'b' are correct
5	Some impurities of MgCl2 are present in NaCl which separation technique can be used to separate the impurities.	A. Filtration B. Crystallization C. Common ion effect D. Chromatography
6	When HCl is added to H2S aqueous solution, Its ionization	A. Decrease B. Increase C. Remains constant D. First increases than decreases
7	When small amount of acid or base is added to buffer, its pH.	A. Remain same B. Always increases C. Always decreases D. slightly increases or decreases
8	pKa of CH3COOH is 4.74. The pKb value of CH3COO- ions will be	A. 7 B. 14 C. 9.26 D. zero
9	pH of buffer is calculated by.	A. Sorenson equation B. Mosley equation C. Henderson equation D. De broglie equation
10	One dm3 of a buffer solution containing 0.01 M NH4Cl and 0.1 M NH4OH having pKa of 3 has pH.	A. 4 B. 6 C. 9 D. 10
11	Buffer action can be explained by	A. Common ion effect B. Law of mass action C. Le Chateller's principle D. All above
12	Sum of pKa and pKb is equal to.	A. 1 B. 7 C. 0 D. 14
13	Which acid has less value of pKa.	A. CH3COOH B. H2S C. H2CO3 D. HCl
14	pH of rain water.	A. 7 B. Slightly basic C. slightly acidic D. Highly basic

15	A solution has pH zero. Its H+ ions concertation will	A. zero B. More than unity C. Less than unity D. Unity only
16	Kw for water at 0 oc is 0.1×10^{-34} and at 100 ^{o}C 7.5×10^{-14}, How many times dissociation of water increase from 0 ^{o}Cto 100 ^{o}C	A. 7.5 times B. 50 times C. 75 times D. 100 times
17	A solution have H+ ions concentration 1×10^{-7} , its pH will	A. Acidic B. Basic C. Neutral D. Zero