

Chemistry 9th Class English Medium Online Test

Bond formation energy of one O-H bond is	Sr	Questions	Answers Choice
2 If the Delta H value is negative then reaction will be choose of these choose of the choose of the entral product of the choose	1	Bond formation energy of one O-H bond is	B. 484 kJ/mol C. 486 kJ/mol
The enthalpy of reaction H2+12 2HI The enthalpy of reaction 2H2 +O22H2O The enthalpy of reaction 2H2 +O22H2O The enthalpy of reaction 2H2 +O22H2O The enthalpy of reaction C+O2 CO2 A - 571.6 kJ B - 110.5 kJ C - 393.5 kJ C - 543.8 kJ D - 110.5 kJ C - 53.8 kJ C - 53.8 kJ D - 110.5 kJ C - 53.8 kJ D	2	If the Delta H value is negative then reaction witll be	B. Exothermic C. May or may not be exothermic or endothermic
4 The enthalpy of reaction 2H2 +O22H2O C. 393.5 kJ D. +53.8 kJ D. +10.5 kJ C. +53.8 kJ D. +10.5 k	3	The enthalpy of reaction H2+I2 2HI	B. +53.8 kJ C. 11 kJ
5 The enthalpy of reaction C+O2 CO2 1. 3-933.5 kJ C. +53.8 kJ D110.5 kJ A Surrounding B. Energy C. System D. Both a and b 7 The word energy is used in physics ofr the firt time. A 1902 B. 1858 C. 1805 D. 1802 8 Who use the word energy for the 1st time A Rutherford B. Bohr C. Thomas Young D. None of these 9 All chemical reaction involves. A Enzymes B. Catalyst C. Energy changes D. All of these 10 When NaOH and HCl are mixed the temperature increases. The reaction A Exothermic with a negative enthalpy change. B. Endothermic with a positive enthalpy change. C. Release D. None of these 11 When new bonds are formed, the energy is. A Release B. Remain same C. Release D. None of these	4	The enthalpy of reaction 2H2 +O22H2O	B110.5 kJ C393.5 kJ
The part of the universe that we want to focus our attention called. C. System D. Both a and b A. 1902 B. 1858 C. 1805 D. 1802 R. Who use the word energy for the 1st time. A. Rutherford B. Bohr C. Thomas Young D. None of these A. Enzymes B. Catalyst C. Energy changes D. All chemical reaction involves. A. Exothermic with a negative enthalpy change. B. Endothermic with a positive enthalpy change D. Exothermic with a positive enthalpy change D. Exothermic with a positive enthalpy change. A. Consume B. Remain same C. Release D. None of these A. Consume B. Remain same C. Release D. None of these A. Release D. None of these A. Release B. Remain same C. Release D. None of these C. Consume C.	5	The enthalpy of reaction C+O2 CO2	B393.5 kJ C. +53.8 kJ
The word energy is used in physics ofr the firt time. B. 1858 C. 1805 D. 1802 A. Rutherford B. Bohr C. Thomas Young D. None of these A. Enzymes B. Catalyst C. Energy changes D. All chemical reaction involves. A. Exothermic with a negative enthalpy change. B. Endothermic with a positive enthalpy change. C. Endothermic with a positive enthalpy change. D. Exothermic with a negative enthalpy change. A. Consume B. Remain same C. Release D. None of these A. Release B. Remain same C. Release D. None of these	6	The part of the universe that we want to focus our attention called.	B. Energy C. System
8 Who use the word energy for the 1st time C. Thomas Young D. None of these 9 All chemical reaction involves. A. Enzymes B. Catalyst C. Energy changes D. All of these 10 When NaOH and HCl are mixed the temperature increases. The reaction A. Exothermic with a negative enthalpy change. B. Endothermic with a positive enthalpy change C. Endothermic with a negative enthalpy change D. Exothermic with a positive enthalpy change D. Exothermic with a positive enthalpy change on the alpy chan	7	The word energy is used in physics ofr the firt time.	B. 1858 C. 1805
9 All chemical reaction involves. B. Catalyst C. Energy changes D. All of these A. Exothermic with a negative enthalpy chaqne. B. Endothermic with a positive enthalpy change. C. Endothermic with a positive enthalpy change. C. Endothermic with a negative enthalpy change. C. Endothermic with a positive enthalpy change D. Exothermic with a positive enthalpy change D. Exothermic with a positive enthalpy change D. None of these 12 When old bonds are broken, the energy is. A. Release B. Remain same C. Consume C. Consume	8	Who use the word energy for the 1st time	B. Bohr C. Thomas Young
When NaOH and HCl are mixed the temperature increases. The reaction When NaOH and HCl are mixed the temperature increases. The reaction C. Endothermic with a positive enthalpy change. D. Exothermic with a positive enthealpy change D. Exothermic with a positive enthealpy change A. Consume B. Remain same C. Release D. None of these A. Release B. Remain same C. Consume	9	All chemical reaction involves.	B. Catalyst C. Energy changes
When new bonds ae formed, the energy is B. Remain same C. Release D. None of these A. Release B. Remain same C. Consume	10	When NaOH and HCl are mixed the temperature increases. The reaction	enthalpy chagne. B. Endothermic with a positive enthaly change. C. Endothermic with a negatie enthalpy change D. Exothermic with a positive
When old bonds are broken, the energy is. B. Remain same C. Consume	11	When new bonds ae formed, the energy is	B. Remain same C. Release
	12	When old bonds are broken, the energy is.	B. Remain same C. Consume