

Chemistry 10th Class English Medium Online Test

Sr	Questions	Answers Choice
1	Reactions which have comparable amounts of reactants and products at equilibrium state have:	A. very small K _c value B. Very large K _c value C. Moderate K _c value D. None of these
2	When the magnitude of K _c is very small it indicates.	A. Equilibrium will never establish B. All reactants will be converted to products. C. Reaction will go to completion D. The amount of products is negligible
3	Which one of the following statements is not correct about active mass?	A. Rate of reaction is directly proportional to active mass. B. Active mass is taken in molar concentration C. Active mass is represented by square brackets D. Active mass means total mass of substances.
4	When a system is at equilibrium states?	A. The concentration of reactants and products becomes equal B. The opposing reactions C. The rate of the reverse reaction becomes very low D. The rates of the forward and reverse reactions becomes equal.
5	The characteristics of reversible reactions are the following except:	A. Products never recombine to form reactants. B. They never complete C. They proceed in both ways D. They have a double arrow between reactants and products.
6	For reactions having large K _c value, the reaction proceeds to:	A. Completion B. Equilibrium state C. back ward D. None of these
7	Equilibrium constant has no unit when number of moles of reactants and products are:	A. same B. Different C. Both a and b D. None of these
8	Which gas is used to manufacture king of chemicals sulphuric acid?	A. N ₂ B. O ₂ C. Cl ₂ D. S
9	Which gas is used to prepare ammonia?	A. N ₂ B. O ₂ C. Cl ₂ D. S
10	The % age of nitrogen and oxygen in our atmosphere is:	A. 80 B. 90 C. 95 D. 99
11	Guldberg and waage put forward law of mass action in:	A. 1860 B. 1869 C. 1870 D. 1879
12	When the reaction causes to produced it is called.	A. Chemical equilibrium state B. Static equilibrium C. Dynamic equilibrium D. All