MDCAT Chemistry Chapter 9 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	Energy of atom in compound is	A. Higher than individual atom B. Lower than individual atom C. equal to individual atom D. Impossible to predict
2	Bond will be covalent when electronegativity difference of bonded atom is	A. Equal to 1.7 B. between 0.5 to I.7 C. Greater to 1.7 D. zero
3	Mostly ionic compounds are produced between elements of group	A. IA and IIA B. IB and VIB C. IA. IIA and VII-A D. IA and IB
4	Which one of the followings has polar covalent bonds hut is overall nom-polar molecule:	A. HF B. CO₂ C. CH₄ D. N₂
5	Geometry of simple molecule with sp2 hybridization	A. Triangular planar B. Trigonal C. Square planner D. Pyramidal
6	Carbon-Carbon double bond length in C3H6	A. 154 pm B. 134 pm C. 120 pm D. 105 pm
7	Polarity of a molecule is expressed in terms of	A. Bond strength B. Dipole moment C. Bond length D. Shape
8	A covalent bond may be	A. 100% covalent B. Partial ionic C. 100% ionic D. Both a and b
9	Bonding in MgO is an example of	A. Ionic bond B. Polar bond C. Covalent bond D. Coordination covalent bond
10	Covalent bonds are	A. directional B. Bidirectional C. Multidirectional D. Non directional
11	Energy required to remove electron from an atom	A. Ionization potential B. Electronegativity C. Electropositivity D. Electron affinity
12	Greater shielding effect corresponds to ionization potential value	A. greater B. lesser C. remain same D. no effect
13	Elements having high ionization potential values are	A. metals B. non- metal C. liquids D. solid
14	Ionic bond is produced after complete transfer of	A. nucleus B. neutrons C. electrons D. protons
15	Elements of group IA and IIA are	A. electronegative B. neutral C. electropositive D. non-metals

16	Total number of valence electrons in phosphonium ion (PH4-) is	A. 8 B. 9 C. 12 D. 10
17	pi-bond can be formed by sideways overlap of	A. s-orbital B. d-orbital C. p-orbital D. sp orbital
18	what is the exact value of angle in BF3	A. 90 B. 104.51 C. 119.5 D. 120°
19	Octet rule is not allowed in the formation of	A. NF3 B. B.CF4 C. CCl4 D. PCl5
20	The ionization energy	A. generally increases from left to right in a period B. increases from top to bottom in a group C. does not change in a period D. does not change in a group
21	In a period the atomic radii	A. increase B. decrease C. remain same D. first increase, then decreased
22	In a group, the atomic radii from top to bottom	A. increase B. decrease C. don't change D. show variable trend
23	A molecule that has polar bonds but is overall non - polar	A. IF B. CCl4 C. PCl3 D. All
24	The no. of lp's on oxygen in CO are	A. 1 B. 3 C. 4 D. 2
25	Which one is a non-polar compound?	A. SnCl2 B. PH3 C. Gecl4 D. H2O
26	What is true for a molecule with standard geometry	A. It lacks a lp B. It can't be a donor C. It can be an acceptor D. All
27	H3O+ can't accept a lp because	A. it has positive charge B. The central atom is not electron deficient C. The shell of oxygen has reached its limit D. it already has a coordinate bond
28	What is not true for NH4Cl	A. It has ionic bond B. It has covalent bond C. It has coordinate bond D. It has hydrogen bond
29	At compromise distance the forces dominating between atoms are	A. repulsive forces B. attractive forces C. Dipole induced dipole force D. H-bonding
30	Low IE is a symbol of	A. high electronegativity B. small size C. High electron affinity D. Metallic character
31	Which one of the following has zero dipole moment	A. NH3 B. CHCl3 C. H2O D. BF3
32	Among the following molecules, which one has coordinate covalent (dative) bond?	A. CCl4 B. CO2 C. CO D. CH4

		D. CH_4
33	Which of the following molecule has zero dipole moment?	A. PCl_3 B. BF_3 C. NH_3 D. H_2O
34	For formation of ionic bond, electronegativity difference should be	A. Equal to zero B. Equal to 0.5 C. More then 1.7 D. Less than 1.7
35	The ionization energy of hydrogen atom is	A. Zero B. 131.3kJmole C. 13.13kJ/mole D. 1313kJ/mole
36	The elements for which the value of ionization energy is low can	A. Gain electrons readily B. Lose electron less readily C. Gain electrons with difficulty D. Lose electron readily
37	The shielding effect of inner electron is responsible for	A. Having no effect on ionization energy B. Decreasing ionization energy C. Increasing ionization energy D. Increasing electronegativity
38	What will be the shape of a molecule which containstwo sigma bond pairs and one lone pair?	A. Linear B. V shape C. Tetragonal D. Triangular
39	A molecule which contains two lone pairs and two bond pairs of electrons in valence shell of central atom, geometrical shape of molecules will be	A. Tetrahedral B. Trigonal pyramidal C. Angular D. Linear