

MDCAT Chemistry Chapter 10 Chemical bonding Online Test

Sr	Questions	Answers Choice
1	Which one of the following elements is not an alkali metal?	A. Na B. Sr C. Cs D. Rb
2	Which one of following property is not true about alkali metals?	A. Strongest bases due to their hydrides B. Low ionization energy C. Oxidation number more than +1 D. Form acidic oxides
3	Lithium differs from rest of members of its group due to which of following reasons	A. High E.N of Li+1 B. Small radius C. High charge density D. All above are correct
4	One of the following metals is the most reactive and form super oxide. Indicate that	A. Mg B. K C. Be D. Li
5	Which is the least reactive of all the alkali metals	A. Li B. Na C. K D. Cs
6	The elements of group I-A react violently with water and make the solution	A. neutral B. amphoteric C. acidic D. alkaline
7	Which of the following does not give flame test?	A. Li B. Ba C. Mg D. Sr
8	Which one of the following elements is most electropositive out of group I -A and II-A group?	A. K B. Mg C. Na D. Ca
9	Which set of elements is good loser of electrons	A. F2, Cl2, Br2 B. N, P, As C. O, S, Se D. Li, Na, K
10	Which of the following are not known to form compounds in more than one oxidation state?	A. Transition metals B. Halogens C. Alkali metals D. Noble gases
11	Which of the following belongs to alkaline earth metals	A. Cu B. Zn C. Sn D. Mg
12	Compared with alkaline earth metals, the alkali metals exhibit.	A. lower ionization energies B. greater hardness C. high boiling point D. smaller ionic radii
13	Li resembles with Mg, because	A. the ratio of their charge to size is nearly the same B. both have nearly same size C. both are metallic in nature D. both are found together in nature
14	Soda lime is often employed to remove both	A. H2O and NO2 B. CO2, and NO2, C. H2O and CO2 D. H2S and CO2
15	Elements of group II-A are called	A. f-block elements B. s-block elements C. p-block elements

		D. d-block elements
16	The alkaline earth metals are called so because they	<p>A. form alkaline solution and are present in earth crust as minerals</p> <p>B. form alkaline solution and are found in nature states</p> <p>C. are present in earth crust</p> <p>D. are present in earth crust as their minerals</p>
17	Which element differs from rest of elements of its group?	<p>A. Ba</p> <p>B. Mg</p> <p>C. Ca</p> <p>D. Be</p>
18	Beryllium differs from other elements of group II-A due to	<p>A. high charge density</p> <p>B. comparatively high nuclear charge</p> <p>C. small radius</p> <p>D. all above</p>
19	Which of the following configurations corresponds to alkaline earth metals?	<p>A. [Ar] 3d10, 4s2</p> <p>B. [Ne] 3d2, 3p2</p> <p>C. [Ar] 4s2</p> <p>D. [Ar], 3d10, 4s1</p>
20	Elements of II-A group are called alkaline earth metals due to the reason that	<p>A. they occur in earth only</p> <p>B. they form divalent cations only</p> <p>C. they have ns2 electronic configuration</p> <p>D. their oxides and hydroxides are alkaline in nature and metals are present in earth crust</p>
21	Plaster of Paris is obtained from	<p>A. marble</p> <p>B. bauxite</p> <p>C. gypsum</p> <p>D. limestone</p>
22	Asbestos is commonly used in making	<p>A. wall board</p> <p>B. black board</p> <p>C. soft board</p> <p>D. hard board</p>
23	One of the following does not give the flame test. Which is that	<p>A. Sr</p> <p>B. Ba</p> <p>C. Be</p> <p>D. Na</p>
24	How magnesium reacts with water?	<p>A. In frozen ice water</p> <p>B. With cold water</p> <p>C. In with steam</p> <p>D. In hot state</p>
25	Which of the following element has high m.p and b.p, it acts as a reducing agent, and can react with bases?	<p>A. Sr</p> <p>B. Ca</p> <p>C. Be</p> <p>D. Mg</p>
26	Which alkaline earth metal makes peroxide?	<p>A. Ba</p> <p>B. Be</p> <p>C. Mg</p> <p>D. Ca</p>
27	Corrundam is ore of which element?	<p>A. Al</p> <p>B. Th</p> <p>C. In</p> <p>D. Mg</p>
28	What is the formula of bauxite?	<p>A. $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$</p> <p>B. Al_2O_3</p> <p>C. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10 \text{H}_2\text{O}$</p> <p>D. $\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 5\text{H}_2\text{O}$</p>
29	What is the formula of cryolite?	<p>A. $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$</p> <p>B. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10 \text{H}_2\text{O}$</p> <p>C. Na_3AlF_6</p> <p>D. $\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 5\text{H}_2\text{O}$</p>
30	What is the formula of clay?	<p>A. Asbestos</p> <p>B. Talc</p> <p>C. $\text{H}_2\text{Al}_2(\text{SiO}_4)_2 \cdot 2\text{H}_2\text{O}$</p> <p>D. Na_2SiO_3</p>
31	Formula of sodium beryllite is	<p>A. $\text{Na}_2\text{B}_4\text{O}_7$</p> <p>B. Na_2BeO_2</p> <p>C. BeONa</p> <p>D. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10 \text{H}_2\text{O}$</p>

A. Diaspore

32	Which is not mineral of Al?	B. Corundum C. Bauxite D. Galena
33	What is the formula of dolomite?	A. $\text{CaMg}_3(\text{SiO}_3)_4$ B. MgCO_3 C. $\text{MgCO}_3 \cdot \text{CaCO}_3$ D. MgSO_4
34	What is the formula of magnesite?	A. PbS B. $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$ C. MgCO_3 D. CaCO_3
35	What is the formula of talc or soapstone?	A. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$ B. $\text{H}_2\text{Mg}_3(\text{SiO}_3)_4$ C. Cu_2S D. NaNO_3
36	About 25% of earth crust mass is made up of element	A. Oxygen B. Silicon C. Aluminium D. Aluminates
37	What is name of hydrated variety of quartz?	A. Rose quartz B. Smokey quartz C. Silica D. Opal
38	What is the formula of silica?	A. Si_2O_3 B. SiO_2 C. Si_3O_4 D. SiO^-
39	The general electronic configuration of group IV-A elements is	A. ns^2, np^6 B. ns^2, np^4 C. ns^2, np^3 D. ns^2, np^2
40	Elements of group IV-A are	A. neither strongly electropositive nor strongly electronegative B. strongly electropositive C. strongly electronegative D. none of these
41	The element which exhibits maximum catenation property is	A. C B. Pb C. Ge D. Sn
42	Which of the following elements is most metallic	A. Bi B. sb C. As D. P
43	Which pair of following pair is metalloid?	A. Antimony and bismuth B. Phosphorous and arsenic C. Nitrogen and phosphorous D. Arsenic and antimony
44	Which one of the following doesn't exhibit allotropy?	A. Bi B. As C. N D. P
45	Which element of group V-A and VII-A does not use d-orbital?	A. Nitrogen B. Sulphur C. Arsenic D. Chlorine
46	Which of the following compounds is not known?	A. SbCl B. NCl_3 C. NI_3 D. NCl_5
47	The reaction between Cu and conc. H_2SO_4 produces	A. Cu^{+2} B. SO_2 C. SO_3 D. H_2
48	Which compound gives carbon when heated with conc. H_2SO_4 .	A. Starch B. Ethyl alcohol C. Oxalic acid D. Formic acid
49	Which noble gas is used in mixture used for breathing by divers?	A. Ge B. Ar C. Kr D. He

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Fluorine is largely used in

- A. rocket fuels
- B. making Teflon
- C. making freon
- D. All