Chemistry Fsc Part 2 Chapter 4 Online Test

Sr	Questions	Answers Choice
1	The most electronegative element of group V-A is	A. N B. P C. Sb D. Bi
2	Out of all the elements of group VA, the highest ionization energy is possessed by	A. N B. As C. Sb D. Bi
3	Laughing gas is chemically	A. NO B. NO₂ C. N₂O D. N₄O₄
4	The brown gas formed, when metal reduce HNO_3	A. NO B. NO₂ C. N₂O₃ D. N₂O₅
5	The oxidation of NO in air produces	A. N₂O₃ B. NO₂ C. N₂O₃ D. N₂O₄
6	Which of the following is a reddish brown gas	A. N₂O₃ B. NO₂ C. N₂O₃ D. N₂O₅
7	Which of the following gives acidic oxide	A. N B. As C. Sb D. Bi
8	Which metal is redered passive by HNO_3due to formation of a film of metal oxide over the metal	A. Pt B. Sn C. CO D. Mn
9	Gold dissolves in "Aqua Regia" due to formation of Halide. Point out correct halide	A. AuF₃ B. AuCl₃ C. AuBr₃ D. AuI₃
10	What is %age of calcium phosphate in bone ash	A. 20 B. 40 C. 80 D. 60
11	Maximum number of unpaired electrons is in	A. O₂ B. O₂⁺ C. O₂⁻ D. O₂²⁻
12	Which catalyst is used in contact process	A. Fe₂O₃ B. V₂O₅ C. SO₃ D. Ag₂O
13	Out of all the elements of Group V-A the highest ionization energy is possessed by	A. N B. P C. Sb D. Bi
14	In group V-A elements the most electronegative elements is	A. Sb B. N C. P D. As
15	Oxidation of NO in air produces	A. N₂O B. N₂O₃ C. N₂O₄ D. N₂O₅

16	The brown gas formed when metal reduces HNO_3	A. N₂O₅ B. N₂O₃ C. NO₂ D. NO
17	Out of all the elements of groups VI-A the highest melting and boiling points is shown by the element	A. Te B. Se C. S D. Po
18	Out of the elements of group VA, the highest energy is possessed by	A. N B. P C. Sb D. Bi
19	In group VA elements the most electronegative elements is.	A. Sb B. N C. P D. As
20	The brown gas formed when metal reduces HNO3 is	A. N2O5 B. N2O3 C. NO2 D. NO
21	Out of the elements of group VIA the highest melting and boiling points is shown by the element.	A. Te B. Se C. S D. Po
22	SO2 is not absorbed in water directly to form H2SO4 because.	A. The reaction does not go to completion B. The reaction is quite slow C. The reaction is exothermic D. SO3 is insoluble in water
23	Which catalyst is used in contact process.	A. Fe2O3 B. V2O5 C. SO3 D. Ag2O
24	Which of the following specie has the maximum number of unpaired electrons.	A. O2 B. O2+ C. O2- D. O2⁻²
25	Lowest oxidation state of nitrogen is present in.	A. NH3 B. NO2 C. NO D. HNO3
26	Which element does not have allotropic form	A. Nitrogen B. Phosphorous C. Arsenic D. Antimony
27	Which one of the following oxide is brown in colour.	A. NO B. NO2 C. N2O D. N2O3
28	NO2 can be obtained by heating.	A. NaNO3 B. KNO3 C. Pb(NO3)2 D. NH3NO3
29	NH4NO3 on heating at 200 ^{o}C changes to	A. N2O B. NO C. NO2 D. N2O4
30	When Cu reacts with conc. HNO3 , which one of the following gases is evolved	A. N2O B. NO C. NO2 D. N2O5
31	Which of the following acids possess oxidizing and reducing properties.	A. HCl B. HNO2 C. HNO3 D. H2SO4
32	Which raw material is used for manufacture eof HNO3 by Birkland eyed process	A. NH3 and CO2 B. Air C. Air and gypsum D. Lime stone and urea
		A. 40% B. 50%

33	Bone ash contain calcium phosphate	C. 70% D. 80%
34	Which form of phosphorus is more stable.	A. White B. Red C. Black D. Both a and b
35	P2O5 is usually used as	A. Drying agent only B. Reducing agent C. Both drying and reducing agent D. Both drying agent and oxidizing agent.
36	Each of the following is true for white and red phosphorus except one.	A. Both are soluble in CCl4 B. Both can be oxidized by heating in air C. Both consists of same kind of atoms D. Both can be converted into each other
37	PCl3 reacts with water to form	A. PH3 B. POCl3 C. H3PO4 D. H3PO5
38	Basicity of ortho phosphoric acid is.	A. 1 B. 2 C. 3 D. 4
39	Which allotrope of phosphorus has layers like graphite.	A. white phosphorus B. Red phosphorus C. Black Phosphorus D. Amorphous phosphorus
40	In aqua regia, the ratio of conc. HCl to Conc. HNO3 is	A. 1 : 1 B. 2 :1 C. 1:2 D. 3 : 1
41	What are the number of the electrons in valence shell of P in PCl3	A. 4 B. 6 C. 8 D. 10
42	In which substance phosphorus is not present.	A. Yolk of egg B. Bones C. Apatite D. Galena
43	Which one is metaphosphoric acid	A. HPO3 B. H3PO3 C. H3PO4 D. H4P2O7
44	Which one of the following group of Periodic table called chalcogen family.	A. Group III A B. Group VA C. Group VI A D. Group VII A
45	An element has oxidation state -2, +4, +6 in its compounds. In which group in the periodic table is this element likely to be.	A. Grooup III A B. Group IV A C. Group V A D. Group VI A
46	Role of H2S in the given chemical reaction is H2S + I2---------- 2HI+S	A. Oxidising agent B. Reducing agent C. Dehydrating agent D. As an acid
47	The element which is present in earth crust about 50% is	A. Oxygen B. sulphur C. Carbon D. Nitrogen
48	Chemical formula of stibnite on.	A. BaSO4 B. Sb2S3 C. FeS2 D. ZnS
49	When concentrated H2SO4 and solid sodium chloride react together at room temperature the product are.	A. Two salts only B. A salt and a base C. A salt and an acid D. A salt and water
		A. An acid

50	The reaction between concentrated H_2SO_4 and glucose give carbon and water. In this reaction H_2SO_4 acts as.	B. An oxidising agent C. Dehydrating agent D. A reducing agent
51	Sulphuric acid acts as dehydrating agent in its reaction with.	A. Sodium chloride B. Potassium nitrate C. Copper D. Ethyl alcohol
52	The composition of oleum is.	A. H_2SO_4 B. $H_2S_2O_3$ C. $H_2S_2O_7$ D. $H_2S_3O_7$
53	In pyrite burner, the gas produced is.	A. SO_3 B. SO_2 C. CO_2 D. NO
54	Which one of the following does not react with dilute sulphuric acid.	A. $Mg(OH)_2$ B. Mg C. MgO D. $Mg(NO_3)_2$
55	Arsenic impurities in contact process are removed.	A. By prolong heating the gases B. By treatment with $Fe(OH)_3$ C. In scrubbing tower D. In absorption tower
56	Most likely product formed when formic acid is dehydrated in the presence of conc. H_2SO_4 is.	A. CO_2 and H_2O B. CO, CO_2 and H_2O C. CO_2 and H_2 D. CO and H_2O
57	The reaction between Cu and conc. H_2SO_4 produces	A. SO_3 B. SO_2 C. H_2 D. Cu + ions
58	Which statement is incorrect about H_2SO_4	A. Dehydration agent B. dibasic acid C. Oxidizing agent D. Reducing agent