

## Chemistry Fsc Part 2 Chapter 4 Online Test

| Sr | Questions   | Answers Choice  |
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| 1  | The most electronegative element of group V-A is  | A. N<br>B. P<br>C. Sb<br>D. Bi  |
| 2  | Out of all the elements of group VA, the highest ionization energy is possessed by                          | A. N<br>B. As<br>C. Sb<br>D. Bi   |
| 3  | Laughing gas is chemically  | A. NO<br>B. NO <sub>2</sub><br>C. N <sub>2</sub> O<br>D. N <sub>4</sub> O <sub>4</sub>  |
| 4  | The brown gas formed, when metal reduce HNO <sub>3</sub>  | A. NO<br>B. NO <sub>2</sub><br>C. N <sub>2</sub> O <sub>3</sub><br>D. N <sub>2</sub> O <sub>5</sub>                             |
| 5  | The oxidation of NO in air produces   | A. N <sub>2</sub> O <sub>3</sub><br>B. NO <sub>2</sub><br>C. N <sub>2</sub> O <sub>3</sub><br>D. N <sub>2</sub> O <sub>4</sub>  |
| 6  | Which of the following is a reddish brown gas   | A. N <sub>2</sub> O <sub>3</sub><br>B. NO <sub>2</sub><br>C. N <sub>2</sub> O <sub>3</sub><br>D. N <sub>2</sub> O <sub>5</sub>  |
| 7  | Which of the following gives acidic oxide   | A. N<br>B. As<br>C. Sb<br>D. Bi   |
| 8  | Which metal is redered passive by HNO <sub>3</sub> due to formation of a film of metal oxide over the metal | A. Pt<br>B. Sn<br>C. CO<br>D. Mn  |
| 9  | Gold dissolves in "Aqua Regia" due to formation of Halide. Point out correct halide                         | A. AuF <sub>3</sub><br>B. AuCl <sub>3</sub><br>C. AuBr <sub>3</sub><br>D. AuI <sub>3</sub>                                      |
| 10 | What is %age of calcium phosphate in bone ash   | A. 20<br>B. 40<br>C. 80<br>D. 60  |
| 11 | Maximum number of unpaired electrons is in  | A. O <sub>2</sub><br>B. O <sub>2</sub> <sup>+</sup><br>C. O <sub>2</sub> <sup>-</sup><br>D. O <sub>2</sub> <sup>2-</sup>        |
| 12 | Which catalyst is used in contact process   | A. Fe <sub>2</sub> O <sub>3</sub><br>B. V <sub>2</sub> O <sub>5</sub><br>C. SO <sub>3</sub><br>D. Ag <sub>2</sub> O             |
| 13 | Out of all the elements of Group V-A the highest ionization energy is possessed by                          | A. N<br>B. P<br>C. Sb<br>D. Bi  |
| 14 | In group V-A elements the most electronegative elements is  | A. Sb<br>B. N<br>C. P<br>D. As  |
| 15 | Oxidation of NO in air produces   | A. N <sub>2</sub> O<br>B. N <sub>2</sub> O <sub>3</sub><br>C. N <sub>2</sub> O <sub>4</sub><br>D. N <sub>2</sub> O <sub>5</sub> |

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| 16 | The brown gas formed when metal reduces $\text{HNO}_3$  | A. $\text{N}_2\text{O}_5$<br>B. $\text{N}_2\text{O}_3$<br>C. $\text{NO}_2$<br>D. NO   |
| 17 | Out of all the elements of groups VI-A the highest melting and boiling points is shown by the element | A. Te<br>B. Se<br>C. S<br>D. Po   |
| 18 | Out of the elements of group VA, the highest energy is possessed by                                   | A. N<br>B. P<br>C. Sb<br>D. Bi  |
| 19 | In group VA elements the most electronegative elements is.  | A. Sb<br>B. N<br>C. P<br>D. As  |
| 20 | The brown gas formed when metal reduces $\text{HNO}_3$ is   | A. $\text{N}_2\text{O}_5$<br>B. $\text{N}_2\text{O}_3$<br>C. $\text{NO}_2$<br>D. NO   |
| 21 | Out of the elements of group VIA the highest melting and boiling points is shown by the element.      | A. Te<br>B. Se<br>C. S<br>D. Po   |
| 22 | $\text{SO}_2$ is not absorbed in water directly to form $\text{H}_2\text{SO}_4$ because.              | A. The reaction does not go to completion<br>B. The reaction is quite slow<br>C. The reaction is exothermic<br>D. $\text{SO}_3$ is insoluble in water |
| 23 | Which catalyst is used in contact process.  | A. $\text{Fe}_2\text{O}_3$<br>B. $\text{V}_2\text{O}_5$<br>C. $\text{SO}_3$<br>D. $\text{Ag}_2\text{O}$   |
| 24 | Which of the following specie has the maximum number of unpaired electrons.                           | A. $\text{O}_2$<br>B. $\text{O}_2^+$<br>C. $\text{O}_2^-$<br>D. $\text{O}_2^{2-}$   |
| 25 | Lowest oxidation state of nitrogen is present in.   | A. $\text{NH}_3$<br>B. $\text{NO}_2$<br>C. NO<br>D. $\text{HNO}_3$  |
| 26 | Which element does not have allotropic form   | A. Nitrogen<br>B. Phosphorous<br>C. Arsenic<br>D. Antimony  |
| 27 | Which one of the following oxide is brown in colour.  | A. NO<br>B. $\text{NO}_2$<br>C. $\text{N}_2\text{O}$<br>D. $\text{N}_2\text{O}_3$   |
| 28 | $\text{NO}_2$ can be obtained by heating.   | A. $\text{NaNO}_3$<br>B. $\text{KNO}_3$<br>C. $\text{Pb}(\text{NO}_3)_2$<br>D. $\text{NH}_4\text{NO}_3$   |
| 29 | $\text{NH}_4\text{NO}_3$ on heating at $200^\circ\text{C}$ changes to                                 | A. $\text{N}_2\text{O}$<br>B. NO<br>C. $\text{NO}_2$<br>D. $\text{N}_2\text{O}_4$   |
| 30 | When Cu reacts with conc. $\text{HNO}_3$ , which one of the following gases is evolved                | A. $\text{N}_2\text{O}$<br>B. NO<br>C. $\text{NO}_2$<br>D. $\text{N}_2\text{O}_5$   |
| 31 | Which of the following acids possess oxidizing and reducing properties.                               | A. HCl<br>B. $\text{HNO}_2$<br>C. $\text{HNO}_3$<br>D. $\text{H}_2\text{SO}_4$  |
| 32 | Which raw material is used for manufacture of $\text{HNO}_3$ by Birkland eyed process                 | A. $\text{NH}_3$ and $\text{CO}_2$<br>B. Air<br>C. Air and gypsum<br>D. Lime stone and urea   |
| 33 |   | A. 40%<br>B. 50%  |

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| 33 | Bone ash contain calcium phosphate   | <p>B. 60%</p> <p>C. 70%</p> <p>D. 80%</p>  |
| 34 | Which form of phosphorus is more stable.   | <p>A. White</p> <p>B. Red</p> <p>C. Black</p> <p>D. Both a and b</p>   |
| 35 | P <sub>2</sub> O <sub>5</sub> is usually used as   | <p>A. Drying agent only</p> <p>B. Reducing agent</p> <p>C. Both drying and reducing agent</p> <p>D. Both drying agent and oxidizing agent.</p>   |
| 36 | Each of the following is true for white and red phosphorus except one.   | <p>A. Both are soluble in CCl<sub>4</sub></p> <p>B. Both can be oxidized by heating in air</p> <p>C. Both consists of same kind of atoms</p> <p>D. Both can be converted into each other</p> |
| 37 | PCl <sub>3</sub> reacts with water to form   | <p>A. PH<sub>3</sub></p> <p>B. POCl<sub>3</sub></p> <p>C. H<sub>3</sub>PO<sub>4</sub></p> <p>D. H<sub>3</sub>PO<sub>5</sub></p>  |
| 38 | Basicity of ortho phosphoric acid is.  | <p>A. 1</p> <p>B. 2</p> <p>C. 3</p> <p>D. 4</p>  |
| 39 | Which allotrope of phosphorus has layers like graphite.  | <p>A. white phosphorus</p> <p>B. Red phosphorus</p> <p>C. Black Phosphorus</p> <p>D. Amorphous phosphorus</p>  |
| 40 | In aqua regia, the ratio of conc. HCl to Conc. HNO <sub>3</sub> is   | <p>A. 1 : 1</p> <p>B. 2 : 1</p> <p>C. 1:2</p> <p>D. 3 : 1</p>  |
| 41 | What are the number of the electrons in valence shell of P in PCl <sub>3</sub>   | <p>A. 4</p> <p>B. 6</p> <p>C. 8</p> <p>D. 10</p>   |
| 42 | In which substance phosphorus is not present.  | <p>A. Yolk of egg</p> <p>B. Bones</p> <p>C. Apatite</p> <p>D. Galena</p>   |
| 43 | Which one is metaphosphoric acid   | <p>A. HPO<sub>3</sub></p> <p>B. H<sub>3</sub>PO<sub>3</sub></p> <p>C. H<sub>3</sub>PO<sub>4</sub></p> <p>D. H<sub>4</sub>P<sub>2</sub>O<sub>7</sub></p>                                      |
| 44 | Which one of the following group of Periodic table called chalcogen family.  | <p>A. Group III A</p> <p>B. Group VA</p> <p>C. Group VI A</p> <p>D. Group VII A</p>  |
| 45 | An element has oxidation state -2, +4, +6 in its compounds. In which group in the periodic table is this element likely to be. | <p>A. Group III A</p> <p>B. Group IV A</p> <p>C. Group V A</p> <p>D. Group VI A</p>  |
| 46 | Role of H <sub>2</sub> S in the given chemical reaction is H <sub>2</sub> S + I <sub>2</sub> ----- 2HI+S                       | <p>A. Oxidising agent</p> <p>B. Reducing agent</p> <p>C. Dehydrating agent</p> <p>D. As an acid</p>  |
| 47 | The element which is present in earth crust about 50% is   | <p>A. Oxygen</p> <p>B. sulphur</p> <p>C. Carbon</p> <p>D. Nitrogen</p>   |
| 48 | Chemical formula of stibnite on.   | <p>A. BaSO<sub>4</sub></p> <p>B. Sb<sub>2</sub>S<sub>3</sub></p> <p>C. FeS<sub>2</sub></p> <p>D. ZnS</p>   |
| 49 | When concentrated H <sub>2</sub> SO <sub>4</sub> and solid sodium chloride react together at room temperature the product are. | <p>A. Two salts only</p> <p>B. A salt and a base</p> <p>C. A salt and an acid</p> <p>D. A salt and water</p>   |
| 50 | The reaction of H <sub>3</sub> PO <sub>4</sub> with NaOH is  | <p>A. An acid</p> <p>B. A base</p>   |

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| 50 | The reaction between concentrated $\text{H}_2\text{SO}_4$ and glucose give carbon and water. In this reaction $\text{H}_2\text{SO}_4$ acts as. | B. An oxidising agent<br>C. Dehydrating agent<br>D. A reducing agent   |
| 51 | Sulphuric acid acts as dehydrating agent in its reaction with.   | A. Sodium chloride<br>B. Potassium nitrate<br>C. Copper<br>D. Ethyl alcohol  |
| 52 | The composition of oleum is.   | A. $\text{H}_2\text{SO}_4$<br>B. $\text{H}_2\text{S}_2\text{O}_3$<br>C. $\text{H}_2\text{S}_2\text{O}_7$<br>D. $\text{H}_2\text{S}_3\text{O}_7$                                      |
| 53 | In pyrite burner, the gas produced is.   | A. $\text{SO}_3$<br>B. $\text{SO}_2$<br>C. $\text{CO}_2$<br>D. $\text{NO}$   |
| 54 | Which one of the following does not react with dilute sulphuric acid.  | A. $\text{Mg}(\text{OH})_2$<br>B. $\text{Mg}$<br>C. $\text{MgO}$<br>D. $\text{Mg}(\text{NO}_3)_2$  |
| 55 | Arsenic impurities in contact process are removed.   | A. By prolong heating the gases<br>B. By treatment with $\text{Fe}(\text{OH})_3$<br>C. In scrubbing tower<br>D. In absorption tower  |
| 56 | Most likely product formed when formic acid is dehydrated in the presence of conc. $\text{H}_2\text{SO}_4$ is.                                 | A. $\text{CO}_2$ and $\text{H}_2\text{O}$<br>B. $\text{CO}$ , $\text{CO}_2$ and $\text{H}_2\text{O}$<br>C. $\text{CO}_2$ and $\text{H}_2$<br>D. $\text{CO}$ and $\text{H}_2\text{O}$ |
| 57 | The reaction between $\text{Cu}$ and conc. $\text{H}_2\text{SO}_4$ produces  | A. $\text{SO}_3$<br>B. $\text{SO}_2$<br>C. $\text{H}_2$<br>D. $\text{Cu}^+$ ions   |
| 58 | Which statement is incorrect about $\text{H}_2\text{SO}_4$   | A. Dehydration agent<br>B. dibasic acid<br>C. Oxidizing agent<br>D. Reducing agent   |