

## Chemistry Fsc Part 2 Chapter 1 Online Test

Sr	Questions	Answers Choice
1	Who gave the law of Triads in 1829?	A. Dobereiner B. Mosely C. Newland D. Mendeleev
2	The concept of atomic number was introduced by	A. Alrazi B. Mendeleev C. Moseley D. Dobereiner
3	The basis of modern periodic table is	A. Electron affinity B. Atomic mass C. Ionization Potential D. Atomic number
4	Elements of Groups IIA are called	A. Alkali metals B. Alkaline earth metals C. Coinage metals D. Halogens
5	In modern periodic table VI period contains elements	A. 8 B. 18 C. 10 D. 32
6	Which of the following are alkaline earth metals?	A. Be, Mg, Ca B. Li, Na, K C. Fe, CO, Ni D. B, Al, Ga
7	Which one is an incomplete period	A. 4th B. 5th C. 6th D. 7th
8	Number of elements in the first period of the periodic table is	A. 2 B. 8 C. 14 D. 18
9	Which is the longest periodic table	A. 4 B. 5 C. 6 D. 7
10	Which of the following statement is correct	A. Na atom is smaller than Na <sup>+</sup> B. Na atom is larger than K atom C. F atom is smaller than F <sup>-</sup> D. F atom is larger than F <sup>-</sup>
11	Which of the following element has lowest ionization energy	A. Beryllium B. Boron C. Carbon D. Oxygen
12	Which element has lowest melting point	A. Beryllium B. Magnesium C. Calcium D. Barium
13	Which the correct statement	A. Cl <sup>-</sup> is smaller than Cl atom B. Cl <sup>-</sup> (lon) and Cl (atom) are equal in size C. Na <sup>+</sup> is smaller than Na atom D. Na <sup>+</sup> is larger than Na atom
14	Correct order according to atomic size in the following is	A. Na > K B. Be > Mg C. ...

		C. O &gt; N D. Cl &gt; F
15	Which of the following has highest M.P	A. Aluminium B. Silicon C. Phosphorus D. Sulphur
16	Which of the following the highest hydration energy	A. $\text{Li}^{+}$ B. $\text{Na}^{+}$ C. $\text{K}^{+}$ D. $\text{Mg}^{++}$
17	The element of 2nd period, which has highest ionization energy from the following is	A. Be B. C C. N D. O
18	Keeping in view size of atoms ,which order is the correct one.	A. $\text{Mg} > \text{Sr}$ B. $\text{Ba} > \text{Mg}$ C. $\text{Lu} > \text{Ce}$ D. $\text{Cl} > \text{I}$
19	Mark the correct statement.	A. $\text{Na}^{+}$ is smaller than Na atom B. $\text{Na}^{+}$ is large than Na atom C. $\text{Cl}^{-}$ is smaller than Cl atom D. $\text{Cl}^{-}$ and Cl are equal in size
20	Which statement is incorrect.	A. All the metals are good conductors of electricity B. All the metals are good conductor of heat C. All the metal form positive ions D. All the metal form acidic oxides
21	Hydrogen resembles in properties with	A. IA, IV A and VII A elements B. III A, IV A and V A elements C. II A, IV A and VI A elements D. II A, III A and VII A elements
22	Ionization energy of calcium is.	A. Lower than that of barium B. Lower than that of magnesium C. Higher than that of beryllium D. Lower than that of strontium
23	Electron affinity is measure of energy	A. Required to remover the electron B. Released by adding an electron C. Required to excite an electron D. Released by removing an electron
24	Mark the incorrect statement	A. Metallic character increase down the group B. Metallic character increase from left to right along a period C. Metallic character decrease from left to right along a period D. Metallic character remains the same down the group
25	In which group of periodic table is the element which has atomic number 14.	A. II B. IV C. III D. VI
26	Which one of the following sets has coinage metaals is it.	A. Cu, Hg, Au B. Cu, Ag, Au C. Ag, Au, Hg D. Cu, Fe, Au
27	Fluorine is in group VII A of periodic table. Its chemistry will most closely resembles that of.	A. Argon B. Boron C. Iodine D. Sulphur
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29	What are the total numbers of periods in the modern periodic table	A. 3 B. 5 C. 7 D. 8
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31	What criteria did Mendeleev use in arranging his periodic table.	A. Atomic mass B. Atomic number C. Mass number D. Density
32	Majority of the elements of the periodic table are.	A. Semi metals B. Non metals C. Metals D. Noble metals
33	Transition elements are those	A. Which becomes before uranium B. Which becomes after uranium C. Which are prepared artificially D. Which are in between s-block and p - block elements
34	In which block of periodic table non metals are present.	A. s B. p C. d D. f
35	In which one of the following sets do all three particles have same number of total electrons	A. F <sup>-</sup> , Cl <sup>-</sup> , Br <sup>-</sup> B. Li <sup>+</sup> , Na <sup>+</sup> , K <sup>+</sup> C. N <sup>3-</sup> , O <sup>2-</sup> , F <sup>-</sup> D. Na <sup>+</sup> , Mg <sup>2+</sup> , K <sup>+</sup>
36	In a period, from left to right in the periodic table, the size of atom generally.	A. Increases B. decreases C. Remains constant D. First increase upto the middle of period and then decreases
37	As you proceed across a period in the periodic table the first ionization energy	A. Decrease B. Increase C. Remains constant D. First increase up the middle of period and then decreases
38	Which class of elements shows low value of first ionization potential.	A. Alkali metals B. Alkaline earth metals C. Halogens D. Noble gases
39	Which one of the following elements has the largest second ionization energy.	A. O B. Na C. F D. Ne
40	Which one of the following is a metalloid.	A. sulphur B. Antimony C. Mercury D. Zinc
41	The first ionization energy of Na, Mg, Al and Si are in the order of.	A. Na < Mg < Al < Si B. Na > Mg > Al > Si C. Na > Mg < Al < Si D. Na < Mg > Al < Si
42	Across a period from left to right in the periodic table, the melting and boiling point.	A. Decrease B. Increase C. Remain constant D. First increase upto the middle of period and then decrease
43	Variable valency is generally exhibited by	A. Transition elements B. Alkali metals C. s-block elements D. Gaseous elements
44	In which group, melting point and boiling point increase downward in a group	A. IA B. IIA C. VIIA D. Both a and b
45	In which compound, oxidation state of sulphur is +6	A. H <sub>2</sub> S B. H <sub>2</sub> SO <sub>4</sub> C. H <sub>2</sub> SO <sub>3</sub> D. SO <sub>3</sub>
46	Which has greater hydration energy.	A. Li <sup>+</sup> B. Na <sup>+</sup> C. K <sup>+</sup> D. Mg <sup>2+</sup>
47	Which element when reacts with chlorine forms polymeric halide.	A. Na B. Be C. Ba D. P

48	Ionic Hydrides react with water to form	A. Proton B. Hydride ions C. Hydroxide ions D. Hydronium ions
49	Which hydride is intermediate in nature.	A. NaH B. BeH <sub>2</sub> C. NH <sub>3</sub> D. HCl
50	Element of which group reacts with hydrogen and form ionic hydrides.	A. II A B. IV A C. V A D. VI A
51	Which one of the following elements burns in air to form an oxide which, when shaken with water, give a solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
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54	Which one of the following oxides is amphoteric in nature.	A. MgO B. Na <sub>2</sub> O C. SO <sub>2</sub> D. ZnO
55	The oxides of metal are generally	A. Acidic B. Basic C. Neutral D. Amphoteric
56	Which property of hydrogen not resemble to alkali metals.	A. Electronic configuration B. Oxidation state C. Reaction with halogen D. Metallic nature
57	Which oxide is more basic in nature.	A. BeO B. MgO C. CaO D. BaO