

Chemistry Fsc Part 1 Chapter 5 Online Test

Sr	Questions	Answers Choice
1	Why can gases be compressed easily.	A. Because they have strong intermolecular forces B. Becasue their particles are lightly packed C. Because most of their volume is empty space D. Because they are denser than solids
2	What is the effect of decresing external pressur eon boiling point of a liquid.	A. It increases B. It remains constant C. It decrease D. It fluctuates
3	What types of force is hydrogen bonding.	A. lonic bond B. A special type of dipole dipole force C. Metalic Bond D. London Dispersion force
4	What i the rearranged form of the idela gas equation to calculate molar mass of a gas.	A. M=PVn/RT B. M = PV /MRT C. M = m RT/PV D. M = RT/PV m
5	According to Avodgaro's Law, volume is directly propertional to.	A. Pressure B. Temperature C. Number of moles D. Density
6	The expasnsion of liqudis increae with temperature is	A. What happens to the kinetic energy of liquid molecules when cooled. B. Increases C. Remain unchanged D. Decreases, and the liquid may turn into solid E. Becomes Zero
7	Which of the following has greater viscosity than water	A. Glycerine B. Hexane C. Acetone D. Methanol
8	What is vapour pressure.	A. Atomospheric pressure. B. Pressure exerted by aliqudi at 0 oC C. pressure exercted by vaporus in equilibrium with their liqudi at a given temperature D. Pressrue inside the liquid only
9	Why is the compression of solids not possible	A. The particles are fixed in place and cannot move closer B. Their particles are widely spaced C. The particles are charged D. Solids are made of gases
10	When does a liqudi start boiling.	A. When temperatur reahes 100 oC B. When its vaporu pressur eequals atmospheric pressure C. When all molecules turn into vapour D. When surface tension becomes zero
11	Which of the following statemetns about ideal gases is true.	A. they have strong intermolecular forces B. Their particles have significant volume C. Their volume is mainly due to

		D. They have negligible intermolecular forces
12	Why is water's surface tension so high.	A. Due to its ioic nature B. Because of hydrogen boding pulling surface molecules downward C. Beacue it has low boiling point D. Beacuse water molecules are very large
13	The ideal gas equation is given as.	A. PV =nRT B. PV=RT C. PV = nRT/V D. None of these
14	Which of the following is a characteristic property of crystalling solids.	A. they have a range of melting points B. They are isotrople C. they have adefinite and sharp melting point D. They lack a regular arrangeness of particles
15	amorphous solids are typically	A. Rigid and hard B. Found in lump or fine powder form C. Have a shrp melting point D. Always crystalline in structure
16	Why does water boil at 69 oC on Mount Everest	A. Due to hgih temperature B. Due to low vapour pressure C. Due to low external pressure D. Due to presence of snow
17	At equilibrium in a closed system, which two processes occur at tha same rate.	A. Mellting and Freezing B. Evporation and condensation C. Sublimation and condensation D. Evaporation and boiling
18	When the vapour pressure of a liquid equals the external pressure, what phenomenon occurs.	A. Sublimation B. Bioling C. Freezing D. condensation
19	What happens to the voue of water when it freezes.	A. Because F is less electnegative B. Beacue it is non polar C. Because hydrogen bonding in HF traps the H+ ion D. Beacuse HF is a gas
20	Why can a water molecule formtwo hydrogen bonds on avarage.	A. Beacuse it is polar B. Due to its temperature C. It has 2 hydrogen atm and 2lone pairs D. Becuse it has a covalent bond

particle size