Atomic Structure

Sr	Questions	Answers Choice
1	Quantum numebr values for 2 p orbitals are.	A. n = 2, , l = 1 B. n = 1, 1= 2 C. n = 1, 1= 0 D. n= 2,1 = 0
2	The quantum numebr values for the fourth electron of Be atom are.	A. 1,0,0 B. 2,0,0 C. 2,1,0 D. 1,1,1
3	s-subshell has.....orbital	A. 1 B. 2 C. 3 D. 4
4	An electron having the set of values n = 4, 1 = 0 ,m = 0 and s =+1/2 lies in	A. 2 s B. 4 s C. 3s D. 5s
5	The shape of d-orbitals are like	A. Complicated B. Polar C. Spherical D. Cloverleaf
6	In which subshell electrons are filled first according to n+1 rule.	A. 2p B. 3p C. 4p D. 4s
7	If we remove electrons from an atom until the nucleus is left, then it is called.	A. Ionization energy B. Ionization potential C. Successive ionization energy D. All
8	d-block elements are also called.	A. Inner transition B. Outer transition C. Typical transition D. None
9	How many electron presetn in Cl are.	A. 17 B. 18 C. 19 D. 35
10	Formula for calculating the value of (Azimuthal quantum number)	A. n+1 B. n-1 C. n+2 D. 2n2
11	Atom having same atomic numebr but different mass numebr are called.	A. Isotopes B. Isotones C. Isobars D. Isoelecrons
12	How many d orbitals are there in a givne energy level.	A. 1 B. 5 C. 3 D. 7
13	A p orbital has a characteritic shape with how many lobes.	A. 1 B. 2 C. 3 D. 4
14	What information does the pricnipal quantum numerb (n) gives us about orbitals	A. size B. Shape C. Spin D. Size and shape
15	No two electrons in an atom have same all four quantum numbers, this is called.	A. Aufbau Principle B. Hund's rule C. Pauli's exclusing

		principle D. None
16	How many d-orbitals are showing their lobes along axis.	A. 1 B. 3 C. 2 D. All
17	Mass of neutron in amu is.	A. 1.0073 B. 1.0067 C. 1.0087 D. 1.0000
18	Two electrons in the same orbital have.	A. Same Spin B. Differece spin C. May same or different D. Parallel spin
19	Atoms having same atomic mass but different atomic numebr are called.	A. Isobars B. Isotopes C. Isotones D. Isoelectrons
20	How many subshells are present in M-Shell	A. 1 B. 2 C. 3 D. 4