Chemistry Fsc Part 1 Chapter 11 Online Test

Sr	Questions	Answers Choice
1	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Fast C. Exothermic D. Endothermic
2	When a chemical reaction is completed the	A. Instantaneous rate > average rate B. Instantaneous rate = average rate C. Instantaneous rate is zero D. Both average and instantaneous rates become zero
3	The unit of the rate constant is same as that of the rate of reaction in	A. First order reaction B. Second order reaction C. Zero order reaction D. Third order reaction
4	The influence of temperature on reaction rate is predicated by	A. Free energy change of reaction B. Arrhenius equation C. Van der weal's equation D. Kinetic equation
5	Unit of rate constant is the same as that of the rate of reaction in	A. Zero order reaction B. 1st order reaction C. 2nd order reaction D. 3rd order reaction
6	The quantitative relationship between rate and concentration is given by.	A. Law of mass action B. Rate law C. Both of these D. Le Chatelier's principle
7	The power of which the concentration of a substance appears in the rate expression is known as	A. order of reaction with respect to that substance B. Rate of reaction C. Order of reaction D. Molecularity of reaction
8	Rate of a chemical reaction generally increase rapidly even for small increase in temperature because of rapid increase in the	A. Collisions frequency B. Activation energy C. Average KE of molecules D. Fraction of molecules with energies more than activation energy
9	In zero order reaction the rate is independent of.	A. Temperature of reaction B. Concentration of reactants C. Concentration of products D. None of these
10	Which statement is incorrect about activated complex	A. Short lived B. Maximum energy C. Unstable combination of atoms D. Less energy than Ea
11	Dilatometer method is useful for the reactions that involve.	A. Ionic species B. Where reactant absorb U.V. visible or infrared radiations C. Small volume changes in solutions D. Change in refractive indices
12	Half life period for 235 U is 92	A. 710 million years B. 810 milion years C. 720 million years D. 820 million years
13	When a reaction occurs in many steps than the slowest step is.	A. Mechanism step B. Rate determining step C. enthalpy determining step D. None of the above
14	a zero order reaction is one is which	A. Reactants do not react B. One reactant is in large excess C. Concentration of reactant do not change with passage of time

		D. Rate is of affected by changing concentration of reactants
15	Half -Life for a given reaction is doubled if initial concentration is doubled. The order of reaction is.	A. 0 B. 1 C. 2 D. 3
16	Question Image	A. 1 B. 2 C. 3 D. None of these
17	The rate of reaction	A. Increases as the reaction proceeds B. Decreases as the reaction proceeds C. Remains the same as the reaction proceeds D. May decrease or increase as the reaction proceeds
18	The mathematical relation between the rate of reaction and the concentrations of the reactants is known as the	A. Rate equation B. Rate law C. Arrhenius equation D. Both a and b
19	A type of meals are usually used as catalyst.	A. Coinage metal B. Alkali metals C. Transition metals D. alkaline earth metals
20	If reactants have very low activation energy it means that reaction is.	A. Slow B. Fast C. Endothermic D. Exothermic
21	Photosynthesis a photochemical reaction has order of reaction	A. 0 B. 1 C. 2 D. Fractional order
22	After 3 half lives of a chemical reaction, the % fraction of the amount left is	A. 6.25 B. 75 C. 12.5 D. 50
23	The unit of the rate constant is the same as that of the rate of reaction in	A. First order reaction B. Second order reaction C. Zero order reaction D. Third order reaction
24	the rate of reaction when concentration of reactants are taken unity is called.	A. Average rate B. Instantaneous rate C. Specific rate D. Rate equation
25	In zero order reaction, the rate is independent of	A. Temperature of reaction B. Concentration of reactants C. Concentration of products D. None of these
26	Which statement is incorrect about order of reaction	A. It cannot be determined experimentally B. It is determined experimentally C. Sum of exponents in rate equation D. It can have fraction value
27	The rate of reaction	A. Increase as the reaction proceeds B. Decreases as the reaction proceeds C. Remains the same as the reactions proceeds D. May decrease or increase as the reaction proceeds
28	After 3 half life of a chemical reaction, amount of reactant un reactive will be.	A. 50% B. 25% C. 12.5% D. 6.25%
29	When a reaction proceeds in a sequence of steps, the overall rate is determined by	A. Fastest step B. Slowest step C. Order of different steps D. Molecularity of all steps
	With increase in 10°C temperature, the rate of reaction double. This increase in rate of	A. Decrease in activation energy of reaction B. Decrease in the number of collisions between reactant molecules

30	With increase in 10 oC temperature, the rate of reaction double. This increase in rate of reaction is due to	collisions between reactant molecules C. Increases in activation energy of reactants D. Increase in number of effect collisions
31	Which technique is used ot determine the absorption of radiations.	A. Spectrometry B. dilatometer method C. Refractometric method D. Optical rotation method
32	With increases of 10 oC temperature the rate of reaction doubles. This increase in rate of reactions is due to.	A. Decrease in activation energy of reaction B. Decrease in the number of collisions between reactant molecules C. Increase in activation energy of reactants D. Increase in number of effective collisions
33	The rate of reaction	A. Increases B. Decreases C. Remains the same D. May decrease of increase
34	When a chemical reaction is completed.	A. Instantaneous rate > average rate B. Instantaneous rate = average rate C. Instantaneous rate is zero D. Both average and instantaneous rates become zero
35	The unit of rate constant depends upon	A. Number of reactants B. Concentration terms C. Order of reaction D. Molecularity of reaction
36	The unit of rate constant is same as that of rate of reaction in	A. first order reaction B. Second order reaction C. Third order reaction D. Zero order reaction
37	Half -Life period of a first order reacting is independent of.	A. Initial concentration of the compound B. Conditions of temperature C. Presence of catalyst D. All the above
38	Which one of the following statements is incorrect.	A. Enzymes are protein in nature B. Enzymes are catalyst C. Enzymes can catalyze any reaction D. Urease is an enzyme
39	All radio active disintegration nuclear reaction are of.	A. First order B. Zero order C. 2nd order D. Third order
40	The reaction that involves gases, its rate does not depend upon	A. Catalyst B. Temperature C. Moles dm⁻³ D. Partial pressure
41	The rate of reaction determined at any given time is called.	A. Average rate B. Instantaneous rate C. Spontaneous rate D. Over all rate
42	Glucose can be converted into ethanol by an enzyme.	A. Lipase B. Zymase C. Sucrose D. Urease
43	The substance which decrease the activity of a catalyst is called.	A. Promoter B. Activator C. Inhibitor D. Positive catalyst
44	The true representation for the units of rate constant K for the first order reaction	A. sec⁻¹ B. mole dm⁻³s C. mole dm⁻³s⁻¹ D. mole⁻¹dm⁺³s⁻¹
45	A substance which itself is not a catalyst but increases the activity of a catalyst is called.	A. Promoter B. Poisoner C. Inhibitor

		D. Enzyme
46	The unit of rate constant is the same as that of the rate of reaction is.	A. First order reaction B. Second order reaction C. Zero order reaction D. Third order reaction
47	Velocity constant is the rate of reaction when the concentrations of reactants are	A. Zero B. Unity C. Two D. Three
48	Which properties of liquid is measured by polarimeter	A. Conductance B. Refractive index C. Optical activity D. Change in volume
49	In zero order reaction, the rate is independent of.	A. Temperature of reaction B. Concentration of reactants C. concentration of products D. None of these
50	A substance which rerareds the rate of a reaction is called	A. Inhibitor B. Activator C. Auto-catalyst D. None of these
51	Chemical reactivity of different substance is controlled by	A. Atomic number B. Electronic arrangement C. Mass number D. Number of isotope of reactant elements