Chemistry Fsc Part 1 Chapter 1 Online Test

Sr	Questions	Answers Choice
1	Isotopes differ in	A. Properties which depend upon mass B. Arrangement of electrons in orbital C. Chemical properties which depend upon weight D. Atomic number
2	Many elements have fractional atomic masses. This is because	A. The mass of the atom is itself fractional B. Atomic masses are average masses of isobars C. Atomic masses are average masses of isotopes D. Atomic masses are average masses of isotopes proportional to their relative abundance
3	The mass of one mole of electrons is	A. 1.008 mg B. 0.55 mg C. 0.184 mg D. 1.673 mg
4	The number of moles of CO_2which contain 8.0 g of oxygen	A. 0.25 B. 0.50 C. 1.0 D. 1.50
5	The volume occupied by 1.4 g on N_2at S.T.P is	A. 2.24 dm³ B. 22.4 dm³ C. 1.12 dm³ D. 112 cm³
6	Calcium has isotopes	A. 7 B. 9 C. 1 D. 6
7	Isotopes are the atoms of same element with similar chemical properties but different	A. Atomic number B. Atomic volume C. Atomic weight D. Atomic structure
8	The separation of the different isotopes in the mass spectrometer is done on the basis of	A. Different amounts of positive changes B. Charge e/m value different amounts of negative changes C. Different m/e value D. Velocities of the ions
9	Tin has Isotopes	A. 7 B. 9 C. 11 D. 5
10	Which of the following compound has the highest % o f oxygen by weight	A. CH₃- OH B. C₂H₅OH C. HCOOH D. H₂O
11	I molar volume of a gas at S.T.P is occupied by	A. 1 g of gas B. 6 x 10²³g of gas C. 22.4 m³of gas D. 1 gram molecular mass of gas
12	The volume occupied by 1.6 g of O_2at STP is	A. 22.4 dm³ B. 2.24 dm³ C. 1.12 dm³ D. 112 dm³
13	One mole of CH_3OH and one mole of C_2H_5OH have	A. Equal number of atoms B. Equal number of molecules C. Equal number of electrons D. Equal number of protons
		A. The reaction can be reversile

14	In stoichiometric calculations	B. Side products can be formed C. Law of conservation of mass may not be obeyed D. Law of definite proportions is definitely obeyed
15	How many moles of AgCl are produced by combination of 1.0 mole of $AgNO_3$and 2.0 mole of NaCl	A. 1.0 B. 2.0 C. 3.0 D. 4.0
16	A limiting reactant is that one which	A. Gives greatest number of moles of products B. Gives least number of moles of products C. Is left behind after the completion of reaction D. Is mostly a cheaper substance as compared to other reactants
17	Actual yield is mostly less than the theoretical yield due to the reason that	A. Rates of reactions are slow B. Loss of the product during handling C. Reactions are never completed 100% D. Law of conservation of mass is not true
18	Cadmium has isotopes.	A. 3 B. 4 C. 5 D. 9
19	Bromine has isotopes	A. 2 B. 4 C. 8 D. 6
20	Nickel has isotopes	A. 3 B. 5 C. 6 D. 11
21	The largest number of molecule are present.	A. 3.6 g of H2O B. 4.6 g of C2H5OH C. 2.8 g of CO D. 5.4 g of N2O5
22	Haemoglobin is a Macro molecule and consists of approximately atoms.	A. 5000 B. 10,000 C. 68000 D. 15000
23	In Al2O3, the ratio between the ions is	A. 1 :2 B. 2:1 C. 2:3 D. 3:2
24	Which is not a molecular ion.	A. He+ B. CH3+ C. NH3- D. CO+
25	Paalladium has isotopes.	A. 6 B. ₇ C. 8 D. 9
26	Tin has isotopes.	A. 9 B. 10 C. 11 D. 12
27	The element nickel has isotopes.	A. 3 B. 2 C. 5 D. 7
28	Isotopes differ in.	A. Properties which depend upon mass B. Arrangement of electrons in orbitals C. Chemical properties D. None of these
29	During combustion analysis, CO2 produced is absorbed in	A. Mg (CLO4)2 B. 50% KOH C. CaCl2 D. P2O5
		A. 0.25

30	The Number of moles of CO2 which contain 8.0 of oxygen.	A. 0.25 B. 0.15 C. 0.35 D. 1.45
31	The mass of one mole of electrons is.	A. 1.008 mg B. 0.55 mg C. 0.184 mg D. 1.673 mg
32	The mass of two moles of electrons is	A. 1.10 mg B. 1.008 mg C. 0.184 mg D. 1.673 mg
33	27g of Al will react completely with how much mass of O2, to produce Al2O3.	A. 8 g of oxygen B. 16 g of oxygen C. 32 g of oxygen D. 24 g of oxygen
34	the volume occupied by 1.4 of N2 at S.T.P is	A. 2.24 dm3 B. 22.4 dm3 C. 1.12 dm3 D. 112 cm3
35	Ascorbic acid is vitamin.	A. A B. B C. C D. D
36	The volume occupied by 16 g of CH4 at S.T.P.	A. 224.14 dm3 B. 22.4 dm3 C. 1.12 dm3 D. 2.24 dm3
37	Many element have fractional atomic masses. This is because.	A. the mass of the atom is itself fractional B. Atomic masses are average mases of Isobars C. Atomic masses are average masses Isotopes D. Atomic masses are average masses of Isotopes proportional to their relative abundance.
38	the number of moles of CO2 which contain 8.0 g of oxygen.	A. 0.25 B. 0.50 C. 1.0 D. 1.50
39	The largest number of molecules are present in	A. 3.6 of H2O B. 4.6 g of C2H5OH C. 2.8 g of CO D. 5.4 g of N2O5
40	One mole of SO2 contains	A. 6.02 x 10²³atoms of oxygen B. 18.1 x 10²³atoms of SO2 C. 6.02 x 10²³atoms of Sulphur D. 4 g atoms of SO2
41	The volume occupied by 1.4 g of N2 at S.T.P is	A. 2.24 dm3 B. 22.4 dm3 C. 1.12 dm3 D. 112 cm3
42	The smallest collection of ions in an ionic compound is called.	A. Formula unit B. Chemical formula C. Formula mass D. Molecular formula
43	the number of Al^{3+} ion sin $AlCl_3$is 2.007 x 10^{23}. The number of Cl- ions are.	A. 6.02 x 10²³ B. 3.01 x 10²³ C. 3.01 x 10²³ D. 1.5 x 10²³
44	One mole of ethanol and one mole of ethane have an equal	A. Masses B. Number of atoms C. Number of electrons D. Number of molecules
45	The ration of number of molecules of 2 g H2 g to number of molecules of 64 g gaseous oxygen is.	A. 1:1 B. 1:2 C. 1:32 D. 1:4
		A. Are chemically identical

46	Chlorine atom and chloride ions.	B. Are allotropes of chlroin C. Have same number of electrons D. Have same number of protons
47	Which statement about a molecule is incorrect.	A. It exist independently B. Molecules of a substances are similar C. Hemoglobin is a homoatomic molecules D. Oxygen molecule is a macromolecule
48	CO+ is an example of.	A. Free radical B. Cationic molecular ion C. Anionic molecular ion D. Stable molecule
49	One molecule of hemoglobin in heavier than one atom of hydrogen.	A. 680 times B. 6800 times C. 68000 times D. 680000 times
50	The total number of protons and neutrons present in the nucleus of an atom is called.	A. Mass number B. Atomic number C. Molecular mass D. Relative atomic mass
51	Which statement is incorrect about Dempster's mass spectrometer.	A. Solid sample can be used in this mass spectrometer B. Ions are detected on the basis of mass to charge ratio C. Atoms or molecules are ionized with beam of electron. D. This spectrometer work at one atmospheric pressure.
52	In mass spectrometry, ions are separated on the basis of.	A. Masses only B. Charge only C. Change to mass ratio D. Mass to charge ratio
53	More abundant isotope of an element is the one with.	A. Even atomic number B. Odd atomic number C. Even mass number D. Odd mass number
54	In mass spectrometry, ions are produced by	A. Heat at high temperature B. Passing gas through high voltage plates C. Throwing fast moving electrons on gas molecules D. All of them
55	If the empirical formula of compound is CH2 and its molecular mass is 56 than what is the actual molecular formula of this compound.	A. CH2 B. C2H4 C. C3H6 D. C4H8
56	Which one of the following substances is used as water absorber in combustion analysis.	A. Mg(ClO4)2 B. 50% KOH C. Lime water D. Dilute solution of NaOH
57	One mole of Carbon -12 has mass	A. 0.012 kg B. 1 kg C. 0.022 kg D. 12 kg
58	The number of moles of hydrogen atoms in 3.2 g of methane CH4.	A. 0.2 B. 0.4 C. 0.6 D. 0.8
59	The number of H2O molecules in 9 grams of ice is	A. 3.01 x 10²³ B. 6.02 x 10²³ C. 6.02 x 10⁸ D. 12.04 x 10²³
60	The mass fo 1.505 x 10^{23}atoms of sulphur is.	A. 0.5 g B. 0.6 g C. 0.7 g D. 0.8 g
61	A glass is full of water and contains 6.02 x 10^{23}molecules of H2O The mass of water molecules is.	A. 18 gm B. 90 g C. 120 g D. 180 g
		A. 0.5 mol

62	How many moles of CO are present having 12.04 x 10^{23}molecules of CO2.	B. 1.0 mol C. 1,5 mol D. 2.0 mol
63	How many atoms are present in half mol of oxygen gas. Gas exist in diatomic state.	A. 3.01 x 10²³ B. 6.02 x 10²³ C. 2 x 10²³ D. 1.003 x 10²³
64	Which statement is incorrect about 64 g of SO2.	A. It is one mole SO2 B. The number of SO2 molecule are 6.02 x 10²³ C. The number of oxygen atoms are 6.02 x 10²³ D. The number sulphur atom are 6.02 x 10²³
65	12.04 x 10^{23}atoms of nitrogen gas is equal to.	A. 1 mol B. 2 mol C. 3 mol D. 4 mol
66	What has a mass equal to that of one mole of water.	A. 22.4 dm3 of water B. One mole of steam C. One molecule of water D. Two moles of hydrogen molecules and one mole of oxygen molecules.
67	Which of the following contains the same number of molecules as 9 g of water.	A. 2 g of hydrogen gas B. 14 g of nitrogen gas C. 32 g of oxygen gas D. 44 g of carbon dioxide gas
68	The Avogadro constant is the number of.	A. Atoms in 1 g of helium ga B. Molecules in 35.5 g of chlorine gas C. Atoms in 6 h graphite D. Atoms in 24 g of magnesium
69	The number of electrons in one mole of hydrogen gas is.	A. 6.02 x 10²³ B. 12.04 x 10²³ C. Only two D. Indefinite
70	Stoichiometric calculations cannot applied to reversible reactions because.	A. Product again changes to reactant B. Less product is formed C. Reaction goes only in one direction D. Products do not disappear.
71	How many moles of water results by burning 4 mole of H2 with excess of oxygen.	A. 1 mol B. 2 mol C. 3 mol D. 4 mol
72	How many molecules of CO2 are formed by burning 12 g carbon with excess of oxygen.	A. 3.01 x 10²³ B. 1 x 10²³ C. 6.02 x 10²³ D. 1.03x 10²³
73	What is the mass of aluminium is 204 g of the aluminum oxide Al2O3.	A. 26 g B. 27 g C. 54 g D. 108 g
74	When one mole of each of the following is completely burned in oxygen, which gives the largest mass of carbon dioxide.	A. Diamond B. C2H6 C. Methane D. CO2
75	If four moles of sulphur dioxide are oxidized to sulphur trioxide, how many moles of oxygen are needed.	A. 0.5 B. 1.0 C. 1.5 D. 2.0
76	A limiting reactant is one.	A. Which is present in least amount B. Which produces minimum number of moles of product C. Which produces maximum number of moles product D. Does not effect the amount of product.