Chemistry 9th Class English Medium Unit 8 Online Test

Sr	Questions	Answers Choice
1	In which period and group yu will place the elemnt whic is an important part of the solar cell?	A. Third period and Sixth a grop Group 16 B. Third priod and forth A group Group 14 C. Second period and forth A group Group D. Third prod and fifth A group Group 15
2	Which is the softtest metal.	A. Zn B. Ca C. Na D. Al
3	A yellow solid element exists in allotropic forms whic is also present in fossil fuel. Indicate the name	A. Iodine B. Carbon C. Sulphur D. Aluminium
4	How many electrons can nitrogen accept in its outermost shell.	A. 2 B. 3 C. 4 D. 5
5	Which element is the most reactive element?	A. Florine B. Oxygen C. Chlorine D. Nitrogen
6	Which element has the highest melting point.	A. K B. Cs C. Na D. Rb
7	The element having less value of ionizatin energy and less value of electron affinity is likely to belong to.	A. Group1 B. Group 13 C. Group 16 D. Group 17
8	When we mvoe form left to right in a period, atomic size.	A. Increases B. Decreases C. First increases then decreased D. None of the above
9	Number of peiod in the periodic table are.	A. 7 B. 8 C. 5 D. 16
10	Which of the following grops contain alkaline earth metals.	A. I A B. II A C. VII A D. VIII A
11	Which of the following element belong to VIII A.	A. Xe B. Mg C. Br D. Na
12	Main group elements are arranged in .....groups.	A. 7 B. 6 C. 8 D. 10
13	Period nuber of $^{27}Al_{13}$ is	A. 1 B. 2 C. 3 D. 4
14	All the elements of Group II A are less reactive than alkali metals. This is because these elements have.	A. Decreased nuclear charge B. Similar electronci configuration C. High ionization energies D. Relatively greatr atomic size.

15	The atomic radii of the elemtns in periodic table.	A. Increase from left to right in a period B. Do not chage from left to right in a period C. Increase from top to bottom in a group D. Decrease from top to bottom in a group
16	4th and 5th priod of the long form of periodic table are called.	A. Short periods B. Normal periods C. Very long peiods D. Long periods
17	Which one of the following halongesn has lowest electronetivity	A. Iodine B. Chlorine C. Fluorine D. Bromine
18	Transition elements are	A. All gases B. All non metals C. All Metals D. All metalloids
19	How many groups are present in the modern periodic table.	A. 8 B. 10 C. 15 D. 18
20	How many periods are present in the modern periodic table	A. 7 B. 8 C. 10 D. 12
21	How many periods are presnet in the modern periodic table.	A. 7 B. 8 C. 10 D. 12
22	How many elements are present in 1st period.	A. 1 B. 2 C. 8 D. 18
23	How many elements are prsent in each 2nd and 3rd period.	A. 2 B. 32 C. 18 D. 8
24	How many elements are present in each 4th and 5th period.	A. 2 B. 8 C. 32 D. 18
25	How many elements are present in 6th period.	A. 2 B. 8 C. 18 D. 32
26	How many elements are present in 7th period.	A. 2 B. 8 C. 18 D. 23
27	How many blocks are presnt in modern periodic table	A. 2 B. 3 C. 4 D. 5
28	Elments re classified into four blocks depending upon	A. Shell B. Atomic mass C. Sub -Shell D. Atomic Number
29	The elementss of group 1 and 2 are placed in which block	A. s B. p C. d D. f
30	Which of the following elemens is presnet in 1st period.	A. Hydrogen B. Helium C. Both a and b D. None of these
31	Second and third periods are called	A. 1st transition series B. Normal periods C. 2nd transiion series D. 3rd transiliton serios

32	Whcih element is presnet in 2nd period.	A. Lithium B. Beryllium C. Boron D. All of these
33	Elements with atomic no .58 to 71 are called.	A. Actinides B. Lanthanides C. Both a and b D. None of these
34	Actinidws belong to period.	A. 4th B. 5th C. 6th D. 7th
35	Lanthanide series starts after the elemetn	A. Osmium B. Actinium C. Lanthanum D. None of these
36	Atomic number of lanthanum is	A. 57 B. 58 C. 59 D. 60
37	Actinide series starts after the element	A. Actinium B. Lanthanum C. Osmum D. Silver
38	Atomic number of actinium is	A. 57 B. 60 C. 89 D. 80
39	Group nuebr tells about the	A. Number of shells B. Number of valence electrons C. Both a and b D. None of these
40	Period nuebr tells abou the	A. No. of valance electrons B. No. of electronic shells C. Both a and b D. None of the above
41	Whcih period of the modern periodic table is considered as incomplete period.	A. 4th B. 5th C. 6th D. 7th
42	Whic period of the moden periodic table is condidered as incommplete period.	A. 5th B. 4th C. 7th D. 6th
43	/which of the followign elements is presnt in group IA.	A. Lithium B. Hydrogen C. Sodium D. All of these
44	Elements of Group1 are called.	A. Alkali Metlas B. Alkali earth metals C. Transition metals D. Halogen
45	How many elecntrons are present in the valence shell of group 1 elements.	A. 1 B. 2 C. 3 D. 4
46	17th group elements ae known as	A. Alkaki metals B. Alkaline earth metals C. Noble gases D. Halogens
47	17th Group of elements contain electrons in their outer most shell	A. 4 B. 5 C. 7 D. 6
48	The elements of group 3 to 12 are clled.	A. Normal elements B. Halogens C. Noble gases D. Transition elements
49	All transition elements belong to	A. s and p block B. d- block C. f-block D. d and f block

50	The vertical columns present in the priodic table are called.	A. Group B. Period C. Both a and b D. None of these
51	The horizontal lines present in the priodic table are called.	A. Groups B. Periods C. Both a and b D. None of these
52	With the increase of atomic numebr , the number of electron in an atom also.	A. Decreases B. First increases then decreases C. Increases D. None of the above
53	Elements o group 13 to 18 have thier valence electrons is subshell	A. s B. p C. f D. d
54	Which is strongest oxidizing agent.	A. Cholorine B. Iodine C. Fluorine D. Bromine
55	Which halogen memebr exists in a liquid stte at room temperature	A. Bromine B. Chlorine C. Fluroine D. Iodine
56	Elements of a period show properties.	A. Same B. Different C. Both a and b D. None of these
57	The elements of a group show properties.	A. Same B. Different C. Both a and b D. None of these
58	The amoutn of energy given out when an electron is added to an eatom is called.	A. Electron affinity B. Lattice energy C. Ionization energy D. Electronegativity
59	Aong the period which one of the followig decreases.	A. Electronegativity B. Ionization energy C. Atomic radius D. Electron affinity
60	Mark the icorrect statement about ionization energy.	A. It is measured in kJmol-1 B. It is absorption of energy C. It decreased in a period D. It decrese in a group
61	Point out the incorrect statement about electron affinity	A. It decreases in a period B. It decreases in a group C. It is measure din kJmol-1 D. None of these
62	Unit of atomic size is	A. pm B. nm C. kJmol-1 D. Both a and b
63	The distance between the neclei of two carbon atoms in its elemetnal from is	A. 150 pm B. 152 pm C. 154 pm D. 156 pm
64	When we move form left to right in a period, atomic number	A. Decreases B. Increases C. First increases then decreases D. None of the above
65	When we mvoe from top to bottom in a group atomic size.	A. Decrease B. Increases C. First increases then decreases D. None of the above
66	The minium amont of energy whcih is required to remvoe an electron from valence shell of the gaseous state of an atom is called.	A. Potential energy B. Ionization energy C. Electron affinity D. Electronegativity
67	The unit of ionization energy is	A. nm and pm B. kJ mol-1 C. Pascal

		D. Newton
68	When we mvoe top to bottom in group, ionization energy.	A. Increases B. No effet C. Decreases D. None of these
69	When we mvoe from left to right in a period, ionization energy.	A. No effect B. Decreases C. Increases D. None of these
70	Uni of electron affinity is.	A. k J mol-1 B. k jmol C. pm D. Newton
71	Electron affinity of fluorine in kJmol-1 is	A. -328 B. 328 C. -330 D. -340
72	The ability of an atom to attract the shared pair of electons towards itself in a molecule is called	A. Ionization energy B. Electronietativity C. Shielding effect D. Electron affinity
73	Which one of the followign halogns has highest electronegativity	A. Iodine B. Fluorine C. Chlorine D. Bromine
74	Electronegativity of oxygen is.	A. 3,1 B. 3,3 C. 3.4 D. 3,2
75	The electronegativity of carbon is	A. 2.5 B. 2.0 C. 1.0 D. 4.0
76	Metals can form ions carrying carges.	A. Uni positive B. Di positive C. Tri postive D. All of these
77	Pure alkalis metals can be cut siply by knife but iron cannot bccause of alkali meals have	A. Non metalic bonding B. Strong metallic bonding C. Weak metallic bonding D. Moderate metqllic bonding
78	Metals lose their elecrons easily because.	A. They are elecrnegativity B. They have electron affinity C. They are electropositive D. Good conductors of heat
79	Metals are teh elements which have.	A. Electronegative character B. Electropositive character C. Both a and b D. None of the above
80	Which are good conductor of heat and electricity	A. Metals B. Non metals C. Metalloids D. All of these
81	All metals bear	A. Positive charge B. Negative charge C. Both a and b D. None of these
82	Metals posses.	A. Ionic bond B. Covalent bond C. Co-ordinate covalent D. Metallic bond
83	Sodium metal has electrons	A. 10 B. 11 C. 12 D. 14
84	Which group elements has low ionization energies.	A. Halogens B. Noble gases C. Alkaline Earth Metals D. Alkali Metals
		A. Gold

85	Platinum alloyed with which metal is used as catalyst in automobiles as atalytic covertor.	B. Rhodium C. Palladium D. Both a and b
86	Which of the following is a metal	A. Magnesium B. Carbon C. Hydrogen D. Nitrogen
87	The haviest metal is	A. Iron B. Lead C. Osmium D. Platinum