

## Chemistry 9th Class English Medium Unit 6 Online Test

Sr	Questions	Answers Choice
1	What will happen if the rates of forward and reverse reactions are very high	A. The reaction will be practically irreversible B. The equilibrium point will reach very soon C. The equilibrium point will reach very late D. The reaction will not attain the state of dynamic equilibrium
2	Predict which components of the atmosphere react in the presence of lightning.	A. N <sub>2</sub> and H <sub>2</sub> O B. O <sub>2</sub> and H <sub>2</sub> O C. N <sub>2</sub> and O <sub>2</sub> D. CO <sub>2</sub> and O <sub>2</sub>
3	An Inorganic chemistry places one mole of PCl <sub>5</sub> in container A and one mole of each Cl <sub>2</sub> and PCl <sub>3</sub> in container B. Both the containers were sealed and heated to the same temperature to reach the state of equilibrium. Guess about the composition of mixtures in both the containers.	A. Both the containers will have zero concentration of its reactants. B. Both the containers will have the same composition of mixtures C. Container A will have more concentration of PCl <sub>3</sub> than B. D. Container A will have less concentration of PCl <sub>3</sub> than B.
4	CaO or lime is used extensively in steel, glass and paper industries. It is produced in an exothermic reversible reaction by the decomposition of lime. Choose the conditions to produce maximum amount of lime.	A. Heating at high temperature in an open vessel B. Heating at high temperature in a closed vessel C. Cooling it in a closed vessel D. Cooling it in an open vessel
5	What condition should be met for the reversible reaction to achieve the state of equilibrium.	A. The concentration of all the reactants and the product should become constant B. All the reactants should be converted into the product C. 50% of the reactant should be converted into products. D. One of the products should be removed from the reaction mixture.
6	Why does the gas start coming out when you open a can of fizzy drink.	A. Because of the solubility of the gas increases B. Because the gas is dissolved under pressure hence it comes out when pressure is decreased C. Because the gas is insoluble in water D. Because the solubility of the gas decreases at high pressure.
7	In an irreversible reaction equilibrium is.	A. The forward reaction will be favoured B. No effect on forward or backward reaction C. No effect on backward reaction D. The backward reaction will be favoured
8	When will a reaction become a reversible one?	A. If the activation energy of the forward reaction is comparable to that of backward reaction B. If the activation energy of the forward reaction is higher than that of backward reaction C. If the activation energy of the forward reaction is lower than that of backward reaction D. If the enthalpy change of both the reactions is zero.
9	If reversible reaction is useful for preparing compounds on a large scale.	A. Yes B. No C. They are useful only when equilibrium lies far to the left side D. They are useful only when

10	What will happen to the concentrations of the product if a reversible reaction at equilibrium is not disturbed.	A. They will keep on increasing B. They will keep on decreasing C. They will remain constant D. They will remain constant for some time and then start decreasing
11	In an irreversible reaction equilibrium	A. Never established B. Established quickly C. Established slowly D. Established when reaction stops
12	The characteristics of reversible reactions are the following except.	A. Product never recombine to form reactants B. They never complete C. They have a double arrow between reactants and products D. They proceed in both ways
13	A reverse reaction is one that	A. Speeds up gradually B. Proceeds from left to right C. In which reactants react to form products D. Slow down gradually
14	The reaction in which the products do not recombine to form reactants are called	A. Addition reactions B. Decomposition reactions C. Irreversible reactions D. reversible reactions
15	The reaction in which the products can recombine to form reactants are called.	A. Reversible Reaction B. Irreversible reactions C. Decomposition reactions D. Addition reactions
16	Which type of reactions speed up gradually?	A. Decomposition reaction B. Forward reaction C. Reverse reactions D. Irreversible reactions
17	Such reaction which continue in both directions are called.	A. Dynamic B. Irreversible C. Reversible D. Non- reactive
18	In chemical reaction, the substances that combine are called.	A. Masses B. Materials C. Products D. Reactants
19	The forward reaction takes place from	A. Right to left B. Left to right C. Both a and b D. None of these
20	A complete reaction is in which	A. Only 10% reactants convert into products B. All the reactants convert into products C. All the reactants do not convert into products D. Half reactants convert into products
21	In the beginning the rate of reverse reaction is.	A. Slow B. very fast C. Moderate D. Negligible
22	The new substance formed in a chemical reaction is.	A. Reverse B. Reactant C. Forward D. Product
23	The colour of anhydrous copper (II) sulphate solid is	A. Pink B. Black C. White D. Blue
24	The colour of hydrated copper (II) sulphate solid is.	A. Black B. Pink C. White D. Blue
25	The colour of anhydrous cobalt(II) chloride solid	A. White B. Black C. Pink D. Blue

26	The colour of hydrated cobalt(II) chloride solid is	A. white B. Black C. Blue D. Pink
27	Which of the following does not happen, when a system is at equilibrium state.	A. Reaction continues to occur in both the directions B. Concentration of reactants and products stop changing C. Forward and reverse reactions stop D. Forward and reverse rates become equal
28	Which is true about the equilibrium state?	A. The forward reaction stops B. Both forward and reverse reactions stop C. Both forward and reverse reactions continue at the same rate D. The reverse reaction stops
29	When system is at equilibrium state.	A. The rate of the forward and reverse reactions become equal B. The concentration of reactants and product becomes equal C. The opposing reactions stop D. The rate of the reverse reaction becomes very low
30	When the rate of the forward reaction takes place at the rate of reverse reaction the composition of the reaction mixture remains constant. It is called.	A. Chemical Equilibrium B. Static equilibrium C. Both a and b D. None of the above
31	Concentration of reactants and product at equilibrium remains unchanged if	A. Concentration of any reactant or product is not changed B. Temperature of the reaction is not changed C. Pressure or volume of the system is not changed D. All of the above are observed
32	At what temperature, rate of ammonia formation and decomposition is the highest.	A. 200 °C B. 300 °C C. 400 °C D. 500 °C
33	Industrially, ammonia is produced by which process.	A. Halogenation B. Solvay process C. Haber Process D. Hydrogenation
34	Formation of ammonia from Nitrogen and hydrogen is an.	A. Exothermic reaction B. Endothermic reaction C. Both a and b D. No heat change
35	How much heat absorbed when NH <sub>3</sub> decomposed into N <sub>2</sub> and H <sub>2</sub> ?	A. 90.4 kJ/mol B. 92.4 kJ/mol C. 94.2 kJ/mol D. 95.2 kJ/mol
36	Which compound is used as a thinner in paint industry?	A. H <sub>2</sub> O B. C <sub>2</sub> H <sub>3</sub> OH C. CH <sub>3</sub> COOC <sub>2</sub> H <sub>5</sub> D. CH <sub>3</sub> COOH