Chemistry 9th Class English Medium Unit 6 Online Test

Sr	Questions	Answers Choice
1	What will happen if the rates of forward and reverse reactions are very high	A. The reaction will be practiclly irreversible B. The equilibrium point will reach very soon C. The equilibiurm point will reach very late D. The reaction will not attain the state of dynamic equilibrium
2	Predict which components of the amosphere react in the presence of lightening.	A. N2 and H2O B. O2 and H2O C. N2 and O2 D. CO2 and O2
3	An Inorganic chemistry places one mle of PCl5 in container A and one mole of each Cl2 and PCl3 in container B. Both the containers were sealed and heated to the same temperture to reach the stte of equilibrium Guess about the composition of mixtures in both the containers.	A. Both the containers wil have zero concentration of its reactants. B. Both the containers wil hae the same composition of mixtures C. Container A will have more concentraion of PCl3 than B. D. Container A will have less concentraion of PCl3 than B.
4	CaO or lime is used extensive in steel, glass and paper industries. It is produced inan exothermic reversitble reaction by the decompositon of lie . Choose the conditions to produce maximum amount of lime.	A. Heating at high temperatur ein an open vessel B. Heating at high temperatur ein a closed vessel C. Cooling it in a closed vessel D. Colling it in an open vessel
5	What conditon Should be met for the reversible reactio to achieve the state of equilibrium.	A. The concentratio of all the reactants and the prodduct should become constant B. all the reactants should be converted into the product C. 50% of the reactant shuld be converted into prodcuts. D. One of the product should be removed from the reaction mixture.
6	Why the gas starts coming out when you open a can of fizzy drink.	A. Because gteh solubiliyt of the gas increases B. Because the gas is dissolved under pressure hence it comes out when pressur eis decreased C. Becaue the gas is insoluble in water D. Because the solubility of hte gas decreases at high pressure.
7	In an irreversibel reaction equiirbrium is.	A. The forward reaction will be fovoured B. No effect on forward or backward reaction C. No effect on bakcwars reaction D. The backward reaction will be favoured
8	When a reaction will be come a reversible one?	A. If the actvation eneryg of the forward reaction is comparrble to that of backward reaction B. If the activation energy of the forward reaction is higher than that of backward reaction C. If the activationenergy of the forward reaction is lower than that of backward reaction D. If the enthalpy change of both the reactions is zero.
9	If reversibe reaction useful for preparing compounds on large scale.	A. Yes B. No C. They are useful only when equilirbium lies far to the left side D. They are useful only when

		equilibrium lies far to the right side
10	What will hapen to the concentrations of the prodcut if a reversible reaction at eqiulibrium is not distrubed.	A. They will keepon increasing B. They will keep on decreasing C. They will remain constant D. They willremain constant for some time and then start decreasing
11	In an irrversibel reaction equilibrium	A. Never establised B. Established quickly C. Established slowly D. Established when reaction stops
12	The characteristics of reversible reactions are the following except.	A. Product nver recombine to form reactants B. They never complete C. They have a double arrow between reactants and products D. The proceed in both ways
13	A reverse reaction is one that	A. Speeds up gradually B. Proceeds from left to right C. In which reactants react to form products D. Slow down gradually
14	The reaction in which the products do not reombine to form reatants are called	A. Addition reactions B. Decomposition reactions C. Irreversitble reactions D. reversible reactions
15	The reaction in which the products can recombine to formreactants are called.	A. Reversible Reaction B. Irreversible reactions C. Decomposition reactions D. Addition reactions
16	Which type of reactions speed up gradully?	A. Decomposition reaction B. Forward reaction C. Reverse reactions D. Irreversibel reactions
17	Such reaction which continue in both directins are called.	A. Dynamic B. Irreversible C. Reversible D. Non- reactive
18	In chemical reaction, the substances that combine are called.	A. Masses B. Materials C. Products D. Reactants
19	The forward reaction takes place from	A. Right to left B. Left to right C. Both a and b D. None of these
20	A complete reaction is in which	A. Only 10% reactants covert into products B. All the reactants covrt into products C. All the reactants do no covrt into products D. Half reactants covert into produts
21	In the beginning the rate of reverse reaction is.	A. Slow B. very fast C. Moderate D. Negligible
22	The new substance formed in a chemcial reaction is.	A. Reverse B. Reactant C. Forward D. Product
23	The colour of anhydrous copper (II)sulphate solid is	A. Pink B. Black C. White D. Blue
24	The colour of hydrated copper (II) sulphate solid is.	A. Black B. Pink C. White D. Blue
25	The colour of anhydrous cobalt(II) cholride solid	A. White B. Black C. Pink D. Blue

26	The colour of hydrated cobalt(II) chloride solid is	A. White B. Black C. Blue D. Pink
27	Which of the following does not happpen, when a system is at equilibrium state.	A. Reaction continues to occur in both the directions B. Concentration of reactants and products stop changig C. Forward and reverse reactions stop D. Forward and reverse rates become equal
28	Whcih is true about the equilibrium state?	A. The forard reaction stops B. Both forward and reverse reactions stop C. Both foward and reverse reactions continue at the same rate D. The reverse reaction stops
29	When system is at quilibrium state.	A. The rate of the forward and reverse rections become equal B. The concentrationof reatants and product becoes equal C. The oposing reactions stop D. The rate of the reverse reactio becomes very low
30	When the rate of the forward reaction takes place at the rate of reverse reaction the composition of the reaction mixgure remains consant. It is called.	A. Chemical Equilibrium B. Static equilirbirum C. Both a and b D. None of the above
31	Concentratin of reactants and product at equilibrium remains unchanged if	A. Concentration of any reactant or product is not changed B. Temperaure of the reation is not changed C. Pressure or volume of the system is not changed D. All of the above are observed
32	At what temperature , rate of ammonia formation and decomposition is the highest.	A. 200 ^oC B. 300^oC C. 400 ^oC D. 500 ^oC
33	Industrialy, ammonia is produced by which process.	A. Halogenation B. Solvay process C. Haber Process D. Hydrogenation
34	Formation of ammonia from Nitrogen and hydrogen is an.	A. Exothermic reaction B. Endothermic reaction C. Both a and b D. No heat change
35	How much heat absorbed when NH3decomposed into N2 and H2?	A. 90.4 kJ/mol B. 92.4 kJ/mol C. 94.2 kJ/mol D. 95.2 kJ/mol
36	Which compound is used a thinner is paint industry?	A. H2O B. C2H3OH C. CH3COOC2H5 D. CH3COOH