

Chemistry 9th Class English Medium Unit 4 Online Test

Sr	Questions	Answers Choice
1	How many atoms are present in one gram of H ₂ O?	A. 1002×10^{23} atom B. 6.022×10^{23} atom C. 0.334×10^{23} atom D. 2.004×10^{23} atom
2	Which is the correct formula of calcium phosphide.	A. CaP B. Ca_3P_2 C. CaP_2 D. Ca_2P_3
3	How many atomic mass units (amu) are there in one gram.	A. 1 amu B. 6.022×10^{23} C. 10 amu D. 6.022×10^{22}
4	How many moles are there in 25 g of H ₂ SO ₄ ?	A. 0.765 Moles B. 0.255 moles C. 0.4 moles D. 0.51 moles
5	A necklace has 6 g of diamonds in it. What are the number of carbon atoms in it?	A. 3.01×10^{23} B. 1.003×10^{23} C. 12.04×10^{23} D. 6.02×10^{23}
6	What is the mass of Al in 204 g of aluminium oxide Al ₂ O ₃	A. 26 g B. 54 g C. 108 g D. 27 g
7	Which one of the following compounds will have the highest percentage of the mass of nitrogen?	A. N ₂ H ₄ B. CO (NH ₂) ₂ C. NH ₃ D. NH ₂ OH
8	When one mole of each of the following compounds is reacted with oxygen, which will produce the maximum amount CO ₂ ?	A. Carbon B. Ethane C. Diamond D. Methane
9	What mass of 95% CaCO ₃ will be required to neutralize 50 cm ³ of 0.5 M HCl solution.	A. 9.5 g B. 1.45 g C. 1.32 g D. 1.25 g
10	Formula of Ozone.	A. O ₂ B. S ₈ C. CO ₂ D. O ₃
11	Avogadro was a scientist	A. Italian B. Greek C. German D. African
12	Which of the following is insoluble salt	A. KCl B. AgCl C. NaCl D. CaCl ₂
13	Stoichiometric calculations are used to prepare.	A. Soaps B. Shampoo C. Perfumes D. All of these
14	Without stoichiometry which industry cannot exist.	A. Metal B. Petroleum C. Leather D. Chemical
15	Which law is obeyed in chemical calculations?	A. Law of mass action B. Law of conservation of mass C. Law of definite proportion D. Both a and b

16	Empirical formula of sand is.	A. SiO ₂ B. SiO ₃ C. SiO ₄ D. SiO
17	Empirical Formula of Glucose is	A. CH ₂ O B. CHO C. CHO ₂ D. C ₂ HO
18	Empirical formula of acetic acid (CH ₃ COOH) is	A. CHO B. CH C. CH ₂ O D. None of these
19	Which of the following represent sand?	A. NaCl B. CaCO ₃ C. H ₂ O D. CH ₂ O
20	Empirical formula of hydrogen peroxide.	A. HO B. CO C. CHO D. CH
21	Empirical formula of Benzene is.	A. CH ₂ O B. CH C. C ₂ H D. CH ₂
22	Which compound has same molecular and empirical formula.	A. C ₆ H ₁₂ O ₆ B. H ₂ O ₂ C. H ₂ O D. C ₆ H ₆
23	Value of Avogadro's number is.	A. 6.6×10^{20} B. 6.02×10^{23} C. 6.00×10^{24} D. 1.32×10^{23}
24	1 gram formula of NaCl contains is grams.	A. 100 g B. 58.5 g C. 32 g D. 49 g
25	1- gram atom of carbon contains how many moles.	A. 1 mole B. 2 mole C. 6 moles D. 12 moles
26	Formula of common salt is	A. NaCl B. AgCl C. LiCl D. KCl
27	Formula mass of K ₂ SO ₄ is.	A. 174 amu B. 110 amu C. 180 amu D. 145 amu
28	Molecular mass of Acetic Acid	A. 70 amu B. 43 amu C. 60 amu D. 80 amu
29	Number of hydrogen atoms present in 18 g of water.	A. $2 \times N_A$ B. N_A C. $3 \times N_A$ D. $\frac{1}{2} N_A$
30	The mass number of sodium is.	A. 19 B. 31 C. 27 D. 23
31	Lime is another name of	A. Sodium hydroxide B. Sodium Carbonate C. Calcium Carbonate D. Silicon dioxide
32	Mass of 3 moles of oxygen atoms is.	A. 64 g B. 16 g C. 32 g D. 48 g
33	Number of moles in 29.25 g NaCl is.	A. 0.50 B. 0.25 C. 0.21

		C. 0.12 D. 0.75
34	How many atom of carbon are present in one molecule of glucose.	A. 11 B. 22 C. 12 D. 6
35	40 g of H_3PO_4 contains numebr of moles	A. 0.58 g B. 4.8 g C. 5.8 g D. 0.408 g
36	A compound with chemicla formula Na_2CX_3 has formula mass 106 ami. Atomicmass of the elemetrn X is.	A. 16 B. 23 C. 12 D. 106
37	How many moels of molecules are there in 16 g oxygen.	A. 0.05 B. 0.1 C. 0.5 D. 1
38	What is the mass of 4 moles of hydrogen gas.	A. 1 g B. 1.008 g C. 8.064 g D. 4.032 g
39	Whcih term is the same for one mole of oxgen and one mole of water.	A. atoms B. mass C. Molecules D. Volume
40	If one mole of carbon contains x atoms what is the number of atoms contained in 12 g of Mg.	A. 1.5 x B. 0.5 x C. x D. 2x
41	The mass of one molecule of water is.	A. 18 g B. 18 mg C. 18 kg D. 18 amu
42	The molecular mass of H_2SO_4 is.	A. 9.8 g B. 98 amu C. 98 g D. 9.8 amu
43	How many number of moles are equivalent to 8 gram of CO_2 .	A. 0.18 B. 0.15 C. 0.24 D. 0.21