Chemistry 9th Class English Medium Unit 3 Online Test

Sr	Questions	Answers Choice
1	When molten copper and olten zinc are mixed together , they give rise to a new substnce called brass. Predict what type of bond is formed between copper and zinc.	A. Ionic bond B. Coordinate Covalent bond C. Metallic bond D. Covalent Bond
2	Which element is capable of forming all the three types of bonds, covalent coordinate covalent or ionic.	A. Carbon B. Silicon C. Magnesium D. Oxygen
3	Why is H_2Oa liquid while H_2S is a gas?	A. Becuase in waer, the atomic size of oxygen in smaller than that of Sulphur B. Because water can easily freeze into ice C. Because water is a polar compound ad tere exists strong forces of attrction betwene its molecules D. Because H₂O molecule is lighter than H₂S
4	Which of the following bond s is expected to b the weakest.	A. Cl-Cl⁻ B. C-C C. F-F D. O -O
5	Which form of carbon is used as a lubricant?	A. Coal B. Diamond C. Charcoal D. Graphite
6	Keeping in view the intermolecular forces of attraction, indicate which compound has the highest boiling point	A. H₂S B. HF C. NH₃ D. H₂O
7	Which metal has the lowest melting point?	A. Li B. Na C. Rb D. K
8	Which ionic compound has the highest melting point.	A. RbCl B. KCl C. LiCl D. NaCl
9	Whcih compound contains both covalent and ionic bonds.	A. MgCl₂ B. PCl5 C. NH₄Cl D. CaO
10	Which among of the followign has a double covalent bond.	A. Ethane B. Methane C. Acetylene D. Ethylene
11	Atoms achieve stability by attainign electonic configuration of.	A. Alkali metals B. Coinage metals C. Inert Gases D. Alkaline earth metals
12	Attaining two electrons in the valence shell is called.	A. Octet rule B. Duplet rule C. Triplet rule D. All of these
13	All the noble gases have their valence electrons.	A. Incomplete B. Partialy filled C. Completely filled D. None of the above
14	Noble gases are non -reactive, because they do not.	A. Gain electrons B. Lose electrons

		C. Share electrons D. All of these
15	Every atom has a natural tendency to accommodate electrons in its valence shell	A. 2 or 6 B. 2 or 4 C. 2 or 8 D. 2 or 10
16	Hydrogen and Helium follow.	A. Octet rule B. Triple rule C. Duplet rule D. None of these
17	Whcih of the following atoms obey duplet rule.	A. O₂ B. Cl₂ C. H₂ D. Li₂
18	Whcih of the following is not true about the formationof Na_2S	A. Each sodium atom loses one electron B. Sodium forms cation C. Each sulphur atom gains one electron D. Sulphur form anion
19	Octet rule is	A. Attainign of eight electrons in its valence shell B. Discription of eight electrons C. Pictur eof electronic configuration D. Pattern of electronic configuration
20	Atoms react with each other because.	A. they are attracted towards each other B. They are short of electrons C. They want to disperse D. They want attain stability
21	An atom having six electonsin its valence shell will achieve noble gas electronic confguration by	A. Gainign one electron B. Gaining two electrons C. Losing all electrons D. Losing two electrons
22	The formaton ionic bond between two ions is due to.	A. Hydrogen bonding B. Metalic force C. Electrostatic forces D. All of the above
23	Which group of the periodic table has the tendency to gain electrons.	A. Group -1 B. Group -17 C. Group-2 D. Group -18
24	Whcih of the following atoms willnot form cation or anion.	A. Atomic no. 16 B. Atomic no. 18 C. Atomic no. 17 D. Atomic No. 19
25	Transfer of electron between elements result in.	A. Coordinate covalent bonding B. Ionic bonding C. Metallic bonding D. Covallent bonding
26	When an electronegative element combines with electrpositive element, the type of bonding. is.	A. Covalent B. Polar Covalent C. Ionic D. Coorinae Covalent
27	How many electron are there in the valence shell of sodium atom.	A. One B. Two C. Three D. Four
28	The electropositive elements have the tendency to	A. Lose electrons B. Gain electrons C. Share electrons D. All of these
29	How many valance shell electrons are there in Na+ ion.	A. 8 B. 9 C. 1 D. 10
30	During the formation of ionic bond heat is.	A. Remains same B. Absorbed C. Released D. Both a and b
		A. Covalent bonds B. Electrostatic forces of attraction

31	Which types of attractive forces are presentin ionic compounds.	B. Electrostatic forces of attraction C. Mtallic bonds D. Coordinate covalent bonds
32	Number of electronsin nitrogen molecule is.	A. 2 B. 4 C. 6 D. 8
33	How many covalent bonds do N2 molecule have	A. 3 B. 4 C. 2 D. 5
34	Silicon belongs to Grou IVA . It has ....electons in the valence shell	A. 2 B. 6 C. 3 D. 4
35	In the formation of AlF_3, aluminum atom loses.....electrons.	A. 1 B. 4 C. 3 D. 2
36	Identify the covalent compund	A. NaCl B. H₂O C. KF D. MgO
37	A bond formed between two non metals is expected to be	A. Ionic B. Coordinate covalent C. Metallic D. Covalent
38	A bond pair is covalent molecules usually has.	A. One electron B. Two electron C. Three electron D. Four electron
39	Covalent Bond involves the	A. Sharing of electrons B. Repulsion of electrons C. Acceptance of electrons D. Donation of electrons
40	How many covalent bonds does C_2H_2 molecule have.	A. Two B. Three C. Four D. Five
41	Triple covalent bond involves how many electrons.	A. Six B. Four C. Eight D. Three
42	Identify the compound whcihis not soluble in water	A. KBr B. MgCl2 C. C6H6 D. NaCl
43	Whcih one of the following is the weakest force among the atoms.	A. Intermolecular force B. Ionic force C. Metallic force D. Covelent forces
44	Covalent bond is most commonly found between the elements of group	A. 1 to 13 B. 16 to 18 C. 13 to 17 D. 15 to 18
45	A bond formed by the mutual sharing an electron pair is called.	A. Ionic bond B. Metallic bond C. Covalent bond D. Coordinate covalent bond
46	A covalent bond formed by the mutual sharing of two pairs of electrons between bonded atoms is called.	A. Signle covalent bond B. Double covalent bond C. Triple covalent bond D. Polar covalent bond
47	Which molecule contains a single vovalent bond.	A. CH4 B. C2H4 C. C2H2 D. O2
48	Nitrogen molecule contains.	A. Polar covalent bond B. Triple Covalent bond C. Double covalent bond D. Single covalent bond

49	How many electrons are involved in the formation of signle covalent bond	A. One B. Two C. Three D. Four
50	A covalent bond formed by two similar stoms is known as.	A. Polar Covalent bond B. Metallic bond C. Double covalent bond D. Non-polar covalent bond
51	Dative covalent bondis also known as	A. Covalent bond B. Ionic Bond C. Metallic Bond D. Coordinate covalent bond
52	Whcih one of the following in as electron deficient molecule.	A. NH3 B. O2 C. BF3 D. N2
53	How many lone pairs are present on nitogenin ammonia molecule.	A. One B. Two C. Three D. Four
54	Whcih types of bond is present between NH3 and BF3	A. Covalend Bond B. Ionic Bond C. Co ordinate covalent bond D. Metallic Bond
55	In metals, the hold of nucieus over the valence shell elecrons is veak due to.	A. Highlectron affinity B. Large sized atoms C. High ionization energies D. All of the above
56	Malleability is the property by virtue of which a metal can be drawn into.	A. Rods B. Plates C. Sheets D. Wires
57	Metal have the tendency to lose elecrons due to.	A. High ionization energies B. Low ionization energies C. Low electron affinity D. None of the above
58	Hydrogen bound ing is always found in	A. Non-polar molecules B. Homoatomic molecules C. Polar Molecules D. All of the above
59	Which of the following is an exampel of polar covalent compound.	A. Cl2 B. H2 C. O2 D. HCl
60	The fore of attraction betwene water molecule is.	A. Ionic bonding B. Covalent bonding C. Hydrgen Bonding D. Co ordinate Covalent bonding
61	The boiling point of water is	A. O ^oC B. 100 ^oC C. 35 ^oC D. 25 ^oC
62	The boiling point of alcohal is	A. ₄₄^oC B. 78^oC C. 53^oC D. 19^oC
63	Water has high boiling point as compared a alcohol due to	A. Low density B. High surface tension C. Hydrogen bonding D. High vapour pressure
64	The compund formed by oppoite charges are known as.	A. Metallic solids B. Ionic compounds C. Non-polar Covalent compound D. None of the above
65	Ionic compund are good conductors electricity in	A. Solution B. Molten state C. Solid state D. both a and b
66	Ionic compund have	A. Low melting and high boiling points B. Low melting and boiling point C. High melting and boiling points D. High meilting and low boiling points

67	None polar compounds are insoluble in	A. Alcohol B. Benzene C. Ether D. Water