

Chemistry 9th Class English Medium Unit 3 Online Test

Sr	Questions	Answers Choice
1	When molten copper and molten zinc are mixed together, they give rise to a new substance called brass. Predict what type of bond is formed between copper and zinc.	A. Ionic bond B. Coordinate Covalent bond C. Metallic bond D. Covalent Bond
2	Which element is capable of forming all the three types of bonds, covalent coordinate covalent or ionic.	A. Carbon B. Silicon C. Magnesium D. Oxygen
3	Why is H_2O liquid while H_2S is a gas?	A. Because in water, the atomic size of oxygen is smaller than that of Sulphur B. Because water can easily freeze into ice C. Because water is a polar compound and there exists strong forces of attraction between its molecules D. Because H_2O molecule is lighter than H_2S
4	Which of the following bonds is expected to be the weakest.	A. $Cl-Cl$ B. $C-C$ C. $F-F$ D. $O-O$
5	Which form of carbon is used as a lubricant?	A. Coal B. Diamond C. Charcoal D. Graphite
6	Keeping in view the intermolecular forces of attraction, indicate which compound has the highest boiling point	A. H_2S B. HF C. NH_3 D. H_2O
7	Which metal has the lowest melting point?	A. Li B. Na C. Rb D. K
8	Which ionic compound has the highest melting point.	A. $RbCl$ B. KCl C. $LiCl$ D. $NaCl$
9	Which compound contains both covalent and ionic bonds.	A. $MgCl_2$ B. PCl_5 C. NH_4Cl D. CaO
10	Which among the following has a double covalent bond.	A. Ethane B. Methane C. Acetylene D. Ethylene
11	Atoms achieve stability by attaining electronic configuration of.	A. Alkali metals B. Coinage metals C. Inert Gases D. Alkaline earth metals
12	Attaining two electrons in the valence shell is called.	A. Octet rule B. Duplet rule C. Triplet rule D. All of these
13	All the noble gases have their valence electrons.	A. Incomplete B. Partially filled C. Completely filled D. None of the above
14	Noble gases are non-reactive, because they do not.	A. Gain electrons B. Lose electrons

		C. Share electrons D. All of these
15	Every atom has a natural tendency to accommodate electrons in its valence shell	A. 2 or 6 B. 2 or 4 C. 2 or 8 D. 2 or 10
16	Hydrogen and Helium follow.	A. Octet rule B. Triple rule C. Duplet rule D. None of these
17	Which of the following atoms obey duplet rule.	A. $O_{2/2}$ B. $Cl_{2/2}$ C. $H_{2/2}$ D. $Li_{2/2}$
18	Which of the following is not true about the formation of Na_2S	A. Each sodium atom loses one electron B. Sodium forms cation C. Each sulphur atom gains one electron D. Sulphur forms anion
19	Octet rule is	A. Attainment of eight electrons in its valence shell B. Description of eight electrons C. Picture of electronic configuration D. Pattern of electronic configuration
20	Atoms react with each other because.	A. they are attracted towards each other B. They are short of electrons C. They want to disperse D. They want to attain stability
21	An atom having six electrons in its valence shell will achieve noble gas electronic configuration by	A. Gaining one electron B. Gaining two electrons C. Losing all electrons D. Losing two electrons
22	The formation of ionic bond between two ions is due to.	A. Hydrogen bonding B. Metallic force C. Electrostatic forces D. All of the above
23	Which group of the periodic table has the tendency to gain electrons.	A. Group -1 B. Group -17 C. Group -2 D. Group -18
24	Which of the following atoms will not form cation or anion.	A. Atomic no. 16 B. Atomic no. 18 C. Atomic no. 17 D. Atomic No. 19
25	Transfer of electron between elements results in.	A. Coordinate covalent bonding B. Ionic bonding C. Metallic bonding D. Covalent bonding
26	When an electronegative element combines with an electropositive element, the type of bonding is.	A. Covalent B. Polar Covalent C. Ionic D. Coordinate Covalent
27	How many electrons are there in the valence shell of sodium atom.	A. One B. Two C. Three D. Four
28	The electropositive elements have the tendency to	A. Lose electrons B. Gain electrons C. Share electrons D. All of these
29	How many valence shell electrons are there in Na^+ ion.	A. 8 B. 9 C. 1 D. 10
30	During the formation of ionic bond heat is.	A. Remains same B. Absorbed C. Released D. Both a and b
		A. Covalent bonds B. Electrostatic forces of attraction

31	Which types of attractive forces are present in ionic compounds.	B. Electrostatic forces of attraction C. Metallic bonds D. Coordinate covalent bonds
32	Number of electrons in nitrogen molecule is.	A. 2 B. 4 C. 6 D. 8
33	How many covalent bonds does N ₂ molecule have	A. 3 B. 4 C. 2 D. 5
34	Silicon belongs to Group IVA. It haselectrons in the valence shell	A. 2 B. 6 C. 3 D. 4
35	In the formation of AlF ₃ , aluminum atom loses.....electrons.	A. 1 B. 4 C. 3 D. 2
36	Identify the covalent compound	A. NaCl B. H ₂ O C. KF D. MgO
37	A bond formed between two non metals is expected to be	A. Ionic B. Coordinate covalent C. Metallic D. Covalent
38	A bond pair in covalent molecules usually has.	A. One electron B. Two electron C. Three electron D. Four electron
39	Covalent Bond involves the	A. Sharing of electrons B. Repulsion of electrons C. Acceptance of electrons D. Donation of electrons
40	How many covalent bonds does C ₂ H ₂ molecule have.	A. Two B. Three C. Four D. Five
41	Triple covalent bond involves how many electrons.	A. Six B. Four C. Eight D. Three
42	Identify the compound which is not soluble in water	A. KBr B. MgCl ₂ C. C ₆ H ₆ D. NaCl
43	Which one of the following is the weakest force among the atoms.	A. Intermolecular force B. Ionic force C. Metallic force D. Covalent forces
44	Covalent bond is most commonly found between the elements of group	A. 1 to 13 B. 16 to 18 C. 13 to 17 D. 15 to 18
45	A bond formed by the mutual sharing of an electron pair is called.	A. Ionic bond B. Metallic bond C. Covalent bond D. Coordinate covalent bond
46	A covalent bond formed by the mutual sharing of two pairs of electrons between bonded atoms is called.	A. Single covalent bond B. Double covalent bond C. Triple covalent bond D. Polar covalent bond
47	Which molecule contains a single covalent bond.	A. CH ₄ B. C ₂ H ₄ C. C ₂ H ₂ D. O ₂
48	Nitrogen molecule contains.	A. Polar covalent bond B. Triple Covalent bond C. Double covalent bond D. Single covalent bond

49	How many electrons are involved in the formation of single covalent bond	A. One B. Two C. Three D. Four
50	A covalent bond formed by two similar atoms is known as.	A. Polar Covalent bond B. Metallic bond C. Double covalent bond D. Non-polar covalent bond
51	Dative covalent bond is also known as	A. Covalent bond B. Ionic Bond C. Metallic Bond D. Coordinate covalent bond
52	Which one of the following is an electron deficient molecule.	A. NH ₃ B. O ₂ C. BF ₃ D. N ₂
53	How many lone pairs are present on nitrogen in ammonia molecule.	A. One B. Two C. Three D. Four
54	Which type of bond is present between NH ₃ and BF ₃	A. Covalent Bond B. Ionic Bond C. Coordinate covalent bond D. Metallic Bond
55	In metals, the hold of nucleus over the valence shell electrons is weak due to.	A. High electron affinity B. Large sized atoms C. High ionization energies D. All of the above
56	Malleability is the property by virtue of which a metal can be drawn into.	A. Rods B. Plates C. Sheets D. Wires
57	Metals have the tendency to lose electrons due to.	A. High ionization energies B. Low ionization energies C. Low electron affinity D. None of the above
58	Hydrogen bonding is always found in	A. Non-polar molecules B. Homoatomic molecules C. Polar Molecules D. All of the above
59	Which of the following is an example of polar covalent compound.	A. Cl ₂ B. H ₂ C. O ₂ D. HCl
60	The force of attraction between water molecules is.	A. Ionic bonding B. Covalent bonding C. Hydrogen Bonding D. Coordinate Covalent bonding
61	The boiling point of water is	A. 0 °C B. 100 °C C. 35 °C D. 25 °C
62	The boiling point of alcohol is	A. 44 °C B. 78 °C C. 53 °C D. 19 °C
63	Water has high boiling point as compared to alcohol due to	A. Low density B. High surface tension C. Hydrogen bonding D. High vapour pressure
64	The compounds formed by opposite charges are known as.	A. Metallic solids B. Ionic compounds C. Non-polar Covalent compound D. None of the above
65	Ionic compounds are good conductors of electricity in	A. Solution B. Molten state C. Solid state D. both a and b
66	Ionic compounds have	A. Low melting and high boiling points B. Low melting and boiling point C. High melting and boiling points D. High melting and low boiling points

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None polar compounds are insoluble in

- A. Alcohol
- B. Benzene
- C. Ether
- D. Water