

## Chemistry 10th Class English Medium Unit 9 Online Test

Sr	Questions	Answers Choice
1	The lives of aquatic plants and animals are indirectly related to concentration of dissolved in water:	A. hydrogen B. oxygen C. chlorine D. nitrogen
2	Such reaction which can be made proceed in either direction depending upon the condition:	A. simple reaction B. reversible reaction C. irreversible reaction D. chain reaction
3	What is the colour of hydrogen iodide in product?	A. Purple B. Yellow C. Blue D. Colorless
4	At equilibrium state, when reaction cases to proceed, it is called:	A. equilibrium constant B. dynamic equilibrium C. static equilibrium D. simple equilibrium
5	At dynamic equilibrium state:	<ul> <li>A. Rate of forward reaction≠ Rate of reverse reaction</li> <li>B. Rate of forward reaction &gt; Rate of reverse reaction</li> <li>C. Rate of forward reaction = Rate of reverse reaction</li> <li>D. Rate of forward reaction &amp;It rate of reverse reaction</li> </ul>
6	Law of Mass Action was put forward by:	A. G.N.Lewis B. Lowry C. Arrhenius D. Gulbderg and Waage
7	Active mass is expressed as:	A. {} B. [] C. () D. II
8	An equilibrium is achievable only in a:	A. big system B. small system C. open system D. closed system
9	When the number of moles of both sides are equal in a reaction then the unit of Kc will be:	A. no unit B. mol-2 dm6 C. mol dm3 D. mo-2 dn
10	The value of equilibrium constant (Kc) depends only on:	A. temperature B. Pressure C. concentration D. density
11	In irreversible reaction, dynamic equilibrium:	<ul> <li>A. Never establishes</li> <li>B. Establishes before the completion of reaction</li> <li>C. Establishes after the completion of reaction</li> <li>D. Establishes readily</li> </ul>
12	The reaction in which the products do not recombine to form reactants are called.	A. Inversible reactions B. Reversible reaction C. Decomposition D. Addition
13	The reaction in which the products can recombine to form reactants are called.	A. Reversible reaction B. Irreversible reaction C. Decomposition D. Addition
14	The color of lodine is:	A. Purple B. Black C. Red

15	The color of hydrogen lodide is:	A. Colourless B. Black C. Red D. Pink
16	When the rate of a the forward reaction takes place at the rate of reverse reaction the composition of the reaction mixture remains constant it is called.	A. Chemical equilibrium B. Dynamic equilibrium C. Static equilibrium D. All
17	When the reaction causes to produced it is called.	<ul> <li>A. Chemical equilibrium state</li> <li>B. Static equilibrium</li> <li>C. Dynamic equilibriun</li> <li>D. All</li> </ul>
18	Guldberg and waage put forward law of mass action in:	A. 1860 B. 1869 C. 1870 D. 1879
19	The % age of nitrogen and oxygen in our atmosphere is:	A. 80 B. 90 C. 95 D. 99
20	Which gas is used to prepare ammonia?	A. N <sub>2</sub> B. O <sub>2</sub> C. Cl <sub>2</sub> D. S
21	Which gas is used to manufacture king of chemicals sulphuric acid?	A. N <sub>2</sub> B. O <sub>2</sub> C. Cl <sub>2</sub> D. S
22	Equilibrium constant has no unit when number of moles of reactants and products are:	A. same B. Different C. Both a and b D. None of these
23	For reactions having large Kc value, the reaction proceeds to:	A. Completion B. Equilibrium state C. back ward D. None of these
24	The characteristics of reversible reactions are the following except:	<ul> <li>A. Products never recombine to form reactants.</li> <li>B. They never complete</li> <li>C. The proced in both ways</li> <li>D. They have a double arrow between reactants and products.</li> </ul>
25	When a system is at equilibrium states?	<ul> <li>A. The concentration of reactants and products becomes equal</li> <li>B. The opposing reactions</li> <li>C. The rate of the reverse reaction becomes very low</li> <li>D. The rates of the forward and reverse reactions becomes equal.</li> </ul>
26	Which one of the following statements is not correct about active mass?	<ul> <li>A. Rate of reaction is directly proportional to activ emass.</li> <li>B. Active mass is taken in molar concentration</li> <li>C. Active mass is represented by square brackets</li> <li>D. Active mass means total mass of substances.</li> </ul>
27	When the magnitude of Kc is very small it indicates.	<ul><li>A. Equilibrium will never establish</li><li>B. All reactants will be converted to products.</li><li>C. Reaction will go to completion</li><li>D. The amount of products is negligible</li></ul>
28	Reactions which have comparable amounts of reactants and products at equilibrium state have:	A. very small Kc value B. Very large Kc value C. Moderate Kc value D. None of these
29	At dynamic equilibrium:	A. The reactions stops to proceed B. The amounts of reactants and products are equal C. The speed of the forward is reverse reactions are equal D. The reaction can no longer be

D. Pink

		reversed
30	In an irreversible reaction dynamic equilibrium:	<ul> <li>A. Never establishes</li> <li>B. Established before the completion of reaction</li> <li>C. Establishes after the completion of reaction</li> <li>D. Establishes readily</li> </ul>
31	For a reaction between $PCL_3$ and $Cl_2$ to form $PCl_5$ the units of Kc are:	A. Mol dm <sup>-3</sup> B. Mol <sup>-1</sup> dm <sup>- 3</sup> C. Mol <sup>- 1</sup> dm <sup>3</sup> D. Mol dm <sup>3</sup>
32	The major components of Atmosphere are:	A. Carbon and Nitrogen B. Nitrogen and Oxygen C. Oxygen and Chlorine D. None of these
33	Which type of reactions do not go to completion?	<ul><li>A. Irreversible reaction</li><li>B. Reversible reactions</li><li>C. Addition reactions</li><li>D. Decomposition reactions</li></ul>
34	Which types of reaction speed up gradually?	<ul><li>A. Irreversible reactions</li><li>B. Reversibel reactions</li><li>C. Forward reactions</li><li>D. Decomposition reactions.</li></ul>
35	Which type of reactions take place in both directions?	<ul><li>A. Decomposition reactions</li><li>B. Irreversibel reactions</li><li>C. Reversible reactions</li><li>D. Addition reactions</li></ul>
36	In a chemical reaction, the substance that combine are called.	A. Reactant B. Products C. Mass D. Material
37	When a reaction ceases to proceed further , it is called.	<ul><li>A. Chemical states</li><li>B. Static state</li><li>C. Physical state</li><li>D. Dynamic equilibrium state</li></ul>
38	Dynamic means, reaction is:	<ul><li>A. In forward direction</li><li>B. Stop</li><li>C. In reverse direction</li><li>D. Still continuing.</li></ul>
39	The forward reaction takes place:	A. Right to left B. Left to Right C. Only to right D. Only to left
40	The unit of molar concentration:	A. mol. dm <sup>-2</sup> B. mol. dm <sup>-1</sup> C. mol. dm D. mol. dm <sup>-3</sup>
41	Equilibrium constant value "K" is equal to:	A. Kt / Kr B. Ki / Kr C. Kc / Qc
42	Which chemical is called king of chemicals?	A. KNO <sub>3</sub> B. H <sub>2</sub> SO <sub>4</sub> C. HCI D. NHO <sub>3</sub>
43	The plant use	A. Carbon di oxide B. Oxygen C. Nitrogen D. Sulphur
44	We exhale gas in the atmosphere during respiration.	A. Carbon dioxide B. Oxygen C. Nitrogen D. Water
45	The rate of reverse reaction in the beginning.	A. Slow B. moderate C. Very fast D. Low
10		A. Never establishes B. Establishes after completion of reaction

46	I he dynamic equilibrium in irreversible reaction.	C. Establishes before completion of reaction D. Establishes very soon
47	The reactions in which products can recombine to form reactants is called.	<ul> <li>A. Irreversible reactions</li> <li>B. Reversible reactions</li> <li>C. Direct reactions</li> <li>D. Indirect reactions</li> </ul>
48	A reverse reaction is that.	<ul><li>A. Which proceed from left to high</li><li>B. In which reactants reacts to form products</li><li>C. Which slow down gradually</li><li>D. Which speed up gradually</li></ul>
49	when a system is in equilibrium state or in dynamic equilibrium state.	<ul><li>A. Reactants and product are equal</li><li>B. Forward reactions stops</li><li>C. Reverse reactions stops</li><li>D. Forward reaction rate and reverse reaction rate</li></ul>
50	At equilibrium state there are possibilities: OR types of chemical equilibrium are.	A. 1 B. 2 C. 3 D. 4
51	A complete reaction is one is which.	<ul> <li>A. All the reactants convert into products.</li> <li>B. All the reactants do not convert into products.</li> <li>C. Half reactants convert into products.</li> <li>D. Only 10% reactants convert into products.</li> </ul>
52	A dynamic equilibrium	<ul> <li>A. Reaction stops to proceed</li> <li>B. Amounts of reactants and products are equal</li> <li>C. Rate of forward and reverse reaction are equil</li> <li>D. Reaction can no longer be reversed.</li> </ul>
53	When $CaCO_3$ , is heated in an open flask, it decomposes to form calcium oxides and .	A. O <sub>2</sub> B. CO C. CO <sub>2</sub> D. CO <sub>3</sub>
54	For which reaction Kf is the rate constant?	A. Forward reaction B. Backward reaction C. Upward reaction D. Downward reaction
55	The substance formed during the chemical reaction are called.	A. Products B. Reactants C. Radical D. Element
56	The substances formed during the chemical reaction are called.	A. Product B. Reactants C. Radicals D. Element
57	Such reactions which continue in both directions are called.	A. Irreversible B. Reversible C. Nonreactive D. Dynamic
58	Which colour of HI is ?	A. Orange B. Purple C. Red D. Colourless
59	In chemical reaction , the substances that combine are called.	A. Reactants B. Products C. Equilbrium D. Numerator
60	In the beginning the rate of reverse reaction is	A. Negligible B. <div>Moderate</div> C. Very fast D. Slow
61	Who proposed "Law of mass action"?	A. Newton B. Boyle C. Guldberg and waage D. Lavoisier
		A. 1889

62	Guldberg and waage out forward law of mass action in:	B. 1879 C. 1869 D. 1859
63	K <sub>c</sub> is always equal to	A. Rf/Rr B. Kf /Kt C. Kf/Kr D. Rr/Rt
64	The unit of molar concentration is:	A. moldm <sup>-3</sup> B. moldm <sup>+3</sup> C. molcm <sup>-3</sup> D. molcm <sup>+3</sup>
65	Formation profess of Ammonia by the combination Hydrogen and Nitrogen was given by:	A. Dalton B. Thomson C. Haber D. Waage
66	The value of $K_{C}$ depends on.	A. Temperature B. Pressure C. Volume D. Atmosphere
67	The reaction in which the number of moles reactants and products not equal in balance chemical equation the units of Kc for this reactions are	A. mol <sup>-2</sup> B. dm <sup>3</sup> C. mol <sup>2</sup> D. mol dm <sup>-1</sup>
68	When the numbers of moles of both sides are equal in a reaction, then the unit of $K_{C}$ will be:	A. No unit B. mol <sup>-2</sup> dm <sup>6</sup> C. mol dm <sup>3</sup> D. mol <sup>-2 </sup> dm
69	Such reationsin which reactants and products are sufficient in quantities the Kc value of equilibrium state will be	A. Very small B. Very large C. Moderate D. None of these
70	When the magnitude of Kc is very large it indicates.	<ul> <li>A. Reaction mixture consist of almost all products.</li> <li>B. Reaction mixture consist of almost all reactants.</li> <li>C. Reaction has not gone to completion</li> <li>D. Reaction mixture has negligible products.</li> </ul>
71	When the magnitude of Kc is very small in indicates.	<ul><li>A. Equilibrium will never establish</li><li>B. All reactants will converted to products.</li><li>C. Reaction will go to completion</li><li>D. The amount of products is negligible</li></ul>
72	When the magnitude of Kc is very large in indicates.	<ul> <li>A. Reaction never go to completion</li> <li>B. Reaction is in equilibrium state</li> <li>C. Reaction will complete after some time</li> <li>D. Reaction has almost to completion</li> </ul>
73	In a Reversible Reaction if Qc = Kc then.	<ul> <li>A. Reaction is occuring in forward direction</li> <li>B. Reaction is occuring in Reverse direction</li> <li>C. Equilibrium has been allained</li> <li>D. Reaction is not at equilibrium</li> </ul>
74	If Qc = Kc the reaction goes in:	A. Forward B. Reverse C. At equilibrium state D. None
75	If Qc < Kc the reaction goes in:	A. Forward B. Reverse C. At equilibrium state D. None
76	The reaction goes from left to right , if:	A. Qc = Kc B. Qc > kc C. Qc <kc D. Qc = 0</kc 
77	If Qc > Kc the reaction will be in	<ul><li>A. Chemical equilibrium</li><li>B. Static equilibrium</li><li>C. Reverse reaction</li><li>D. Forward reaction</li></ul>

78	Reaction which have comparable amount of reactants and products at equilibrium state have.	A. Very small Kc value B. Very large Kc value C. Moderate Kc value D. None of these
79	The two major components of atmosphere are:	<ul> <li>A. Hydrogen and oxygen</li> <li>B. Nitrogen and Hydrogen</li> <li>C. Nitrogen and oxygen</li> <li>D. Oxygen and water.</li> </ul>
80	The reaction will attain the equilibrium if:	A. Qc <kc B. Qc &gt; Kc C. Qc = Kc D. Qc = 0</kc 