

ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	Which one of the following is weakest inter molecular force?	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatic forces b/w ions D. Dipole dipole forces
2	Ionic Solids are characterized by	A. Low melting points B. Good conductivity in solid state C. High vapour pressure D. Solubility in polar solvents
3	Chloroform and acetone are soluble in each other due to:	A. Instantaneous dipole interactions. B. Dipole-dipole interactions. C. Inter molecular hydrogen bonding. D. All of above
4	The boiling point of NH_3 is maximum among the hydrides of group V elements due to	A. Enhanced electronegative character of Nitrogen B. Pyramidal structure of NH_3 C. Very small size of Nitrogen D. Enhanced electropositive character of Nitrogen
5	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
6	The gases can be converted into liquids by:	A. Increasing the pressure only. B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical points D. Lowering temperature only
7	Heat of vapourization for liquids with strong dipole-dipole forces will have	A. Negligible Values B. Reasonably high values C. Very high values D. very low values
8	The gases can be converted into liquids by	A. increasing the pressure only B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical point D. Lowering temperature only
9	Hydrocarbon molecule with large chain lengths experience:	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
10	In a group on going downward, polarizability generally:	A. Decreases B. Increases C. Remain constant D. Negligible
11	London dispersion forces are also called:	A. Hydrogen bonding. B. Debye forces C. Van der Waal's forces D. Instantaneous dipole-induced dipole forces.
12	HF has exceptionally low acidic strengths due to	A. Smaller size of fluorine B. Strong polar bond between H and F C. Electronegativity of fluorine D. Strong hydrogen bonding
13	In a group on going downward, polarizability generally	A. Decreases B. Increases C. Remains constant D. Negligible
		A. 1000 mm Hg

14	At sea level and at 100°C the vapor pressure of water in an open system is:	B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg
15	Evaporation of water is possible at	A. Above 100°C B. 0°C C. 100°C D. At all temperature
16	Which of the following is a pseudo solid?	A. Atoms of He is gaseous state at high temperature. B. Molecules of water in liquid state. C. Molecules of solid I ₂ D. Molecular of hydrochloric acid gas
17	H ₂ S is a gas while H ₂ O is a liquid at room temperature. It is due to:	A. Less inter-molecular forces in water. B. Covalent bonding H-O in water molecule. C. Hydrogen bonding in water molecules D. Ionic character in water molecules
18	Which of the following liquids has low vapour pressure at 25°C:	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
19	The molecules of CO ₂ in dry ice form the :	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystals
20	The only forces are London dispersion forces among the	A. Atoms of He in gaseous state at high temperature B. Molecules of water in liquid state C. Molecules of solid I ₂ D. Molecular of hydrochloric acid gas