

ECAT Chemistry Chapter 4 Liquids & Solids Online Test

Sr	Questions	Answers Choice
1	Ionic Solids are characterized by	A. Low melting points B. Good conductivity in solid state C. High vapour pressure D. Solubility in polar solvents
2	Amorphous solids	A. Have sharp melting point B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms
3	The molecules of CO ₂ in dry ice form the	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystal
4	Which of the following is a pseudo solid	A. CaF ₂ B. Glass C. NaCl D. All
5	Diamond is a bad conductor because	A. It has a tight structure B. It has a high density C. there is no free electron present in the crystal of diamond of conduct electricity D. Is transparent to light
6	Chloroform and acetone are soluble in each other due to	A. Instantaneous dipole interactions B. Dipole-dipole interactions C. Intermolecular hydrogen bonding D. All of above
7	Force of attraction between atoms of He is	A. London dispersion forces B. Hydrogen bonding C. Coordinate covalent bond D. Covalent bond
8	The only forces are London dispersion forces among the	A. Atoms of He in gaseous state at high temperature B. Molecules of water in liquid state C. Molecules of solid I ₂ D. Molecular of hydrochloric acid gas
9	The density of water decreases, when it is freeze at 0°C because of	A. Change of bond lengths B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of ice
10	Which one of the following is the weakest intermolecular force	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatics forces between ions D. Dipole-dipole forces
11	The gases can be converted into liquids by	A. increasing the pressure only B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical point D. Lowering temperature only
12	The intermolecular forces in liquids are	A. Negligible B. Very weak C. Very strong D. Reasonably strong
13	London dispersion forces are also called	A. Hydrogen bonding B. Debye forces C. Van de Waal's forces D. Instantaneous dipole-induced dipole forces
		A. Decreases

14	In a group on going downward, polarizability generally	B. Increases C. Remains constant D. Negligible
15	Hydrocarbon molecules with large chain lengths experience	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
16	HF has exceptionally low acidic strengths due to	A. Smaller size of fluorine B. Strong polar bond between H and F C. Electronegativity of fluorine D. Strong hydrogen bonding
17	Polypeptide chains are coiled about one another into a spiral by	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds
18	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
19	Debye forces are present in one of the following pairs	A. Na ⁺ ion and water B. Argon and water C. Argon and Na ⁺ ion D. Ne and Water
20	Hydrogen bonding is present between the molecules of	A. NH ₃ B. H ₂ O C. HF D. All of above
21	H ₂ S is a gas which H ₂ O is liquid at room temperature. it is due to	A. Less intermolecular forces in water B. Covalent bond in H-O in water molecule C. Hydrogen bonding in water molecules D. Ionic characters in water molecules
22	Evaporation of water is possible at	A. Above 100 ⁰ B. 0 ⁰ C. 100 ⁰ D. At all temperature
23	Rate of evaporation and rate of condensation at equilibrium	A. Become very low B. Become very high C. Become equal D. Can never be equal
24	At sea level and at 100 ⁰ C the vapour pressure of water in an open system is	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg
25	A liquid on evaporation causes	A. Heating effect B. Cooling effect C. Suffication D. All of above
26	Escape of high energy molecules from the surface of a liquid is called	A. Sublimation B. Distillation C. Condensation D. Evaporation
27	Which of the following liquids has low vapour pressure at 25 ⁰ C	A. Diethyl ether B. Acetone C. Water D. Ethyl alcohol
28	The boiling point of NH ₃ is maximum among the hydrides of group V elements due to	A. Enhanced electronegative character of Nitrogen B. Pyramidal structure of NH ₃ C. Very small size of Nitrogen D. Enhanced electropositive character of Nitrogen
29	Heat of vapourization for liquids with strong dipole-dipole forces will have	A. Negligible Values B. Reasonably high values C. Very high values D. very low values
30	Trend of boiling point of halogens from fluorine to Iodine is that it	A. Decreases B. Is negligible C. Increases D. Remains constant

31	Ionic solids are characterized by :	<p>A. Low meeting points</p> <p>B. Good conductivity in solid state</p> <p>C. High vapor pressure</p> <p>D. Solubility in polar solvents</p>
32	Amorphous solids:	<p>A. Have a sharp melting point</p> <p>B. Undergo clean cleavage when cut with knife</p> <p>C. Have a perfect arrangement of atoms</p> <p>D. Can possesses small regions of orderly arrangement of atoms</p>
33	The molecules of CO ₂ in dry ice form the :	<p>A. Ionic crystals</p> <p>B. Coverlet crystals</p> <p>C. Molecular crystals</p> <p>D. Any type of crystals</p>
34	Which of the following is a pseudo solid?	<p>A. CaF_2</p> <p>B. Glass</p> <p>C. NaCl</p> <p>D. All</p>
35	Diamond is a bad conductor because:	<p>A. It has tight structure</p> <p>B. It has a high density</p> <p>C. There is no free electron present in the crystal of diamond of conduct electricity is transparent to light</p>
36	Chloroform and acetone are soluble in each other due to:	<p>A. Instantaneous dipole interactions.</p> <p>B. Dipole-dipole interactions.</p> <p>C. Inter molecular hydrogen bonding.</p> <p>D. All of above</p>
37	Force of attraction b/w atoms of He is :	<p>A. London dispersion forces</p> <p>B. Hydrogen bonding</p> <p>C. Coordinate covalent bond</p> <p>D. Covalent bond</p>
38]Which of the following is a pseudo solid?	<p>A. Atoms of He is gaseous stat at high temperate.</p> <p>B. Molecules of water in liquid state.</p> <p>C. Molecules of solid I₂</p> <p>D. Molecular of hydrochloric acid gas</p>
39	The density of water decreases, When it is freezed at 0°C	<p>A. Change of bond length</p> <p>B. Change of bond angles</p> <p>C. Cubic structure of ice</p> <p>D. Empty spaces present in the structure of ice</p>
40	Which one of the following is weakest inter molecular force?	<p>A. Dipole induced dipole forces</p> <p>B. Ionic dipole forces</p> <p>C. Electrostatic forces b/w ions</p> <p>D. Dipole dipole forces</p>
41	The gases can be converted into liquids by:	<p>A. Increasing the pressure only.</p> <p>B. Lowering temperature and increasing pressure</p> <p>C. Increasing pressure and bringing temperature below critical points</p> <p>D. Lowering temperature only</p>
42	The inter-molecular forces in liquids are:	<p>A. Negligible</p> <p>B. Very weak</p> <p>C. Very strong</p> <p>D. Reasonably strong</p>
43	London dispersion forces are alos called:	<p>A. Hydrogen bonding.</p> <p>B. Debye forces</p> <p>C. Van der Waal's forces</p> <p>D. Instantaneous dipole-induced dipole forces.</p>
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45	In a group on going downward, polarizability generally:	<p>A. Decreases</p> <p>B. Increases</p> <p>C. Remain constant</p>

		D. Negligible
46	Hydrocarbon molecule with large chain lengths experience:	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
47	HF has exceptionally low acidic strength due to:	A. Smaller size of fluorine. B. Stronger polar bond between H and F. C. Electronegativity of fluorine D. Strong hydrogen bonding
48	Polypeptide chains are coiled about one another into a spiral by:	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds
49	The strongest forces are:	A. Debye forces B. London dispersion C. Dipole-dipole attraction D. Hydrogen bonding
50	Debye forces are present in one of the following pairs:	A. Na ⁺ ion and water B. Argon and water C. Argon and Na ⁺ ion D. Ne and water
51	Hydrogen bonding is present in one of the following pairs:	A. NH ₃ B. H ₂ O C. HF D. All of above
52	H ₂ S is a gas while H ₂ O is a liquid at room temperature. It is due to:	A. Less inter-molecular forces in water. B. Covalent bonding H-O in water molecule. C. Hydrogen bonding in water molecules D. Ionic character in water molecules
53	Evaporation of water is possible at:	A. Above 100°C B. 0°C C. 100°C D. At all temperature
54	Rate of evaporation and rate of condensation at equilibrium:	A. Become very low B. Become very high C. Become equal D. Can never be equal.
55	At sea level and at 100°C the vapor pressure of water in an open system is:	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg
56	A liquid on evaporation causes:	A. Heating effect. B. Cooling effect. C. Suffocation . D. All of above
57	Escape of high energy molecules from the surface of liquid is called:	A. Sublimation. B. Distillation. C. Condensation. D. Evaporation.

58	Which of the following liquids has low vapor pressure at 25°C:	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
59	Which of the following liquids has low vapour pressure at 25°C:	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
60	The boiling point of NH_3 is maximum among the hydrides of group V elements due to:	A. Enhanced electronegative character of Nitrogen B. Pyramidal structure of NH_3 C. Very small size of Nitrogen. D. Enhanced electropositive character of Nitrogen