Competition for Matric Class

Sr	Questions	Answers Choice
1	The lives of aquatic plants and animals are indirectly related to concentration of dissolved ------ in water:	A. hydrogen B. oxygen C. chlorine D. nitrogen
2	Such reaction which can be made proceed in either direction depending upon the condition:	A. simple reaction B. reversible reaction C. irreversible reaction D. chain reaction
3	What is the colour of hydrogen iodide in product?	A. Purple B. Yellow C. Blue D. Colorless
4	At equilibrium state, when reaction cases to proceed, it is called:	A. equilibrium constant B. dynamic equilibrium C. static equilibrium D. simple equilibrium
5	At dynamic equilibrium state:	A. Rate of forward reaction≠ Rate of reverse reaction B. Rate of forward reaction > Rate of reverse reaction C. Rate of forward reaction = Rate of reverse reaction D. Rate of forward reaction < rate of reverse reaction
6	Law of Mass Action was put forward by:	A. G.N.Lewis B. Lowry C. Arrhenius D. Gulbderg and Waage
7	Active mass is expressed as:	A. { } B. [ ] C. ( ) D. | |
8	An equilibrium is achievable only in a:	A. big system B. small system C. open system D. closed system
9	When the number of moles of both sides are equal in a reaction then the unit of Kc will be:	A. no unit B. mol-2 dm6 C. mol dm3 D. mo-2 dn
10	The value of equilibrium constant (Kc) depends only on:	A. temperature B. Pressure C. concentration D. density
11	Binary compounds of oxygen such as CO2 and SO2 were names as acids by;	A. Jabir Bin Hayan B. Lavoisier C. Al- Jahiz D. Sir Humphrey Davy
12	Latin word 'Acidus' means:	A. sweet B. salty C. sour D. bitter
13	All water soluble metallic oxides have characteristic of:	A. acidic B. basic C. neutral D. amphoteric
14	Which one is not a characteristic of an acid:	A. it turns red litmus blue B. it reacts with base to form salt and water C. Its aqueous solution conduct electric current D. it has sour taste

15	The product of any Lewis acid-base reaction is a single specie called:	A. free radical B. adduct C. molecular ion D. conjugate
16	Which of one the following is Lewis base?	A. BF3 B. H+ C. NH3 D. Ag+
17	Which acid is found in sour milk?	A. `Formic acid B. Lacitc acid C. Citric acid D. Butric acid
18	Water is a;	A. strong electrolyte B. non-electrolyte C. weak electrolyte D. natural compound
19	Sum of PH and POH of a solution is:	A. 7 B. 14 C. 16 D. 18
20	Soda time is a mixture of:	A. CaO and NaOH B. CaCl2 and NaOH C. CaO and ca(OH)2 D. CaCl2 and KOH
21	Acid used for cleaning metals , tanning and in printing industries, is:	A. H2SO4 B. HNO3 C. CH3COOH D. HCl
22	Who put forward " the vital force theory"?	A. Kolbe B. Wohler C. Berzellius D. Jabir - Bin - Hayan
23	Chemical formula for urea is:	A. NH4CNO B. NH4CH C. NH2CONH2 D. NH4Cl
24	Open chain compounds are also called:	A. aliphatic compounds B. alicyclic compounds C. aromatic compounds D. hydrocarbons
25	Aromatic compounds are given this name because of;	A. smell they have B. slippery touch they have C. bitter taste they have D. sour taste they have
26	Organic compounds contain:	A. ionic bond B. Covalent bond C. Metallic bond D. Co-ordinate covalent
27	What is the percentage of carbon in anthracite?	A. 60% B. 70% C. 80% D. 90%
28	The black residue of the coal tar is called:	A. peat B. lignite C. pitch D. matte
29	Alkanes are also called:	A. olefins B. Paraffins C. aliphatic compound D. cyclic compounds
30	Petroleum is refined by:	A. destructive distillation B. fractional distillation C. simple distillation D. dry distillation
31	Which one of the following is a synthetic fiber:	A. cotton B. wool C. nylon D. silk
32	Which one of the following is gas at room temperature:	A. C2H6 B. C6H14 C. C6H6 D. C8H18

		D. C_6H_6
33	Volatility of hydrocarbons deceases with:	A. decrease of molecular mass B. increase of molecular mass C. increase of double or triple bonds D. decrease of triple bonds
34	Marsh gas is mostly :	A. ethane B. methane C. butane D. propane
35	Hydrogenation of alkenes and alkynes may be carried out at room temperature in the presence of catalyst:	A. Pt or Pd B. Ni C. Zn D. Fe_2O_3
36	Which one of these is used as anaesthesia?	A. carbon black B. Mehtane C. Ethane D. Chloroform
37	Olefins is a Latin word meaning:	A. inert B. less reactive C. oil forming D. most reactive
38	Ethen is prepared by heating a mixture of ethanol and excess of concentrated sulphuric acid at:	A. 150°C B. 180°C C. 250°C - 300°C D. 400°C - 450°C
39	Hydrogenation of vegetable oil to convert it into banaspatic ghee is carried out in the presence of catalyst:	A. Ni-metal B. Zn-metal C. Na - Metal D. Fe_2O_3
40	Which of one of the following is used as as fumigant:	A. Ethyl alcohol B. ethylene oxide C. ethylene glycol D. diethyl ether
41	Which one of the following is polmerized to from benzene?	A. ethane B. ethene C. methane D. ethyne
42	When bromine water is added to acetylene its red-brown colour is changed into:	A. red B. green C. violet D. discharged
43	Which one of these protects us against the disease;	A. Carbohydrates B. Lipids C. Proteins D. all of these
44	Which are responsible for transmitting genetic information form generation to generation:	A. vitamins B. lipids C. proteins D. Nucleic acids
45	The general formula of carbohydrate is:	A. $C_{2n}(H_2O)$ B. $CN(H_2O)n$ C. $Cn(H_2O)_{2n}$ D. $C_{2n}(H_2O)n$
46	Monosaccharides consist of:	A. 2 to 6 carbon atoms B. 3 to 9 carbon atoms C. 4 to 10 carbon atoms D. 10 to 15 carbon atoms
47	Glucose and fructose are;	A. trioses B. tetroses C. pentoses D. hexoses
48	Which one is a pentahydroxy ketone?	A. sucrose B. glucose C. fructose D. lactose
49	Which one of the following is non- reducing in nature;	A. glucose B. fructose C. starch D. sucrose
50	Lactose and maltose are;	A. monosaccharides B. disaccharides C. trisaccharides

		C. trisaccharides D. tetrasaccharides
51	Hypoglycemia result due to:	A. high cholesterol level B. low cholesterol level C. low sugar level in body D. high sugar level in body
52	Dextrose is a crystallized:	A. sucrose B. lactose C. glucose D. fructose
53	Enzymes are chemically;	A. lipids B. carbohydrates C. proteins D. vitamins
54	Methyl butanoate smells like;	A. a pine apple B. an apple C. an orange D. a lemon
55	Atmosphere is divided into;	A. 3 regions B. 4 regions C. 5 regions D. 6 regions
56	Percentage of nitrogen in atmosphere is:	A. 20.94% B. 78.09% C. 0.93% D. 0.03%
57	How much amount of sunlight reaches up to earth and absorbed:	A. 26% B. 32% C. 50% D. 18%
58	In the troposphere (lowest layer), the temperature decreases form 17°C to;	A. -58°C B. -93°C C. -61°C D. -49°C
59	Mesophere extend forms 50 km to:	A. 50 km B. 12 km C. 60 km D. 85 km
60	The major constituents of troposphere are:	A. carbon dioxide and argon B. nitrogen and oxygen C. argon and oxygen D. carbon dioxide and oxygen
61	Almost all aircrafts fly in:	A. troposphere B. mesosphere C. stratosphere D. themosphere
62	ozone depletion (i.e O3----------> O2 + O) occurs in:	A. lower stratosphere B. middle stratosphere C. upper stratosphere D. troposphere
63	A strange bitter smell noticed near photocopiers is of:	A. SO2 B. H2S C. O2 D. O3
64	Which gas like a glass wall of g green house:	A. CO B. O2 C. CO2 D. N2
65	Normal rain water is:	A. weakly acidic B. strongly acidic C. weakly basic D. strongly basic
66	Ozone id beneficial for us as it:	A. absorbs infrared radiations B. absorbs ultraviolet radiation C. absorbs chlorofluorocarbons D. absorbs air pollutants
67	The percentage of water in human body is about:	A. 60% B. 50% C. 70% D. 80%
		A. -4°C B. 0°C

68	Freezing point of water is:	B. 0°C C. -36°C D. -58°C
69	Density of water at 4°C is:	A. 0.976 gem-3 B. 1 gem-3 C. 0.956 gem-3 D. 1.1 gem-3
70	The surface tension of water is:	A. very low B. moderate C. high D. zero
71	Which process is responsible for ascending of water in roots:	A. evaporation B. transpiration C. condensation D. capillary action
72	Water molecule had a structure;	A. non-polar B. polar C. ionic D. tetrahedral
73	The water which produces good lather with soap is called:	A. soft water B. hard water C. heavy water D. typical water
74	Water become hard because of:	A. Ca+2 B. Mg+2 C. SO-24 D. all of these
75	The chemical used in Clark's method is:	A. CaO B. slaked lime(Ca(OH)s) C. washing soda D. sodium zeolite
76	Mercury poisoning cause:	A. neurological damage B. high blood pressure C. kidney damage D. gastro
77	Fertilizers are used to make up the deficiency of;	A. oxygen and carbon B. iron and magnesium C. nitrogen and phosphorus D. hydrogen and calcium
78	Hookworm is a parasite that infects the:	A. large intestine B. small intestine C. stomach D. liver
79	Which one of the following is manufactured by Solvar's process:	A. caustic soda B. washing soda C. urea D. sulphuric acid
80	Chalco-pyite is an ore of:	A. iron B. aluminium C. silver D. copper
81	Chemical formula of copper glance is:	A. Cu2O B. Cu2S C. CuFeS2 D. CuSO4
82	Blonde hair contains compounds of:	A. molybdenum B. iron C. titanium D. copper
83	The process of removal of gangue form the ore is technically known as:	A. concentration B. purification process C. metallurgy D. refining
84	In forth flotation process the ore particles are preferentially wetted by:	A. water B. oil C. benzene D. petrole
85	When copper pyrite (CuFeS2) is strongly heated in excess of airit converted into mixture of:	A. Cu2S + Cu2O B. Cu2O +FeS C. Cu2S + FeS D. FeO + SO2 + Cu2S

86	In smelting process a small amount of coke is required because:	A. it is highly endothermic process B. it is highly exothermic process C. it is very fast process D. it is very slow process
87	Chemical formula of matte will:	A. Cu2S.FeS B. Cu2S C. Cu2O D. FeS
88	Bister copper is pure about;	A. 70% B. 80% C. 90% D. 98%
89	Manufacturing of urea involves:	A. 2 steps B. 3 steps C. 4 steps D. 5 steps
90	Catalyst used in Haber's process is:	A. Ni B. Fe C. Cu D. Zn
91	In irreversible reaction, dynamic equilibrium:	A. Never establishes B. Establishes before the completion of reaction C. Establishes after the completion of reaction D. Establishes readily
92	Which gas is evolved when acids react with carbonates and bicarbonates?	A. Carbon monoxide B. Carbon dioxide C. Hydrogen D. Hydrogen chloride
93	General formula of alcohols is:	A. RCHO B. ROH C. O= R - C -OH D. O = R - C- R
94	Formula of acetaldehyde is:	
95	The reaction in which the products do not recombine to form reactants are called.	A. Inversible reactions B. Reversible reaction C. Decomposition D. Addition
96	The reaction in which the products can recombine to form reactants are called.	A. Reversible reaction B. Irreversible reaction C. Decomposition D. Addition
97	The color of Iodine is:	A. Purple B. Black C. Red D. Pink
98	The color of hydrogen Iodide is:	A. Colourless B. Black C. Red D. Pink
99	When the rate of a the forward reaction takes place at the rate of reverse reaction the composition of the reaction mixture remains constant it is called.	A. Chemical equilibrium B. Dynamic equilibrium C. Static equilibrium D. All
100	When the reaction causes to produced it is called.	A. Chemical equlibrium state B. Static equilibrium C. Dynamic equilibriiun D. All
101	Guldberg and waage put forward law of mass action in:	A. 1860 B. 1869 C. 1870 D. 1879
102	The % age of nitrogen and oxygen in our atmosphere is:	A. 80 B. 90 C. 95 D. 99
103	Which gas is used to prepare ammonia?	A. N₂ B. O₂ C. Cl₂ D. S
		A. N₂ B. O₂

104	Which gas is used to manufacture king of chemicals sulphuric acid?	B. O₂ C. Cl₂ D. S
105	Equilibrium constant has no unit when number of moles of reactants and products are:	A. same B. Different C. Both a and b D. None of these
106	For reactions having large Kc value, the reaction proceeds to:	A. Completion B. Equilibrium state C. back ward D. None of these
107	The characteristics of reversible reactions are the following except:	A. Products never recombine to form reactants. B. They never complete C. The proced in both ways D. They have a double arrow between reactants and products.
108	When a system is at equilibrium states?	A. The concentration of reactants and products becomes equal B. The opposing reactions C. The rate of the reverse reaction becomes very low D. The rates of the forward and reverse reactions becomes equal.
109	Which one of the following statements is not correct about active mass?	A. Rate of reaction is directly proportional to activ emass. B. Active mass is taken in molar concentration C. Active mass is represented by square brackets D. Active mass means total mass of substances.
110	When the magnitude of Kc is very small it indicates.	A. Equilibrium will never establish B. All reactants will be converted to products. C. Reaction will go to completion D. The amount of products is negligible
111	Reactions which have comparable amounts of reactants and products at equilibrium state have:	A. very small Kc value B. Very large Kc value C. Moderate Kc value D. None of these
112	At dynamic equilibrium:	A. The reactions stops to proceed B. The amounts of reactants and products are equal C. The speed of the forward is reverse reactions are equal D. The reaction can no longer be reversed
113	In an irreversible reaction dynamic equilibrium:	A. Never establishes B. Established before the completion of reaction C. Establishes after the completion of reaction D. Establishes readily
114	For a reaction between PCL_3 and Cl_2 to form PCl_5 the units of Kc are:	A. Mol dm⁻³ B. Mol⁻¹ dm⁻³ C. Mol⁻¹dm³ D. Mol dm³
115	The major components of Atmosphere are:	A. Carbon and Nitrogen B. Nitrogen and Oxygen C. Oxygen and Chlorine D. None of these
116	Which type of reactions do not go to completion?	A. Irreversible reaction B. Reversible reactions C. Addition reactions D. Decomposition reactions
117	Which types of reaction speed up gradually?	A. Irreversible reactions B. Reversibel reactions C. Forward reactions D. Decomposition reactions.
118	Which type of reactions take place in both directions?	A. Decomposition reactions B. Irreversibel reactions C. Reversible reactions D. Addition reactions
119	In a chemical reaction, the substance that combine are called.	A. Reactant B. Products C. Mass D. Material
120	When a reaction ceases to proceed further , it is called.	A. Chemical states B. Static state C. Physical state D. Dynamic equilibrium state

		D. Dynamic equilibrium state
121	Dynamic means, reaction is:	A. In forward direction B. Stop C. In reverse direction D. Still continuing.
122	The forward reaction takes place:	A. Right to left B. Left to Right C. Only to right D. Only to left
123	The unit of molar concentration:	A. mol. dm⁻² B. mol. dm⁻¹ C. mol. dm D. mol. dm⁻³
124	Equilibrium constant value "K" is equal to:	A. Kt / Kr B. Ki / Kr C. Kc / Qc D. Qc /Kc
125	Which chemical is called king of chemicals?	A. KNO₃ B. H₂SO₄ C. HCl D. NHO₃
126	A base is a substance which neutralizes and acid. Which of these substances is not a base?	A. aqueous ammonia B. Sodium chloride C. Sodium carbonate D. Calcium oxide
127	Lewis acid-base concept have the following characteristics except:	A. formation of an adduct B. Formation of a co-ordinate covalent bound C. Donation and acceptance of an electron pair D. Donation and acceptance of a proton
128	Acetic acid is a weak acid because it.	A. Is used in cooking and flavouring food. B. Has very low pH C. Is not fully ionized D. Does not contain any hydrogen ions.
129	A salt is not composed of	A. A metallic cation B. Non -metallic anion C. an anion of a base D. An anion of an acid
130	If a liquid has a pH of 7 then it must.	A. Be a colourless and odourless liquid B. Freeze at 0 ^oC and boil at 100 ^oC C. Be natural D. Be a solution containing water
131	A salt always	A. Contain ions B. Contains water of crystallization C. Dissolves in water D. Forms crystals which conduct electricity
132	Dilute acids react with carbonates to produce the given products except.	A. Salt B. Water C. Hydrogen D. Carbon dioxide
133	In the preparation of insoluble salts, which one of the facts is incorrect?	A. Two soluble salts are mixed B. Two in soluble salts are mixed . C. One of the salt produced is insoluble D. Both of the salts produced are insoluble
134	The reaction between an acid and a base produces.	A. Salts and water B. Salt and gas C. Salt and an acid D. Salt and an base
135	What is the POH of a 0.02 M $Ca(OH)_2$?	A. 1.698 B. 1.397 C. 12.31 D. 12.61
136	Which one of the following species is not amphoteric?	A. H₂O B. NH₃ C. HCO₃ D. SO₄
137	The product of Lewis acid-base reaction is called adduct. The bond between the adduct species is.	A. Ionic B. Covalent C. Metallic D. Co-ordinate covalent
		A. Melting points of crystals B. Boiling points of crystals

138	The water of crystallization is responsible for the.	B. Boiling points of crystals C. Shapes of crystals D. Transition point of crystals
139	You want to dry a gas which one of the following salt you will use?	A. CaCl₂ B. NaCl C. CaO D. Na₂SIO₃
140	Which one of the following is Lewis base?	A. AlCl₃ B. H⁺ C. NH₃ D. BF₃
141	According to the Lewis concept acid is a substance which can.	A. Donate a proton B. Donate a pair of electron C. Accept a proton D. Accept a pair of electron.
142	Jabir Bin Haiyan prepared.	A. Nitric Acid B. Hydrochloric acid C. Sulphueric Acid D. All of these
143	Lavoisier named binary compounds of oxygen acids in .	A. 1787 B. 1790 C. 1815 D. 1828
144	Who proved that the presence of hydrogen as the main constituent of all acids.	A. Lavoisier B. Humphrey Davy C. Dalton D. Arrhenius
145	The word acid is derived from the.	A. Greek word B. Latin word C. English word D. Arabic word.
146	Acid means	A. Sour B. Bitter C. Sweet D. Salty
147	Which acid is present in our stomach.	A. Nitric acid B. Hydrochloric acid C. Sulphuric acid D. All of these
148	All acids turn blue litmus.	A. Red B. Blue C. Pinck D. White
149	All bases turn red litmus	A. Red B. Blue C. Pink D. White
150	Arrhenius presented his concept about acids and bases in.	A. 1785 B. 1787 C. 1923 D. 1930
151	According to Arrhenius concept acid is a substance which dissociates in aqueous solution to give.	A. Hydrogen ions B. Hydroxide ions C. Both a and b D. None of these
152	|Which one is not an Arrhenius acid?	A. HCl B. H2SO₄ C. CO₂ D. HNO₃
153	Which one is not an Arrhenius base?	A. NaOH B. KOH C. Ca(OH)₂ D. NH₃
154	Bronsted and Lowry presented their theories of acids and bases in	A. 1785 B. 1787 C. 1923 D. 1925
155	According to Bronsted and Lowry concept an acid is a substance that can donate.	A. Proton B. Electron pair C. Neutron D. Electron

156	A conjugate acid is a specie formed by accepting a.	A. Proton B. Electron pair C. Neutron D. Electron
157	According to Bronsted and Lowry concept a base is a substance that can accept.	A. Proton B. Electron pair C. Neutron D. Electron
158	A conjugate base is a specie formed by donating a.	A. Proton B. Electron pair C. Neutron D. Electron
159	A substance which can behave as an acid as well as a base is called.	A. Acid B. Base C. Amphoteric D. Neutral
160	According to Lewis concept a base is a substance which can donate.	A. Proton B. Electron pair C. Neutron D. Electron
161	A product of any Lewis acid base reaction is a single specie called.	A. Salt B. Water C. Adduct D. None of these
162	Which one is Lewis Acid?	A. BF₃ B. AlCl₃ C. FeCl₃ D. All of these
163	When acids react with metals which gas is evolved?	A. H₂ B. O₂ C. Cl₂ D. N₂
164	When acid react with carbonates and bicarbonates which gas is evolved?	A. H2 B. CO₂ C. Cl₂ D. N₂
165	When acid reacts with sulphites and Bi sulphates which gas is evolved?	A. H2 B. CO₂ C. SO₂ D. NH₃
166	Which one of the mineral acid.	A. HCl B. H₂SO₄ C. HNO₃ D. All of these
167	Which acid is used an electrolyte in lead storage battery?	A. H₂SO₄ B. HNO₃ C. HCl D. CH₃COOH
168	Which acid is used for etching designs on copper plates?	A. H₂SO₄ B. HNO₃ C. HCl D. CH₃COOH
169	Which acid used for food preservation?	A. H₂SO₄ B. HNO₃ C. HCl D. CH₃COOH
170	Citric acid is present in	A. Citrus fruits B. Sour milk C. Rancid butter D. Apple
171	Formic acid is present in	A. String of bees B. Sour milk C. Apple D. Fats
172	Mallic acid is present in	A. Apple B. Feats C. String of bees D. Urine
173	Uric Acid is present in	A. Apple B. Fats C. Urine

		D. Grapes
174	Stearic acid present in	A. Apple B. Fats C. Urine D. Grapes
175	Alkalis react with ammonium salt to liberate.	A. SO₂ B. CO₂ C. NH₃ D. H₂
176	Which is used to manufacture of soap?	A. NaOH B. Ca(OH)₂ C. KOH D. Mg(OH)₂
177	Which one is used for alkaline batteries?	A. NaOH B. Ca(OH)₂ C. KOH D. Mg(OH)₂
178	Which is used to remove the grease stains from clothes?	A. NaOH B. KOH C. Al(OH)₃ D. NH₄OH
179	The value of constant of ionic product of water Kw at 25 C.	A. 1.0 x 10⁻⁴ B. 1.0 x 10¹⁴ C. 1.0 x 10⁻⁴ D. 1.0 x 10⁴
180	pH value normally varies from.	A. 0-14 B. 1-14 C. 7-14 D. 10-14
181	pH of neutral solution is always.	A. 6 B. 5 C. 7 D. 10
182	Acidic solutions have pH value.	A. Less than 7 B. Greater than 7 C. equal to 7 D. None of these
183	Basic solution have pH value.	A. Less than 7 B. Greater than 7 C. equal to 7 D. None of these
184	Indicators are the	A. Inorganic compounds B. Organic compounds C. Ionic compounds D. Covalent compounds
185	Phenolphthalein produces red colour in	A. Acid B. Base C. Both a and b D. None of these
186	Methyl orange produces which colour in basic solution	A. Red B. Yellow C. Pink D. White
187	Which salt is used as a table salt?	A. NaCl B. Na₂CO₃ C. Na₂SiO₃ D. NaCl
188	Which salt is used for the manufacture of detergents, pulp and paper?	A. NaCl B. Na₂CO₃ C. Na₂SiO₃ D. NaCl
189	Which is used for cleaning agent for domestic and commercial purpose?	A. NaCl B. Na₂CO₃ C. NaHCO₃ D. Na₂SiO₃
190	The ability of carbon atoms to form chains is called.	A. Isomerism B. Catenation C. Resonance D. Condensation
191	Coal having 90% carbon contents is called	A. Peat B. Lignite

	Coal having 90% carbon contents is called.	C. anthracite D. bituminous
192	Main component of natural gas is	A. Methane B. Propane C. Butane D. Propene
193	The strone heating of coal in retorts in the absence of air is called.	A. Fractional distillation B. sublimation C. Roasting D. Destructive distillation
194	Pitch is black residue of	A. Coke B. Coal-tar C. Coal D. Coal gas
195	Natural gas is 85% methane, It is used to make the follwing except.	A. Carbon black B. Ethane C. Propane D. Both b and c
196	Which one of the following does not contain starch	A. Sugar cane B. Maize C. Barley D. Potatoes
197	Petroleum is refined by	A. Destructive distillation B. Fractional distillation C. Simple distillations D. Dry distillation
198	In laboratory urea was prepared by	A. Wohler B. Rutherford C. Berzelius D. Dalton
199	Main component of natural gas is	A. Methane B. propane C. Butane D. Propene
200	The strong heating of coal in retoris in the absence of air is called.	A. Fractional distillation B. Sublimation C. Roasting D. Destructive distillation
201	General formula of alkyl radical is	A. C_nH_{2n+2} B. C_nH_{2n-}2 C. C_nH_{2n+1} D. C_nH_{2n}
202	Identify which one of the following compounds is a ketron?	A. (CH₃)₂ CHOH B. (CH₃)₂CHCl C. (CH₃)₂CO D. (CH₃)₂CHCl
203	Which of the following statement s is not true about fossil fuels?	A. they all contain carbon B. They are renewable C. they produce pollutants when burnt D. they cause acid rain
204	The functional group- COOH is found in	A. Carboxylic acid B. aldehydes C. alcohals D. easter
205	Which of the following is the hardest coal?	A. Pear B. Lignite C. Bituminous D. Anthracite
206	In which of the following groups, oxygen is attached on both sides with carbon atoms?	A. Ketone B. Ether C. Aldehyde D. Ester
207	Carbonization process is the conversion of	A. Coal into coal gas B. Coal into wood C. Wood into coal D. Wood into coal tar
208	Coal gas is mixture of	A. CO and CH₄ B. CO . CH₄. CO₂ C. CO, CH₄, H₂ D. CO, H₂ and CO₂

209	Which one of the following is a synthetic fibre?	A. Cotton B. Wool C. nylon D. silk
210	Which one of the following is not a fossil fuel?	A. Coal B. Natural gas C. Bio gas D. Petroleum
211	Which one of the following does ot contain protein	A. Pulses B. Potatoes C. Beans D. eggs
212	Conversion of dead plants into coal by the action of bacteria and heat is called.	A. Carbonization B. Catenation C. Hydrogenation D. Cracking
213	Which one of the following compounds is an aldehyde?	A. CH₃-CH₂-OH B. CH₃-COOH C. CH₃CHO D. CH₃COCH₃
214	Who put forward the vital force theory?	A. Berzelius B. Wohler C. Dalton D. Lavoisier
215	The 1st organic compound prepared in laboratory.	A. Urea B. NaCl C. Thiourea D. Pyridine
216	Who was prepared acetic acid in laboratory?	A. Berzelius B. Wohler C. Kolbe D. Dalton
217	The branch of chemistry which deals with the study of hydrocarbons and their derivatives is known as.	A. Organic chemistry B. Inorganic chemistry C. Biochemistry D. Nuclear chemistry
218	The formula which represents the actual number of atoms in one molecule of organic compound is called.	A. Molecular formula B. Structural formula C. Condensed formula D. Dot and cross formula
219	Which one is homocyclic compound?	A. Benzene B. Cyclobutane C. Cyclohexane D. All
220	Which one is heterocyclic compound?	A. Benzene B. Cyclobutane C. Thiophene D. Nephthalene
221	Silicon occurs in the form of	A. both a and b B. Silicates C. None of these D. Silica
222	Which one contains double covalent bound?	A. Pentane B. Ethylene C. Acetylene D. All
223	Which one contains triple covalent bond?	A. Pentane B. Ethylene C. Acetylene D. All
224	Coal is blackish complex mixture of compounds of	A. Carbon B. Hydrogen C. oxygen D. All
225	Conversion of wood into coal is called.	A. Carbonization B. Destructive distillation C. Fractional distillation D. All
226	The % age of carbon is coal is	A. 40-60 B. 50-70 C. 40-80

		D. 40-90
227	The % age of carbon in peat is :	A. 60% B. 70% C. 80% D. 90%
228	The % age of carbon is lignite is	A. 60 B. 70 C. 80 D. 90
229	The % age of carbon is bituminous is	A. 60 B. 70 C. 80 D. 90
230	The strong heating of coal in the absence of air is called.	A. Carbonization B. Destructive distillation C. Fractional distillation D. All
231	Coal is the mixture of	A. Hydrogen B. Methane C. Carbon monoxide D. All
232	Coal tar contains compounds	A. Benzene B. Phenol C. Toluene D. All
233	The % age of carbon in coke is	A. 60 B. 70 C. 90 D. 98
234	Natural gas contains.	A. Methane B. Ethane C. Propane D. All
235	The general formula of alkane is	A. C_nH_{2n+2} B. C_nH_{2n-2} C. C_nH_{2n} D. C_nH_{2n+1}
236	Alkyl radical is derivative of	A. Alkane B. Alkene C. Alkyne D. All
237	Which one of these hydrocarbon molecules would have no effect on an aqueous solution of bromine?	A. CH₄ B. C₁₀H₂₀ C. C₂H₄ D. C₂H₂
238	If an organic compound has 4 carbon atoms, all singly bonded, it will have the following characteristics except one.	A. It will be saturated hydrocarbon B. It will hav 8 hydrogen atoms C. Its name will be n-butane. D. It will be least reactive.
239	The reduction of alkyl halides takes place in the presence of	A. Zn/HCl B. Na/HCl C. Mg/HCl D. Cu/HCl
240	Halogenation of methane produces following valuable chemical compounds used as solvents except.	A. Carbonterachloride B. Chloroform C. Carbon Black D. Chloromethane
241	Incomplete combustion of alkanes produce.	A. Carbon dixoide only B. Carbon monoxide only C. Carbon monoxide carbon black and water D. Carbon dioxide and carbon black
242	Dehydrohalogenation takes place in the presence of .	A. NaOH aqueous B. Alcoholic KOH C. Aqeous KOH D. Alcoholic NaOH
243	Oxidation of ethene with $KMnO_4$produces.	A. oxalic acid B. glyoxal C. ethene glycol D. Propene glycol
		A. C₂H₄ B. C₃H₆

244	Which one of these is a saturated hydrocarbon?	B. C₃H₆ C. C₄H₈ D. C₅H₁₂
245	A hydrocarbon has molecular formula C_8H_{14}. What is the molecular formula of the next number of the same homologous series.	A. C₉H₁₈ B. C₉H₁₆ C. C₉H₂₀ D. C₉H₁₂
246	The molecular formula of the first three members of the alkane hydrocarbons are CH_4,C_2H_6 and C_3H_8. What is the molecular formula for the eight alkane member, octane, which is found is petrol?	A. C₈H₈ B. C₈H₁₆ C. C₈H₁₈ D. C₈H₂₀
247	One of the hydrocarbons reacts with one mole of hydrogen to form a saturated hydrocarbon. What formula could be of the X.	A. C₃H₈ B. C₆H₁₂ C. C₄H₁₀ D. C₇H₁₆
248	Dehydration of alcohols can be carried out with.	A. NaOH B. KOH C. H₂SO₄ D. HCl
249	The end product of oxidation of acetylene is .	A. Oxalic acid B. Glyoxal C. Glycol D. None of these
250	Dehalogenation of tetrahalides produces acetylene. This reaction takes place in the presence of	A. Sodium metal B. Zinc metal C. Magnesium metal D. Potassium metal
251	Substitution reaction is the characteristics of.	A. Alkanes B. Alkenes C. Alkynes D. None of these
252	Halogenation of alkanes in the presence of diffused sunlight takes place.	A. Suddenly, only in one step B. Slowly in one step C. In a series of step D. Fastly in two steps
253	In which one of the following is a substitution reaction?	A. Halogenations of alkynes B. Halogenations of alkenes C. Halogenations of alkanes D. Bromination of alkenes
254	The order of reactivity of hydrogen halides with alkenes is.	A. HI>HBr B. HBr>HI C. HCl>HBr D. HBr < HCl
255	Oxidation of alkenes produce.	A. Glyoxal B. Glycol C. Oxalic acid D. Formic acid
256	Which is simplest alkane?	A. CH₄ B. C₂H₈ C. C₂H₂ D. C₂H₄
257	Carbon black is used in the manufacture of.	A. Dry cleaning B. Shoe polishes. C. Fertilizer D. None of these
258	Alkanes give reaction only	A. Addition B. Decomposition C. Substitution D. Displacement
259	Chemical formula of chloroform is :	A. CH₂Cl₂ B. CH₃Cl C. CHCl₃ D. CCl₄
260	Alkanes are produced in large amounts by cracking of	A. Natural gas B. Petroleum C. Benzene D. Xylol
261	Traces of acetylene are present to coal gas about.	A. 0.06 % B. 0.08% C. 1.1 % D. 90%
		A. Methane

262	Which of the following gas is used in warfare?	A. Methane B. Ethane gas C. Mustard gas D. None of these
263	Condensed formula of ethane is	A. C₃Cl₈ B. C₂Cl₆ C. H₃CCH₃ D. None of these
264	The general formula of alkynes is	A. C_nH_{2n} B. C_nH_{2n+2} C. C_nH_{2n-2} D. C_n H_{2n+1}
265	Alkanes do not react in	A. Diffused sunlight B. Dark C. Bright sunlight D. None of these
266	Carbon Tetra chloride is used in .	A. Fertilizer B. Dry cleaning C. Metallurgy D. Anesthesia
267	The alkanes consisting of C_5 to C_{10}	A. gases B. liquids C. solid D. plasma
268	Chloroform is used for	A. Anesthesia B. Fever C. Link D. Toys
269	Molecular formula of butyne is .	A. C₄H₆ B. C₃H₄ C. C₄H₇ D. C₄H₈
270	Alkanes are least reactive compounds because they are.	A. Saturated hydrocarbons. B. Unsaturated hydrocarbons. C. Both a and b D. None of the above.
271	Which is present 85% in natural gas?	A. Ethane B. Propane C. Methane D. Butane
272	Hydrogenation of alkenes and alkynes takes place at room temperature in the presence of.	A. Ni B. Pt C. Pd D. Both a and b
273	Which one is the formula of chloromethane.	A. CH₂Cl₂ B. CCl₄ C. CHCl₃ D. CH₃Cl
274	Which gas creates suffocation and causes death?	A. CO B. CO₂ C. SO₃ D. SO₂
275	In shoe polished, which chemical is used.	A. Ethanol B. Methanol C. Carbon black D. Formaldehyde
276	In dry cleaning, which chemical is used.	A. Chloroform B. Carbon tetachloride C. Ecetaldehyde D. Ethanol
277	Some orchids attracts bees for pollination by producing.	A. alkanes B. almenes C. alkynes D. all the above.
278	Dehydration means removal of.	A. Water B. Halogen C. Hydrogen D. All above.
279	Benzene is formed by the polymerization of.	A. Alkene B. Alkane C. Acetylene D. Methane.

280	Ethylene is present is natural gas sometimes to the extent of .	A. 10% B. 20% C. 30% D. 40%
281	Carbohydrates are synthesized by plants through photosynthesis process which requires the following except.	A. CO₂ and water. B. Sunlight C. O₂ D. Chlorophyll
282	Which of the followings is a disaccharide?	A. Glucose B. Fructose C. Sucrose D. Starch
283	Photosynthesis process produces.	A. Starch B. Cellulose C. Sucrose D. Glucose
284	Which of the following is tasteless?	A. Starch B. glucose C. Fructose D. Sucrose
285	When glucose and fructose combine they produce	A. Starch B. Cellulose C. Sucrose D. None of these
286	Glucose is:	A. Hexabydroxy aldehyde B. Hexahydroxy ketone C. Polyhydroxy aldehyde D. Pentahydroxy ketone.
287	Thousand of amino acids polymerize to form.	A. Carbohydrates B. Proteins C. Lipids D. Vitamins
288	Enzymes are proteins which have the following properties except.	A. They catalyze reaction B. They are highly non-spcific C. They are highly efficient D. They are pro
289	Which one of the following vitamins is water soluble?	A. Vitamin A B. Vitamin C C. Vitamin D D. Vitamin E
290	Which one of the following is a fat soluble vitamin?	A. A B. E C. K D. All of these
291	Which one of the following is not the characteristics of monosaccharide?	A. White crystalline solids B. Soluble in water C. Hydrolysable D. Reducing is nature
292	Which one of the following statements about glucose and sucrose is incorrect.	A. Soluble in water B. Naturally occuring. C. Carbohydrates D. Disaccharides
293	Which of the following is reducing sugar.	A. Glucose B. Fructose C. Sucrose D. Starch
294	The most important oligosaccharide is:	A. Sucrose B. Glucose C. Fractose D. Maltose
295	Night blindness is because of deficiency of:	A. Vitamin A B. Protein C. Vitamin C D. Vitamin D
296	The organic compound used as drugs to control bleeding are.	A. Vitamins B. Proteins C. Lipids D. Glycerides
297	Deficiency of Vitamin E causes.	A. rickets B. Scurvy C. Anemia in babies

		D. Night blindness.
298	Lipids are macromolecules. They have characteristics except one of the following.	A. They are high energy foods. B. They are soluble in water. C. They are poor conductor of heat. D. They are esters of fatty acids.
299	Vitamin are accessory Growth factors they play important role in our body like.	A. Provide energy to the body. B. Insulate our body from electric shock C. Build brain cells D. Regulate metabolism
300	General formula of carbohydrate is.	A. C_n(H₂O)₆ B. CH C. C_nH_{2n} D. C₆H_{2n}O
301	Carbohydrates are synthesized by plants through.	A. Respiration B. Photosynthesis C. Dehydrateion D. Evaporation
302	Which one of the following cannot by hydrolyzed?	A. Polysacchrides B. Monosacchrides C. Oligosacchrides D. All of these
303	Glucose and fructose are.	A. Pentose B. Triose C. Hexoses D. None of these
304	Fructose contain group.	A. Ketone B. Aldehyde C. Alcoholic D. Alkyl
305	Monosaccharides are crystalline solids	A. Grey B. Crimson C. Silver D. White
306	Which is not a reducing sugar?	A. Glucose B. Fructorse C. Cellulose D. All of them
307	Tetra saccharides are classified under.	A. Monosaccharides. B. Oligosaccharides C. Polysaccharides D. All of them
308	The most important disaccharide is .	A. Sucrose B. Glucose C. Cellulose D. None of them
309	Which one of the following are amorphous solids?	A. Monosacchrides B. Oligosaccharides C. Polysaccharides D. All of them
310	On hydrolysis sucrose produces one unit of glucose and one unit of.	A. Fructose B. Starch C. Cellulose D. None of them
311	The source of galactose.	A. Fruits B. Vegetables C. Cereals D. All of them
312	Which is the essential sugar found in milk?	A. Maltose B. Lactose C. Galactose D. Starch
313	Maltose is commonly found in .	A. Cereal B. Milk C. Cotton D. Honey
314	Human body uses carbohydrates in the form of	A. Glucose B. Maltose C. Fructose D. Galatose
		A. Galactose

315	Which carbohydrate is used directly by muscles for energy?	B. Lactos C. Glucose D. Fructose
316	Low sugar level in human body result in.	A. Hyperglycemia B. Hypoglycemia C. Anemia D. All of them
317	Which help to keep the bowel functioning property	A. Dietary fiber B. Vitamins C. Lipids D. Carbohydrates
318	Which helps in lowering of cholesterol level?	A. Vitamin B. Fiber C. Carbohydratres D. All of them
319	The energy provided by carbohydrates in per gram.	A. 17 KJ B. 21 KJ C. 35 KJ D. 10 KJ
320	Which is the natural sugar found in starchy food.	A. Pentose B. Dextrose C. Hexose D. All of them
321	50% dextrose contains approximate amount of energy.	A. 250 calories B. 16 Calories C. 170 calories D. 120 calories
322	Amino acid are the building blocks of .	A. Proteins B. Carbohydrates C. Vitamins D. fats
323	The percentage of protein in dry weight of cell.	A. 20% B. 40% C. 50% D. 70%
324	Protein is not found in:	A. Muscles B. Skin C. Cotton D. Hair
325	Chemical formula of amino group is	A. NH₃ B. NH₂ C. NH₄ D. COOH
326	Out of twenty how many amino acids can by synthesized by human body?	A. Five B. Ten C. Seven D. Twelve
327	A bound formed between two amino acids is:	A. Peptide linkage B. Covalent bound C. Hydrogen bound D. Glycosidic linkage
328	The nature of enzyme is.	A. Protein B. Vitamin C. Fats D. Carbohydrate
329	When bones are heated they give.	A. Starch B. Gelatin C. Fats D. Oils
330	Lipids are macromolecules made up of.	A. Fetty acids B. Amino acids C. Nucleotides D. None of them
331	Oils and fats are esters of large chain fatty acids with.	A. Glycogen B. Glucose C. Starch D. Glycerol
332	Triglycerides are fatty acids.	A. Unsaturated B. Saturated C. Both of them D. None of them
		A. C₁₅H₃₁COOH

333	Chemical formula of Stearic acid.	A. C₁₅H₃₁COOH B. C₁₇H₃₅COOH C. C₁₅H₃₇COOH D. None of them
334	In hydrogenation of vegetable oil catalyst employed is.	A. Ni B. Pt C. ZnO D. Cr₂O₃
335	Alkenes are produced in large amounts by cracking of.	A. Natural gas B. Petroleum C. Benzene D. xylol
336	Traces of acetylene are present in coal gas about.	A. 0.06 % B. 0.08% C. 1.1% D. 90%
337	Which one the following is more reactive?	A. Methane B. Ethane C. Ethene D. Acetylene
338	Condensed formula of ethane is.	A. C₃H₈ B. C₂H_{2n+2} C. C₂H_{2n-2} D. C_nH_{2n+1}
339	Margarine is produced by adding hydrogen to vegetable oil at.	A. 2000 ^oC B. 100^oC C. 200^oC D. 1000^oC
340	Runcid butter has	A. Foul smell B. Rotten egg smell C. Pungent smell D. No smell
341	Smell of rancid butter is due to the presence of.	A. Propanoic acid B. Butanoic acid C. Acitic acid D. Citric acid
342	Methyl butanoate smell like.	A. Apple B. Mange C. Lemon D. Grapes
343	Ethyl butanoate smells like.	A. Apple B. Pine apple C. Lemon D. Melon
344	Nucleic acids made up of long chain of.	A. Nucleotide B. Fatty acids C. Amino acid D. None of them
345	DNA was discovered by	A. J .Watson B. Funk C. Robert brown D. Hopkins
346	RNAconsists of	A. Ribos B. Pentose C. Hexose D. Trioses
347	Synthesis of protein is directed by.	A. DNA B. RNA C. Both of them D. None of them
348	The accumulation of which vitamin causes bone like deposits in the kidney.	A. Vitamin D B. Vitamin E C. Vitamin B D. Vitamin A
349	Eggs oils and fats contain vitamin .	A. A B. B C. C D. D
350	Which is a hereditary material?	A. DNA B. RNA C. Protein D. All of them

351	Denaturing of protein is caused by.	A. Heating B. Changing pH C. Both of them D. None of these
352	White viscous fluid present in an egg is.	A. Proten B. Fats C. Vitamin D. Carbohydrates.
353	About 99% atmosphere's lies within:	A. 30 kilometre B. 35 kilometre C. 15kilometre D. 11 kilometre
354	Just above the Earth's surface is	A. Mesosphere B. Stratosphere C. Theremosphere D. Troposphere
355	A group of gases that maintains temperature of atmosphere is.	A. Carbon dioxide and water vapours B. Nitrogen and carbon dioxide C. Thermosphere D. Troposphere
356	The Earth's atmosphere is getting hotter because of.	A. Increasing concentration of CO B. Increasing concentration of CO₂ C. Increasing concentration of O₃ D. Increasing concentration of SO₂
357	Which one of the following is not a greenhouse Effect?	A. Increasing atmosphere temperature. B. Increasing food chains C. increasing flood risks D. Increasing sea-level
358	Normally rain water is weakly acidic because of.	A. SO₃ gas. B. CO₂gas. C. SO₂gas. D. NO₂gas.
359	Building are being damaged by acid rain because it attacks.	A. Calcium sulphate B. Calcium nitrate C. Calcium carbonate D. Calcium oxalate
360	Acid rain affects the aquatic life by clogging fish gills because of:	A. Lead metal B. Chromium metal C. Mercury metal D. Aluminium metal
361	Ozone is beneficial for us as it	A. Absorb infrared radiation. B. Absorb Ultraviolet radiations. C. Absorb Chlorofluorocarbons D. Absorb Air pollulant
362	Which one of the following is not an air pollutants?	A. Nitrogen B. Carbon dioxide C. Nitrogen oxide D. Ozone
363	Iron and steel structure are damaged by.	A. Carbon monoxide B. Sulphur dioxide C. Methane D. Carbon dioxide
364	Infrared radiations emitted by the Earth are absorbed by.	A. CO₂ and H₂O B. N₂ and O₂ C. CO₂and N₂ D. O₂and CO₂
365	Global warming causes rising of the sea level. The cause of global warming is.	A. CO₂ gas. B. SO₂gas. C. NO₂gases. D. O₃gas.
366	Which gas protects the Earth's surface from ultraviolet radiations?	A. CO₂ B. CO C. N₂ D. O₃
367	Effects of ozone depletion are following except the one.	A. Increases infectious diseases B. Increases crops production C. Can cause skin cancer D. Can cause climatic changes.
368	Which one of these pollutants are not found in car exhaust fumes?	A. CO₂ B. O₃

368	Which one of these pollutants are not found in car exhaust fumes?	C. NO₂ D. SO₂
369	The process by which atmospheric nitrogen is turned into nitrates in the soil is called.	A. Nitration B. Fixing C. Oxidation D. Reduction
370	Global warming is because of.	A. Absorption of infrared radiation emitted by earth surface. B. Absorption of infrared radiation coming from sun. C. Absorption of infrared Ultraviolet coming from the sun. D. Emission of ultraviolet radiation from the earth's surface.
371	Carbon monoxide is harmful to us because.	A. It paralyses the lungs. B. It damages lungs tissues. C. It reduces oxygen carrying ability of haemoglobin D. It makes the blood coagulate.
372	Earth has natural systems.	A. One B. Two C. Three D. Four
373	Atmosphere has regions.	A. One B. Two C. Three D. Four
374	The envelope of different gases around the earth is called.	A. Atmosphere B. Brosphere C. Lithosphere D. Hydrosphere
375	The percentage by volume of nitrogen in dry gas is.	A. 78.09% B. 20.94% C. 0.93% D. 0.03%
376	The percentage of sun light absorbed by atmospheric gases is.	A. 2% B. 10% C. 18% D. 25%
377	Atmospheric region found between 50-58 km from the earthis.	A. Thermosphere B. Stratosphere C. Mesophere D. Thermosphere
378	Mesosphere has a temperature range.	A. 17 ^oC - 58^oC B. 58^oC - 2^oC C. 2^oC - 93^oC D. - 93^oC
379	What gas is the major constituent of troposphere?	A. Nitrogen B. Oxygen C. Hydrogen D. Both a and b
380	Which gas is responsible in warming the atmosphere?	A. Nitrogen B. Hydrogen C. Helium D. Fluorine
381	At which region all weather occurs?	A. Troposphere B. Stratosphere C. Mesosphere D. Thermosphere
382	Almost all air crafts fly in which reion?	A. Troposphere B. Stratosphere C. Mesosphere D. Thermosphere
383	Major portion of ozone layer is foun in.	A. Troposphere B. Stratosphere C. Mesosphere D. Thermosphere
384	The region of ozone decomposition in stratosphere is.	A. 20 km B. 30 km C. 40 km D. 50 km
385	The recombination of O and O_2 in mid stratosphere is an	A. Exothermic reaction B. Endothermic reaction C. Heat absorbing process

		D. None of these
386	The percentage of SO_2 released by the combustion of coal and petroleum product.	A. 40% B. 60% C. 70% D. 80%
387	Which is not a character of SO_2	A. it is a colourless gas B. It has irritating smell C. It causes suffocation D. It do not form sulphuric acid
388	Which gas is produce by the electrical lightening of air.	A. NO B. SO₂ C. SO₃ D. CO₂
389	Which of the following can be used as a fuel?	A. Methanol B. Ethanol C. Bio-diesel D. All of them
390	The pH of water containing CO_2 is	A. 4-6 B. 5.6-6 C. 6-7 D. 7-8
391	The pH of acid rain.	A. 2 B. 3 C. 4 D. 5
392	High concentration of which metal clogs the fish gills.	A. Zinc B. Aluminium C. Sodium D. Copper
393	Ozone is an allotropic form of	A. Carbon B. Oxygen C. Sulphur D. Phosphorous
394	Ultraviolet radiations can causes.	A. Hepatitis B. Asthma C. Skin cancer D. Night blindness
395	Which gas is in involved in ozone depletion?	A. Nitrogen B. CFC's C. Chlorine D. All of them
396	The region in which ozone layere depletes is called.	A. Ozone hole B. Black hole C. Both of them D. None of them
397	Ozone depletion was first noticed is.	A. 1970's B. 1980's C. 1990's D. 1960's
398	Which is not an air pollutant?	A. CO₂ B. CO C. SO₂ D. NH₃
399	The gas use by plants to perform photosynthesis.	A. O₂ B. CO₂ C. N₂ D. CO
400	The gas used by animals to perform respiration.	A. O₂ B. N₂ C. SO₂ D. Cl₂
401	Which is not a poisonous gas?	A. Ozone B. Chlorine C. Carbon di oxide D. All of them
402	Which gas acts as a glass wall of a green house?	A. Oxygen B. Carbon dioxides C. sulphur dioxide D. Hydrogen
		A. Decrease in heat energy

403	By the increase in the concentration of CO_2 in air.	B. Increase in heat energy. C. Heat energy remains same D. None of them
404	The green house effect is proportional to the amount of which gas in air.	A. CO₂ B. O₂ C. N₂ D. All of them
405	Which is the major effect of global warming?	A. Increase in temperature. B. Rise in sea level C. Melting of glaciers D. All of them
406	Higher concentration of CO causes.	A. Fatique B. Headache C. Both of them D. None of them
407	Catalytic converters convert.	A. CO to CO₂ B. N₂ to NO C. CO₂ to CO D. N₂to NO₂
408	Which gas is also known as life gas for plants.	A. CO B. CO₂ C. O₂ D. NO₂
409	In the bacterial decay, the compound of which element are emitted.	A. Sulphur B. Carbon C. Nitrogen D. All of them
410	Ozone layer is not found is	A. Upper stratosphere B. Mid stratosphere C. Lower stratosphere D. All of them
411	A pollutant is a waste material that pollutes.	A. Air B. Water C. Soil D. All of them
412	Which factor determines the severity of a pollutant?	A. Chemical nature B. concentration C. Persistence D. All of them
413	Which pollutant is responsible for changing weather?	A. Air pollutant B. Water pollutant C. Soil pollutant D. All of them
414	Ozone has a smell	A. Bitter B. Rotten egg C. Sweat D. None of them
415	Which of the following is a poisonous gas.	A. Oxygen B. Ozone C. Nitrogen D. Carbon dioxide
416	The waste products driven out because of the combustion of fossil fuels.	A. Primary pollutant B. Secondary pollutant C. Tertiary pollutant. D. None of them
417	The smell of photocopies machine is due to the presence of.	A. Chlorine gas B. Neon gas C. Helium gas D. Ozone gas.
418	PAN stands for	A. Poly aniline nitrate B. Proxy acetyl nitrate C. Poly acetyl nitrite D. Proxy acetyl nitrite
419	Which of the following is a secondary pollutant.	A. CO₂ B. CO C. SO₃ D. HF
420	99% of atmosphere consists of .	A. N₂and H₂ B. N₂and O₂ C. N₂and CO₂ D. O₂and CO₂

421	Which gas is emitted due to volcanic eruption?	A. CO₂ B. SO₃ C. NO₂ D. H₂
422	Fossil fuel means.	A. Coal B. Petroleum C. Natural gas D. All of them
423	Forest fires and burning of wood emit.	A. CO₂ B. NO₂ C. SO₂ D. Cl₂
424	The range of temperature in burning solid waste burning in incinerators is.	A. 650 ^oC - 1000 ^oC B. 650^oC - 11000^oC C. 1000^oC - 2000^oC D. 650^oC - 1000^oC E. 5000^oC - 1000^oC
425	Incinerator reduces solid waste into .	A. Ash B. Flue gas C. Heat D. All of them
426	Which is not a part of flue gas?	A. Furans B. Dioxins C. HCl D. H₂SO₄
427	Thermosphere lies beyond.	A. Stratosphere B. Troposphere C. Mesosphere D. Biosphere
428	The solid particle deposit on the filter paper during filtratin is called.	A. Precipitates B. Residue C. Crystals D. All of them
429	Which one of the properties of water is responsible for rising of water plants?.	A. Specific heat capacity B. Surface tension C. Excellent solvent action D. Capillary action.
430	Specific heat capacity of water is.	A. 4.2 KJg⁻¹ K⁻¹ B. 4.2 Jg⁻¹K⁻¹ C. 2.4 KJg⁻¹K⁻¹ D. 2.4 Jg⁻¹K⁻¹
431	Water dissolves non-ionic compound by	A. Ion-ion forces B. Ion-dipole forces C. Dipole -dipole forces D. Hydrogen bonding.
432	Temporary hardness is because of .	A. Ca(HCO₃)₂ B. CaCO₃ C. MgCO₃ D. MgSO₄
433	Temporary hardness is removed by adding.	A. Quick lime B. Slaked lime C. Lime stone D. Lime water.
434	Permanent hardness is removed by adding.	A. Na₂ zeolite. B. Soda lime C. Lime water D. Quick lime
435	Which one of the following salts makes the water permanently hard?	A. NaCO₃ B. NaHCO₃ C. CA(HCO₃)₂ D. CaSO₄
436	Rapid growth of algae in water bodies is because of detergent having.	A. Carbonate salts. B. Sulphonic acid salts. C. Sulphate salts. D. Phosphate salts.
437	Which one of the following diseases causes liver inflammation?	A. Typhoid B. Jaundice C. Cholera D. Hapatitis
438	Which one of the diseases causes severe diarrhea and can be fatal?	A. Jaundice B. Dysentery C. Cholera

		D. Typhoid
439	Which one of the following diseases severe diarrhea and can be fatal?	A. Joundice B. Cholera C. Dysentery D. Typhoid
440	Which one of the following gases is used to destroy harmful bacteria in water.	A. Iodine B. Chlorine C. Fluorine D. Bromine
441	The percentage of water in human body is.	A. 40% B. 50% C. 60% D. 70%
442	The percentage of ocean in world water is.	A. 50% B. 67% C. 97% D. 25%
443	Inland water includes.	A. River B. Lakes C. Streams D. All of them
444	Sea water in unfit for drinking purpose die to the presence of.	A. Salts B. Algae C. Fishes D. All of them
445	The percentage of potable water on earth is.	A. 2% B. 0.2% C. 0.02% D. 0.002%
446	The freezing point of water is:	A. 10 ^oC B. 100^oC C. 0^oC D. 46^oC
447	The boiling point of water is.	A. 100^oC B. 4^oC C. 0^oC D. 25^oC
448	Water has a maximum density at .	A. 10^oC B. 0^oC C. 4^oC D. 100^oC
449	The heat capacity of water is.	A. 4.2 Jg⁻¹ K⁻¹ B. 2.1 Jg⁻¹K⁻¹ C. 3.2 Jg⁻¹K⁻¹ D. 5.9 Jg⁻¹K⁻¹
450	Water has a maximum density at 4 ^{o}C	A. 12 cm⁻³ B. 2 g cm⁻³ C. 1 g cm⁻³ D. 4 gcm⁻³
451	How many times the heat capacity of water is greater than that of rocks.	A. Two B. Three C. Four D. Six
452	The process by which water rises up from the roots of plants to leaves is called.	A. Photosynthesis B. Respiration C. Surface tension D. Capillary action
453	Which of the following salt is soluble in water.	A. NaCl B. KCl C. Na₂SO₄ D. All of them
454	Which of the following in insoluble in water.	A. Benzene B. NaCl C. KCl D. All of them
455	One H_2O molecule can form hydrogen bonding with how many other H_2O molecules?	A. One B. Two C. Three D. Four
		A. Tetrahedral B. Trigonal

456	Water molecule show structure.	B. Trigonal C. Pentagonal D. All of them
457	Some organic compound are soluble in water due to the presence of .	A. -OH B. H⁺ C. Botha a and b D. Covalent bond
458	Which of the following in soluble in water.	A. Organic acids B. Glucose C. Alcohal D. All of them
459	Water which produces good lather with soap is called.	A. Soft water B. Hard water C. Heavy water D. All of them
460	Chemical form of gypsum.	A. MgSO₄. 5H₂O B. CaSO₄. 2H₂O C. FeSO₄. 5H₂O D. CuSO₄. 5H₂O
461	Gypsum in water is.	A. Sparingly soluble B. Insoluble C. Highly soluble D. None of them
462	Temporary hardness is due to the presence of bicarbonates of.	A. Calcium B. Magnsium C. Both of them D. None of them
463	the removal of which ion causes water softening.	A. Na⁺ B. Mg² C. Li⁺ D. K⁺
464	Calcium carbonate is in water.	A. Insoluble B. Sparingly C. None of them D. Soluble
465	Temporary hardness in water can he removed by.	A. Boiling Method B. Using washing soda C. Using sodium zeolite D. All of them
466	Sodium zeolite is naturally occurring reason of .	A. NaAl (SiO₃)₂ B. Na₂CO₃ C. CaCO₃ D. Na₂SlO₃
467	Hard water can cause.	A. Stomach disorder B. Boiler blasts C. Inefficiency of engine D. All of them
468	soap is the sodium salt of long chain.	A. Amino acids B. Fatty acids C. Nucleotides D. None of them
469	Mg^{2+} and Ca^{2+} ions react with soap to form calcium and magnesium salts of fatty acids called.	A. Gelatin B. Scum C. Paste D. None of them
470	Industrial effluents are highly	A. Toxic organic compounds. B. Inorganic salts. C. Heavy metals D. All of them
471	Which is not a heavy metal.	A. Cadmium B. Lead C. Zinc D. Murcury
472	Acute lead poisoning causes dysfunction of.	A. Kidney B. Liver C. CNS D. All of them
473	Neurological damage is caused by the poisoning of.	A. Lead B. Cadimium C. Mercury D. All of them

474	The salts of which element are present in detergent that causes the rapid growth of algae in water bodies is.	A. Phosphate B. Calcium C. Sodium D. All of them
475	The depletion of which gas results in the death of aquatic life.	A. Oxygen B. Carbon di oxide C. Boath of them D. Nome of them
476	Example of pest is	A. Weeds B. Herbs C. Insect D. All of them
477	Which element protects teeth from decay?	A. Potassium B. Fluorine C. Sodium D. Calcium
478	Which disease is caused by polluted water.	A. Cholera B. Typhoid C. Diarrhea D. All of them
479	Which element do not causes toxicity in water?	A. Lead B. Arsenic C. Sodium D. Mercury
480	Vibrious cholera causes.	A. Choleera B. Dysentery C. Fluorsis D. Hepatitis
481	Which hepatitis is caused by contaminated water?	A. Hepatitis A B. Hepatitis B C. Hepatitis C D. Hepatitis D
482	Hook worm infects.	A. Liver B. Small intestine C. Large intestine D. Stomach
483	Hook warm larvae enter the body through.	A. Food B. Water C. Skin D. All of them
484	A disease is caused by excess of bile pigments in the blood is.	A. Typhoid B. Jaundice C. Cholera D. Dysentery
485	Which organ cause to function during Jaundice?	A. Liver B. Kindney C. Stomach D. Large intestine.
486	Swimming pools are cleaned by the process.	A. Chlorination B. Hydrogenations C. Saponification D. None of them
487	Chemical formula of hypochlorous acid is.	A. HCl B. HOCl C. H₂CO₃ D. HF
488	There are _____ types of salts.	A. 2 B. 3 C. 6 D. 8
489	Formula of Potassium ferrocyanide.	A. K₄[Fe(CN)₆]⁻⁴ B. K₃[Fe(CN)₄]⁻³ C. K₂[Fe(CN)₆]⁻⁴ D. K₄[Fe(CN)₅]⁻⁴
490	Formula of Methyl alcohol	A. CH₅-OH B. CH₃-CH₂-OH C. All of them D. None of these

491	Extraction of metals from its ores is called.	A. Metallurgy B. Mining C. Griding D. All
492	At the time of partition, How many industries wee present in Pakistan.	A. 30 B. 32 C. 34 D. 40
493	Which one of the ore of copper?	A. Copper glance B. Chalcopyrite C. Both a and b D. None of these
494	Brown hair contains.	A. Iron compound B. Copper compound C. titanium compound D. Both a and b
495	Blonde hair contains compound of.	A. Iron B. Copper C. titanium D. Molybdenum
496	Red hair contains compound of.	A. Iron B. copper C. titanium D. Molybdenum
497	Process of heating the concentrated ore to high temperature in excess of air is called.	A. Roasting B. Smelting C. Bessemerization D. All
498	Which one of the not metal.	A. Copper B. Carbon C. chromium D. Iron
499	The elements that do not conduct heat and electricity are called.	A. Metallurgy B. Non metal C. Metalloid D. Alloy
500	Metallurgy involves which of the following steps?	A. Mining and enrichment B. Reduction C. Refining and casting. D. All of these
501	Blast furnace usually used for the metallurgy of.	A. Iron B. Copper C. Aluminum D. Both a and b
502	The process of roasting during metallurgy of copper is carried out in a special furnace called.	A. Blast furnace B. Fire furnace C. Bessemer converter D. Reverberatory furnace
503	Froth flotation process is used to concentrate.	A. Copper ore B. Iron ore C. Chromium ore D. Aluminum ore
504	Compound of metal exist under earth crust are called.	A. Ore B. Gangue C. Minral D. None of these
505	Which contains sufficient amount of metal?	A. Mineral B. Ores C. Rocks D. Soil
506	A saturated solution of sodium chloride is called.	A. Brine B. Suspension C. Colloidal D. None of these
507	Raw materials used in Solvay's process.	A. Brine B. Lime stone C. Ammonia gas D. All
508	Formula of baking soda is.	A. Na₂CO₃ B. NaHCO₃ C. Na₂SO₄ D. Na₂PO₄

509	Formula of soda ash is.	A. Na₂CO₃ B. NaHCO₃ C. Na₂SO₄ D. Na₃PO₄
510	Imperial chemical industries was established in.	A. 1942 B. 1944 C. 1950 D. 1990
511	Sindh alkalies limited was established near Karachi in.	A. 1965 B. 1966 C. 1970 D. 2000
512	How many % age of nitrogen in urea fertilizers?	A. 40.6 B. 45.6 C. 46.6 D. 50
513	The raw materials for the manufacturing of urea are.	A. Ammonia B. Carbondioxide C. Limestone D. a & b
514	Ammonia is prepared by the process .	A. Ostwald B. Haber C. Clark D. All
515	How many % age of nitrogen present in air by volume?	A. 70% B. 75% C. 78% D. 80%
516	How many % age of urea is used as fertilizers?	A. 80% B. 90% C. 95% D. 98%
517	Formula of Urea is.	A. KCNO B. H₂N-CO-NH₂ C. HN-CO₂-NH D. H₃N-CO-NH₃
518	The number of carbon atoms present in petroleum gas.	A. 1-2 B. 1-3 C. 1-4 D. 1-5
519	The number of carbon atoms present in petroleum ether.	A. 1-5 B. 2-5 C. 3-7 D. 5-7
520	The number of carbon atoms present in gasoline or petrol.	A. 5-10 B. 6-10 C. 7-10 D. 8-10
521	The number of carbon atoms present in kerosene oil.	A. 8-12 B. 9-12 C. 10-12 D. 11-12
522	The number of carbon atoms present in diesel oil.	A. 10-15 B. 11-15 C. 12-15 D. 13-15
523	The number of carbon atoms present in fuel oil.	A. 14-18 B. 15-18 C. 16-18 D. 17-18
524	Concentration is a separating technique in which mineral is separated from.	A. Gangue B. Silicates C. Aluminates D. All
525	Sodium carbonate is manufactured by.	A. Haber's process B. Ostwald's process C. Solvay's process D. All
526	Ammonical brine is prepared by dissolving ammonia gas in.	A. NaCl B. CaCO₃ C. CaCl₂

		D. Na₂SO₄
527	The residual oil is heated above 400 c to produce.	A. Lubricants B. Paraffin wax C. Asphalt D. All
528	Concentration is a .	A. Mixing technique B. Separating technique C. Boiling technique D. cooling technique
529	Froth flotation process is used to concentrate the ore on:	A. Density basis B. Concentration basis C. Wetting basis D. Magnetic basis
530	Matte is a mixture of:	A. FeS and CuS B. Cu₂O and FeO C. Cu₂S and FeS D. CuS and FeO
531	In the bessemerization process.	A. Roasted ore is beated. B. Molten matte is removed. C. Molten matte is heated D. Molten matte is added
532	Concentration of the copper ore is carried out by:	A. Calcinations B. Roasting C. Forth flotation D. Distillation
533	When CO_2 is passed through the ammonical brine the only salt that precipitates is:	A. NaHCO₃ B. NH₄HCO₃ C. Na₂CO₃ D. (NH₄) ₂CO₃
534	In solvay's process slaked lime is used to:	A. Prepare CO₂ B. Prepare quick lime C. Recover ammonia D. Form Na₂CO₃
535	When $NaHCO_3$ is heated it forms.	A. CO₂ B. Ca(OH)₂ C. CaCO₃ D. CaO
536	Crude oil is heated to the fractionating furnace upto.	A. 300 ^oC B. 350^oC C. 400^oC D. 450^oC
537	When crude oil is heated in the fractionating tower:	A. Vapour of higher boiling point fraction condense first in the lower part of the tower. B. Vapours of lower boiling point fraction condense first in the lower part of tower. C. Vapour of higher boiling point condense lather in the upper part of tower. D. Vapours of higher boiling point never condense.
538	Which one of the following is used as jet fuel.	A. Kerosene oil B. Lubricating oil C. Fuel oil D. Diesel oil
539	Which one of the following is not faction of crude oil.	A. Paraffin wax B. Asphalt C. Fuel oil D. Petroleum coke
540	Which one of the following is not a fraction of petroleum?	A. Kerosene oil B. Diesel oil C. Alcohol D. Petrol
541	The nitrogen present in urea is used by plants to synthesize.	A. Sugar B. Proteins C. Fats D. DNA
542	Which one of the following organic compound is found is gasoline?	A. C₂H₄ B. C₃H₈ C. C₇H₁₀ D. C₁₂H₂₆
543	The plant use	A. Carbon di oxide B. Oxygen

		C. Nitrogen D. Sulphur
544	We exhale gas in the atmosphere during respiration.	A. Carbon dioxide B. Oxygen C. Nitrogen D. Water
545	The rate of reverse reaction in the beginning.	A. Slow B. moderate C. Very fast D. Low
546	The dynamic equilibrium in irreversible reaction.	A. Never establishes B. Establishes after completion of reaction C. Establishes before completion of reaction D. Establishes very soon
547	The reactions in which products can recombine to form reactants is called.	A. Irreversible reactions B. Reversible reactions C. Direct reactions D. Indirect reactions
548	A reverse reaction is that.	A. Which proceed from left to high B. In which reactants reacts to form products C. Which slow down gradually D. Which speed up gradually
549	when a system is in equilibrium state or in dynamic equilibrium state.	A. Reactants and product are equal B. Forward reactions stops C. Reverse reactions stops D. Forward reaction rate and reverse reaction rate
550	At equilibrium state there are possibilities: OR types of chemical equilibrium are.	A. 1 B. 2 C. 3 D. 4
551	A complete reaction is one is which.	A. All the reactants convert into products. B. All the reactants do not convert into products. C. Half reactants convert into products. D. Only 10% reactants convert into products.
552	A dynamic equilibrium	A. Reaction stops to proceed B. Amounts of reactants and products are equal C. Rate of forward and reverse reaction are equl D. Reaction can no longer be reversed.
553	When $CaCO_3$,.is heated in an open flask, it decomposes to form calcium oxides and .	A. O₂ B. CO C. CO₂ D. CO₃
554	For which reaction Kf is the rate constant?	A. Forward reaction B. Backward reaction C. Upward reaction D. Downward reaction
555	The substance formed during the chemical reaction are called.	A. Products B. Reactants C. Radical D. Element
556	The substances formed during the chemical reaction are called.	A. Product B. Reactants C. Radicals D. Element
557	Such reactions which continue in both directions are called.	A. Irreversible B. Reversible C. Nonreactive D. Dynamic
558	Which colour of HI is ?	A. Orange B. Purple C. Red D. Colourless
559	In chemical reaction , the substances that combine are called.	A. Reactants B. Products C. Equlibrium D. Numerator
560	In the beginning the rate of reverse reaction is	A. Negligible B. Moderate C. Very fast D. Slow
		A. Newton

561	Who proposed "Law of mass action"?	B. Boyle C. Guldberg and waage D. Lavoisier
562	Guldberg and waage out forward law of mass action in:	A. 1889 B. 1879 C. 1869 D. 1859
563	K_C is always equal to	A. Rf/Rr B. Kf /Kt C. Kf/Kr D. Rr/Rt
564	The unit of molar concentration is:	A. moldm⁻³ B. moldm⁺³ C. molcm⁻³ D. molcm⁺³
565	Formation profess of Ammonia by the combination Hydrogen and Nitrogen was given by:	A. Dalton B. Thomson C. Haber D. Waage
566	The value of K_C depends on.	A. Temperature B. Pressure C. Volume D. Atmosphere
567	The reaction in which the number of moles reactants and products not equal in balance chemical equation the units of Kc for this reactions are	A. mol⁻² B. dm³ C. mol² D. mol dm⁻¹
568	When the numbers of moles of both sides are equal in a reaction, then the unit of K_C will be:	A. No unit B. mol⁻² dm⁶ C. mol dm³ D. mol⁻²dm
569	Such reationsin which reactants and products are sufficient in quantities the Kc value of equilibrium state will be	A. Very small B. Very large C. Moderate D. None of these
570	When the magnitude of Kc is very large it indicates.	A. Reaction mixture consist of almost all products. B. Reaction mixture consist of almost all reactants. C. Reaction has not gone to completion D. Reaction mixture has negligible products.
571	When the magnitude of Kc is very small in indicates.	A. Equilibrium will never establish B. All reactants will converted to products. C. Reaction will go to completion D. The amount of products is negligible
572	When the magnitude of Kc is very large in indicates.	A. Reaction never go to completion B. Reaction is in equilibrium state C. Reaction will complete after some time D. Reaction has almost to completion
573	In a Reversible Reaction if Qc = Kc then.	A. Reaction is occuring in forward direction B. Reaction is occuring in Reverse direction C. Equilibrium has been allained D. Reaction is not at equilibrium
574	If Qc = Kc the reaction goes in:	A. Forward B. Reverse C. At equilibrium state D. None
575	If Qc < Kc the reaction goes in:	A. Forward B. Reverse C. At equilibrium state D. None
576	The reaction goes from left to right , if:	A. Qc = Kc B. Qc > kc C. Qc<Kc D. Qc = 0
577	If Qc > Kc the reaction will be in	A. Chemical equilibrium B. Static equilibrium C. Reverse reaction D. Forward reaction
578	Reaction which have comparable amount of reactants and products at equilibrium state have.	A. Very small Kc value B. Very large Kc value C. Moderate Kc value D. None of these

579	The two major components of atmosphere are:	A. Hydrogen and oxygen B. Nitrogen and Hydrogen C. Nitrogen and oxygen D. Oxygen and water.
580	The reaction will attain the equilibrium if:	A. Qc<Kc B. Qc > Kc C. Qc = Kc D. Qc = 0
581	Acids turn	A. Blue litmus red B. Red litmus blue C. Blue litmus green D. Blue litmus blue
582	The meaning of latin word acidus is.	A. Sweet B. Tasteless C. Salty D. Sour
583	The colour of litmus paper in strong acidic solution.	A. Red B. Blue C. Yellow D. Colourless
584	Acids have taste	A. Bitter B. Seetish C. Sour D. Saltish
585	Bases gave taste:	A. Bitter B. Sweetsh C. Sour D. Saltish
586	Which base is more corrosive?	A. NH₄OH B. NaOH C. Ca(OH)₂ D. AL(OH)₂
587	Word "acid" derived from:	A. Urdu B. English C. Latin D. Greek
588	Arrhenius put forward acid base concept in.	A. 1878 B. 1786 C. 1787 D. 1790
589	_____ is not an acid:	A. HCl B. NH₃ C. H₂CO₃ D. H₂SO₄
590	A base is a substance which neutralizes an acid which of these substances is not a base?	A. Aqueous ammonia B. Sodium chloride C. Sodium Hydroxide D. Calcium Hydroxide
591	conjugate base of HCl is:	A. HCl⁻ B. CH C. Cl D. NH₃
592	Which of following is not amphoteric?	A. H₂O B. NH₃ C. HCO₃ D. SO₄
593	Which of following is not conjugate base.	A. S⁻² B. F⁻¹ C. Na⁺ D. SO₄
594	When Bronsted and Lowry put forward acid base concept?	A. 1913 B. 1923 C. 1933 D. 1943
595	The bond in the product of lewis acid base reaction.	A. Ionic B. Covalent C. Metallic D. Coordinate covalent
596	Which of the following thing is a lewis base?	A. NH₃ B. BF₃ C. H⁺

		D. AlCl₃
597	According to Lewis concept, acid is a substance which.	A. Can donate protons B. Can donate an electron pair C. Can accept protone D. Can accept electron pair
598	Which is a Lewis base?	A. H⁺ B. NH₄ C. BF₃ D. AlCl₃
599	Dilute acid react with carbonates the produce product except _____	A. Salt B. Water C. Carbon di oxide D. Hydrogen gases
600	Uric acid is found in	A. Urine B. Fats C. Apple D. Grapes
601	Acid occurring in sour milk	A. Citric acid B. Lactic acid C. Bytyric acid D. Malic acid
602	Lactic acid founds in:	A. Citrus fruits B. Sour milk C. Rancid Butter D. Apple
603	Malic acid founds in:	A. Apple B. Fats C. Rancid Butter D. Oranges
604	Acetic acid is used for:	A. Elching designs B. Clearing metals C. Flavouring food D. Making explosives
605	The acid which is called king of chemicals.	A. Sulphuric Acid B. Nitric Acid C. Hydrochloric acid D. Acetic acid
606	Natural source of citric acid is:	A. Rancid butter B. Fats C. Lemon D. Sour milk
607	_____ is not mineral acid.	A. HCl B. CH₃COOH C. H₂SO₄ D. HNO₃
608	Acid reacting with metal sulphides, liberate gas.	A. Oxygen B. Hydrogen C. Hydrogen Sulphide D. Hydrogen Oxide
609	Which acid is found in Ant sting?	A. Citric Acid B. Formic Acid C. Uric Acid D. Sulphuric acid
610	_____ acid is used in lead storage batteries as electrolyte.	A. CH₃COOH B. HCl C. HNO₃ D. H₂SO₄
611	_____ is the king of chemicals.	A. Hydrochloric acid B. Nitric Acid C. Sulphuric acid D. Phosphoric acid
612	Which acid is used for food preservation?	A. Hydrochloric acid B. Benzoic acid C. Sulphuric acid D. Nitric acid.
613	Which one gas is liberated when alkalies react with ammonium salts?	A. O₂ B. CO₂ C. H₂ D. NH3
		A. White B. Brown

614	The color of $Fe(OH)_2$ ppt is.	B. Brown C. Muddy green D. Blue
615	Grease stains from clothes are removed by sing.	A. Ammonium nitrate B. Aluminium hydroxide C. Ammonium hydroxide D. Aluminum chloride
616	The colour of litmus in a strong basic solution become.	A. Yellow B. Blue C. Red D. Colourless
617	The base which is used in alkaline batteries is.	A. NaOH B. Al(OH)₂ C. KOH D. Mg(OH)₂
618	What is pOH of 0.01 M solution of KOH?	A. 3 B. 11 C. 2 D. 4
619	What is pOH of 0.01 M solution of HCl?	A. 1 B. 4 C. 12 D. 13
620	What is the pOH of Ca(OH)2, 0.02 M Solution.	A. 1.698 B. 1.397 C. 12.31 D. 12.61
621	A reaction between an acid and base produce.	A. Salt and water B. Salt and gas C. Salt and acid D. Salt and bases
622	Which one Ion is not present in salt.	A. Metallic cation B. Anion of Base C. Anion of acid D. None metallic anion
623	Potassium ferrocyanide salt is.	A. Normal B. complex C. Basic D. Acidic
624	Water of crystallization is responsible for.	A. Melting points of crystals2 B. Boiling point of crystal C. Shape of crystal D. Transition points of crystals
625	When HCl and KOH is reacted the salt formed is.	A. Acidic B. Basic C. Normal D. Complex
626	Example of complex salt is.	A. Zinc sulphate B. Potash alum C. Potassium ferrocynide D. Sodium Phosphate
627	Potassium Ferrocyanide is a:	A. Normal salt B. Mixed salt C. Complex salt D. Double salt
628	A neutral salt is not composed of.	A. Metallic ion B. Non metallic anion C. Anion of base D. Anion of acid
629	Soda lime is a mixture of..	A. CaCl, + KOH B. NaOH + CaO C. NaOH + CaCl₂ D. Ca(OH)₂ + CaO
630	_____ acid cause acidity of stomach.	A. Oxalic acid B. Hydrochloric acid C. Sulphuric acid D. Nitric Acid
631	Which acid is used for etching glass?	A. HF B. HCl C. HBr D. HI
		A. Jabir Bin Hayan

632	Vital force theory was rejected by.	A. Jabir Bin Hayan B. Drawin C. Wohler D. Berzellius
633	Hydrogen atoms preset in pentane are.	A. 10 B. 12 C. 14 D. 16
634	Molecular formula of butane is	A. C₄H₈ B. C₄H₁₀ C. C₄H₁₂ D. C₄H₆
635	Formula of Decane is.	A. C₁₀H₂₀ B. C₁₀H₂₂ C. C₁₀ H₄ D. C₁₀H₁₆
636	The example of heterocyclic compound is.	A. Benzene B. Hexane C. Cyclohexane D. Pyridine
637	Number of discovered element till today are.	A. 140 B. 118 C. 90 D. 16
638	The ability of carbon atoms to form chains or ring called.	A. Hydrogenation B. Chlorination C. Cantenation D. Halogenation
639	Bond energy of C-C is.	A. 200 KJmol⁻¹ B. 452 KJmol⁻¹ C. 300 KJmol⁻¹ D. 355 KJmol⁻¹
640	The ability of carbon atom to form chain is called.	A. Isomerism B. Catenation C. resonance D. Condensation
641	What one of the following does not contain protein?	A. Pulses B. Potatoes C. Beans D. Eggs
642	The carbon contents in lignite is.	A. 80% B. 50% C. 60% D. 70%
643	Percentage of carbon in peat is.	A. 50% B. 60% C. 70% D. 80%
644	The strong heating of coal in the absence of air called.	A. Fractional Distilation B. Destructive Distiation C. Carbonization D. Catenation
645	Carbon content in coke is.	A. 80% B. 98% C. 70% D. 88%
646	The coal in which the percentage of carbon is 60%	A. Peat B. Lignite C. Bituminous D. Anthracit
647	Wood contain the amount of carbon.	A. 70% B. 80% C. 60% D. 40%
648	Pitch is black residue of.	A. Coke B. Coal tar C. Coal D. Coal gas
649	Which one of the followoing is the Hardest coal.	A. Peat B. Lignite C. Bituminous D. Anthracite

650	Main component of Natrual gas.	A. CH₃ B. CH₄ C. C₂H₆ D. C₂H₂
651	Which is the main component of natrual gas.	A. Methane B. Ethane C. Propane D. Charcoal
652	Percentage of methane present in natural gas is.	A. 75% B. 80% C. 85% D. 90%
653	_____ is synthetic fiber.	A. Cotton B. Woal C. Nylon D. Silk
654	General formula of Alkane is.	A. C₁₀H_{2n-2} B. C₆H₂₆ C. C₁₁H_{2n-2} D. C_nH_{2n-2}
655	General formula of Alkynes is.	A. C₆H_{2n-2} B. C_nH_{2n+2} C. C₁₁H_{2n-2} D. C_nH_{2n-2}
656	General formula of Alkyl radical is..	A. C₆H_{2n-2} B. C_nH_{2n+2} C. C₁₁H_{2n-1} D. C_nH_{2n-2}
657	In which of the following groups , oxygen is attached on both sides with carbon atoms.	A. Ketone group B. Ether group C. Aldehyde group D. Easter group
658	The functional group -COOH is found in.	A. Aldehydes B. Esters C. Carboxylic acids D. alcohols
659	Functional group of Alcohols is	A. -COOH B. -C=0 C. C-O-C D. -OH
660	Incomplete combustion of alkanes produces.	A. Carbon dioxide only B. Carbon monoxide only C. Carbon monoxide and carbon black D. Carbon dioxide and carbon black
661	Which is the substitution reaction in the following?	A. Halogenation of alkenes B. Halogenation of alkynes C. Bromination of alkene D. Halogenation of alkanes
662	Marsh gas is mostly consisted of.	A. Butane B. Propane C. Ethane D. Methane
663	Which hydrocarbon has no effect on an aqueous solution of bromine.	A. CH₄ B. C₁₀H₂₀ C. C₂H₂ D. C₂H₄
664	Amont the followign substitution reaction is characteristics of.	A. Alkanes B. Alkenes C. Alkynes D. None of these
665	Halogenation of methane on the presence of diffused sunlight takes place.	A. Only in one step B. Slowly in one step C. Freshly in two steps D. In a series of four steps
666	Which compound do not produce in the halogenation of methane.	A. Chloroform B. Carbon tetrachloride C. Carbon black D. Chloromethane
667	Dichloromethane reacts with halogens in the presence of diffused sun light and produces	A. Carbon tetrachloride B. Chloroform C. Chloromethane D. No reaction

		D. No reaction
668	Chemical formula of chloroform.	A. CH₃Cl B. CH₂O₂ C. CHCl₂ D. CCl₄
669	Which hydrocarbon molecule would have no effect on the aqueous solution of bromine?	A. CH₄ B. C₁₀H₂₀ C. C₂H₄ D. C₂H₂
670	How many percent of natural gas is consisted of methane.	A. 60% B. 70% C. 80% D. 85%
671	Which are called paraffines?	A. Alkyle B. Alkynes C. Alkenes D. Alkanes
672	The reduction of alkyl halides takes place in the presence of.	A. Zn/HCl B. Na/HCl C. Mg/HCl D. Cu/HCl
673	Which gas in produced during ripening of bnabanas.	A. Methane B. Ethene C. Acetylene D. Carbon dioxide
674	One of the hydrocarbons reacts with one mole of hydrogen to form a saturated hydrocarbon. What is the formula could be of the x:	A. C₃H₈ B. C₆H₁₂ C. C₄H₁₀ D. C₂H₁₆
675	General formula of alkenes.	A. C_n H_{2n-2} B. C_nH_{2n} C. C_nH_{2n+2} D. C_nH_{2n-1}
676	Dehydrohalogenation take place in the presence of.	A. Aqueous NaOH B. Alcoholic KOH C. Aqueous KOH D. Alcoholic NaOH
677	Which hydrocarbon decolorizes the pink colour of acidic solution of potassium permanganate?	A. CH₄ B. C₂H₄ C. C₂H₆ D. C₂H₅
678	The order of reactivity of hydrogen halides with alkenes is:	A. HI>HBR B. HBr>HI C. HCl>HBr D. HBr>HCl
679	By which process alkenes are prepared from alcohols:	A. Dehelogenation B. Dehydrogenation C. Dehydration D. Dehydro-halogenation
680	Which of following dehydration of alcohols take place:	A. NaOH B. KOH C. H₂SO₄ D. HCl
681	The process which convert vegetable oil into banaspati ghee called.	A. Hydrogenation B. Halogenation C. a and b D. None of these
682	Oxidation of alkenes forms:	A. Glycol B. Glyoxal C. Formic acid D. Oxalic acid
683	_____ use for ripening of fruits.	A. Ethene B. Ethane C. Ethyne D. Propane
684	The order or reactivity of hydrogen halides with alkenes is.	A. HI/Br B. HBr > HI C. HCl > HBr D. HBr>HCl
		A. CH = CH B. CH₃ - CH₃

685	Ethylene glycol can be prepared with the reaction of $KMnO_4$ and _____	C. CH₃CH₂CH₃ D. CH₂= CH₃
686	Oxidation of Ethene with $KMnO_4$ produces.	A. Oxalic acid B. Glyoxal C. Ethene Glycol D. Propene glycol
687	General formula of alkynes is:	A. C_nH_n B. C_nH_{2n} C. C_nH_{2n+2} D. C_nH_{2n-2}
688	Ethyne is oxidized by alkaline $KMnO_4$ then ..... hydroxyl group add to the triple bond.	A. Two B. Three C. Four D. Five
689	The end product of oxidation of acetylene is.	A. Propen glycol B. Ethan glycol C. Glyoxal D. Oxalic acid
690	How trace amount of acetylene present in the coat gas.	A. 0.06% B. 0.07% C. 0.08% D. 0.09%
691	Dehydrohalogenation of vicinal-dihalides take place in the presence of.	A. NaOH B. KOH C. NaCl D. None of these
692	Dehalogenation of Tetrahalides takes place in the presence of .	A. K B. Mg C. Na D. Zn dust
693	Boiling point of alcohol in centigrade is.	A. 68 B. 78 C. 118 D. 128
694	Alkynes are also called.	A. Olefine B. Ethene C. Paraffins D. Acetylenes
695	General formula of carbohydrates is.	A. C_n(H₂O)₆ B. C_n(H₂O₂)₆ C. C_n(H₃O)₆ D. C_n(H₂O₃)₆
696	Which one of the following is crystalline solid.	A. Glucose B. Starch C. Cellulose D. Glycogen
697	Plants produces by the photosynthesis process.	A. Glucose B. Fructose C. Sucrose D. Maltose
698	Photosynthesis process produce.	A. Carbondioxide B. Glucose and oxygen C. Carbon di oxide D. glucose and carbondoxide
699	Which of the following is reducing sugar?	A. Glucose B. Maltose C. Sucrose D. Starch
700	Which part of digestive system glucose absorb.	A. Stomach B. Liver C. Small intestine D. Large intestine.
701	Monosaccharide consists of number of carbon atoms.	A. 2 to 4 B. 4 to 8 C. 3 to 9 D. 5 to 10
702	Which one is the simplest sugar which can not be hydrolyzed?	A. Glucose B. Sucrose

702	Which one is the simplest sugar which can not be hydrolyzed?	C. Starch D. Cellulose
703	Which is Pentahydroxy aldehydes of the following?	A. Starch B. Glucose C. Fructose D. Sucrose
704	Pentahydroxy ketone is called as:	A. Glucose B. Fructose C. Starch D. Sucross
705	When glucose and fructose combine and forms.	A. Starch B. Cellulose C. Sucrose D. None of these
706	Which is disaccharide?	A. Glucose B. Fructose C. Sucrose D. Starch
707	Which one of the following is testeless?	A. Sucrose B. Glucose C. Fructose D. Starch
708	Which protect our muscles from cramping?	A. Carbohydrats B. Vitamins C. Lipids D. Proteins
709	How many percentage of protein is present in dry weight of animal?	A. More then 20% B. More then 30% C. More then 40% D. More then 50%
710	Proteins are polymers of.	A. Polysaccharide B. Oligosaccharids C. Amino acid D. Nucleic acid
711	Amino acid are linked to each other through.	A. Hydrogen link B. Ionic link C. Gelatin link D. Peptide link
712	Which one of the following does not contain proteins.	A. In pulses B. In Potatoes C. In fruits D. In eggs
713	Gelatin protein is found in.	A. Blood B. Skin C. Heart D. Bones
714	Which one is triglyceride.	A. carbohydrates B. Proteins C. Lipids D. Vitamins
715	Building block of lipids are.	A. Fatty acids B. Carboxylic acids C. Mineral acids D. Alcohols
716	Molecular formula of stearic acid is.	A. C₁₅H₃₁ COOH B. C₁₂H₁₅COOH C. C₁₅H₃₀COOH D. C₁₇H₃₄COOH
717	Formula of palmitic acid is.	A. C₁₅H₃₁COOH B. C₁₂H₃₅COOH C. C₁₅H₃₀COOH D. C₁₇H₃₄COOH
718	These are safe source of food and energy for body during emergency.	A. Proteins B. Vitamins C. Carbohydrats D. Lipids
719	The body reactions are catalized by.	A. Amino acids B. Lipids C. Enzymes D. Fatty acids

720	Essential components of every living all are.	A. Nucleic acids B. Enzymes C. Lipids D. Cell wall
721	Vitamin B_1 (Thiamin) was discovered by.	A. Hopkins B. Funk C. J.Watson D. Davy
722	Cause of Night blindness is.	A. Deficiency of vitamin B B. Deficiency of vitamin A C. Deficiency of vitamin D D. Deficiency of vitamin E
723	Eye inflammation is caused by the deficiency of vitamin.	A. Vitamin D B. Vitamin C C. Vitamin B D. Vitamin A
724	Which of the following is fat soluble vitamin?	A. Vitamin A B. Vitamin E C. Vitamin K D. All of these
725	Which disease cause by deficiency of vitamin D?	A. Rickets B. Anemia in babies C. Scurvy D. Night blindness
726	Which diseases cause by the deficiency of vitamin E?	A. Rickets B. Anemia in babies C. Scurvy D. Night blindness
727	Rickets disease is caused by the deficiency of.	A. Vitamin D B. Vitamin A C. Vitamin E D. Vitamin C
728	Which vitamin is soluble in water?	A. Vitamin A B. Vitamin E C. Vitamin D D. Vitamin C
729	Number of vitamins in vitamin B complex is.	A. 10 B. 8 C. 6 D. 12
730	About 99% of atmospheric mass lies in.	A. 35 km B. 30 km C. 15 km D. 16 km
731	What percentage of nitrogen and oxygen is in the atmosphere.	A. 69% B. 79% C. 9% D. 89%
732	Which gas is not present in atmosphere?	A. Nitrogen B. Oxygen C. Helium D. Carbon dioxide
733	The sphere just above the Earth's surface is.	A. Mescosphere B. Stratosphere C. Thermosphere D. Troposphere
734	A strange bitter smell noticed near photo copier machine is of.	A. H₂S B. SO₂ C. O₂ D. O₃
735	Our planet earth consists of _____ natural Spheres.	A. 1 B. 4 C. 6 D. 8
736	The layer above troposphere extend upto 50 kilometers called.	A. Mesosphere B. Hydrosphere C. Stratosphere D. Thermosphere
737	75% atmosphere mass lies in how many kilometers.	A. 11 km B. 30 km C. 50 km D. 85 km

738	Depending upon temperature variation, atmosphere ins divided into how many regions?	A. 1 B. 2 C. 3 D. 4
739	On which base atmosphere is divided into four regions?	A. Change in pressure B. Change in rediation C. Change in temperature D. Change in volume
740	Infrared radiation emitted by the earth are absorbed by?	A. CO₂,H₂O B. N₂,O₂ C. CO₂,N₂ D. O₂,CO₂
741	The stratosphere layer is at height above the earth's surface.	A. 0-12 km B. 12-50 km C. 50-85 km D. 65-120 km
742	Ozone is formed in:	A. Troposphere B. Stratosphere C. Mesosphere D. Thermosphere
743	At the height 880120 km from earth's surface is.	A. Troposphere B. Mesosphere C. Stratosphere D. Thermosphere
744	Ozone is beneficial for us as it absorbs.....	A. Infrared Radiations B. Ultraviolet radiations C. Chlorofluorocarbons D. Air pollutants
745	Which one of the following is not air pollutants?	A. N₂ B. CO C. NO₂ D. O₂
746	How many percentage of sunlight is absorbed by atmosphere gases.	A. 12% B. 18% C. 24% D. 30%
747	The earth atmosphere is getting hotter because of :	A. SO₂ B. O₃ C. CO₂ D. CO
748	Which one of the following gas has greater retaining capacity?	A. O₂ B. N₂ C. CO D. CO₂
749	Which pollutant is not found in car exhaust gases?	A. CO B. O₃ C. NO₂ D. SO₃
750	_____ is not green house effect.	A. Increase of food chain B. increase atmospheric pressure C. Increase of sea level D. Increase of flood risks
751	How many atmospheric temperature increase every year due to accumulation of CO_2 in air.	A. 0.03 ^oC B. 0.05 ^oC C. 1^oC D. 2^oC
752	Cause of global warming is :	A. CO₂ gas B. SO₂gas C. NO₂gas D. O₂gas
753	Cause of global warming is:	A. CO₂gas B. SO₂gas C. NO₂gas D. O₂gas
754	Carbon monoxide is harmful to use because it:	A. Paralyses lungs B. Damages lungs C. Reduce oxygen carrying ability of hemoglobin D. Make the blood coagulate
755	Building and Monuments lose their beauty and shine due to	A. U.V radiations B. Chlorofluoro carbons C. Acid rain

		D. IR radiations.
756	Which metal present in acid rain affect the aquatic life by clogging fish gills?	A. Lead B. Chromium C. Aluminium D. Mercury
757	pH of acid rain is:	A. 7-5 B. 6-7 C. 1-2 D. 4-5
758	Normaily rain water is weakly acidic because.	A. SO₃ B. CO₂ C. CO₃ D. NO₂
759	Acid rain damages the building it is due to reaction on.	A. Calcium sulphate B. Calcium nitrate C. Calcium carbonate D. Calcium oxilat
760	Which one of the following is not air pollutant.	A. CO₂ B. CO C. NO₂ D. O₂
761	The formula of ozone is.	A. O₂ B. O₃ C. O D. CO
762	Which gas saves the surface of earth from ultra violet radiation?	A. CO₂ B. CO C. N₂ D. O₃
763	The % age of drinkable water on earth is	A. 2.0% B. 0.02% C. 0.2% D. 5.0%
764	How many percentage to water contains oceans?	A. 67% B. 77% C. 87% D. 97%
765	Specific Heat capacity of water is about.	A. 4.0 Jg⁻¹ K⁻¹ B. 4.1 Jg⁻¹K⁻¹ C. 4.2 Jg⁻¹K⁻¹ D. 4.3 Jg⁻¹K⁻¹
766	The heat capacity of water is greater than rocks.	A. 4 times B. 5 times C. 6 times D. 7 times
767	The density of water is ..... at 4 ^{o}C	A. 1 gcm⁻¹ B. 2 gcm⁻¹ C. 3 gcm⁻¹ D. 4 gcm⁻¹
768	At which temperature o water shows mxamimum density.	A. 0 ^oC B. 100^oC C. 4^oC D. -4^oC
769	Boiling point of water is:	A. 0^oC B. 25^oC C. 80^oC D. 100^oC
770	Which one of the following properties of water is responsible for rising of water in plants?	A. Specific heat capacity B. Viscosity C. Excellent solvent action D. Capillary action
771	Which process is responsible for ascending of water in plants from roots to leaf?	A. Condensatin B. Transpiration C. Capillary action D. Evaporation
772	A sea level the boiling point of water is.	A. 0^oC B. 98^oC C. 100^oC D. 110^oC
		A. Ionic

773	Water molecule has a structure:	B. Non polar C. Tetrahedral D. Polar
774	Ionic compounds are soluble in water due to:	A. Hydrogen bonding B. ion-dipole forces C. Dipole-dipole forces D. Dipole -induced dipole forces
775	The bond angle between H-O-H in water is:	A. 104.5^o B. 104.6^o C. 104.7^o D. 104.8^o
776	Water dissolves Non-ionic compounds by:	A. Ion-Ion forces B. Dipole forces C. Dipole - Dipole forces D. Hydrogen Bonding
777	Which ion cause water hardness?	A. Al² B. Mg² C. Fe² D. Na²
778	Temporary hardness is because of.	A. Ca(HCO₃)₂ B. CaCO₃ C. MgCO₃ D. MgSO₄
779	Temporary hardness can be removed by_____.	A. Quick lime B. Slaked lime C. Lime stone D. HCl
780	Permanent hardness can be removed by using .	A. Soda lime B. Sodium Zeolite C. Quick lime D. Lime water
781	Which one of the following salts makes the water permanent hard.	A. NaCO₃ B. NaHCO₃ C. Ca(HCO₃)₂ D. CaSO₄
782	Salt makes the water permanently hard.	A. CaSO₄ B. Ca(HCO₃)₂ C. NaHCO₃ D. NaCO₃
783	Which salt cause permanent hardness in water is:	A. Ca(HCO₃)₂ B. Mg(HCO₃)₂ C. CaCl₂ D. KCl
784	Clark's method is used to remove the hardness of water, in this method which is used.	A. Ca(HCO₃)₂ B. Na-Zeolite C. Ca-Zeolite D. Ca(OH)₂
785	Which one of the following ion does not cause hardness in water?	A. Ca⁺² B. Mg⁻² C. SO⁻² D. Na⁺
786	Sodium zeolite is resin of:	A. NaAl(SiO₃)₂ B. KAl(SiO₃)₂ C. LiAl(SiO₃)₂ D. RbAl(SiO₃)₂
787	The removal of Mg^{+2} and Ca^{+2} Ion which are responsible for the hardness of water is called.	A. Temporary hardness B. Permanent hardness C. Water softening D. Hydrogen bonding
788	The process of removing temporary hardness of water.	A. Clark's method B. Washing soda method C. Sodium zeolite D. Filteration method
789	Temporary hardness of water is removed by adding.	A. Ca(OH)₂ B. NaOH C. KOH D. CaSO₄
790	The lives of aquatic plants and animals are indirectly related to concentration of dissolved gas in water.	A. Nitrogen B. Hydrogen C. Oxygen D. Carbon

791	Which disease cause when humans use water of industrial effluents.	A. Cancer B. Asthma C. Jaundice D. cholera
792	Which disease cause when humans use water of industrial effuents.	A. Cancer B. Asthma C. Jaundica D. Cholera
793	Rapid growth of algae in water bodies in because of detergent having.	A. Carbonat salt B. Sulphonic acid C. Sulphat salt D. Phosphate salt
794	Cause of diarrhoea among following is:	A. Diarrhea B. Cholera C. Cryptosporidium D. Typhoid
795	Which disease causes bone and tooth damage?	A. Fluorosis B. Hepatits C. Cholera D. Jaundice
796	Which compound protect teeth from diseases?	A. Fluorine compound B. Chlorine compound C. Iodine compound D. Bromine compound
797	Which disease cause liver inflammation?	A. Typhoid B. Jaundice C. cholera D. Hepatitis
798	Vibrious cholera bacteria causes the disease	A. Cholera B. Diarrhea C. Jaundice D. Hepatitis
799	Which gas is used to destroy harmful bacteria in water?	A. Iodine B. Chlorine C. Flourine D. Bromine
800	Cholera is caused by	A. Protozoa B. Virus C. Bacteria D. Fungi
801	Which one of the following disease causes severe diarrhea and can be fatal:	A. Jaundic B. Dysentery C. Cholera D. Typhoid
802	Which of the following gas is used to destroy harmful bacteria in water.	A. Iodine B. Chlorine C. fluorine D. Bromine
803	Swimming pools are cleaned by a process:	A. Hydrogenation B. Bromination C. Chlorination D. Nitration
804	Concentration of copper are is carried out by process.	A. Roasting B. Forth floatation C. Gravity sepration D. Electrostatic separation
805	the chemical formula of chalcopyrite is:	A. Cu₂S B. CuFe₂S C. CuS D. FeS
806	Chalco-Pyrite is an ore of:	A. Copper B. Silver C. Iron D. Aluminium
807	Concentration is process of:	A. Mixting technique B. Separating technique C. Boiling technique D. Cooling techique
808	How many percentage blister copper is pure?	A. 95% B. 75% C. 98%

		D. 100%
809	On which bases froth floatation process take place?	A. Density bases B. Concentration basis C. Watting basis D. None of these
810	The impurities associated with the minerals are known as:	A. Metallurgy B. Ores C. Gangue D. Compounds
811	When $NaHCO_3$ is heated if forms:	A. CO₂ B. Ca(OH)₂ C. CaCO₃ D. CaO
812	Ammonia is prepared by:	A. Solvay's process B. Haber's process C. Floatation process D. Bayer's process
813	Urea is nitrogenous fertilizer. It consist of nitrogen.	A. 26.6 % B. 36.6% C. 46.6% D. 56.6%
814	How many units are in Pakistan for manufacturing of Urea.	A. 3 B. 6 C. 9 D. 12
815	Composition of carbon in kerosene oil is.	A. C₁₀-C₁₂ B. C₁₂-C₁₂ C. C₁₆-C₁₈ D. C₃₀-C₃₀
816	Composition of carbon in the diesel oil.	A. C₇-C₁₀ B. C₁₀-C₁₂ C. C₁₃-C₁₅ D. C₁₅-C₁₈
817	Which one is not the fraction of residual oil?	A. Paraffin wax B. Asphalt C. Fuel oil D. Coke
818	Petroleum fraction having composition C_7 to C_{10} IS called	A. Petroleum gas B. Petroleum Ether C. Kerosine oil D. Gasoline or petrol
819	Petroleum fraction having composition C_5 to C_7 is.	A. Petroleum Gas B. Petroleum Ether C. Kerosine oil D. Gasoline or petrol