Chemistry General Science Test Easy Mode

Sr	Questions	Answers Choice
1	Matter is present in neon signs in the state of	A. Supereritical fluid B. Gas C. Liquid Crystal D. Plasma
2	Hazardous effects of shopping bags are studied in	A. Geochemistry B. Inorganic chemistry C. Environmental chemistry D. analytical chemistry
3	The man made polymer is	A. Polystyrene B. Starch C. Protein D. Cellulose
4	The crystals of whic substance has rhombic shape?	A. Brass B. Bronze C. Sulphur D. Grahite
5	Which liquid among the followingis a colloidal solution.	A. Milk B. Slaked lime used for white wash C. Vinegar solution D. Mixtur eof AgCl in water
6	Which of the following is a heterogeneous mixture	A. Concrete mixture B. a solution of potassium nitrate in water C. a solution of calcium hydroxide in water D. Hot chocolate
7	A state of matter whose properties are between those of liqids and crystalline solids.	A. Plasma B. dark matter C. Liquid crystal D. Supercriticl fluid
8	When the tiny particles of a substance are dispersed athough a medicium, the mixture is named as.	A. True solution B. Suspension C. Colloid D. Heterogeneios mixture
9	You are styding the rate of hydrolysis of organic compund staarch under different conditonns of temperature. In which branch of chemistry this topic will fall	A. Organic Chemistry B. analytical chemistry C. Biochemistry D. Physical Chemistry
10	Which branch of chemistry is athe study elements and their compounds except for organic compounds?	A. Physical Chemistry B. Geochemisty C. Organic chemistry D. Inorganic chemistry
11	Which branch of chemistry helps to protectwater that bas been poisoned by soil.	A. Organic chemistry B. inorganic chemistry C. Environmenal chemistry D. Geochemistry
12	Which area of chemistry imoroves to gauge the behavior of pollutants and develp techniques for polution control?	A. analytical chemistry B. geochemistry C. Organic chemistry D. Environmental Chemsitry
13	The branch of chemistry that helps to treat dseases and synthesize new medicines.	A. Physical Chemistry B. Environmental Chemistry C. Organic chemistry D. Inorganic chemistry
14	The branch of science helps to understand chemcial product andprocess that reduce the use of hazadous substances.	A. Green chemistry B. Astrochemistry C. Analytical chemistry D. Physical Chemistry
15	To Identify the concentration of a particular solution through titration is and application of.	A. astrochemistry B. analytical chemistry C. geochemistry

		D. Organic chemistry
16	The batteries in our vehicles are bilt on the principl of electronchemistry It is the application.	A. Astrochemistry B. Physical chemistry C. Analyticla chemistry D. Organic chemistry
17	The branch of chemistry that is concerned with the large scale production of chemical substances is.	A. Environemantal chemistry B. Inorganic chemistry C. Physical Chemistry D. Industrial chemistry
18	The branch of chemistry that focuses on the study of polymers, their types properties, uses is called.	A. Industrial chemistry B. Polymer chemistry C. Organic chemistry D. Astrochemistry
19	The study of theinteraction between drugs and biological targets, as wellas the development of new medicinal agents.	A. Inorganic Chemistry B. Organic chemistry C. Medicinal Chemistry D. Environmental chemistry
20	Which is deal with compositon, structure, properties, behavior and changes of matter and energy.	A. Technology B. Chemistry C. Engineering D. Science
21	Whcih branches of chemistry deals with the changes that occur in atomic nuclei.	A. Environmental chemistry B. nuclear chemistry C. Astrochemistry D. Bio chemistry
22	Which branch of chemistry deals with elements and compounds is earth's crust.	A. Polymer chemistry B. Organic chemistry C. Geochemistry D. Physical chemistry
23	Which branch of chemistry deals with the study of stars, planets, comets and interstellar space.	A. Physical chemistry B. Medcinal chemistry C. Astrochemistry D. Geochemistry
24	Which branch of chemistry tells us sulphuric acid is extremely corrosive to skin.	A. Organic chemistry B. Physical chemistry C. Inorganic chemistry D. Biochemistry
25	Which branch of chemistry give information about starch syntehsis in plants?	A. Biochimstry B. Organic chemistry C. Environmental chemistry D. Inorganic Chemistry
26	Polymers are sometimes called	A. Monomers B. Micromolecules C. Macromolecules D. None of these
27	Any thing that has mass and occupies space is called.	A. Liquid B. Gas C. Solid D. Matter
28	Following are state of matter.	A. Liquid B. Solid C. Gas D. All of these
29	Macroscopic propertues are properties that can be visualized by	A. Telescope B. The naked eye C. Microscope D. Electron microscope
30	Matter can be described by both is	A. Chemical Properties B. Physical Properties C. Physical propeties and chemcial properties D. None of these
31	A substance formed when two or more different elements combine chemically.	A. Compund B. Solution C. Element D. Atom
32	How many state of matter exist.	A. One B. Two C. Three D. Four
		A. Liquid

33	In whcih state of matter does not hae definite shape and volume.	A. Liquid B. Gas C. Solid D. All of these
34	Pressure is a significant property of.	A. Gas B. Solid C. Liquid D. None of these
35	Which state of matter has fixed shape and volume?	A. Liquid B. Solid C. Gas D. All of these
36	The simplist form of matter	A. Liquid B. Solid C. Gas D. Both b and c
37	Building block of all matter is.	A. element B. Compound C. Mixture D. Atom
38	Mixture can be separated by means of	A. Nuclear method B. Phyical medhod C. Chemical Method D. All of these
39	Example of heterogeneious mixture is	A. Soil B. Gasoline C. AIR D. Ice Cream
40	A good example of homogeneious mixture is.	A. Soil B. Ice-Creem C. Rock D. Wood
41	Allotropes of oxygen are	A. 2 B. 3 C. 4 D. 5
42	Graphite is .....conductor of electiricity	A. Bad B. Moderate C. Good D. None of these
43	In diamond, each C-atom is covaently bonded wiht.....C-atom	A. Two B. Three C. Four D. Five
44	Diamond is.....	A. conductor B. non -conductor C. both a and b D. None of these
45	The structure of C atom in diamond is.	A. Pentagonal B. Triangle C. Hexahedral D. Tetrahedral
46	Buckyballs also knowns as.	A. Fullerenes B. Graphite C. Both a and b D. None of these
47	Whcih one of the following wil show negligible effect of termperture on its solubility.	A. KCl B. KNO3 C. NaCl D. NaNO3
48	Which one of the following is heterogeneious mixture.	A. Milk B. Ink C. Sugar solution D. Milk of magnesia
49	Solubility of whcih salt increases with the increase of temperture.	A. KNO3 B. KCL C. NaNO3 D. All of These
50	The solubility of which salt decreases with the increase of temperture.	A. KNO3 B. Li2SO4 C. KCl D. NaNO3

51	Which one of the example of colloid.	A. Paints B. Milk of magnesia C. Jelly D. None of these
52	Which one of the suspension.	A. Chalk in water B. Tootpast C. Ink D. Blood
53	How many electrons can be accommodated at the msot in the third shell of the elements.	A. 8 B. 12 C. 10 D. 18
54	What information was obtained fom discharge tube expriments?	A. Electrons and protons were discovered B. Structur eof atom was discovered C. Neutrons and protons were discovered D. Presence of nucleus in an atom was discovered
55	Why has isotopes not been shown int he periodic table.	A. Isotopes do not show periodic behavior B. Periodic tble cannot accomodate a large number of isotopes of different elements C. All the isotops have same atomic number so there is no need to give them separate places D. Someof the isotopes are unstble and they give rise to different elements.
56	Which particle is present in differente number in the isotopes.	A. Proton B. Electron C. Neutron D. Both neutron and electron
57	In which isotope of oxygen there are the equal numbr of protons electons and neutorns.	A. O¹⁶ B. ¹⁷O C. ¹⁸O D. None of these
58	What will be the relative atomic mass of nitrogen given the abundance of its two isotopes, ^{14}N and ^{15}N are 99.64 and 0.35	A. 14.0210 B. 14.2100 C. 14.0021 D. 14.1200
59	How is radio carbon dating useful for archeolgists.	A. It helps detemine whether the matter is radioactie or not B. It helps determine the age of organic matter C. It helps determine the composition of matter D. It helps deterime the usefulness of matter
60	What does keep the particles present in the nucleus intact.	A. Particles are held together by dipolar force B. Particles are held together by wee nuclear force C. Particles are held together by strong nuclear force D. Particles are held together by electrostatic force
61	How doe electrons keep themselves away from the oppositely charged nucleus.	A. A agnetic field around the nucleus keeps them away B. By keeping themselves stationary C. By revolving around the nucleus D. Due to their wave like nature
62	M shell has sub sheels.	A. 1s , 2s B. 1s,2s,3s C. 2s,2p D. 3s,3p,3d
63	A sub sheel that cna accommodate 6 electrons is	A. b B. s C. f D. p
64	John Dalton put forward his atomic theory.	A. 1800 B. 1803 C. 1805 D. 1903

65	Rutherford used a gold foil in his experiment, which had a tickness of	A. 0,0002 cm B. 0,0001 cm C. 0.001 cm D. 0.00004cm
66	Who performed fist experiment to split atom	A. Bohr B. Newton C. Rutherford D. Soddy
67	According to Rutherford's atomic theory, atom should produce.	A. Line spectrum B. Continuous spectrum C. Both a and b D. None of these
68	Quantum means.	A. Virable energy B. Fixed energy C. High energy D. Minimum energy
69	Proton are feflected toward plate.	A. Positive B. Negative C. Both a and b D. None of these
70	The nucleus of an atom is composed of	A. Electrons B. Electons and protons C. Protons and neutrons D. Electrons and neutrons
71	How many electons can be accommodated in S subshell?	A. 2 B. 6 C. 14 D. 10
72	Number of electrons that an be accomodated in f - subshell	A. 6 B. 10 C. 2 D. 14
73	Whcih subshells are present in L - shell?	A. S and P B. Only s -sub shell C. Only p - sub shell D. Sub shell
74	How many subshells are there in M shell?	A. 2 B. 4 C. 3 D. 5
75	N-shell contains number of subshells.	A. 1 B. 3 C. 4 D. 2
76	An element has 5 electrons in M shell. Its tomic number is.	A. 5 B. 10 C. 15 D. 20
77	d- subshell can accommodate maximum electrons.	A. 2 B. 6 C. 10 D. 14
78	The removal of electron from a neutral atom gives rise to.	A. Molecular anion B. Anion C. Cation D. Molecular Cation
79	How many electrons can be accommodated at the most in the third shell of the elements.	A. 8 B. 10 C. 18 D. 32
80	Number of neurtons in $^{27}M_{13}$are	A. 13 B. 14 C. 27 D. 15
81	Number of protons in the nucleus of an atom is called.	A. Atomic number B. Mass Number C. Mass Unit D. Electron Number
82	Atom is electrically	A. Positive particle B. Negaitve particle C. Neutral particle

		D. None of these
83	Atomic number is represented by	A. P B. A C. At D. Z
84	$^{238}U_{92}$ has number of neutrons.	A. 92 B. 146 C. 238 D. 330
85	Mass Number is repreented by	A. Z B. S C. A D. M
86	Which of the following statement is not correct about isotopes.	A. they have same atomic number B. They have same number of protons C. They have same physical properties D. They have same chemical propeties
87	Which isotope is used in nuclear reactors.	A. U-234 B. U-235 C. U-238 D. All of these
88	Chlorine has two isotopes, both of which have	A. Same mass Number B. Same number of electrons C. Same number of neurtons D. Different number of protons
89	Which Isotopes is commonly used to irradiate cancer cells.	A. Cobalt -60 B. Iodine-23 C. Carbon -14 D. Iodine-131
90	Number of isotopes of hydrogen is	A. 2 B. 5 C. 4 D. 3
91	^{13} C and ^{14} C are both present in nature.	A. 0,1 % B. 1.1 % C. 0.9 % D. 1.5 %
92	The percentage of ^{238} U $_{92}$ found in nature.	A. 97% B. 0.72% C. 98% D. 1.5%
93	Whcih Isotopes is used for diagonosis of goiter?	A. Iodine-131 B. Cobalt -60 C. P-32 D. Sr-90
94	Carbon -14 is used for the	A. Growth of bones B. Diagonosis of goiter C. Age determination of old objects D. All of these
95	When molten copper and olten zinc are mixed together , they give rise to a new substnce called brass. Predict what type of bond is formed between copper and zinc.	A. Ionic bond B. Coordinate Covalent bond C. Metallic bond D. Covalent Bond
96	Which element is capable of forming all the three types of bonds, covalent coordinate covalent or ionic.	A. Carbon B. Silicon C. Magnesium D. Oxygen
97	Why is H_2Oa liquid while H_2S is a gas?	A. Becuase in waer, the atomic size of oxygen in smaller than that of Sulphur B. Because water can easily freeze into ice C. Because water is a polar compound ad tere exists strong forces of attrction betwene its molecules D. Because H₂O molecule is lighter than H₂S
98	Which of the following bond s is expected to b the weakest.	A. Cl-Cl⁻ B. C-C C. F-F

		D. O -O
99	Which form of carbon is used as a lubricant?	A. Coal B. Diamond C. Charcoal D. Graphite
100	Keeping in view the intermolecular forces of attraction, indicate which compound has the highest boiling point	A. H₂S B. HF C. NH₃ D. H₂O
101	Which metal has the lowest melting point?	A. Li B. Na C. Rb D. K
102	Which ionic compound has the highest melting point.	A. RbCl B. KCl C. LiCl D. NaCl
103	Whcih compound contains both covalent and ionic bonds.	A. MgCl₂ B. PCl5 C. NH₄Cl D. CaO
104	Which among of the followign has a double covalent bond.	A. Ethane B. Methane C. Acetylene D. Ethylene
105	Atoms achieve stability by attainign electonic configuration of.	A. Alkali metals B. Coinage metals C. Inert Gases D. Alkaline earth metals
106	Attaining two electrons in the valence shell is called.	A. Octet rule B. Duplet rule C. Triplet rule D. All of these
107	All the noble gases have their valence electrons.	A. Incomplete B. Partialy filled C. Completely filled D. None of the above
108	Noble gases are non -reactive, because they do not.	A. Gain electrons B. Lose electrons C. Share electrons D. All of these
109	Every atom has a natural tendency to accommodate electrons in its valence shell	A. 2 or 6 B. 2 or 4 C. 2 or 8 D. 2 or 10
110	Hydrogen and Helium follow.	A. Octet rule B. Triple rule C. Duplet rule D. None of these
111	Whcih of the following atoms obey duplet rule.	A. O₂ B. Cl₂ C. H₂ D. Li₂
112	Whcih of the following is not true about the formationof Na_2S	A. Each sodium atom loses one electron B. Sodium forms cation C. Each sulphur atom gains one electron D. Sulphur form anion
113	Octet rule is	A. Attainign of eight electrons in its valence shell B. Discription of eight electrons C. Pictur eof electronic configuration D. Pattern of electronic configuration
114	Atoms react with each other because.	A. they are attracted towards each other B. They are short of electrons C. They want to disperse D. They want attain stability
115	An atom having six electonsin its valence shell will achieve noble gas electronic confguration by	A. Gainign one electron B. Gaining two electrons C. Losing all electrons

	configuration by	C. Losing all electrons D. Losing two electrons
116	The formaton ionic bond between two ions is due to.	A. Hydrogen bonding B. Metalic force C. Electrostatic forces D. All of the above
117	Which group of the periodic table has the tendency to gain electrons.	A. Group -1 B. Group -17 C. Group-2 D. Group -18
118	Whcih of the following atoms willnot form cation or anion.	A. Atomic no. 16 B. Atomic no. 18 C. Atomic no. 17 D. Atomic No. 19
119	Transfer of electron between elements result in.	A. Coordinate covalent bonding B. Ionic bonding C. Metallic bonding D. Covallent bonding
120	When an electronegative element combines with electrpositive element, the type of bonding. is.	A. Covalent B. Polar Covalent C. Ionic D. Coorinae Covalent
121	How many electron are there in the valence shell of sodium atom.	A. One B. Two C. Three D. Four
122	The electropositive elements have the tendency to	A. Lose electrons B. Gain electrons C. Share electrons D. All of these
123	How many valance shell electrons are there in Na+ ion.	A. 8 B. 9 C. 1 D. 10
124	During the formation of ionic bond heat is.	A. Remains same B. Absorbed C. Released D. Both a and b
125	Which types of attractive forces are presentin ionic compounds.	A. Covalent bonds B. Electrostatic forces of attraction C. Mtallic bonds D. Coordinate covalent bonds
126	Number of electronsin nitrogen molecule is.	A. 2 B. 4 C. 6 D. 8
127	How many covalent bonds do N2 molecule have	A. 3 B. 4 C. 2 D. 5
128	Silicon belongs to Grou IVA . It has ....electons in the valence shell	A. 2 B. 6 C. 3 D. 4
129	In the formation of AlF_3, aluminum atom loses.....electrons.	A. 1 B. 4 C. 3 D. 2
130	Identify the covalent compund	A. NaCl B. H₂O C. KF D. MgO
131	A bond formed between two non metals is expected to be	A. Ionic B. Coordinate covalent C. Metallic D. Covalent
132	A bond pair is covalent molecules usually has.	A. One electron B. Two electron C. Three electron D. Four electron
		A. Sharing of electrons

133	Covalent Bond involves the	B. Repulsion of electrons C. Acceptance of electrons D. Donation of electrons
134	How many covalent bonds does C_2H_2 molecule have.	A. Two B. Three C. Four D. Five
135	Triple covalent bond involves how many electrons.	A. Six B. Four C. Eight D. Three
136	Identify the compound whcihis not soluble in water	A. KBr B. MgCl2 C. C6H6 D. NaCl
137	Whcih one of the following is the weakest force among the atoms.	A. Intermolecular force B. Ionic force C. Metallic force D. Covelent forces
138	Covalent bond is most commonly found between the elements of group	A. 1 to 13 B. 16 to 18 C. 13 to 17 D. 15 to 18
139	A bond formed by the mutual sharing an electron pair is called.	A. Ionic bond B. Metallic bond C. Covalent bond D. Coordinate covalent bond
140	A covalent bond formed by the mutual sharing of two pairs of electrons between bonded atoms is called.	A. Signle covalent bond B. Double covalent bond C. Triple covalent bond D. Polar covalent bond
141	Which molecule contains a single vovalent bond.	A. CH4 B. C2H4 C. C2H2 D. O2
142	Nitrogen molecule contains.	A. Polar covalent bond B. Triple Covalent bond C. Double covalent bond D. Single covalent bond
143	How many electrons are involved in the formation of signle covalent bond	A. One B. Two C. Three D. Four
144	A covalent bond formed by two similar stoms is known as.	A. Polar Covalent bond B. Metallic bond C. Double covalent bond D. Non-polar covalent bond
145	Dative covalent bondis also known as	A. Covalent bond B. Ionic Bond C. Metallic Bond D. Coordinate covalent bond
146	Whcih one of the following in as electron deficient molecule.	A. NH3 B. O2 C. BF3 D. N2
147	How many lone pairs are present on nitogenin ammonia molecule.	A. One B. Two C. Three D. Four
148	Whcih types of bond is present between NH3 and BF3	A. Covalend Bond B. Ionic Bond C. Co ordinate covalent bond D. Metallic Bond
149	In metals, the hold of nucieus over the valence shell elecrons is veak due to.	A. Highlectron affinity B. Large sized atoms C. High ionization energies D. All of the above
150	Malleability is the property by virtue of which a metal can be drawn into.	A. Rods B. Plates C. Sheets D. Wires

151	Metal have the tendency to lose elecrons due to.	A. High ionization energies B. Low ionization energies C. Low electron affinity D. None of the above
152	Hydrogen bound ing is always found in	A. Non-polar molecules B. Homoatomic molecules C. Polar Molecules D. All of the above
153	Which of the following is an exampel of polar covalent compound.	A. Cl2 B. H2 C. O2 D. HCl
154	The fore of attraction betwene water molecule is.	A. Ionic bonding B. Covalent bonding C. Hydrgen Bonding D. Co ordinate Covalent bonding
155	The boiling point of water is	A. O ^oC B. 100 ^oC C. 35 ^oC D. 25 ^oC
156	The boiling point of alcohal is	A. ₄₄^oC B. 78^oC C. 53^oC D. 19^oC
157	Water has high boiling point as compared a alcohol due to	A. Low density B. High surface tension C. Hydrogen bonding D. High vapour pressure
158	The compund formed by oppoite charges are known as.	A. Metallic solids B. Ionic compounds C. Non-polar Covalent compound D. None of the above
159	Ionic compund are good conductors electricity in	A. Solution B. Molten state C. Solid state D. both a and b
160	Ionic compund have	A. Low melting and high boiling points B. Low melting and boiling point C. High melting and boiling points D. High meilting and low boiling points
161	None polar compounds are insoluble in	A. Alcohol B. Benzene C. Ether D. Water
162	How many atoms are present in one gram of H2O?	A. 1002 x 10²³ atom B. 6.022 x 10²³atom C. 0.334 x 10²³atom D. 2.004 x 10²³atom
163	Which is the correct formula of calcium phosphide.	A. CaP B. Ca₃P₂ C. CaP₂ D. Ca₂P₃
164	How many atoic mass units (amu) are there in one gram.	A. 1 amu B. 6.022 x 10²³ C. 10 amu D. 6.022 x 10²²
165	How many moles are there in 25 g of H_2SO_4?	A. 0.765 Moles B. 0.255 moles C. 0.4 moles D. 0.51 moles
166	A nceklaee has 6 g of diamonds in it . What are the numberof carbon atoms in it?	A. 3.01 x 10²³ B. 1.003 x 10²³ C. 12.04 x 10²³ D. 6.02 x 10²³
167	What is the mass of Al in 204 g of aluminium oxide Al_2O_3	A. 26 g B. 54 g C. 108 g D. 27 g
168	Which one of the following compounds will hae the highest percentage of the mass of nitrogen?	A. N2H4 B. CO (NH2)2 C. NH3 D. NH2OH

169	When one mole of each of the following compounds is reacted with oxygen, which wil produce the maximum amount CO2?	A. Carbon B. Ethane C. Diamond D. Methane
170	What mass of 95% CaCO3 will be requred to neutralize 50 cm 3 of 0.5 M HCl solution.	A. 9.5 g B. 1.45 g C. 1.32 g D. 1.25 g
171	Formula of Ozone.	A. O₂ B. S₈ C. CO₂ D. O₃
172	Avogadrao was a scientist	A. Italian B. Greek C. Garman D. African
173	Which of the following is insoluble salt	A. KCl B. AgCl C. NaCl D. CaCl2
174	Stochiometric calculators are used to prepare.	A. Soaps B. Shampo C. Perfumes D. All of these
175	Without stoichiometry which industry cannot exist.	A. Meta B. Petroleium C. Leather D. Chemical
176	Whcih law is obeyed in chemicla calculations?	A. Law of mass action B. Law of conservationof mass C. Law of definite proportion D. Both a and b
177	Empricial formula of sand is.	A. SiO2 B. SiO3 C. SiO4 D. SiO
178	Emperical Formula of Glucose is	A. CH2O B. CHO C. CHO2 D. C2HO
179	Empirical formula of acetic acid (CH3COOH) is	A. CHO B. CH C. CH2O D. None of these
180	Whof the following represent sand?	A. NaCl B. CaCO3 C. H2O D. CH2O
181	Empirical formula of hydrogen peroxide.	A. HO B. CO C. CHO D. CH
182	Empirical formula of Benzene is.	A. CH2O B. CH C. C2H D. CH2
183	Which compund has same molecular and emiprical formula.	A. C6H12 O6 B. H2O2 C. H2O D. C6H6
184	Value of Avagadro's number is.	A. 6.6 x 10⁻²⁰ B. 6.02 x 10²³ C. 6.00 x 10²⁴ D. 1.32 x 10²³
185	1 gram formula of NaCl contan is grams.	A. 100 g B. 58.5 g C. 32 g D. 49 g
186	1- gram atom of carbon contain hwo many moles	A. 1 mole B. 2 mole

	1 gram atom of carbon contain how many moles.	C. 6 moles D. 12 moles
187	Formula of common salt is	A. NaCl B. AgCl C. LiCl D. KCl
188	Formula mass of K2SO4 is.	A. 174 amu B. 110 amu C. 180 amu D. 145 amu
189	Molecular mass of Acetic Acid	A. 70 amu B. 43 amu C. 60 amu D. 80 amu
190	Numbr of hydrogen atoms present in 18 g of water.	A. 2 x NA B. N_A C. 3 x N_A D. 1/2 N_A
191	The mass number of sodium is.	A. 19 B. 31 C. 27 D. 23
192	Limes is another name of	A. Sodium hydroxide B. Sodium Carbonate C. Calcium Carbonate D. Silicon dioxide
193	Mass of 3 moles of oxygen atoms is.	A. 64 g B. 16 g C. 32 g D. 48 g
194	Number of moels in 29.25 g NaCl is.	A. 0.50 B. 0.25 C. 0.21 D. 0.75
195	How many atom of carbon are present in one molecule of glucose.	A. 11 B. 22 C. 12 D. 6
196	40 g of H_3PO_4 contains numebr of moles	A. 0.58 g B. 4.8 g C. 5.8 g D. 0.408 g
197	A compound with chemicla formula Na_2CX_3 has formula mass 106 ami. Atomicmass of the elemetn X is.	A. 16 B. 23 C. 12 D. 106
198	How many moels of molecules are there in 16 g oxygen.	A. 0.05 B. 0.1 C. 0.5 D. 1
199	What is the mass of 4 moles of hydrogen gas.	A. 1 g B. 1.008 g C. 8.064 g D. 4.032 g
200	Whcih term is the same for one mole of oxgen and one mole of water.	A. atoms B. mass C. Molecules D. Volume
201	If one mole of carbon contains x atoms what is the number of atoms contained in 12 g of Mg.	A. 1.5 x B. 0.5 x C. x D. 2x
202	The mass of one molecule of water is.	A. 18 g B. 18 mg C. 18 kg D. 18 amu
203	The molecular mass of H2SO4 is.	A. 9.8 g B. 98 amu C. 98 g D. 9.8 amu
		A. 0.18

204	How many number of moles are equivalent to 8 gram of CO_2.	B. 0.15 C. 0.24 D. 0.21
205	When old bonds are broken, the energy is.	A. Release B. Remain same C. Consume D. None of these
206	When new bonds ae formed, the energy is	A. Consume B. Remain same C. Release D. None of these
207	When NaOH and HCl are mixed the temperature increases. The reaction	A. Exothermic with a negative enthalpy chagne. B. Endothermic with a positive enthaly change. C. Endothermic with a negatie enthalpy change D. Exothermic with a positive enthealpy change
208	All chemical reaction involves.	A. Enzymes B. Catalyst C. Energy changes D. All of these
209	Who use the word energy for the 1st time	A. Rutherford B. Bohr C. Thomas Young D. None of these
210	The word energy is used in physics ofr the firt time.	A. 1902 B. 1858 C. 1805 D. 1802
211	The part of the universe that we want to focus our attention called.	A. Surrounding B. Energy C. System D. Both a and b
212	The enthalpy of reaction $C+O_2$ ------ CO_2	A. -571.6 kJ B. -393.5 kJ C. +53.8 kJ D. -110.5 kJ
213	The enthalpy of reaction $2H_2 + O_2$-----$2H_2O$	A. -571.6 kJ B. -110.5 kJ C. -393.5 kJ D. +53.8 kJ
214	The enthalpy of reaction H_2+I_2--- 2HI	A. -571.6 k J B. +53.8 kJ C. 11 kJ D. -393.5 kJ
215	If the Delta H value is negative then reaction witll be	A. Endotermic B. Exothermic C. May or may not be exothermic or endothermic D. None of these
216	Bond formation energy of one O-H bond is.....	A. 488 kJ/mol B. 484 kJ/mol C. 486 kJ/mol D. 489 kJ/mol
217	Bond dissocialation for H_2 is	A. 435 kJ/mol B. 440 kJ/mol C. 430 kJ/mol D. 445 kJ/mol
218	Bond dissociation for O_2 is	A. 505 kJ/mol B. 705 kJ/mol C. 605 kJ/mol D. 498 kJ/mol
219	Formation of NO is	A. Exothrmic B. Endothermic C. No Heat Change D. None of these
220	Activation energy of a chemical reaction must be..... the everage kinetic energy of reacting molecules.	A. Equal to B. Grether than C. Lower than D. None of these

221	No reaction occurs if the energy of reacting particles.....activation energy.	A. Lower than B. Greather than C. Nearest to D. Equal to
222	Washing clothes at 140 ^{O}F uses almost the energy as at 140 ^{O}F wash	A. Half B. Thrice C. Twice D. None of the above
223	---------- of the energy used by traditional electric bulb is wasted in producing heat.	A. 60% B. 50% C. 70% D. 90%
224	Which is not produced in an aerobic respiration.	A. Carbon dioxide B. Lactic acid C. Water D. Energy
225	----------acts are reserve energy sources.	A. Enzymes B. Vitamins C. Proteins D. Lipids
226	Which is released in anacrobic respiration.	A. Stearic acid B. Citric acid C. Lactic acid D. Amino Acid
227	Aerobic respiration releases.....energy than anaerobic respiration.	A. Equal B. Less C. More D. None of these
228	---------- acts a catalyst promoting the breakdown of ozone.	A. I2 B. Br2 C. Cl2 D. None
229	During the glycolysis net ATP produced are.	A. 2 B. 4 C. 6 D. 8
230	What will happen if the rates of forward and reverse reactions are very high	A. The reaction will be practiclly irreversible B. The equilibrium point will reach very soon C. The equilibiurm point will reach very late D. The reaction will not attain the state of dynamic equilibrium
231	Predict which components of the amosphere react in the presence of lightening.	A. N2 and H2O B. O2 and H2O C. N2 and O2 D. CO2 and O2
232	An Inorganic chemistry places one mle of PCl5 in container A and one mole of each Cl2 and PCl3 in container B. Both the containers were sealed and heated to the same temperture to reach the stte of equilibrium Guess about the composition of mixtures in both the containers.	A. Both the containers wil have zero concentration of its reactants. B. Both the containers wil hae the same composition of mixtures C. Container A will have more concentraion of PCl3 than B. D. Container A will have less concentraion of PCl3 than B.
233	CaO or lime is used extensive in steel, glass and paper industries. It is produced inan exothermic reversitble reaction by the decompositon of lie . Choose the conditions to produce maximum amount of lime.	A. Heating at high temperatur ein an open vessel B. Heating at high temperatur ein a closed vessel C. Cooling it in a closed vessel D. Colling it in an open vessel
234	What conditon Should be met for the reversible reactio to achieve the state of equilibrium.	A. The concentratio of all the reactants and the prodduct should become constant B. all the reactants should be converted into the product C. 50% of the reactant shuld be converted into prodcuts. D. One of the product should be removed from the reaction mixture.
		A. Because gteh solubiliyt of the gas increases B. Because the gas is dissolved under pressure hence it comes out when

235	Why the gas starts coming out when you open a can of fizzy drink.	pressur eis decreased C. Becaue the gas is insoluble in water D. Because the solubility of hte gas decreases at high pressure.
236	In an irreversibel reaction equiirbrium is.	A. The forward reaction will be fovoured B. No effect on forward or backward reaction C. No effect on bakcwars reaction D. The backward reaction will be favoured
237	When a reaction will be come a reversible one?	A. If the actvation eneryg of the forward reaction is comparrble to that of backward reaction B. If the activation energy of the forward reaction is higher than that of backward reaction C. If the activationenergy of the forward reaction is lower than that of backward reaction D. If the enthalpy change of both the reactions is zero.
238	If reversibe reaction useful for preparing compounds on large scale.	A. Yes B. No C. They are useful only when equilirbium lies far to the left side D. They are useful only when equilibrium lies far to the right side
239	What will hapen to the concentrations of the prodcut if a reversible reaction at eqiulibrium is not distrubed.	A. They will keepon increasing B. They will keep on decreasing C. They will remain constant D. They willremain constant for some time and then start decreasing
240	In an irrversibel reaction equilibrium	A. Never establised B. Established quickly C. Established slowly D. Established when reaction stops
241	The characteristics of reversible reactions are the following except.	A. Product nver recombine to form reactants B. They never complete C. They have a double arrow between reactants and products D. The proceed in both ways
242	A reverse reaction is one that	A. Speeds up gradually B. Proceeds from left to right C. In which reactants react to form products D. Slow down gradually
243	The reaction in which the products do not reombine to form reatants are called	A. Addition reactions B. Decomposition reactions C. Irreversitble reactions D. reversible reactions
244	The reaction in which the products can recombine to formreactants are called.	A. Reversible Reaction B. Irreversible reactions C. Decomposition reactions D. Addition reactions
245	Which type of reactions speed up gradully?	A. Decomposition reaction B. Forward reaction C. Reverse reactions D. Irreversibel reactions
246	Such reaction which continue in both directins are called.	A. Dynamic B. Irreversible C. Reversible D. Non- reactive
247	In chemical reaction, the substances that combine are called.	A. Masses B. Materials C. Products D. Reactants
248	The forward reaction takes place from	A. Right to left B. Left to right C. Both a and b D. None of these
		A. Only 10% reactants covert into products B. All the reactants covrt into products

249	A complete reaction is in which	B. All the reactants covrt into products C. All the reactants do no covrt into products D. Half reactants covert into produts
250	In the beginning the rate of reverse reaction is.	A. Slow B. very fast C. Moderate D. Negligible
251	The new substance formed in a chemcial reaction is.	A. Reverse B. Reactant C. Forward D. Product
252	The colour of anhydrous copper (II)sulphate solid is	A. Pink B. Black C. White D. Blue
253	The colour of hydrated copper (II) sulphate solid is.	A. Black B. Pink C. White D. Blue
254	The colour of anhydrous cobalt(II) cholride solid	A. White B. Black C. Pink D. Blue
255	The colour of hydrated cobalt(II) chloride solid is	A. White B. Black C. Blue D. Pink
256	Which of the following does not happpen, when a system is at equilibrium state.	A. Reaction continues to occur in both the directions B. Concentration of reactants and products stop changig C. Forward and reverse reactions stop D. Forward and reverse rates become equal
257	Whcih is true about the equilibrium state?	A. The forard reaction stops B. Both forward and reverse reactions stop C. Both foward and reverse reactions continue at the same rate D. The reverse reaction stops
258	When system is at quilibrium state.	A. The rate of the forward and reverse rections become equal B. The concentrationof reatants and product becoes equal C. The oposing reactions stop D. The rate of the reverse reactio becomes very low
259	When the rate of the forward reaction takes place at the rate of reverse reaction the composition of the reaction mixgure remains consant. It is called.	A. Chemical Equilibrium B. Static equilirbirum C. Both a and b D. None of the above
260	Concentratin of reactants and product at equilibrium remains unchanged if	A. Concentration of any reactant or product is not changed B. Temperaure of the reation is not changed C. Pressure or volume of the system is not changed D. All of the above are observed
261	At what temperature , rate of ammonia formation and decomposition is the highest.	A. 200 ^oC B. 300^oC C. 400 ^oC D. 500 ^oC
262	Industrialy, ammonia is produced by which process.	A. Halogenation B. Solvay process C. Haber Process D. Hydrogenation
263	Formation of ammonia from Nitrogen and hydrogen is an.	A. Exothermic reaction B. Endothermic reaction C. Both a and b D. No heat change
264	How much heat absorbed when NH3decomposed into N2 and H22	A. 90.4 kJ/mol B. 92.4 kJ/mol

264	How much heat absorbed when NH3decomposed into N2 and H2?	C. 94.2 kJ/mol D. 95.2 kJ/mol
265	Which compound is used a thinner is paint industry?	A. H2O B. C2H3OH C. CH3COOC2H5 D. CH3COOH
266	Which acid is not used as a food or mixed with food?	A. Tartaric Acid B. Formic Acid C. Ascorbic acid D. Citric Acid
267	While baking which gas is responsible for raising the bread and making it soft?	A. Oxygen B. Carbon monoxide C. Carbon di oxide D. Nitrogen
268	Predict themain characteristics of the reactions of metals with acids.	A. Metals are dissolved B. Hydrogen gas is envolved C. Metals are coverted into salts D. All the above mentioned characteristics are true
269	How many hydroxide ions, calium hydroxide will release in water.	A. 2 B. 1 C. 3 D. zero
270	In a neutralization reaction between KOH and H3PO4how many molecules of KOH will react with one molecule of H3PO4	A. 1 B. 2 C. 3 D. 4
271	Which acid is used in he preparation of soaf.	A. Oxalic Acid B. Citric Acid C. Stearic Acid D. Tartaric Acid
272	Which compound formed when SO2 is dissolved in water	A. H2SO3 B. SO3 C. H2SO D. H2S2O7
273	Which of the following contains oxalic acid	A. Orange B. Sour Milk C. Tomato D. Tamarind
274	When a chemical reaction is carried out with a substancZ; a gas is produced which turns red litmus paper blue. What is the reaction?	A. Reaction of an acid with a meal carbonate B. Reactin of an alkali with ammonium salt C. Reaction of an acid with ammonium salt D. None of these
275	A base is a substance which neutralizes an acid. Which of these substances is not aa base.	A. Aqueious ammonia B. Calcium oxide C. sodium Chloride D. Sodium carbonate
276	Acidic Acid is used for	A. Cleaning metal B. Etching designs C. Flavouring food D. Making explosives
277	Acids means	A. Bitter B. Salty C. Sweet D. Sour
278	All Acids turn blue litmus.	A. Pink B. White C. Colourless D. Red
279	All bases turn red litmus.	A. Blue B. White C. Pink D. Colourless
280	Which of the following cannot be classified as Arrhenius acid	A. CO2 B. HNO3 C. H2CO3 D. H2SO4
	Milk of magnesia contains Mg(OH)2.It is used as antacid.It neutralizes excess stomachacid	A. MgCl2 B. MgSO4

281	Milk of magnesia contains Mg(OH)2 it is used as antacid it neutralizes excess stomachaacid. Which salt is formed in this reaction.	B. MgSO4 C. MgCO3 D. MgO
282	According to Arrhenius concept acid is a substance which dissociates in aqueious solutin to give.	A. Proton B. Pair of Electron C. Hydrogen Ions D. Hydroxide ion
283	According to Arrhenius concept base is a substance which dissociates in aqueios solution to give	A. Hydroxide ions B. Hydrogen ions C. Pair of Electrons D. Proton
284	Which one is not an Arrhenius base.	A. KOH B. NaOH C. NH3 D. Ca(OH)2
285	Which one is not an Arrhenious acid?	A. HCl B. H2SO4 C. CO2 D. HNO3
286	Which of the following is Bronsted base?	A. HCl B. CH3COOH C. H2O D. NH3
287	Ammonia is a base, because it	A. Ionizes in water to give OH- Ions B. Can accept proton C. Contains OH group D. Can accept an election pair
288	According to Bronsted and Lowry concept, an acid is a substance that can donate.	A. Proton B. Electron pair C. Neutron D. Electron
289	A substance whcih can behave as an acid as well as a base is called	A. Amphoteric specie B. Acid C. Base D. Neutral specie
290	A reaction between an acid and a base produces	A. Salt and an acid B. Sald and base C. Salt and water D. Salt and gas
291	Whcih acid is present in our stomach.	A. Nitirc Acid B. Hydrochloric acid C. Sulphuric acid D. All of the above
292	When acids react with metals which gas is evolved?	A. O2 B. N2 C. Cl2 D. H2
293	When acids react with metal with carbonates and bicarbonates which gas is produced.	A. N2 B. H2 C. Cl2 D. CO2
294	Alkalis react with ammonium salts to librate.	A. CO2 B. SO2 C. H2 D. NH3
295	Which is used to manufacture soap?	A. NH4OH B. Ca(OH)2 C. NaOH D. Mg(OH)2
296	Acid rain has pH less than.	A. 6 B. 7 C. 7.4 D. 5.6
297	Caustic chemical drain cleaner are captable of dissolving.	A. Hair B. Food C. Grease D. All of these
298	In which period and group yu will place the elemnt whic is an important part of the solar cell?	A. Third period and Sixth a grop Group 16 B. Third priod and forth A group Group 14 C. Second period and forth A group

		Group D. Third prod and fifth A group Group 15
299	Which is the softtest metal.	A. Zn B. Ca C. Na D. Al
300	A yellow solid element exists in allotropic forms whic is also present in fossil fuel. Indicate the name	A. Iodine B. Carbon C. Sulphur D. Aluminium
301	How many electrons can nitrogen accept in its outermost shell.	A. 2 B. 3 C. 4 D. 5
302	Which element is the most reactive element?	A. Florine B. Oxygen C. Chlorine D. Nitrogen
303	Which element has the highest melting point.	A. K B. Cs C. Na D. Rb
304	The element having less value of ionizatin energy and less value of electron affinity is likely to belong to.	A. Group1 B. Group 13 C. Group 16 D. Group 17
305	When we mvoe form left to right in a period, atomic size.	A. Increases B. Decreases C. First increases then decreased D. None of the above
306	Number of peiod in the periodic table are.	A. 7 B. 8 C. 5 D. 16
307	Which of the following grops contain alkaline earth metals.	A. I A B. II A C. VII A D. VIII A
308	Which of the following element belong to VIII A.	A. Xe B. Mg C. Br D. Na
309	Main group elements are arranged in .....groups.	A. 7 B. 6 C. 8 D. 10
310	Period nuber of $^{27}Al_{13}$ is	A. 1 B. 2 C. 3 D. 4
311	All the elements of Group II A are less reactive than alkali metals. This is because these elements have.	A. Decreased nuclear charge B. Similar electronci configuration C. High ionization energies D. Relatively greatr atomic size.
312	The atomic radii of the elemtns in periodic table.	A. Increase from left to right in a period B. Do not chage from left to right in a period C. Increase from top to bottom in a group D. Decrease from top to bottom in a group
313	4th and 5th priod of the long form of periodic table are called.	A. Short periods B. Normal periods C. Very long peiods D. Long periods
314	Which one of the following halongesn has lowest electronetivity	A. Iodine B. Chlorine C. Fluorine D. Bromine
		A. All gases

315	Transition elements are	B. All non metals C. All Metals D. All metalloids
316	How many groups are present in the modern periodic table.	A. 8 B. 10 C. 15 D. 18
317	How many periods are present in the modern periodic table	A. 7 B. 8 C. 10 D. 12
318	How many periods are presnet in the modern periodic table.	A. 7 B. 8 C. 10 D. 12
319	How many elements are present in 1st period.	A. 1 B. 2 C. 8 D. 18
320	How many elements are prsent in each 2nd and 3rd period.	A. 2 B. 32 C. 18 D. 8
321	How many elements are present in each 4th and 5th period.	A. 2 B. 8 C. 32 D. 18
322	How many elements are present in 6th period.	A. 2 B. 8 C. 18 D. 32
323	How many elements are present in 7th period.	A. 2 B. 8 C. 18 D. 23
324	How many blocks are presnt in modern periodic table	A. 2 B. 3 C. 4 D. 5
325	Elments re classified into four blocks depending upon	A. Shell B. Atomic mass C. Sub -Shell D. Atomic Number
326	The elementss of group 1 and 2 are placed in which block	A. s B. p C. d D. f
327	Which of the following elemens is presnet in 1st period.	A. Hydrogen B. Helium C. Both a and b D. None of these
328	Second and third periods are called	A. 1st transition series B. Normal periods C. 2nd transiion series D. 3rd transiliton serios
329	Whcih element is presnet in 2nd period.	A. Lithium B. Beryllium C. Boron D. All of these
330	Elements with atomic no .58 to 71 are called.	A. Actinides B. Lanthanides C. Both a and b D. None of these
331	Actinidws belong to period.	A. 4th B. 5th C. 6th D. 7th
332	Lanthanide series starts after the elemetn	A. Osmium B. Actinium C. Lanthanum D. None of these

333	Atomic number of lanthanum is	A. 57 B. 58 C. 59 D. 60
334	Actinide series starts after the element	A. Actinium B. Lanthanum C. Osmum D. Silver
335	Atomic number of actinium is	A. 57 B. 60 C. 89 D. 80
336	Group nuebr tells about the	A. Number of shells B. Number of valence electrons C. Both a and b D. None of these
337	Period nuebr tells abou the	A. No. of valance electrons B. No. of electronic shells C. Both a and b D. None of the above
338	Whcih period of the modern periodic table is considered as incomplete period.	A. 4th B. 5th C. 6th D. 7th
339	Whic period of the moden periodic table is condidered as incommplete period.	A. 5th B. 4th C. 7th D. 6th
340	/which of the followign elements is presnt in group IA.	A. Lithium B. Hydrogen C. Sodium D. All of these
341	Elements of Group1 are called.	A. Alkali Metlas B. Alkali earth metals C. Transition metals D. Halogen
342	How many elecntrons are present in the valence shell of group 1 elements.	A. 1 B. 2 C. 3 D. 4
343	17th group elements ae known as	A. Alkaki metals B. Alkaline earth metals C. Noble gases D. Halogens
344	17th Group of elements contain electrons in their outer most shell	A. 4 B. 5 C. 7 D. 6
345	The elements of group 3 to 12 are clled.	A. Normal elements B. Halogens C. Noble gases D. Transition elements
346	All transition elements belong to	A. s and p block B. d- block C. f-block D. d and f block
347	The vertical columns present in the priodic table are called.	A. Group B. Period C. Both a and b D. None of these
348	The horizontal lines present in the priodic table are called.	A. Groups B. Periods C. Both a and b D. None of these
349	With the increase of atomic numebr , the number of electron in an atom also.	A. Decreases B. First increases then decreases C. Increases D. None of the above
350	Elements o group 13 to 18 have thier valence electrons is subshell	A. s B. p

350	Elements o group 13 to 18 have thier valence electrons is subshell	C. f D. d
351	Which is strongest oxidizing agent.	A. Cholorine B. Iodine C. Fluorine D. Bromine
352	Which halogen memebr exists in a liquid stte at room temperature	A. Bromine B. Chlorine C. Fluroine D. Iodine
353	Elements of a period show properties.	A. Same B. Different C. Both a and b D. None of these
354	The elements of a group show properties.	A. Same B. Different C. Both a and b D. None of these
355	The amoutn of energy given out when an electron is added to an eatom is called.	A. Electron affinity B. Lattice energy C. Ionization energy D. Electronegativity
356	Aong the period which one of the followig decreases.	A. Electronegativity B. Ionization energy C. Atomic radius D. Electron affinity
357	Mark the icorrect statement about ionization energy.	A. It is measured in kJmol-1 B. It is absorption of energy C. It decreased in a period D. It decrese in a group
358	Point out the incorrect statement about electron affinity	A. It decreases in a period B. It decreases in a group C. It is measure din kJmol-1 D. None of these
359	Unit of atomic size is	A. pm B. nm C. kJmol-1 D. Both a and b
360	The distance between the neclei of two carbon atoms in its elemetnal from is	A. 150 pm B. 152 pm C. 154 pm D. 156 pm
361	When we move form left to right in a period, atomic number	A. Decreases B. Increases C. First increases then decreases D. None of the above
362	When we mvoe from top to bottom in a group atomic size.	A. Decrease B. Increases C. First increases then decreases D. None of the above
363	The minium amont of energy whcih is required to remvoe an electron from valence shell of the gaseous state of an atom is called.	A. Potential energy B. Ionization energy C. Electron affinity D. Electronegativity
364	The unit of ionization energy is	A. nm and pm B. kJ mol-1 C. Pascal D. Newton
365	When we mvoe top to bottom in group, ionization energy.	A. Increases B. No effet C. Decreases D. None of these
366	When we mvoe from left to right in a period, ionization energy.	A. No effect B. Decreases C. Increases D. None of these
367	Uni of electron affinity is.	A. k J mol-1 B. k jmol C. pm D. Newton
		A. -328

368	Electron affinity of fluorine in kJmol-1 is	B. 328 C. -330 D. -340
369	The ability of an atom to attract the shared pair of electons towards itself in a molecule is called	A. Ionization energy B. Electronietativity C. Shielding effect D. Electron affinity
370	Which one of the followign halogns has highest electronegativity	A. Iodine B. Fluorine C. Chlorine D. Bromine
371	Electronegativity of oxygen is.	A. 3,1 B. 3,3 C. 3.4 D. 3,2
372	The electronegativity of carbon is	A. 2.5 B. 2.0 C. 1.0 D. 4.0
373	Metals can form ions carrying carges.	A. Uni positive B. Di positive C. Tri postive D. All of these
374	Pure alkalis metals can be cut siply by knife but iron cannot bccause of alkali meals have	A. Non metalic bonding B. Strong metallic bonding C. Weak metallic bonding D. Moderate metqllic bonding
375	Metals lose their elecrons easily because.	A. They are elecrnegativity B. They have electron affinity C. They are electropositive D. Good conductors of heat
376	Metals are teh elements which have.	A. Electronegative character B. Electropositive character C. Both a and b D. None of the above
377	Which are good conductor of heat and electricity	A. Metals B. Non metals C. Metalloids D. All of these
378	All metals bear	A. Positive charge B. Negative charge C. Both a and b D. None of these
379	Metals posses.	A. Ionic bond B. Covalent bond C. Co-ordinate covalent D. Metallic bond
380	Sodium metal has electrons	A. 10 B. 11 C. 12 D. 14
381	Which group elements has low ionization energies.	A. Halogens B. Noble gases C. Alkaline Earth Metals D. Alkali Metals
382	Platinum alloyed with which metal is used as catalyst in automobiles as atalytic covertor.	A. Gold B. Rhodium C. Palladium D. Both a and b
383	Which of the following is a metal	A. Magnesium B. Carbon C. Hydrogen D. Nitrogen
384	The haviest metal is	A. Iron B. Lead C. Osmium D. Platinum