Chemistry Fsc Part 2 Online Test

Sr	Questions	Answers Choice
1	Who gave the law of Triads in 1829?	A. Dobereiner B. Mosely C. Newland D. Mendeleev
2	The concept of atomic number was introduced by	A. Alrazi B. Mendeleeve C. Moseley D. Dobereiner
3	The basis of modern periodic table is	A. Electron affinity B. Atomic mass C. Ionization Potential D. Atomic number
4	Elements of Groups IIA are called	A. Alkali metals B. Alkaline earth metals C. Coinage metals D. Halogens
5	In modern periodic table VI period contains elements	A. 8 B. 18 C. 10 D. 32
6	Which of the following are alkaline earth metals?	A. Be, Mg, Ca B. Li, Na, K C. Fe, CO, Ni D. B, Al, Ga
7	Which one is an incomplete period	A. 4th B. 5th C. 6th D. 7th
8	Number of elements in the first period of the periodic table is	A. 2 B. 8 C. 14 D. 18
9	Which is the longest periodic table	A. 4 B. 5 C. 6 D. 7
10	Which of the following statement is correct	A. Na atom is smaller than Na⁺ B. Na atom is larger than K atom C. F atom is smaller than F⁻ D. F atom is larger than F⁻
11	Which of the following element has lowest ionization energy	A. Beryllium B. Boron C. Carbon D. Oxygen
12	Which element has lowest melting point	A. Beryllium B. Magnesium C. Calcium D. Barium
13	Which the correct statement	A. Cl⁻is smaller than Cl atom B. Cl⁻(Ion) and Cl (atom) are equal in size C. Na⁺is smaller than Na atom D. Na⁺is larger than Na atom
14	Correct order according to atomic size in the following is	A. Na > K B. Be > Mg C. O > N D. Cl > F
15	Which of the following has highest M.P	A. Aluminium B. Silicon C. Phosphorus D. Sulphur

16	Which of the following the highest hydration energy	A. Li⁺ B. Na⁺ C. K⁺ D. Mg⁺⁺
17	The word alkali is derived from which language	A. Arabic B. Greek C. French D. German
18	Which one does not belong to the alkaline earth metals	A. Be B. Ba C. Ra D. Rn
19	Which one of the following is not an alkali metal	A. Francium B. Caesuym C. Rubidium D. Radium
20	Chile saltpeter has the chemical formal	A. NaNO₃ B. KNO₂ C. Na₂B₄O₇ D. Na₂CO₃.H₂O
21	---------- does not belong to Alkaline-Earth metals	A. Be B. Ra C. Ba D. Rn
22	Dolomite is a carbonate of	A. Be B. Mg C. Na D. Ba
23	The mineral $CaSO_4.2H_2O$ has the general name	A. Dolomite B. Gypsum C. Calcite D. Epsom Salt
24	Which is the least reactive of all the alkali metals	A. Li B. Na C. K D. Cs
25	Point out the element which forms super oxide	A. Li B. Na C. K D. C
26	Which metal oxide is insoluble in water	A. MgO B. CaO C. SrO D. BaO
27	The most metallic elements from the following is	A. Arsenic B. Oxygen C. Antimony D. Bismuth
28	The Milk of magnesia is used for the treatment of	A. Basicity B. Rancidity C. Acidity D. Jaundance
29	The oxides of beryllium are	A. Basicity B. Rancidity C. Amphoteric D. none of all
30	Which of the following gas will turn lime water milky	A. Cl₂ B. NO₂ C. CO D. CO₂
31	Nelson's cell is used to prepare	A. NaOH B. Na₂CO₃ C. Na metal D. NaCl
32	Which elements are deposited at the cathode during the electrolysis of brine in diaphragm cell	A. H₂ B. Ba C. Ra D. Rn
33	The chief of aluminum is	A. Na₃AlF₆ B. Al₂O₃.2H₂O C. Al₂O₃

		D. Al₂O₃.H₂O
34	Tincal is a mineral of	A. Al B. C C. Si D. B
35	Kaolin is a mineral of	A. Carbon B. Magnesium C. Silicon D. Aluminium
36	The aqueous solution of Borax	A. Acidic B. alkaline C. Amphoteric D. manual
37	Which is used in the leather industry	A. Borax B. Boric acid C. Boric oxide D. Tetra Boric acid
38	Boric acid cannot be used	A. As antiseptic in medicine B. For washing eyes C. In soda bottles D. For Enamals and Glazes
39	Aluminum reacts with nitrogen to form	A. AlN B. Al₂N C. Al₂N₃ D. Al₄N₆
40	Aluminium oxide is	A. Acidic oxide B. Baric oxide C. Amphoteric oxide D. None of these
41	---------- element forms an ion with charge 3^+	A. Beryllium B. Aluminium C. Carbon D. Silicon
42	---------- metal is used in the Thermite process because of its reactivity	A. Tron B. Copper C. Aluminium D. Zinc
43	Which one of following is used in cosmetics	A. Talc B. Asbestos C. Sodium sulphate D. Aluminium Sulphate
44	The element of 2nd period, which has highest ionization energy from the following is	A. Be B. C C. N D. O
45	Which one of the following does not belong to alkaline earth metals	A. Be B. Ra C. Ba D. Rn
46	Aluminum oxide is	A. acidic oxide B. basic oxide C. amphoteric oxide D. none of these
47	Chemical composition of colemanite is	A. Ca₂B₆O₁₁. 5H₂O B. CaB₄O₇. 4H₂O C. Na₂B₄O₇.4H₂O D. CaNaB₅O₉. 8H₂O
48	Which element form an ion with charge +3	A. Beryllium B. Aluminum C. Carbon D. Sillicon
49	Which electronic configuration corresponds to an element of group III-A of the Periodic Table	A. 1s²,2s²,2p⁶,3s²,3p¹ B. 1s²,2s²,2p⁶,3s²,3p⁶,4s⁶ C. 1s²,2s²,2p⁶ D. 1s²,2s²,3s²,3p³
50	Which of the following elements is not present abundantly in earth's crust	A. Silicon B. Aluminium C. Sodium D. Oxygen

51	The most electronegative element of group V-A is	A. N B. P C. Sb D. Bi
52	Out of all the elements of group VA, the highest ionization energy is possessed by	A. N B. As C. Sb D. Bi
53	Laughing gas is chemically	A. NO B. NO₂ C. N₂O D. N₄O₄
54	The brown gas formed, when metal reduce HNO_3	A. NO B. NO₂ C. N₂O₃ D. N₂O₅
55	The oxidation of NO in air produces	A. N₂O₃ B. NO₂ C. N₂O₃ D. N₂O₄
56	Which of the following is a reddish brown gas	A. N₂O₃ B. NO₂ C. N₂O₃ D. N₂O₅
57	Which of the following gives acidic oxide	A. N B. As C. Sb D. Bi
58	Which metal is redered passive by HNO_3due to formation of a film of metal oxide over the metal	A. Pt B. Sn C. CO D. Mn
59	Gold dissolves in "Aqua Regia" due to formation of Halide. Point out correct halide	A. AuF₃ B. AuCl₃ C. AuBr₃ D. AuI₃
60	What is %age of calcium phosphate in bone ash	A. 20 B. 40 C. 80 D. 60
61	Maximum number of unpaired electrons is in	A. O₂ B. O₂⁺ C. O₂⁻ D. O₂²⁻
62	Which catalyst is used in contact process	A. Fe₂O₃ B. V₂O₅ C. SO₃ D. Ag₂O
63	Out of all the elements of Group V-A the highest ionization energy is possessed by	A. N B. P C. Sb D. Bi
64	In group V-A elements the most electronegative elements is	A. Sb B. N C. P D. As
65	Oxidation of NO in air produces	A. N₂O B. N₂O₃ C. N₂O₄ D. N₂O₅
66	The brown gas formed when metal reduces HNO_3	A. N₂O₅ B. N₂O₃ C. NO₂ D. NO
67	Out of all the elements of groups VI-A the highest melting and boiling points is shown by the element	A. Te B. Se C. S D. Po
68	Which one of halogens is a liquid	A. F₂ B. Cl₂ C. Br₂

		D. I₂
69	Which is the second most abundant element in the universe	A. H B. He C. CO D. C
70	Which one is perchloric acid	A. HClO B. HClO C. HClO₃ D. HClO₄
71	Hydrogen bond is the strongest between the molecules of	A. HF B. HCl C. HBr D. HI
72	Which halogen will react spontaneously with Au(s) to produce Au^{3+}	A. Br₂ B. F₂ C. I₂ D. Cl₂
73	The anhydride of $HClO_4$is	A. ClO₃ B. ClO₂ C. Cl₂O₅ D. Cl₂O₇
74	Which is the strongest acid	A. HClO B. HClO₂ C. HClO₃ D. HClO₄
75	Which one is chlorous acid	A. HClO B. HClO₂ C. HClO₃ D. HClO₄
76	_____ is use as a cooling medium for nuclear reactors	A. Ne B. He C. Ar D. Kr
77	Which of the following noble gas is used for arc welding and cutting	A. Helium B. Argon C. Xenon D. Radon
78	Which of the following hydrogen halide is the weakest acid in solution	A. HF B. HBr C. HI D. HCl
79	Chlorine heptoxide (Cl_2O_7) reacts with water to form	A. Hypochlorous acid B. Chloric acid C. Perchloric acid D. Chlorine and oxygen
80	Which halogen will react spontaneously with $Au_{(s)}$to produce Au^{3+}	A. Br₂ B. F₂ C. I₂ D. Cl₂
81	The anhydride of $HClO_4$is	A. ClO B. ClO₂ C. ClO₃ D. Cl₂O₇
82	Bleaching powder may be produced by passing chlorine over	A. calcium carbonate B. hydrated calcium sulphate C. calcium hydroxide D. magnesium hydroxide
83	Which halogen occurs naturally in a positive oxidation state	A. Fluorine B. Chlorine C. Bromine D. Iodine
84	An element that has high ionization energy and tends to be chemically inactive would most likely to be	A. an alkali metal B. a transition element C. a noble gas D. a halogen
85	The total number of transition element is	A. 10 B. 14 C. 40 D. 58
	Total number of d-block	A. 10 B. 20

86	Total number of d-block elements are	B. 20 C. 30 D. 40
87	Which of the following is non-typical transition metal	A. Fe B. Mn C. Zn D. Ni
88	Typical transition element is	A. Sc B. CO C. Ra D. Y
89	Group VI B to transition elements contains	A. Zn, Cd, Hg B. Fe, Ru, OS C. Cr, MO, W D. Mn, Te, Re
90	The colour of transition metal complexes	A. d-d transitions of electrons B. paramagnetic nature of transition elements C. ionization D. loss of s-electron
91	The strength of binding energy of transition elements depends upon	A. number of electron pairs B. number of unpaired electron pairs C. number of neutrons D. number of protons
92	Co-ordination number of Pt in Pt $Cl(NO_2)(NH_3)_4$	A. 2- B. 4 C. 1 D. 6
93	Co-ordination number of Cu in	A. Zero B. Two C. Four D. Six
94	Which is not an ore of iron	A. haematite B. Magnetite C. limonite D. Cassiterite
95	Mild steel contains carbon percentage	A. 0.1 - 0.2% B. 0.3 - 0.7% C. 0.7 - 1.5% D. 1.6 - 2.0%
96	Which of the following is non-typical transition element	A. Cr B. Mn C. Zn D. Fel
97	Which of the following is a typical transition metal	A. Sc B. Y C. Ra D. Co
98	f-block elements are also called	A. non typical transition elements B. outer transition elements C. normal transition elements D. inner transition elements
99	The strength of binding energy of transition elements depend upon	A. number of electron pairs B. number of unpaired electron pairs C. number of neutrons D. number of protons
100	Group VI-B of transition elements contains	A. Zn, Cd, Hg B. Fe, Ru, Os C. Cr, Mo, W D. Mn, Te, Re
101	The percentage of carbon in different types of iron products is in the order of	A. cast iron > wrought iron > steel B. wrought iron > steel > cast iron C. cast iron > steel > wrought iron D. cast iron = steel > wrought iron
102	The chemist who synthesized urea from ammonium cyanate was	A. Berzelius B. Kolbe C. Wholer D. Lavoisier
103	Formula of marsh gas is	A. CH₄ B. C₂H₆ C. C₃H₆ D. C₄H₁₀
		A. Thermal Cracking

104	The process used to improve quality of gasoline	A. Thermal Cracking B. Reforming C. Combination D. Steam Cracking
105	---------- is Alcohol in the following	A. CH₃.CH₂.OH B. CH₃.O.CH₃ C. CH₃COOH D. CH₃.CH₂.SH
106	CO_2H is a functional group as	A. Alkoxy B. Carbonyl C. Carboxyl D. Hydroxyl
107	Which one is alcohol in the following	A. CH₃.CH₂.OH B. CH₃.O.CH₃ C. CH₃COOH D. CH₃.CH₂.Br
108	Which one of the following is an amide	A. (NH₂)₂CO B. NH₂.CH₃ C. C₆H₅NH₂ D. N(CH₃)₃
109	Select from the following the one which is Alcohol	A. CH₃CH₂OH B. CH₃OCH₃ C. CH₃COOH D. CH₃.CH₂.Br
110	Which one is the heterocyclic compound of oxygen	A. Pyrridine B. Parrole C. Furan D. Thiophene
111	-SH Functional group is called	A. Cyano B. Mercapto C. Nitro D. Carboxyl
112	Linear shape is associated with set of hybrid orbitals	A. SP B. sp² C. dsp² D. sp³
113	The state of hybridization of "C" in ethane is	A. SP B. sp² C. dsp² D. sp³
114	The state of hybridization of carbon atom in Ethyne	A. sp B. sp² C. dsp² D. sp³
115	The bond angle between any two SP^2Hydridized orbitals is of	A. 180° B. 109.5° C. 120° D. 107.5°
116	Which set of Hybrid orbital has planner triangular shape	A. sp B. sp² C. dsp² D. sp³
117	The state of hybridization in ethene molecule is	A. dsp² B. sp³ C. sp² D. sp
118	Ethers show the phenomenon of	A. Position Isomerism B. Functional group isomerism C. Metamerism D. Chain isomerism
119	The presence of a double bond in a compound is the sign of	A. Saturation B. Un-saturation C. Substitution D. None of these
120	The catalytic oxidation of methane produces	A. CO + H₂O B. CO₂+ H₂O C. C₂+ H₂O D. H₃C - OH
121	The general formula for Alkanes is	A. C_nH_{2n+1} B. C_nH_{2n} C. C_nH_{2n-2} D. C_nH_{2n+2}

122	Formula of choloroform is	A. CH₃Cl B. CCl₄ C. CH₂Cl₂ D. CHCl₃
123	Preparation of vegetable ghee involves	A. Halogenation B. Hydrogenation C. Hydroxylation D. Dehydrogenation
124	The general formula for Alkene having one double bond is	A. C_nH_{2n+1} B. C_nH_{2n} C. C_nH_{2n-2} D. C_nH_{2n+2}
125	Which one is not property or uses of mustard gas	A. Used in 1st world war B. Powerful vesicant C. High boiling liquid D. High boiling gas
126	Vinyl acetylene reach with HCl to form	A. Polycetylene B. Benzene C. Chloroprene D. Divinylacetylene
127	Which compound is the most reactive	A. Benzene B. Ethene C. Ethane D. Ethyne
128	Synthetic rubber is made by polymerization of	A. Vinylaecetate B. Acetylene C. Divinylacetylene D. Chloroprene
129	Which gas is used for artificial ripening of fruits	A. Ethene B. Metheane C. Propane D. Ethyne
130	Formula of chloroform is	A. CH₃Cl B. CCl₄ C. CH₂Cl₂ D. CHCl₃
131	Vinyl acetylene combines with HCl to form	A. Polyacetylene B. Benzene C. Chloroprene D. Divinyl acetylene
132	The addition of unsymmetrical reagent to an unsymmetrical alkene is in accordance with the rule	A. Hund's rule B. Markownikov's rule C. Pauli's Exclusion Principle D. Aufbau Principle
133	When methane reacts with Cl_2 in the presence of diffused light the products obtained are	A. Chloroform only B. Carbon tetrachloride only C. Chloromethane and dichloromethane D. Mixture of a, b, c
134	Which one of the following gases is used for artificial ripening of fruits	A. Ethene B. Ethyne C. Methane D. Propane
135	The presence of a double bond in a compound is the sign of.	A. Saturation B. Unsaturation C. Substitution D. None of these
136	Aromatic hydrocarbons are the derivatives of	A. Normal series of paraffins B. Alkene C. Benzene D. Cyclohexane
137	The resonating contributing structures of Benzene are	A. 2 B. 3 C. 5 D. 7
138	The structure of Benzene is	A. Hexagonal irregular B. Tetrahedral C. Trigonal planner D. Hexagonal planner
139	During Nitration of Benzene the active nitrating agent is	A. NO₃ B. NO₂⁺ C. NO₂

	the active nitrating agent is	C. NO₂⁺ D. HNO₃
140	Which of the following is Ortho and Para directing group	A. -I B. -CHO C. -COR D. -NH₂
141	Which compound is the most reactive	A. Benzene B. Ethene C. Ethane D. Ethyne
142	The benzene molecule contains	A. three double bonds B. two double bonds C. one double bond D. delocalized pie-electron charge
143	Which of the following acid can be used as a catalyst in Friedel Craft's reactions	A. AlCl₃ B. HNO₃ C. BeCl₂ D. NaCl
144	Benzene cannot undergo	A. Substitution reactions B. addition reactions C. oxidation reactions D. elimination reactions
145	Which catalyst is used Friedel Crafts reactions	A. AlCl₃ B. BeCl₂ C. NaCl D. HNO₃
146	The electrophile in Aromatic sulphonation is	A. H₂SO₄ B. HSO₄¹⁻ C. SO₃ D. SO₃¹⁺
147	m-choronitro benzene is prepared by	A. Nitration of chlorobenzene B. Nitration of Benzene C. Chlorination of Nitrobenzene D. Nitration of m-chlorobenzene
148	Amongst the following, the compound of that can be most readily sulphonated is	A. Toluene B. Benzene C. Nitro-benzene D. Chloro-benzene
149	Which one is not a meta directing group	A. -COOH B. -CHO C. -COR D. -NH₂
150	Amongst the following, the compound that can be most readily sulphonated is	A. toluene B. benzene C. nitrobenzene D. chlorobenzene
151	During nitration of benzene, the active nitrating agent is	A. NO₃ B. NO₂⁺ C. NO₂⁻ D. HNO₃
152	Which compound is the most reactive one	A. benzene B. ethene C. ethane D. ethyne
153	In primary alkyl halides, the halogen atom is attached to a carbon which is further attached to how many carbon atoms	A. Two B. Three C. One D. Four
154	The most reactive Alkyl halide is	A. Alkyl Iodide B. Alkyl Bromide C. Alkyl fluoride D. Alkyl Chloride
155	S_N2 reactions can be carried out with	A. Primary alkylhalide B. Secondary alkyhalide C. Tertiary alkylhalide D. All of these
156	---------- is not a nucleophile	A. H₂O B. NO₃ C. BF₃ D. NH₃

157	Elimination Bimolecular reactions involve	A. Second order kinetics B. First order kinetics C. Third order kinetics D. Zero order kinetics
158	S_N 2 mechanism involves	A. 1st order kinetics B. 2nd order kinetics C. 3rd kinetics D. zero order kinetics
159	For ---------- Mechanism, the first step involved is the same	A. E1 and E2 B. E2 and S_N2 C. S_N1 and S_N2 D. E1 and S_N1
160	Cyanogen chloride reacts with ethyl magnesium bromide to give	A. CH₃CH₂Cl B. CH₃CH₂Br C. C₄H₁₀⁺ D. CH₃CH₂CN
161	When $CO_2$is made to react with ethyl-magnesium iodide followed by acid hydrolysis, the product formed is	A. Propane B. Propanoic acid C. Propanal D. Propanol
162	The reactivity of Grignard's regent is due to	A. Polarity of Mg-x bond B. Polarity of C-Mg bond C. Electro negativity of halogen atom D. Presence of Mg-atom
163	Which compound is formed, when CH_3OH reach with CH_3-Mg -Br	A. Ethane B. Methane C. Ethanol D. Acetone
164	The reactivity order of alkyl halides for a particular alkyl group is	A. Fluoride > Chloride > Bromide > iodide B. Chloride > Bromide > Fluoride > iodide C. Bromide > iodide > chloride > Fluoride D. Iodide > Bromide > Chloride > Fluoride
165	When $CO_2$is made to react with ethyl magnesium iodide, followed by acid hydrolysis, the product formed is	A. propane B. propanoic acid C. propanal D. propanol
166	Grignard's reagent is reactive due to	A. the presence of halogen atom B. the presence of Mg atom C. the polarity of C-Mg bond D. none of the above
167	SN_2reactions can be best carried out with	A. primary alkyl halides B. secondary alkyl halides C. tertiary alkyl halides D. All the three
168	For which mechanisms, the first step involved is the same	A. E₁and E₂ B. E₂and SN₂ C. E₁and E₂ D. E₁and SN₁
169	Alkyl halides are considered to be very reactive compounds towards nucleophile because	A. They have an electrophilic carbon B. They have an electrophilic carbon and a good leaving group C. They have an electrophilic carbon and a bed leaving group D. They have a nucleophilic carbon and a good leaving group
170	Which compound is called universal solvent	A. CH₃OH B. C₂H₅OH C. CH₃O CH₃ D. H₂O
171	In t-butyl alcohol, the tertiary carbon is bonded	A. Three hydrogen atoms B. Two hydrogen atoms C. One hydrogen atom D. No hydrogen atom
172	Which compound shows maximum hydrogen bonding with water	A. CH₃OH B. C₂H₅OH C. CH₃- O - CH₃ D. C₆H₅OH
173	Which compound will have the maximum repulsion with water	A. C₆H₆ B. C₂H₅OH C. C₃H₇OH D. CH₃OCH₃
174	Alcohol obtained by fermentation is only upto	A. 10% B. 12% C. 20%

		D. 95%
175	Methyl alcohol is not used	A. As a solvent B. As an anti freezing agent C. As a substitute for petrol D. For denaturing of ethyl alcohol
176	Rectified spirit contains alcohol about	A. 80% B. 85% C. 90% D. 95%
177	---------- compound shows extensive hydrogen bonding with water	A. C₂H₆ B. H₂S C. C₂H₅OH D. CH₃Cl
178	Which enzyme is not involved in fermentation of starch	A. Zymase B. Urease C. Invertase D. Diastase
179	Which compound show hydrogen bonding	A. C₂H₆ B. C₂H₅Cl C. CH₃OCH₃ D. C₂H₅OH
180	Ethanol can be converted into ethanoic acid by	A. Hydrogenation B. Hydration C. Oxidation D. Fermentation
181	Which one the following a dihydric alcohol	A. Ethanol B. Cyclo hexanol C. Glycerol D. Glycol
182	Which compound is more soluble in water	A. C₂H₅OH B. C₆H₅OH C. CH₃COCH₃ D. n - hexanol
183	Isopropyl alcohol on oxidation gives	A. Acetaldehyde B. Acetone C. Ether D. Propene
184	The most reactive alcohol when O-H bond breaks is	A. Tertiary alcohol B. Secondary alcohol C. Primary alcohol D. Methyl alcohol
185	Bakelite is obtained from phenol by reacting with	A. Acetal B. Ethanal C. Formaldehyde D. Methanol
186	According to Lewis concept, ethers behave as	A. Acid B. Base C. Nucleophile D. Solvent
187	The carbon atom of carbonyl group is hybridized	A. Sp B. Sp² C. Sp³ D. dsp²
188	Which of the following has highest boiling point	A. Methanal B. Ethanal C. Propanal D. 2-hexanone
189	Formalin is a 40% solution of	A. CH₃CHO B. CH₃OH C. HCHO D. CH₃OCH
190	Which of the following compounds will not give Iodoform test on treatment with I_2/ NaOH	A. Acetaldehyde B. Acetone C. Butanone D. 3-Pentanone
191	Cannizzaro's reaction is not given by	A. Formaldehyde B. Acetaldehyde C. Benzaldehyde D. Trithylacetaldehyde
		A. Grignard reagent

192	Which reagent will react with both aldehyde and ketones	B. Tollen's reagent C. Fehling's reagent D. Benedict's reagent
193	Acetone reacts with HCN to form cyanohydrin it is an example of	A. Electrophilic addition B. Electrophilic substitution C. Nucleophilic addition D. Nucleophilic subtitution
194	Aldehyde react with hydroxyl amine in acidic solution to give	A. An oxime B. Aldol C. Polymer D. Acetic acid
195	Which is most difficult to be oxidized	A. CH₃CHO B. CH₃COCH₃ C. HCHO D. C₂H₅CHO
196	Cannizzaro's reaction is given by	A. Acetaldehyde B. Formaldehyde C. Propanal D. Propanone
197	Which reaction is disproportionate reaction	A. Aldol Condensation B. Cannizzaros's reaction C. Haloform reactions D. Acid-catalyzed reactions
198	Which reaction is disproportionate reaction	A. Aldol Condensation B. Cannizzaros's reaction C. Haloform reactions D. Acid Catalyzed reactions
199	Silver mirror test is given by	A. Ethers B. Ketones C. Acids D. Aldehydes
200	Aldehydes and ketones can be defected by	A. 2, 4 DNPH test B. Tollen's test C. Sodium Nitro prusside test D. Benedicts solution test
201	Which one has yellow or orange cyrstalline ppt	A. Acetone hydrazone B. 2, 4 DNPH C. Ethanal oxime D. Bisulphite addition product
202	The carbon atom of a carbonyl group is	A. sp hybridized B. sp²hybridized C. sp³hybridized D. none of these
203	Formalin is	A. 10% solution of formaldehyde in water B. 20% solution of formaldehyde in water C. 40% solution of formaldehyde in water D. 60% solution of formaldehyde in water
204	The solution of which acid is used for seasoning of food	A. Benzoic acid B. Butanoic acid C. Formic acid D. Acetic acid
205	Catalyst used to reduce carboxylic acid to alcohol is	A. H₂/ Ni B. H₂/ pt C. NaBH₄ D. LiAlH₄
206	The flavour of amylacetate is	A. Orange B. Apricot C. Banana D. Pineapple
207	Which of the following is not directly prepared from CH_3COOH	A. Ethyl acetate B. Acetyl chloride C. Acetic anhydride D. Acetamide
208	The flavor of octylacetate is	A. Orange B. Apricot C. Banana D. Jasmine
209	Organic compound having fruity smell are	A. Caboxylic acid B. Alcohols C. Ethers D. Esters

210	Banana flavour is given by the ester	A. Octyl acetate B. Amyl butyrate C. Amyl acetate D. Ethyl butyrate
211	Acetic acid manufactured by	A. Distillation B. Fermentation C. Ozonolysis D. Esterification
212	The nature of lysine amino acid is	A. Acidic B. Basic C. Amphoteric D. Natural
213	Which of the following is not an amino acid	A. Aspartic acid B. Lysin C. Alanine D. Aniline
214	Which one is neutral amino acid	A. Lysine B. Histidine C. Glutamic acid D. Valine
215	Acetic acid is manufactured by	A. distillation B. fermentation C. ozonalysis D. esterification
216	A caboxylic acid contains	A. a hydroxyl group B. a caboxyl group C. a hydroxyl and carboxyl group D. a carboxyl and an aldehyde group
217	Which acid is used in the manufacture of synthetic fibre	A. Formic acid B. Oxalic acid C. Carbonic acid D. Acetic acid
218	Which of the following derivative cannot be prepared directly from acetic acid	A. Acetamide B. Acetyl choride C. Acetic anhydride D. Ethyl acetate
219	Which reagent is used to reduce a carboxylic group to an alcohol	A. N₂/ Ni B. H₂/Pt C. NaBH₄ D. LiAlH₄
220	Which of the following is not a fatty acid	A. Propanoic acid B. Acetic acid C. Phthalic acid D. Butanoic acid
221	Which of these polymers is a synthetic polymer	A. Animal fat B. Starch C. Cellulose D. Polyester
222	Which one of the following is a condensation polymer	A. Polystyrene B. Polyester C. Polyethene D. Nylon 6,6
223	Which of the following is an addition polymer	A. Polyester B. Polystyrene C. Nylon 6,6 D. Terylene
224	The fiber which is made from acrylonitrile as monomer	A. PVC B. Polyester fiber C. Rayon fiber D. Acrylic fiber
225	Nylon 6,6 is obtained by the reaction of hexamethylene diamine with	A. Acetic acid B. Adipic acid C. Viny chloride D. Acetyl chloride
226	Starch is	A. Monosaccharide B. Disaccharide C. Polysaccharide D. Oligosaccharide
227	Which one is a disaccharide	A. Glucose B. Sucrose C. Fructose

		D. Cellulose
228	Which of the following are mono-saccharides	A. Fructose B. Sucrose C. Stach D. Cellulose
229	Which of the following element is not present in all proteins	A. Carbon B. hydrogen C. Nitrogen D. Sulphur
230	Vegetable oils are	A. Polyesters B. Glycerides of unsaturated fatty acids C. Essential oils D. Fatty acids
231	The reaction between fat and NaOH is	A. Estrification B. Hydrogenolysis C. Fermentation D. Saponification
232	Which of the following enzymes brings about the hydrolysis of fats	A. Urease B. Maltase C. Zymase D. Lipase
233	In which of following processes are small organic molecules made into macromolecules	A. the cracking of petroleum fractions B. the fractional distillation of crude oil C. the polymerization of ethene D. the hydrolysis of proteins
234	Which of these polymers is an addition polymer	A. Nylon 6, 6 B. polystyrene C. Terylene D. epoxy resin
235	Which of these polymers is a synthetic polymer	A. Animal fat B. Starch C. Cellulose D. polyester
236	Plastics are pollution problem because many plastics	A. are made from petroleum B. are very inflammable C. burn to produce toxic fumes D. decompose to produce toxic products
237	A polymeric substance that is formed in the liquid state and then hardened to a rigid solid is called a	A. fibre B. plastic C. varnish D. polyamide resin
238	Micro-nutrient is required in quantity for plant growth ranging from	A. 4-40 gm B. 6-200 gm C. 6-200 kg D. 4-40 kg
239	Which elements are needed for healthy growth of plants	A. N,S,P B. N,Ca,P C. N,P,K D. N,K,C
240	The macronutrients are required in quantities ranging from	A. 4-40 kg per acre B. 10-100 kg per acre C. 5-100 kg per acre D. 5-200 kg per acre
241	Ammonium Nitrate fertilizer is not useful for	A. Wheat B. Cotton C. Sugan cane D. Paddyrice
242	Phosphorus helps the growth of	A. Root B. Leaf C. Steam D. Seed
243	Which is not a calcarious material	A. Clay B. Limestone C. Marble D. Chalk
244	One of following is argillaceous material	A. Marble B. Clay C. Lime D. Marine Shell
245	The word paper is derived from the name of which reedy	A. Rose B. Sun flower

	plant	C. Papyrus D. Water
246	Woody raw material for paper pulp is obtained from	A. Cotton B. Bagasse C. Poplar D. Rice straw
247	Which three elements are needed for the healthy growth of plants	A. N,S,P B. N,Ca,P C. N,P,K D. N,K,C
248	Which woody raw material is used for the manufacture of paper pulp	A. Cotton B. Biogases C. Rice straw D. Poplar
249	The nitrogen present in some fertilizers helps plants	A. to fight against diseases B. to produce fat C. to undergo photosynthesis D. to produce protien
250	Phosphorous helps the growth of	A. root B. leave C. stem D. seed
251	Micronutrients are required in quantity ranging from	A. 4g-40g B. 6g-200g C. 6kg-200kg D. 4kg-40kg
252	During the manufacturing process of cement the temperature of the decomposition zone goes up to	A. 600°C B. 800°C C. 1000°C D. 1200°C
253	The word paper is derived from the name of which reedy plant	A. Rose B. Sun flower C. Papyrus D. Water Hyacinth
254	How many zones through which the charge passes in a rotary kiln	A. 4 B. 3 C. 2 D. 5
255	Ecosystem is smaller unit of	A. Lithosphere B. Hydroshpere C. Amosphere D. Biosphere
256	Thickness of atmosphere is about how much kilometer above the surface of warth	A. 100 km B. 1000 km C. 10,000 km D. unlimited
257	Which is a secondary pollutant	A. Carbonic acid B. CO₂ C. SO₂ D. CO
258	A single chloride free radical can destroy how many ozone molecules	A. 100 B. 100,000 C. 100,00 D. unlimited
259	The pH of unpolluted rain water should be	A. 5.00 B. 5.60 C. 6.50 D. 7.00
260	The pH of truly acidic rain is	A. 7-6.8 B. 6.5-6 C. 6-5.6 D. less than 5
261	Peroxyacetyl nitrate (PAN) is a irritant to human beings and it affects	A. Eyes B. Ears C. Stomach D. Nose
262	Which gas is cause of Asthma	A. O₃ B. O₂ C. SO₂ D. CO₂

263	The main pollutant of leather tanneries in the waste water is due to	A. Lead B. Chromium VI C. Copper D. Chromium III
264	In water the concentration of dissolved O_2should be	A. 1-3 ppm B. 2-4 ppm C. 4-8 ppm D. 8-12 ppm
265	Water is disinfected by a substance to avoid toxification	A. KMnO₄ B. Alums C. O₃ D. Cl₂
266	The news paper can be recycled again and again as many times as	A. 5 B. 3 C. 4 D. 2
267	The pH range of the acid rain is	A. 7-6-5 B. 6.6-6 C. 6-5.6 D. less than 5
268	Peroxyacetylnitrate (PAN) is an irritant to human beings and it affects	A. eyes B. ears C. stomach D. nose
269	The avoid the formation of toxic compounds with chlorine which substance is used for disinfecting water	A. KMnO₄ B. O₃ C. Alums D. Chloramines
270	Fungicides are the pesticides which	A. Control the growth of fungus B. Kill insects C. Kill plants D. Kill herbs
271	Ecosystem is a smaller unit of	A. lithospher B. hydroshper C. atmosphere D. biosphere
272	Keeping in view size of atoms ,which order is the correct one.	A. Mg > Sr B. Ba > Mg C. Lu > Ce D. Cl > I
273	Mark the correct statement.	A. Na+ is smaller than Na atom B. Na+ is large than Na atom C. Cl- is smaller than Cl atom D. cl- and Cl are equal in size
274	Which statement is incorrect.	A. All the metals are good conductors of electricity B. All the metals are good conductor of heat C. All the metal form positive ions D. All the metal form acidic oxides
275	Hydrogen resembles in properties with	A. IA, IV A and VII A elements B. III A, IV A and V A elements C. II A, IV A and VI A elements D. II A, III A and VII A elements
276	Ionization energy of calcium is.	A. Lower than that of barium B. Lower than that of magnesium C. Higher than that of beryllium D. Lower than that of strontium
277	Electron affinity is measure of energy	A. Required to remover the electron B. Released by adding an electron C. Required to excite an electron D. Released by removing an electron
278	Mark the incorrect statement	A. Metallic character increase down the group B. Metallic character increase from left to right along a period C. Metallic character decrease from left to right along a period D. Metallic character remains the same down the group
279	In which group of periodic table is the element which has atomic number 14.	A. II B. IV C. III D. VI
280	Which one of the following sets has coinage metaals is it.	A. Cu, Hg, Au B. Cu, Ag, Au C. Ag, Au, Hg D. Cu, Fe, Au

281	Fluorine is in group VII A of periodic table. Its chemistry will most closely resembles that of.	A. Argon B. Boron C. Iodine D. Sulphur
282	Fluorine is in group VIIA of periodic table. Its chemistry will most closely resembles that of.	A. Argon B. Boron C. Iodine D. Sulphur
283	What are the total numbers of periods in the modern periodic table	A. 3 B. 5 C. 7 D. 8
284	What are the total numbers of periods in the modern periodic table.	A. 3 B. 5 C. 7 D. 8
285	What criteria did Mendeleev use in arranging his periodic table.	A. Atomic mass B. Atomic number C. Mass number D. Density
286	Majority of the elements of the periodic table are.	A. Semi metals B. Non metals C. Metals D. Noble metals
287	Transition elements are those	A. Which becomes before uranium B. Which becomes after uranium C. Which are prepared artificially D. Which are in between s-block and p - block elements
288	In which block of periodic table non metals are present.	A. s B. p C. d D. f
289	In which one of the following sets do all three particles have same number of total electrone	A. F- ,Cl-,Br- B. Li+ , Na+ , K+ C. N^{--3,} O-2, F-, D. Na+, Mg+2, K+
290	In a period, from left to right in the periodic table, the size of atom generally.	A. Increases B. decreases C. Remains constant D. First increase upto the middle of period and then decreases
291	As you proceed across a period in the periodic table the first ionization energy	A. Decrease B. Increase C. Remains constant D. First increase up the middle of period and then decreases
292	Which class of elements shows law value of first ionization potential.	A. Alkali metals B. Alkaline earth metals C. Halogens D. Noble gases
293	Which one of the following elements has the largest second ionization energy.	A. O B. Na C. F D. Ne
294	Which one of the following is a metalloid.	A. sulphur B. Antimony C. Mercury D. Zinc
295	The first ionization energy of Na, Mg, Al and Si are in theorder of.	A. Na < Mg < Al < Si B. Na > Mg> Al> Si C. Na > Mg< Al< Si D. Na< Mg> Al< Si
296	Across a period from left to right in the periodic table, the melting and boiling point.	A. Decrease B. Increase C. Remain constant D. First increase upto the middle of period and then decrease
297	Variable valency is generally exhibited by	A. Transition elements B. Alkali metals C. s-block elements D. Gaseous elements
298	In which group, melting point and boiling point increase	A. IA B. II A C. VII A

	downward in a group	C. VII A D. Both a and b
299	In which compound, oxidation state of sulphur is +6	A. H2S B. H2SO4 C. H2SO3 D. SO3
300	Which has greater hydration energy.	A. Li+ B. Na+ C. K+ D. Mg+2
301	Which element when react with chlorine form polymeric halide.	A. Na B. Be C. Ba D. P
302	Ionic Hydrides react with water to form	A. Proton B. Hydride ions C. Hydroxide ions D. Hydronium ions
303	Which hydride is intermediate in nature.	A. NaH B. BeH2 C. NH3 D. HCl
304	Element of which group reacts with hydrogen and form ionic hydrides.	A. II A B. IV A C. V A D. VI A
305	Which one of the following elements burns in air to form an oxide which, when shaken with water, give sa solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
306	Which one of the following elements burns in air to form an oxide which, when shaken with water, give a solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
307	Which one of the following elements burns in air to form an oxide which, when shaken with water, give a solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
308	Which one of the following oxides is amphoteric in nature.	A. MgO B. Na2O C. SO2 D. ZnO
309	The oxides of metal sare generally	A. Acidic B. Basic C. Neutral D. Amphoteric
310	Which property of hydrogen not resemble to alkali metals.	A. Electronic configuration B. Oxidation state C. Reaction with halogen D. Metallic nature
311	Which oxide is more basic in nature.	A. Beo B. MgO C. CaO D. BaO
312	Which one of the following does not belong to alkaline earth metals.	A. Be B. Ra C. Ba D. Rn
313	The oxide of beryllium is	A. Acidic B. Basic C. Amphoteric D. None of these
314	Which ion will have the maximum value of heat of hydration.	A. Na+ B. Ca+ C. Ba+2 D. Mg+2
315	Which one of the following is not an alkali metal.	A. Francium B. Cesium C. Rubidium

		D. Radium
316	Which of the following sulphates is not soluble in water.	A. Sodium sulphate B. Potassium sulphate C. Zinc sulphate D. Barium Sulphate
317	The element cesium bears resemblance with	A. Ca B. Cr C. Both of the above D. None of the above
318	Chile saltpeter has the chemical formula	A. NaNO3 B. KNO3 C. Na2B4O7 D. Na2CO3.H2O
319	The ore CaSO4.2H2O has the general name.	A. Gypsum B. dolomite C. Calcite D. Epsom salt
320	Down's cell is used to prepare.	A. Sodium carbonate B. Sodium bicarbonate C. sodium metal D. Sodium hydroxide
321	Ca Mg3 (SiO3)2 is the composition of.	A. Dolomite B. Gypsum C. Calcite D. Asbestos
322	In which ore carbon is present	A. Magnesite B. Epsom salt C. Barite D. Sylvite
323	Which element is present in chlorophyll	A. Ca B. Mg C. K D. Be
324	Which meal is least reactive	A. Li B. NaOH C. K D. Pb
325	Which hydroxide decomposes on heating.	A. LiOH B. NaOH C. KOH D. CaOH
326	Which one of the following pairs of metal reactive directly with nitrogen	A. Na and Mg B. Li and Mg C. Mg and Ca D. Li and Be
327	Alkali metal whose carbonate is is relatively less stable and decomposes giving its oxides is.	A. Li B. Ba C. K D. Rb
328	Which one of the following has lowest first ionization energy.	A. Li B. Na C. Rb D. Cs
329	Which is more basic	A. RbOH B. NaOH C. KOH D. Li OH
330	Which is more basic	A. RbOH B. NaOH C. KOH D. Li OH
331	Which gas is evolved when CO2 is passed through KO2.	A. O2 B. CO C. O3 D. None of these
332	The crystals of caustic soda, NaOH, are hygroscopic when these crystals are exposed to air.	A. Gain water and remains solid B. Gain water and becomes liquid C. Lose mass and remain solid D. Remain unchanged on heating.
333	Which element reacts with	A. Li B. Na

	carbon to form metal carbide.	C. K D. All of these
334	Magnesium metal does not burn in the jar of	A. N2 B. O2 C. Ne D. N2 and O2
335	When carbon di oxide is bubbled through lime water, white precipitate is formed This precipitate is.	A. COCl2 B. H2CO3 C. CaO D. CaCO3
336	Oxide of Alkaline earth metal are	A. Acid B. Basic C. Neutral D. Amphoteric
337	Oxides of alkaline earth metal are	A. Acidic B. Basic C. Neutral D. Amphoteric
338	Which element reacts with alkali and H2 gas is produced.	A. Be B. Mg C. Ca D. Sr
339	Setting of Plaster of Paris is accompanied with	A. Hydrogen bonding B. Hydration C. Dehydration D. None of these
340	Which one of the following elements has the same oxidation number in all its known compounds.	A. Beryllium B. Nitrogen C. Sulphur D. Chorine
341	Which on eof the following alkaline earth metal form peroxide when heated with oxygen at 600 ^{o}C	A. Magnesium B. Calcium C. Strontium D. Barium
342	Which oxide is amphoteric in nature.	A. BeO B. MgO C. CaO D. BaO
343	Which is called milk of magnesia.	A. MgCO3 B. Mg(OH)2 C. MgSO4 D. MgCl2
344	Which is insoluble in water.	A. BeSO4 B. MgSO4 C. CaSO4 D. BaSO4
345	Sodium is collected over molten sodium chloride in Down's cell, because.	A. It does not react with NaCl B. It is lighter in weight and floats on molten NaCl C. Its reactivity is greater than NaCl D. None of these
346	In Down's cell anode is made up of.	A. Iron B. Copper C. Graphite D. Platinum
347	Which substance is used to lower the melting point of NaCl in Down's cell	A. CaCO3 B. CaCl2 C. MgSO4 D. AlCl3
348	Which of the following compounds is used to make spray, which has fungicidal action.	A. BaSO4 B. Na2CO3 C. CaO D. CaSO4.2H2O
349	Sulphur is essential constituent for plants. Which is not role of sulphur	A. Chlorophyll development B. Development of root system C. Constituents of some proteins D. Increase transpiration force
350	When water is added to plaster of paris, it changes to a hard mass. Its volume also expands upto	A. 1% B. 2% C. 3% D. 5%
		A. Extraction of metals

351	Which one is not use of lime	B. Ceramic industry C. Preparation of bleaching powder D. Preparation of gypsum
352	Which metal is used in the thermite process because of its activity.	A. Iron B. Copper C. Aluminium D. Zinc
353	Aluminium oxide is.	A. Acidic oxide B. Basic oxide C. Amphoteric oxide D. None of these
354	Which element forms on ion wiht charge 3+	A. Beryllium B. Aluminium C. Carbon D. silicon
355	Which element among the following belongs to Group IV A of the Periodic table.	A. Barium B. Iodine C. Lead D. Oxygen
356	Boric acid can not be used.	A. As antiseptic in medicine B. For washing eyes C. In soda bottle D. For enamels and glazes
357	Which of the following elements is not present abundantly in earth's crust.	A. Silicon B. Aluminium C. Sodium D. Oxygen
358	Tincal is a mineral of.	A. Al B. B C. Si D. C
359	Corundum is	A. Al_2O_3 B. Na_2AlF_6 C. Quartz D. Calcium
360	Which one of the followings is not ore of aluminium	A. Coruridum B. Bauxite C. Colemanite D. Kaolin
361	Which one of the following elements exhibits s-inert pair effect.	A. B B. Al C. Pb D. Sc
362	Third abundant element on earth crust is.	A. Boron B. Oxygen C. Aluminium D. Silicon
363	Boron usually exist in nature as	A. Borides B. Oxborates C. Free element D. It is a artificial element
364	Boric acid is used as eye wash due to its	A. Weakly acidic property B. Antiseptic nature C. Antibiotic nature D. Weakly basic properties
365	Ethyl alcohol on reacting with orthoboric acid form	A. Ethyl borate B. Boric acid C. Boron acetate D. Borax
366	In which compound, the oxidation state of boron in not +3	A. H_2BO_3 B. $Ca_2B_2O_{11}$ C. $Na_2B_4O_7$ D. Mg_3B_2
367	A solution of borax in water is.	A. Acidic B. Alkaline C. Neutral D. None of these
368	Which is not used of Borax	A. Softening of water B. As medicine for washing eyes C. As flux in metallurgical operations D. To make quartz

369	Ortho boric acid on strong beating gives.	A. Borax B. Boron oxide C. Metaboric acid D. Tetraboric acid
370	Aluminium does not corrode because.	A. It is a semi metal B. It is a silver shining metal C. It does not react with O2 D. It forms a protective layer of Al2O3
371	Al reacts with caustic soda to form	A. Sodium aluminates B. Aluminium hydroxide C. Aluminium oxide D. sodium tetra aluminate
372	Which element burns in nitrogen atmosphere to form nitride.	A. Mg B. Al C. Both a and b D. None of these
373	Aluminium is used for making petrol and milk storage tanks because it is	A. Conductor B. Non magnetic C. Excellent reflected D. Corrosion resistant
374	Which are correct oxidation states for lead	A. +1, +2 B. +3 , +5 C. +2 ,+4 D. +4 only
375	In which substance silicon is not present.	A. Talc B. Asbestos C. Dolomite D. Zirocon
376	The shape of SiO2 IS	A. Tetrahedral B. Trigonal bipyramid C. Linear D. Cubic
377	silicones are	A. Some salts of sodium B. Allotropes of Sl C. Inorganic polymers D. Coloured compounds of Si
378	Which substance is water repellant and used in ceramic insulators.	A. Asbestos B. Lead compounds C. Silicon carbide D. silicones
379	Which one of the following compounds of lead is red in colour	A. PbO B. Pb2O C. Pb3O4 D. PbCO3
380	The composition of white lead is	A. PbCO3.Ph(OH)2 B. PbCO3.2Pb(OH)2 C. 2PbCO3.Pb(OH)2 D. PbCO3
381	Litharge is chemically	A. PbO B. PbO2 C. Pb3O4 D. Pb(CH3COO)2
382	Lead monoxide is	A. Amphoteric B. Neutral C. Acidid D. Basic
383	Which substances is called chrome yellow.	A. Pb3O4 B. PbO C. 2PbCO3.Pb(OH)2 D. PbCrO4
384	Out of the elements of group VA, the highest energy is possessed by	A. N B. P C. Sb D. Bi
385	In group VA elements the most electronegative elements is.	A. Sb B. N C. P D. As
386	The brown gas formed when metal reduces HNO3 is	A. N2O5 B. N2O3 C. NO2

		D. NO
387	Out of the elements of group VIA the highest melting and boiling points is shown by the element.	A. Te B. Se C. S D. Po
388	SO2 is not absorbed in water directly to form H2SO4 because.	A. The reaction does not go to completion B. The reaction is quite slow C. The reaction is exothermic D. SO3 is insoluble in water
389	Which catalyst is used in contact process.	A. Fe2O3 B. V2O5 C. SO3 D. Ag2O
390	Which of the following specie has the maximum number of unpaired electrons.	A. O2 B. O2+ C. O2- D. O2⁻²
391	Lowest oxidation state of nitrogen is present in.	A. NH3 B. NO2 C. NO D. HNO3
392	Which element does not have allotropic form	A. Nitrogen B. Phosphorous C. Arsenic D. Antimony
393	Which one of the following oxide is brown in colour.	A. NO B. NO2 C. N2O D. N2O3
394	NO2 can be obtained by heating.	A. NaNO3 B. KNO3 C. Pb(NO3)2 D. NH3NO3
395	NH4NO3 on heating at 200 ^{o}C changes to	A. N2O B. NO C. NO2 D. N2O4
396	When Cu reacts with conc. HNO3 , which one of the following gases is evolved	A. N2O B. NO C. NO2 D. N2O5
397	Which of the following acids possess oxidizing and reducing properties.	A. HCl B. HNO2 C. HNO3 D. H2SO4
398	Which raw material is used for manufacture eof HNO3 by Birkland eyed process	A. NH3 and CO2 B. Air C. Air and gypsum D. Lime stone and urea
399	Bone ash contain calcium phosphate	A. 40% B. 50% C. 70% D. 80%
400	Which form of phosphorus is more stable.	A. White B. Red C. Black D. Both a and b
401	P2O5 is usually used as	A. Drying agent only B. Reducing agent C. Both drying and reducing agent D. Both drying agent and oxidizing agent.
402	Each of the following is true for white and red phosphorus except one.	A. Both are soluble in CCl4 B. Both can be oxidized by heating in air C. Both consists of same kind of atoms D. Both can be converted into each other
403	PCl3 reacts with water to form	A. PH3 B. POCl3 C. H3PO4 D. H3PO5
	Basicity of ortho phosphoric	A. 1 B. 2

404	Basicity of ortho phosphoric acid is.	B. 2 C. 3 D. 4
405	Which allotrope of phosphorus has layers like graphite.	A. white phosphorus B. Red phosphorus C. Black Phosphorus D. Amorphous phosphorus
406	In aqua regia, the ratio of conc. HCl to Conc. HNO3 is	A. 1 : 1 B. 2 :1 C. 1:2 D. 3 : 1
407	What are the number of the electrons in valence shell of P in PCl3	A. 4 B. 6 C. 8 D. 10
408	In which substance phosphorus is not present.	A. Yolk of egg B. Bones C. Apatite D. Galena
409	Which one is metaphosphoric acid	A. HPO3 B. H3PO3 C. H3PO4 D. H4P2O7
410	Which one of the following group of Periodic table called chalcogen family.	A. Group III A B. Group VA C. Group VI A D. Group VII A
411	An element has oxidation state -2, +4, +6 in its compounds. In which group in the periodic table is this element likely to be.	A. Grooup III A B. Group IV A C. Group V A D. Group VI A
412	Role of H2S in the given chemical reaction is H2S + I2---------- 2HI+S	A. Oxidising agent B. Reducing agent C. Dehydrating agent D. As an acid
413	The element which is present in earth crust about 50% is	A. Oxygen B. sulphur C. Carbon D. Nitrogen
414	Chemical formula of stibnite on.	A. BaSO4 B. Sb2S3 C. FeS2 D. ZnS
415	When concentrated H2SO4 and solid sodium chloride react together at room temperature the product are.	A. Two salts only B. A salt and a base C. A salt and an acid D. A salt and water
416	The reaction between concentrated H2SO4 and glucose give carbon and water. In this reaction H2SO4 acts as.	A. An acid B. An oxidising agent C. Dehydrating agent D. A reducing agent
417	Sulphuric acid acts as dehydrating agent in its reaction with.	A. Sodium chloride B. Potassium nitrate C. Copper D. Ethyl alcohol
418	The composition of oleum is.	A. H2SO4 B. H2S2O3 C. H2S2O7 D. H2S3O7
419	In pyrite burner, the gas produced is.	A. SO3 B. SO2 C. CO2 D. NO
420	Which one of the following does not react with dilute sulphuric acid.	A. Mg (OH)2 B. Mg C. MgO D. Mg(NO3)2
421	Arsenic impurities in contact process are removed.	A. By prolong heating the gases B. By treatment with Fe(OH)3 C. In scrubbing tower

		D. In absorption tower
422	Most likely product formed when formic acid is dehydrated in the presence of conc. H_2SO_4 is.	A. CO_2 and H_2O B. CO, CO_2 and H_2O C. CO_2 and H_2 D. CO and H_2O
423	The reaction between Cu and conc. H_2SO_4 produces	A. SO_3 B. SO_2 C. H_2 D. Cu + ions
424	Which statement is incorrect about H_2SO_4	A. Dehydration agent B. dibasic acid C. Oxidizing agent D. Reducing agent
425	Which of the following hydrogen halide is the weakest acid in solution.	A. HF B. HBr C. HI D. HCl
426	Chlorine heptoxide reacts with water to form	A. Hypochlorous acid B. Chloric acid C. Perchloric acid D. Chlorine and oxygen
427	Hydrogen bond is the strongest between the molecules of.	A. HF B. HCl C. HBr D. HI
428	Which halogen will react spontaneously with Au to produce Au^{3+}	A. Br_2 B. F_2 C. I_2 D. Cl_2
429	The anhydride of $HClO_4$ is	A. ClO_3 B. ClO_2 C. Cl_2O_5 D. Cl_2O_7
430	Bleaching powder may be produced by passing chlorine over.	A. Calcium carbonate B. Hydrated calcium sulphate C. Anhydrous calcium sulphate D. Calcium hydroxide
431	Which halogen occurs naturally in a positive oxidation state.	A. Fluorine B. Chlorine C. Bromine D. Iodine
432	An element that has a high ionization energy and tends to be chemically inactive would most likely to be	A. An alkali metal B. A transition element C. A noble gas D. A halogen
433	Which statement is correct about the given reaction. $2NaOH + Cl_2$ ---------- $NaCl + NaClO + H_2O$	A. Cl is oxidized and O is reduced B. Cl is reduced and O is oxidized C. Cl is oxidized as well as reduced D. Neither Cl nor oxygen is reduced or oxidized
434	Stability of halogen molecules decreases from	A. F_2 to I_1 B. Cl_2 to I_2 C. I_2 to F_2 D. I_2 to Cl_2
435	Colour of which halogen is not correctly related.	A. F_2 --- colourless gas B. Cl_2---- greenish yellow gas C. Br_2----- Reddish brown liquid D. I_2-------- grayish Black solid
436	Which one of the following uses is not correctly related with the halogen.	A. fluorine -------- Teflon B. Chlorine ---------Bleaching powder C. Bromine ----------PVC plastics D. Iodine ----------Iodex
437	The halogens ae best described by which of the following statements.	A. Their outer shell is complete B. Most of them are colourless C. They all are oxidizing agent D. They all are gases at room temperature
438	Bromine can be liberated from KBr solution by the action of.	A. I_2 solution B. Chlorine C. NaCl D. KI
		A. HF

439	Which one of the following has highest melting and boiling points.	A. HF B. HBr C. HCl D. HI
440	Which acid can not be stored in glass bottles.	A. HCl B. HF C. H2SO4 D. HNO3
441	The most ionic is	A. HF B. HCl C. HBr D. HI
442	The chemical formula of Sodium Bromite is.	A. NaBrO B. NaBrO2 C. NaBrO3 D. NaBrO4
443	Which one of the following acids acts as oxidizing agent but never a reducing agent.	A. HClO B. HClO2 C. HClO3 D. HClO4
444	Bleaching powder contains available chlorine approximately	A. 100% B. 70-80% C. 35-40% D. 10-20%
445	Which one of the following is not use of chlorine.	A. Formation PVC B. Formation of mustard gas C. Disinfectant and bleaching agent D. Formation of sodium chloride
446	Iodine deficiency in diet is known to cause.	A. Beri Beri B. Goiter C. Rickets D. Night blindness
447	Which raw material is used for preparation of bleaching powder.	A. Cl2 and H2O B. Cl2 and Lime C. Cl2 and HOCl D. HCl and Lime
448	Which substance is used in photography	A. AgCl B. AgBr C. AgI D. Ag3PO4
449	Which element form maximum compounds with Xenon	A. F B. Cl C. Br D. I
450	What is the oxidation state of Xenon in XeOF2	A. 0 B. +2 C. +4 D. +6
451	In fluorescent tube, the gas filled is.	A. He B. Ne C. Ar D. Xe
452	The gas used in bactericidal lamps is	A. Be B. Ar C. Kr D. Xe
453	Which gas has highest boiling points.	A. He B. Ne C. Ar D. Kr
454	Which one of the following gases exist in monoatomic form.	A. Ozone B. Nitrogen C. Krypton D. Phosphine
455	Which one of the following noble gases is least polarizable	A. He B. Ne C. Ar D. Kr
456	XeF6 n hydrolysis produces.	A. XeOF2 B. XeOF3 C. XeOF4 D. XeF2

457	Refrigeration capacity of liquid neon is greater than liquid helium by	A. 80 times B. 50 times C. 40 times D. 10 times
458	Which noble gas is used in radiotherapy	A. Neon B. Argon C. Krypta D. Radon
459	Which of the following is a non typical transition element.	A. Cr B. Mn C. Zn D. Fe
460	Which of the following is a typical transition metal.	A. Sc B. Y C. Ra D. CO
461	f - block elements are also called.	A. Non typical transition elements B. Outer transition elements C. Inner transition elements D. None of true
462	Group VI B of transition elements contains.	A. Zn, Cd, Hg B. Fe, Ru,Os C. Cr, Mo, W D. Mn, Te, Re
463	The percentage of carbon is different types of iron products is in the order of.	A. Cast iron > wrought iron > Steel B. Wrought iron >Steel >Cast iron C. Cast iron > Steel > Wrought iron D. Cast iron > Steel > Wrought iron
464	The colour of transition metal complexes is due to	A. d-d transition of electrons B. Paramagnetic nature of transition elements C. Ionization D. Loss of s -electrons
465	The total number of translation elements is	A. 10 B. 14 C. 40 D. 65
466	Which one of the following elements shows variable valency, can act as a catalyst and form coloured compounds.	A. Carbon B. Chlorine C. Sulpur D. Iron
467	Which one of the following elements has no variable valency.	A. Zinc B. Iron C. cobalt D. Manganese
468	Which one of the following elements commonly exhibits oxidation states of +6 and +3 in aqueous solution.	A. Na B. Cr C. Mg D. C
469	The variation pattern in ionic radii of first transition series shows	A. A regular increase B. A regulars decrease C. No regular pattern D. A regular decrease and than alight increase
470	Which form interstitial compounds.	A. Fe B. Ni C. CO D. All of those
471	The transition elements which are present in 4th period of periodic table have atomic number.	A. 22 to 30 B. 21 to 30 C. 21 to 29 D. 21 to 31
472	d -block elements which show anomalous configuration in first series are	A. Cr and Ni B. Cr and Cu C. Cu and CO D. Fe and Ni
473	Element of which group are called non typical transition elements.	A. IB B. II B C. II A D. VII B
	In [Co(NH3)6]+3 the	A. Zero B. Two

474	coordination number of cobalt is.	B. Two C. Four D. Six
475	In IUPAC system, the name of K4[Fe(CN)6] is	A. Potassium ferricynaide B. Potassium ferrocynide C. Potassium Hexacyanoferrate (II) D. Poatssium hexacaynoferrate (III)
476	The type of hybridization in PCl3 is.	A. dsp2 B. sp3 C. dsp3 D. d2sp3
477	Which one of the following complexes is chelate.	A. Potassium hexacyanoferrate (II) B. Diammine silver (I) Chloride C. Tetracarbonylnikel (0) D. Sodium dioxalatoplatinate (II)
478	Which complex shows zero oxidation state of the transition metal.	A. [Fe(CO)5] B. K3[Fe(CN)6] C. K2[Fe(CN)6] D. [Cu(NH3)4]SO4
479	Geometrical shape of [CO(NH3)6Cl3	A. linear B. square planar C. Octahedral D. Trigonal hypyramid
480	Stainless steel is	A. Compound B. An element C. Mixture D. 100% pure iron
481	Which is not coloured ion	A. SO4⁻² B. MnO4-3 C. CrO4⁻² D. Cr2O7⁻²
482	Which furnace is used to prepared steel	A. Blast furnace B. Pudding furnace C. Bessemer converter D. Pyrite furnace
483	The substance which is added to remove impurities is known as	A. Slag B. Flux C. Ore D. Gangue
484	To prevent corrosion, Iron pipes carrying drinking water are covered with zinc by	A. alley formation B. Electroplating C. Galvanizing D. Soldering
485	Galvanized iron is protected by a thin layer of	A. Cr B. Zn C. Sn D. Pb
486	Which element shows highest oxidation state among these	A. Zn B. Fe C. Mn D. Sc
487	Chlomyl chloride test is used for the confirmation of	A. Cl- B. CO3⁻² C. NO3- D. Cu2+
488	Colour of K2Cr2O7 is	A. Red B. Orange C. Green D. Yellow
489	The chemical composition of pyrolusite is.	A. KMnO4 B. K2MnO4 C. MnO2 D. MnO
490	The aqueous solution of which substances is green in colour	A. K2CrO4 B. K2CrO7 C. KMnO4 D. K2MnO4
491	The conversion of potassium magnate to potassium permanganate by passing Cl2 Through aqueous solution of K2MnO4 is called.	A. Contact process B. Open hearth process C. Stadeler's process D. Thermite process

492	First organic compound prepared in laboratory was.	A. Glucose B. Methane C. Urea D. Alcohol
493	Which one of the following compound has octane number 100.	A. 2,2,4-trimethyl petane B. n- pentane C. 2,4-dimethyl pentane D. 2- methyl pentane
494	The quality of petroleum is determined by	A. Decane number B. Octane number C. Hexane number D. Gold number
495	Wohler synthesized first organic compound in laboratory from	A. Heating cyanogen's B. Cyanogen and ammonium chloride solution C. Cyanogen and HNO2 D. Heating ammonium cyanate
496	Kerosene oil is a mixture of hydrocarbon having carbon	A. 11 to 13 B. 10 to 12 C. 11 to 12 D. 8 to 9
497	Which type of coal has greater percentage of carbon.	A. Peat B. Lignite C. Bituminous D. Anthracite
498	which one is not fossil fuel	A. Petroleum B. Natural gas C. Coal D. Alcohol
499	which one of the general formula of alkene	A. CnH2n B. Cn2n+2 C. CnH2n-2 D. CnH2n+1
500	An atom or group of atoms, which confers characteristic properties to organic compounds, are called.	A. Isomerism B. Metamerism C. Ligands D. Functional groups
501	Which one of the following is not a heterocyclic compound.	A. Furan B. Thiophene C. Pyridine D. Aniline
502	Which is an aromatic compound	A. Anthracene B. Naphthalene C. Toluene D. All of the these
503	Catalyst used in thermal cracking	A. Platinum B. Nichel C. Al2O3 and SiO2 D. Fe2O3 and CuO
504	Which compound is alicyclic in nature.	A. Cyclobatane B. Iso batane C. n butane D. Toluene
505	In ethene molecule, the number of atoms which are present in the same plane are.	A. 2 B. 6 C. 3 D. 4
506	An sp3- hybrid orbital contains.	A. 25% s- characters B. 50% s- characters C. 75% s- characters D. 100% s - characters
507	In which hybridization bond angle is maximum	A. sp3 B. sp2 C. sp D. sp3 and sp have same angles
508	In ethene molecule how many carbon orbitals are equivalent and degenerate in nature.	A. 3 B. 4 C. 5 D. 6
509	Boiling point range of petroleum ether	A. 5- 20 ^oC B. 10- 30^oC C. 20- 60^oC

	petroleum ether.	C. 20- 60^oC D. 30- 90^oC
510	The isomers having same functional group but different alkyl group on either side of functional group are called.	A. Metamers B. Polymers C. Monomers D. Homologous series
511	Which one of the following compounds is a heterocyclic.	A. Anthracene B. Phenol C. Pyridine D. Aniline
512	Which class of compounds can not show positional isomerism.	A. Alkanes B. Alkene C. Alkynes D. Alcohol
513	Which one of the following compounds show geometrical isomerism in it.	A. 1- pentene B. 1,1 dichloro ethane C. all of these D. 2- Pentene
514	The isomers always have same	A. Chemical properties B. Structural formula C. Molecular formula D. Physical properties as well as chemical properties
515	Which isomerism is not found in alkenes.	A. Chain isomerism B. Positional isomerism C. Geometrical isomerism D. Metamerism
516	Geometrical isomerism in alkene is due to.	A. C = C free rotation of bond B. No C = C free rotation of bond C. Presence of multiple bond only D. Opticla rotation due to multiple bond
517	Preparation of vegetable ghee involves.	A. Halogenation B. Hydrogenation C. Hydroxylation D. Dehydrogenation
518	Venyl acetylene combines with HCl in	A. Polyacetylene B. Benzene C. Chloroprene D. divinyl acetylene
519	The addition of unsymmetrical reagent to an unsymmetrical alkene is in accordance with the rule	A. Hund's rule B. Markowikov's rule C. Pauli's Exclusion Principle D. Aufbau Principle
520	Which one of the following gases is used for artificial ripening of fruits.	A. Ethene B. Ethyne C. Methane D. Propane
521	Synthetic rubber is made by polymerization of.	A. Chloroform B. Acetylene C. Divinylactylene D. Chloroprene
522	When methane reacts with Cl2 is commonly known as	A. Mustard gas B. Laughing gas C. Phosgene gas D. Bio gas
523	Saturated hydrocarbon are also called.	A. Olefins B. Acetylenes C. Paraffins D. Alicyclic
524	The IUPAC name of C(CH3)4 is	A. Iso Propyl methane B. 2-Methylbutane C. Iso bytylmethane D. 2,2 dimethylpropane
525	Which is called marsh gas	A. S2Cl2 B. SOCl2 C. CH4 D. CHBr3
526	An aldehyde is reduced to alkane with hydrazine is the presence of	A. KOH B. NaOH C. CaO and NaOH D. Ca(OH)2
	Which one of the following	A. CH4

527	gases is prepared by heating a mixture of sodium acetate and sodium hydrozide.	B. CH2-CH3 C. CO2 D. CO
528	In which reactions alkane is not produced	A. Subatier's and Sendern reaction B. Koibe's reaction C. Wolf -Kishner's reduction D. Dow's process
529	Which compound is least reactive	A. CH3-CH3 B. CH2=CH2 C. CH=CH D. C6H6
530	Which substance is formed by the catalytic oxidation of methane at 100 ^{o}C, 200 atmospheric pressure and copper catalyst	A. Methanol B. Methanal C. Methanoic acid D. All of these
531	Which one of the followings is not observed in the combustion of pure methane in a plentiful supply of air	A. Water in produced B. CO2 is produced C. The flame is smoky D. Energy is released
532	Which one is used for manufacture of fertilizers	A. Methane B. Ethane C. Ethene D. Ethyne
533	When ethyl alcohol is heated with conc. H2SO4 it produces ethene. The temperature required is as proximately	A. 100 ^oC B. 78 ^oC C. Above 200 ^oC D. 140- 170 ^oC
534	Which gas act as 'Blistering agent'	A. Acetylene B. Phosphine C. Phosgene D. Mustard gas
535	Which one of the following compounds will decolorized both acidified KMnO4 and aqueous bromine.	A. Benzene B. Ethane C. Ethene D. Methane
536	Which one of the following can best be used to distinguish between samples of ethane and ethene	A. Aqueous BaCl2 B. Aqueous bromine C. Lime water D. Litmus solution
537	What type of reaction occurs between ethene and hydrogen.	A. Addition B. Substitution C. Oxidation D. Dehydration
538	Which one of the followings is major product when HBr reacts with 2-butene	A. 2- bromobutane B. 1- bromobutane C. 1-1 di bromobutane D. 1,2 di bromobutane
539	A gas decolorizes alkaline KMnO4 but has no action with ammoniacal AgNO3, this gas may be	A. C2H2 B. C2H4 C. C2H6 D. CH4
540	Sodalime is	A. NaOH B. Mixture of Na and Ca(OH)2 C. KOH D. Mixture of CaO and NaOH
541	Raney Nickel is	A. Nickel compound B. Naturally occurring nickel C. Spongy form of a nickel D. Alloy of nickel
542	Ranney nickel is prepared by reacting dilute NaOH solution with.	A. Nickel B. Brass C. Nickel and aluminum alloy D. Nickel oxide
543	Vinyl chloride when boiled with alcoholic KOH , gives	A. Acetylene B. Ethylene C. Ethene D. Ethyl alcohol
544	Which one of the following compounds will form red precipitate with ammoniacal	A. Acetylene B. Ethylene C. Benzene

	cuprous chloride	D. Methane
545	Which one of the following gases is used in welding purpose usually	A. Methane B. Ethane C. Ethene D. Acetylene
546	Which one of the following gases is used welding purpose usualy.	A. Methane B. Ethane C. Ethene D. Acetylene
547	Which one is alcylic compound	A. Pyridine B. Toluene C. Ethyl Benzene D. Ethylcyclobutane
548	Acetylene reacts with ammoniacal cuprous chloride a red ppt is formed. This red ppt is of	A. Copper chlorite B. Coper chlorate C. Cupric oxide D. Copper acetylide
549	Acetylene polymerized in the presence of ammonium chloride and cuprous chloride ot form.	A. Benzene B. PVC C. Di vieiyl acetylene D. Polyethene
550	Which one of the following substances have garlic odour and a colourless gas.	A. CH3OH B. HCOOH C. CH2=CH2 D. HC= CH
551	Which one of the following is formed when ethyne is heated in copper tube at 300 ^{O}C	A. Ethene B. Ethane C. Benzene D. Cyclohexane
552	The number of acidic hydrogen present in 1-Propyne is	A. 1 B. 2 C. 3 D. 4
553	Which has reddish brown colour.	A. silver acetylide B. Copper acetylide C. BaSO4 D. Aqueous KMnO4 solution
554	Ethyne on oxidation with strong alkaline KMnO4 changes to	A. Ethyl alcohol B. Acetaldehyde C. Vinyl alcohol D. Glyoxal
555	Which alkyne reacts with water and form aldehyde	A. Ethyne B. Propyne C. 1- Butyne D. 2- Butyne
556	The benzene molecule contains.	A. Three double bonds B. Two double bonds C. One double bonds D. Delocalized sigma electron charge
557	Aromatic hydro carbon are the derivatives of	A. Normal series of paraffins B. Alkene C. Benzene D. Cyclohexane
558	Benzene can not undergo.	A. Substitution reactions B. Addition reactions C. Oxidation reactions D. Elimination reactions
559	Among the following the compound that can be most readily suphonted is.	A. Toluene B. Benzene C. Nitrobenzene D. Chlorobenzene
560	Which compound is the more reactive	A. Benzene B. Ethene C. Ethane D. Ethyne
561	Aromatic compounds burns with sooty flame because.	A. They have high percentage of hydrogen B. The have a right structure C. The have high percentage of carbon D. The resist reaction with air
		A. Isomerization

562	The conversation of n-hexane into benzene by heating in in the presence of Pt is called	B. Aromatization C. Dealkylation D. Rearrangement
563	The number of possible isomers of xylene are	A. 2 B. 3 C. 4 D. 5
564	Biphenyl is	A. Monocyclic aromatic compound B. Polyclic aromatic compound C. Polycyclic fused ring compound D. Alicyclic compound
565	Effect of substituent on benzene ring is due to	A. Resonance B. Inductive effect C. Both a and b D. Neither a nor b
566	Cyclic structure of benzene was proposed by	A. Dewar B. Faraday C. Down D. Kekule
567	Benzene molecule is made up of.	A. 3- pi and 9 sigma bonds B. 6- pi and 6 sigma bonds C. 6 -pi and 12 sigma bonds D. 4 -pi and 12 sigma bonds
568	Structure of benzene is	A. Tetrahedral B. Trigonal planar C. Hexagonal planar D. Linear
569	The difference of actual and theoretical heat of hydrogenation of compound is called.	A. Lattice energy B. Resonance energy C. Ionization energy D. Enthalpy of formation
570	Which makes a molecule more stable.	A. Greater localization of pi electrons B. Less delocalization of pi electrons C. Greeter delocalization of pi electrons D. Less delocalization for pi electrons
571	Hydrolysis of benzene sulphonic acid with superheated steam or by boiling with dil HCl gives.	A. Toluene B. Benzene C. Xylene D. Chlorobenzene
572	Phenol when distilled with zinc dust gives.	A. Toluene B. Benzaldehyde C. Benzoic acid D. Benzene
573	Main source of aromatic compound is	A. Petroleum B. Coaltar C. Living organisms D. Dead marine animals
574	When acetylene is heated at 300 ^{O}C in copper tube, the product obtained is.	A. Benzene B. Alkyl benzene C. Ether D. Alcohol
575	The conversion of benzene to chlorobenzene is a	A. Addition reaction B. Elimination reaction C. Substitution reaction D. Dehydration process
576	Which compound form benzoic acid on oxidation with strong oxidizing agent.	A. Toluene B. Ethyl benzene C. n propyl benzene D. All
577	The aromatic ring of Benzene can be hydrogenated in the presenc eof.	A. Pt B. Rh C. Sunlight D. O3
578	Which of the following species is ortho and para director.	A. CHO B. SO3H C. NO2 D. Cl
579	Which one of the following species is meta director if present at benzene ring.	A. -NO2 B. -Cl C. -CH3 D. -OH

580	Chlorobenzene on nitration with conc. HNO3 and Conc. H2SO4 gives.	A. m -chloronitrobenzene B. o and p chloronitrobenzene C. o and m chloronitrobenzene D. mixture of O2 m and p chloronitrobenzene
581	Which group when attached with benzene increases electron density of ring.	A. -COOH B. -NO C. -CHO D. -CH3
582	Which one is not electrophile	A. BF3 B. SO3 C. AlCl3 D. NH3
583	The compound which can not be nitrated easily.	A. Benzene B. Nitrobenzene C. Phenol D. Toluene
584	In primary alkyl halides, the halogen atom is attached to a carbon which is further attached to how many carbon atoms.	A. One B. Two C. Three D. Four
585	The reactivity order of alkyl halides for a particular alkyl group is.	A. Fluoride > Chloride > Bromide > Iodide B. Chloride > Bromide >Chloride > Fluoride C. Iodide > Bromide > Chloride > Fluoride D. Bromide > Iodide > Chloride > Fluoride
586	When CO2 is made to react with ethyl magnesium iodide, followed by acid hydrolysis, the product formed is.	A. Propane B. Propanoic acid C. Propanal D. Propanol
587	Grignard reagent is reactive due to	A. The presence of halogen atom B. The presence of Mg atom C. The polarity of C -Mg bond D. None of the above
588	SN2 reactions can be best carried out with	A. Primary alkyl halides B. Secondary alkyl halides C. Tertiary alkyl halides D. All the three
589	Elimination biomolecular reactions involve.	A. First order kinetics B. Second order kinetics C. third order kinetics D. Zero order kinetics
590	For which mechanisms, the first step involved is the same.	A. E2 and E2 B. E2 and SN2 C. SN1 and E2 D. E1 and SN1
591	Which one of the following is not a nucleophile.	A. H2O B. H2S C. BF3 D. NH3
592	Secondary alkyl halides are those in which halogen atom is attached with a carbon atom which is further attached to.	A. One beta carbon B. Two beta carbon C. Three beta carbon D. Four beta carbon
593	The reacts with halogen acids to form alkyl halide the process is known as.	A. Halogenation B. Hydrohalogenation C. Hydrogenation D. Dehydrohalogenation
594	Which substance is used to convert alcohol to alkyl halide.	A. SOCl2 B. PCl3 C. HCl +ZnCl2 D. All of these
595	In which process, alkyl halide is not produced.	A. Reaction of alcohol with halogen acid B. Reaction of Grignard reagent with water C. Reaction of alcohol with phosphorous pentachloride D. Action of alkene on halogen acid
596	Which products is not formed when ethyl alcohol reacts with SOCl2 in the presence of pyridine.	A. Ethyl chloride B. Hydrogen chloride C. Sulphur di oxide D. Sulphur tri oxide
597	Nucleophilic substitution reactions, which are completed in two steps are called as	A. SN1 B. SN2 C. E1

	in two steps are called as.	D. E2
598	SN1 reaction usually occurs in	A. Primary alkyl halides B. Secondary alkyl halides C. Tertiary alkyl halides D. All of these
599	SN2 mechanism involves	A. 1st order kinetic B. 2nd order kinetic C. 3rd order kinetic D. Zero order kinetic
600	Which one of the following species is a nucleophile	A. CH3 B. (CH3)2 C C. BF3 D. OH-
601	In unimolecular reactions, the reaction completes in	A. _{One step} B. Two steps C. Three steps D. None of these
602	Which one of the following species is not an electrophile.	A. HN3 B. Br C. H+ D. BF3
603	An alkyl halide may be converted to alcohol by	A. Addition B. Substitution C. Dehydrohalogenation D. Elimination
604	The reaction of alkyl halides with sodium metal in the presence of ether to from alkane is known as.	A. Wortz reaction B. Frankland reaction C. Sabatier sendron D. Kolbe's synthesis
605	Which one of the following products will be formed in Wurtz reaction when sodium metal reacts with ethyl chloride in anhydrous ether.	A. Methane B. Ethane C. Propane D. Butane
606	Which one of the following reactants will be required to form ethyl alcohol form ethyl bromide.	A. Alcoholic KOH B. Aqueous KOH C. Alkaline KMnO4 D. Sodium metal in ether
607	Which one of the following will be required to form ethene from ethyl chloride.	A. Alcoholic KOH B. Aqueous KOH C. Alkaline KMnO4 D. Bromine
608	Which alkyl halide does not form Grignard's reagent.	A. CH3-Br B. CH3-Cl C. CH3- F D. CH3-I
609	The general representation for Grignard reagent is.	A. RMgX B. ReMgX C. RXMg D. RMgX2
610	Acetic acid can be obtained from CH3MgI by treatment with.	A. H2O B. ClNH2 C. CO2 D. HCHO
611	Which one of the following alkanes will be formed by the hydrolysis of ethyl magnesium bromide	A. Methane B. Ethane C. Butane D. do not hydrolysed
612	Which one of the following alcohols will be formed when ethyl magnesium bromide reacts with acetone.	A. Primary alcohol B. Secondary alcohol C. Tertiary alcohol D. Dehydrin alcohol
613	Which one of the following molecules does not form alcohol when reacts with a Grignard reagent.	A. Formaldehyde B. Acetaldehyde C. Propanone D. CO2
614	Which substance is used to convert Grignard reagent to alkane.	A. H2O B. NH3 C. Ethyl alcohol D. All of these

615	What products is formed when ethyl bromide reacts with magnesium to form Grignard's reagent.	A. Pyridine B. Anhydrous ether C. Ethyl alcohol D. Carbon tetrachloride
616	The product of fermentation of sucrose is	A. Ethanol and H2O B. Ethanol and CO C. Ethanol and CO2 D. Glucose and CO2
617	Absolute alcohol is that which is	A. 100% B. 95% C. Ethanol mixed with methanol D. Ethanol mixed with H2O
618	Ethanol on oxidation in the presence of K2Cr2O7 and Conc. H2SO4 changes to.	A. Acetaldehyde B. Ethane C. Ethene D. CO2 and H2O
619	Ethanol on dehydration can be changed to	A. Ethene B. Diethyl ether C. Both 'a' and 'b' D. None of these
620	Ethanol reacts with Na metal to form sodium ethoxide. What product will be formed when C2H5ONa reacts with methyl bromide.	A. C2H5OC2H5 B. C2H5OCH3 C. CH3COC2H5 D. C2H5Br and NaBr
621	Which compound shows hydrogen bonding.	A. C2H6 B. C2H5Cl C. CH3-O-CH3 D. C2H5OH
622	Which compound shows maximum hydrogen bonding with water.	A. CH3OH B. C2H5OH C. CH3-O-CH3 D. C6H5OH
623	Which compound is more soluble of water	A. C2H5OH B. C6H5OH C. CH3OCH3 D. n- Hexanol
624	Which compound will have maximum repulsion with H2O	A. C6H6 B. C2H5OH C. CH2CH2CH2OH D. CH5-O-OH3
625	Ethanol can be converted into ethanoic acid by.	A. Hydrogenation B. Hydration C. Oxidation D. Fermentation
626	Which enzyme is not involved in fermentation of starch.	A. Diastase B. Zymase C. Urease D. Invertase
627	Which compound is called a universal solvent.	A. H2O B. CH2OH C. C2H5OH D. CH5-O-CH3
628	According to Lewis concept ethers behave as	A. Acid B. Base C. Acid as well as a base D. Noe of them
629	Which one of the following compounds is the isomer of ethyl alcohol.	A. CH3OH B. CH4OCH3 C. CH5-CH(OH)CH3 D. CH3OC2H5
630	Which substance is used for denaturing of ethanol	A. Methanol B. Acetone C. Pyridine D. All
631	Methyl alcohol can be represented by all of the following words or symbols except.	A. CH3OH B. Wood spirit C. Methanol D. Grain alcohol
632	Phenol is also known as	A. Citric acid B. Carbonic acid C. Carbolic acid

		C. Carbolic acid D. Maleic acid
633	The IUPAC name of CH3OCH6H5 is	A. Methyl phenyl ether B. Methaoxy benzene C. Phenoxy methane D. methoxy phenyl
634	Which ether is symmetrical in nature.	A. Methyl ethyl ether B. Diphenyl ether C. Methyl n propyl other D. Methoxy benzene
635	Which substance shows very weak hydrogen bonding with water.	A. Methanol B. Ethanol C. Diethyl ether D. Benzene
636	Absolute alcohol can be obtained from rectified spirit by	A. By adding sodium metal B. By extraction C. By predistillation in the presence of CaO D. Not possible because of azeotropic mixture
637	The conversion of ethene to ethanol is an example of.	A. Hydration B. Dehydration C. Neutralization D. Esterification
638	Zymase can be used to convert glucose to	A. Carbon and steam B. CO2 and hydrogen C. CO2 and Ethanol D. Ethanol and water
639	Which condition are not suitable for the growth of enzymes.	A. Temperature between 25 ^oC to 37 ^oC B. Solution must be dilute C. Environment must be aerated D. Some preservative should be present in solution
640	The correct name of CH3-CH= CH2-OH is	A. 2-buten -4 -ol B. 3-buten-I-ol C. 2-Buten -I-ol D. Ethylene glycol
641	When ethyl alcohol is heated, with NH3 in presence of ThO2 then	A. O-H bond is broken B. C-O bond is broken C. Ethene is formed D. Ethane is formed
642	The conversion of ethanol to ethene is an example of.	A. Dehydration B. Hydration C. Hydrogenation D. Fermentation
643	Methyl alcohol can be distinguished from ethyl alcohol by	A. Action of Cl2 B. Action of NH3 C. Dissolving in H2O D. NaOH +I2
644	Primary, Secondary and tertiary alcohols can be distinguish by.	A. Iodoform test B. Lucas test C. Fehling solution D. Ammoniacal silver nitrates
645	Which one is used as debydrating agent for alcohol.	A. H2SO4 B. Al2O3 C. H3PO4 D. All of these
646	How much does of methanol can cause death	A. 10-15 ml B. 15-20 ml C. 100- 250 ml D. has no effect
647	Which substance is used to convert ethanol to ethyl chloride	A. SOCl2 B. PCl3 C. PCl5 D. All of these
648	Phenol is the derivative of	A. Alkane B. Aromatic hydrocarbon C. Aliphatic hydrocarbon D. Alkene
649	Phenol after reduction with hydrogen changes to	A. Picric acid B. Benzene C. Cyclohexane D. Cyclohexanol
	Phenol can be prepared from	A. Williamson synthesis B. Dawn's process

650	Phenol can be prepared from chlorobenzene by	B. Down's process C. Kolbe reaction D. Cannizzaro reaction
651	Phenol on heating with concentrated nitric acid forms	A. o-nitrophenol B. T.N.T C. Na2CO3 D. Cyclohexanol
652	Conversion of phenol to benzene is known as.	A. Oxidation B. Reduction C. Hydrolysis D. Hydration
653	When ethyl bromide is heated with Ag2O the product formed is.	A. Ethanol B. Ethene C. Ethanol D. Di ethyl ether
654	Which one of the following methods is used for the preparation of ether.	A. Kolbe's reaction B. Frankland reaction C. Williamson synthesis D. Down's process
655	Di ethyl ether can be converted to alcohol by heating with.	A. HI B. NaOH C. Water D. KMnO4
656	A snake was preserved in a solution and was placed in biology laboratory. The solution is.	A. De ionized water B. Fehling solution C. Formalin D. Chloroform
657	Structure of carbonyl is	A. Tetrahedral B. Linear C. Octahederal D. Trigonal planar
658	Predict the product or reaction. Acetaldehyde +NaOH	A. An aldol B. Acetic acid C. Ethanol D. Paraldehyde
659	Which substance reacts with sodium nitroprusside.	A. Ethanol B. Acetaldehyde C. Dimethyl ketone D. Methanol
660	Which of the following will have the highest boiling point.	A. Methanal B. Ethanal C. Propanal D. 2-Hexanone
661	Ketones are prepared by the oxidation of.	A. Primary alcohol B. Secondary alcohol C. Tertiary alcohol D. None of these
662	Acetone reacts with HCN to form a cyanohydrin. It is an example of.	A. Electrophilic addition B. Electrophilic substitution C. Nucleophilic addition D. Nucleophilic substitution
663	Which of the following compounds will not give iodoform test on treatment with I2/NaOH	A. Acetaldehyde B. Acetone C. Butanone D. 3-Pentaneone
664	Which of the following reagents will reacts with both aldehydes and ketones.	A. Grignard reagent B. Tollens's reagent C. Fehling's reagent D. Benedict 's reagent
665	Common names of aldehydes are derived from the common names of.	A. Alcohol B. Ketones C. Carboxylic acids D. Alkenes
666	Which one is symmetrical ketone	A. Acetone B. Methyl ethyl ketone C. Methyl n propyl ketone D. 2- pentonone
667	What product is formed by the dry distillation of calcium acetate.	A. CH3CH2COOH B. CH3COOH C. CH3COCH3 D. CH3CH2CHO

668	Acetaldehyde can be prepared by the oxidation of.	A. Acetic acid B. Ethanol C. 1- Proponal D. Ethanonic acid
669	Industrially ethanal is prepared by air oxidation of.	A. Ethane B. Ethene C. Ethyne D. Acetic acid
670	Acetaldehyde reacts usually with	A. Electrophiles only B. Nucleophiles only C. Electrophiles and nucleophiles D. Free radicals only
671	A cyanohydrin is formed by the reaction of	A. Alcohol and HCN B. Ketone and NH3 C. Aldehyde and NH2OH D. Aldehyde and HCN
672	Acetaldehyde react, with Grignard reagent to form	A. Primary alcohol B. Secondary alcohol C. Ter alcohol D. Carboxylic acids
673	Formation of cyanohydrin from an aldehyde in an example of.	A. Nucleophilic substitution B. Nucleophilic addition C. Electrophilic addition D. Electrophilic substitution
674	Which substance usually undergo nucleophilic addition reactions.	A. Benzenes B. Aldehydes C. Alkenes D. All of these
675	Which one of the following compound does not react with NaOH and I2 and also does not form Iodoform.	A. C2H5OH B. CH3CHO C. CH3-CO-CH3 D. C2H5-O-C2H5
676	Which one of the following compounds undergoes Cannizzaro's Reaction.	A. Acetaldehyde B. Benzaldehyde C. Acetone D. Propionaldehyde
677	Acetaldehyde undergoes, aldol condensation in the presence of.	A. dil HNO3 B. DIL NaOH C. dil HCl D. Conc. H2SO4
678	When benzaldehyde is converted to benzyl alcohol by reading with NaOH the reaction is known as.	A. Cannizzaro reaction B. Wurtz reaction C. Wol;ficisthner reaction D. Aldol reaction
679	Which can produce ketone	A. Sec alcohol B. Calcium acetate C. Propyne D. All of these
680	All of the followings react with Fehling solution except	A. CH2CH2CHO B. CH3-CH2COCH3 C. HCHO D. CH3-CHO
681	when aldehydes react with Tollen's reagent.	A. A ketone is produced B. An alcohol is produced C. Ag ions are produced D. Ag ions are reduced
682	To product aldehyde group against alkaline oxidizing agent.	A. It is reduced in the presence of catalyst B. An acetal is formed C. It is oxidized D. It is treated with aqueous NaBH4
683	Oxime is an addition product obtained when aldehydes react with.	A. HCN B. NH2OH C. Phenyl hydrazine D. H2O
684	On heating aldehydes with Fehling's solution we get a precipitate whose colour is	A. Pink B. Black C. Yellow D. Brick red
685	Which one of the following compounds is not derivative of NH3.	A. Aniline B. Hydrazine C. Phenyl hydrazine D. Picric acid

		B. Picric acid
686	Paraldehyde is polymer of.	A. HCHO B. CH3CHO C. CH3COCH3 D. CH3CH2-CH2OH
687	All of the following tests are used to identify aldehydes except.	A. Tollen's test B. Fehling test C. Baeyer's test D. Benedict test
688	Air of the following tests are used to identify aldehydes except.	A. Tollen's test B. Fehling test C. Baeyer's test D. Benedict test
689	An aldehyde compound can be reduced to alkane by	A. Wurtz reaction B. Grignard reaction C. Wolf Kishner reaction D. Kolbe's reaction
690	Paraldehyde is used as a	A. Medicine B. Poison C. Polymer D. Dye
691	Which substance is used for silvering of mirror.	A. Acetaldehyde B. Ethanol C. Ethylene glycol D. Acetone
692	What is effect of solubility of carboxylic acid in water by increasing their molecular masses.	A. Decreases B. Increase C. Remain constant D. Non effected
693	Which one is cyclic amino acid	A. Glycine B. Alanine C. Proline D. Aspartic acid
694	Which one is basic amino acid	A. Lysine B. Alanine C. Glycine D. Aspartic acid
695	Which chemicals used for detection of amino acids.	A. NaOH + I2 B. Phenyl hydrazine C. Ninhydrin D. Benedict's solution
696	Number of peptide bonds in tripeptide is	A. 1 B. 2 C. 3 D. 4
697	Acetic acid is manufactured by.	A. Distillation B. Fermentation C. Ozonolysis D. Esterification
698	A carboxylic acid contains	A. A hydroxyl group B. A carboxyl group C. A hydroxyl and carboxyl group D. A carboxyl and an aldehydic group
699	Which acid is used in the manufacture of synthetic fiber	A. Formic acid B. Oxalic acid C. Carbonic acid D. Acetic acid
700	Which of the following derivative connot be prepared directly from acetic acid.	A. Acetamide B. Acetyl chloride C. Acetic anhydride D. Ethyl acetate
701	The solution of which acid is used for seasoning of food.	A. Formic acid B. Acetic acid C. Benzoic acid D. Butanoic acid
702	Which of the follow8ing is not a fatty acid.	A. Propanoic acid B. Acetic acid C. Phthalic acid D. Butanoci acid
703	Acetamide is prepared by	A. Heating ammonium acetate B. Heating methyl cyanide

	Acetamide is prepared by	C. Heating ethyl acetate D. The hydrolysis of methyl cyanide
704	Malonic acid is	A. Aromatic monocarboxylic acid B. Aromatic dicarboxylic acid C. Aliphatic monocarboxylic acid D. Aliphatic di carboxylic acid
705	Systematic name of phthalic acid is	A. Benzenedicarboxylic acid B. Benzene doioc acid C. 1,2 benzenedicarboxylic acid D. o, carboxylic benzoic acid
706	Formic acid is present in.	A. Butter B. Vinegar C. Ant D. Sunflower
707	Which one of the following comp9ounds on hydrolysis will not produce a carboxylic acid.	A. Alkyl halide B. Alkyl Nitrite C. Addition product of Grignard reagent and CO2 D. Ester boiled with sodium hydroxide
708	Acetic acid is commercially prepared form	A. Ethene B. Ethane C. Ethyne D. Ethanol
709	Zwitter ion is	A. Dipolar ion B. Amino acid with two amino groups C. Amino acid with two carboxylic acid group D. A synthetic amino acid
710	The process in when alkene are converted into carboxylic acid is known as	A. Oxidation B. Reduction C. Hydrolysis D. Hydration
711	What product are formed by the oxidative cleavage of 2-butene by alkaline KMnO4.	A. Propionic acid and formic acid B. Ethanoic acid only C. Ethanal only D. Water and ethanol
712	Which one of the following metal cannot evolve hydrogen from acetic acid.	A. Sodium B. Potassium C. Magnesium D. Zinc
713	Ester are formed the reaction of carboxylic acid with.	A. Alcohol B. Ethers C. Aldehydes D. Alkyl halides
714	Which one of the following liberates CO2 from an aqueous solution of NaHCO3.	A. Acetic acid B. Ethyl alcohol C. Phenol D. Acetyl chloride
715	Reverse of esterification is known as	A. Trans esterification B. Saponification C. Hydrolysis D. Neutralization
716	An ester can be prepared by the reaction of.	A. Two alcohols B. Alcohol and an aldehyde C. An alcohol and an organic acid D. an acid and a ketone
717	Which one is fatty acid	A. Benzoic acid B. Malonic acid C. Phtalic acid D. Palmitic acid
718	Which one of the following products is not formed when acetic acid reacts with PCl5	A. C2H5Cl B. HCl C. POCl3 D. CH3COCl
719	Which one of the following substance does not react with Na.	A. CH3COOH B. CH3OCH3 C. CH3OH D. C2H5OH
720	Which one of the following product is not formed when acetic acid is reacted with HI and red phosphorus.	A. I2 B. H2O C. CH3CH3 D. CH3CH2OH
		A. Coagulant for rubber latex

721	Which one of the following is not use of acetic acid.	B. Local irritant C. Formation of rayon and silk D. Formation of alcohol
722	Acetic acid form a dimer in liquid phase because.	A. Low ionization constant of acid B. High solubility in water C. Hydrogen bonding D. Greater polarity
723	Which substance is used to consulate rubber latex	A. Ethyl alcohol B. Acetaldehyde C. Acetic acid D. Water
724	What is Glacial acetic acid	A. Pure acetic acid B. 95% acetic acid C. a mixture of acetic acid and glycerol D. Vinegar
725	Flavour of ethyl butyrate is	A. Orange B. Pine apple C. Banana D. Apricot
726	Which one of the following is not amino acid.	A. Alanine B. Glycine C. Aspartic acid D. Aniline
727	In which process, an amino acid is produced	A. Wurtz synthesis B. Strecker synthesis C. Kolbe synthesis D. Cannizzaro reaction
728	The number of peptide bonds in dipeptide is	A. 0 B. 1 C. 2 D. 3
729	Which is simplest amino acid	A. Alanine B. Protein C. Lysine D. Glycine
730	Essential amino acids are those which	A. Are present in every protein B. Must be supplied to body through diet C. Contain two carbocyclic acid and one amino group D. Is syntherised by our body
731	Which amino acid is present in cheese	A. Glycine B. Alanine C. Tyrosine D. Valine
732	Fats are	A. Saturated glycerides B. Unsaturated glycerides C. Polyhydroxy ketose D. Polyhydroxy aldose
733	Which process is used to convert vegetable oil to vegetable ghee.	A. Hydrolysis B. Oxidation C. Esterification D. Hydrogenation
734	Which property of triglycerides is used to determine its molecular mass.	A. acid number B. Saponification number C. Iodine number D. gold number
735	In which these process are small organic molecules made into macromolecules.	A. The cracking of petroleum fractions B. The fractional distillation of crude oil C. The polymerization of ethene D. The hydrolysis of proteins
736	Plastics are a pollution problem because many plastics	A. Are made from petroleum B. Are very inflammable C. Burn to produce toxic fumes D. Decompose to produce toxic products
737	A polymeric substance that is formed in the liquid state and then hardened to a right solid is called a	A. Fiber B. Plastic C. Varnish D. Polyamid resin
738	Which one of the following polymers is called a polyamide	A. Nylon B. Rayon C. Terylene D. Orlon

739	Which one of the following macromolecules contains carbon, hydrogen, nitrogen and oxygen in it.	A. Nylon-6,6 B. Terylene C. Starch D. Bakelite
740	Which one of the following compound is not of a polymer.	A. Starch B. Glucose C. Protein D. Nylon -6, 6
741	The Plastic which become soft and melt on heating and can be molded and remolded are called.	A. Thermoplastic B. Thermosetting plastic C. Resin D. Melamine
742	Which one of the following plastic is a thermosetting plastic.	A. PVC B. Polystyrene C. Polyethene D. Bakelite
743	The fiber in which monomer isCH2=CH-Cl is known as	A. Saran fiber B. PVC C. Rayon fiber D. Arcylic fiber
744	Which one of the following polymers has no peptide linkage in it.	A. Terlene B. Nylon -6,6 C. Protein D. None of these
745	Which of the following is an ester.	A. Soap B. Starch C. PVC D. Dacron
746	Which one of the following is inorganic polymer	A. Graphite B. Rubber C. DNA D. Protein
747	The length of the polymer chain is specified by the number of repeating units which is called.	A. Condensation B. Co-polymerization C. Iodine number D. Degree of polymerization
748	Which one of the following is co polymer	A. PVC B. Polyvinyl acetate C. Nylon -6,6 D. Polyethene
749	Which is not polymer	A. Diamond B. Starch C. Sand D. Nucleotide
750	Starch is polymer is	A. Fructose B. a -d Glucose C. Sucrose D. B-D-Glucose
751	Which one of the following carbohydrates give blue colour with iodine.	A. Glucose B. fructose C. Sucrose D. Starch
752	A carbohydrate that cannot be acid hydrolysed is called.	A. Monosaccharides B. Di saccharides C. Poly saccharides D. Starch
753	Mono saccharides contain carbon atoms.	A. 3 to 6 B. 3 to 7 C. 3 to 9 D. only six
754	The oligosaccharides contain number of hexose unit.	A. 2 to 7 B. 2 to 8 C. 2 to 9 D. 2 to 100
755	Which sugar is called milk sugar	A. Glucose B. Fructose C. Lactose D. Maltose
756	Which carbohydrate can be used for silvering of mirror	A. Glucose B. Fructose C. Maltose

	used for silvering of mirror?	C. Maltose D. All
757	Which one of the following compounds is most abundant is nature.	A. glucose B. Starch C. Cellulose D. Fructose
758	Glycoside linkage is present in	A. Proteins B. Nylon -6,6 C. Starch D. DNA
759	Lactose has same molecular formula as	A. Glucose B. Fructose C. Ribose D. Maltose
760	In which form, glucose is stored in the liver	A. Lactic acid B. Maltose C. Ribose D. Glycogen
761	Which carbohydrate is called animal starch	A. Glucose B. Fructose C. Glycogen D. Starch
762	When hydrolyzed, protein yields.	A. Fatty acid B. Glycerol C. Amino acid D. Nucleosides
763	Which one of the following proteins transports oxygen in blood stream.	A. Insulin B. Albumin C. Hemoglobin D. Globulin
764	Which one of the following is a compound protein or conjugate protein.	A. Legumin B. Albumin C. Collagen D. Phosphoprotein
765	Dennturing of protein is	A. Hydrolysis of protein B. Unfolding of protein C. Three dimensional twisting and folding of peptide chain D. Developing hydrogen bonding in peptide chain
766	Some non protein portion attached to the protein is called.	A. Prosthetic group B. Secondary protein C. Transport protein D. All of these
767	To which class of organic compounds soap belongs.	A. Ether B. Ketones C. Aldehyde D. Salt of an organic acid
768	The degree of unsaturation of fat is measured by	A. Iodine number B. Oxidation number C. Reduction number D. Saponification value
769	Cholesterol is an important precursor in the biosynthesis of	A. Sex harmonies B. Adrenal hormones C. Vitamin D D. All of these
770	The amount of free fatty acid in fats and oil is determined by	A. Iodine number B. Acid number C. Saponification number D. Gold number
771	Which one of the following Lipids does not have glycerol backbone.	A. Cholesterol B. Oil C. glycogen D. Vitamin D
772	Which of the following lipids does not have glycerol backbone.	A. Fat B. Oil C. Cholesterol D. Phospholipid
773	Which one of the following is not steroid	A. Cholesterol B. Ergosterod C. Female sex Harmons D. Globulin
	The substance that retard the	A. Co enzyme B. Eno enzyme

774	The substance that retard the activity of enzyme is called.	B. Epo enzyme C. Activity D. Substrate
775	The enzyme which bring about exchange of functional group between two compounds is called.	A. Hydrolases B. Transferase C. Lyases D. Ligases
776	Which one of the following enzyme is used for the treatment of blood cancer in children.	A. Thrombin B. Asparaginase C. Glucokinase D. Fumarase
777	Cholesterol is a	A. Glyceride B. Wax C. Steroid D. Fat
778	Starch mixture of	A. Amylose and xylose B. Amylopectin and lactose C. Lactose and sucrose D. amylose and amylopectin
779	Which chemical reduces the acidity of soil.	A. Lime B. Urea C. Ammonium nitrate D. Ammonium sulphate
780	Which nitrogen fertilizer make the soil acdic.	A. Calcium nitrate B. sodium nitrate C. Potassium nitrate D. Ammonium nitrate
781	Which one is not an argillaceous material.	A. Clay B. Shale C. State D. Gypsum
782	Pulp is washed to remove lignin from it. Due to the presence of lignin in pulp paper becomes.	A. Soft B. Brittle C. Acidic D. Colourless
783	In Pakistan pulp is usually bleached with Cl2. The residual chlorine is removed from pulp by an antichlor which is	A. dil HCl B. dil NaOH C. Water D. NaCl
784	Which three elements are needed for the healthy growth of plants.	A. N,S,P B. N, Ca, P C. N,P,K D. N,K,C
785	Which woody raw material is used for the manufacture of paper pulp.	A. Cotton B. Bagasse C. Poplar D. Rice straw
786	The nitrogen present to some fertilizers helps plants to	A. Fight against diseases B. Produce fat C. Undergo photosynthesis D. Produce protein
787	Phosphorus helps in the growth of	A. Root B. Leave C. Stem D. Seed
788	Micro nutrients are required in quantity ranging from	A. 4 g - 40 g B. 6 g -200 g C. 6kg - 200 kg D. 4 kg - 40 kg
789	During the manufacturing process of cement the temperature of the decomposition zone goes up to.	A. 600 ^oC B. 900 ^oC C. 1000 ^oC D. 1200 ^oC
790	The word paper is derived from the name of which reed plant.	A. Rose B. Sun flower C. Papyrus D. Water Hyacinth
791	Which is not a calcarious material.	A. Lime B. Clay C. Marble D. Marine shell

		D. Marine shell
792	Ammonium nitrate fertilizers is not used for which crop	A. Cotton B. Wheat C. Sugar cane D. Paddy rice
793	Which one of the following is the macronutrient for plants.	A. Boron B. Zinc C. Calcium D. Nitrogen
794	Micronutrients for plant are those which	A. Are required in very large amount B. Are required in very small amount C. Are produced from plants in very small amount D. Retard the growth of plants
795	Which one of the following set of raw material is most suitable for manufacture of urea.	A. CH_4, N_2 and CO_2 B. H_2, N_2 and CO C. H_2,CO_2 and H_2O D. H_2O, N_2 and H_2
796	Ammonia is prepared industrially by	A. Contact process B. Ostwald process C. Birkland Eyed process D. Heber process
797	The percentage of nitrogen in urea is.	A. 36% B. 46% C. 56% D. 65%
798	Which one of the following is inorganic fertilizer.	A. Manure B. Urea C. Ammonium nitrate D. All of these
799	The macro nutrients are required in ranging from	A. 5 kg to 200 kg B. 5 kg to 200 g C. 6g to 200 g D. 1 kg to 100 kg
800	the percentage of nitrogen in NH_3 is	A. 46% B. 60% C. 82% D. 100%
801	Prilling is a process in which	A. Concentration of urea is increased B. Water is removed from urea C. Molten urea is converted to solid drop D. Some other ingredients are added to urea
802	Potassium fertilizers are especially useful for	A. Tobacco B. Coffee C. Potato D. All of these
803	Conversion of ammonium carbamate into urea is	A. Hydrolysis B. Hydrogenation C. Hydration D. Dehydration
804	Cement is a mixture of	A. Clay and clinker B. Clay, lime stone and gypsum C. Lime stone and gypsum D. Lime stone and clay
805	What is clinker	A. Roasted calcareous material B. Roasted argillaceous material C. Roasted calcareous and argillaceous material D. Roasted gyposum
806	Which one of the following raw materials is not present in the cement.	A. Lime stone B. Gypsum C. Blast furnace slag D. Red lead
807	Which one of the following is calcareous material	A. Marine shells B. clay C. Shale D. Blast furnace slag
808	The composition mixture of clay and lime stone in the raw material.	A. 75% lime stone and 25% clay B. 25% lime stone and 75% clay C. 15% lime stone and 55% clay D. 55% lime stone and 15% clay
809	Setting process of cement is based upon	A. Hydrolysis B. Hydration C. Dehydration

	based upon	C. Dehydration D. Botha a and b
810	The percentage of clay and liem stone in cement is in the ratio of.	A. 1 : 1 B. 1 :2 C. 1:3 D. 3:1
811	During manufacturing of cement , the temperature of pre heating zone is.	A. 500 ^oC B. 900^oC C. 1000^oC D. 1500^oC
812	In which country first of all paper was invented	A. USA B. China C. Egypt D. Germany
813	Which one of the followings is woody raw material used for making pulp and paper	A. Eucalyptus B. Wheat straw C. Bagasse D. Cotton linter
814	Which process of pulp making is mostly used in Pakistan.	A. Kraft process B. Sulphite process C. Neutral sulphite semi chemical process D. Wet process
815	The temperature of digester is main tained at	A. 100^oC B. 160-180^oC C. 200^oC D. 200-240^oC
816	Which substance is used ot bleach the pulp.	A. Na2SO3 B. NaCl C. NaClO D. NaOH
817	What is the function of Head box in paper making machine.	A. It reduces the thickness of paper B. It dry the paper C. It discharge teh pulp at the screen of fourdrinier table D. Web structure is consolidated
818	Usually the percentage of moisture is paper is	A. 1- 3% B. 4-6% C. 6-8% D. 5%
819	In Pakistan, the total consumption of paper per person per year is.	A. 2 kg B. 5 kg C. 7 kg D. 10 kg
820	The diameter of rotary kiln in the manufacture of Portland cement is.	A. 1 to 2 feet B. 2 to 4 feet C. 4 to 8 feet D. 8 to 15 feet
821	During pulp making, the pH of digester is maintained at.	A. 1-2 B. 3-5 C. 6-8 D. 7-9
822	The region of earth capable of supposing life is called.	A. Atmosphere B. Bisphere C. Dithosphere D. Hydrosphere
823	Which one is primary pollutant.	A. Peroxyacetyl nitrate B. Sulphuric acid C. Carbonic acid D. Carbon monoxide
824	The minimum temperature of troposphere is.	A. -2^oC B. -56^oC C. -100^oC D. 15^oC
825	The disease can be eradicated by using pesticides.	A. Malaria B. Yellow fever C. Sleeping sickness D. All of these
826	Water will be considered polluted if it has dissolved oxygen.	A. 3ppm B. 4ppm C. 5ppm D. 6 ppm
	Disinfection of water by	A. NH2Cl B. NCl3

827	chlorine is done by the production of.	B. NCl2 C. HOCl D. NHCl2
828	The pH range of the acid rain is.	A. 7-6.5 B. 6.5-6 C. 6-6.5 D. Less than 5
829	Peroxyacetyl nitrate is an irritant to humab beings and it affects	A. Eyes B. Ears C. Stomach D. Nose
830	The avoid teh formation of toxic compounds with chlorine which substance is used for disinfecting water.	A. KMnO4 B. O3 C. Alum D. Chloramines
831	A single chloride free radical can destroy how many ozone molecules.	A. 100 B. 100000 C. 10000 D. 10
832	Fungicides are teh pesticides which	A. Control the growth of fungus B. Kill insects C. Kill plant D. Kill herbs
833	Ecosystem is smaller units of.	A. Lithosphere B. Hydrosphere C. Atmosphere D. Biosphere
834	The main pollutant of leather tanneries in the waste water is due the salt of.	A. Lead B. Chromium (VI) C. Copper D. Chromium (III)
835	In purification of potable water the coagulant used is	A. Nickel sulphate B. Copper sulphate C. Barium sulphate D. Alum
836	The temperature in the incineration of industrial and hazardous waste process has a range.	A. 900 to 1000 ^oC B. 250 to 500^oC C. 950 to 1300^oC D. 500 to 900^oC
837	Half mass of atmosphere gases is present in.	A. 5-6 km distance above the surface of earth B. 10 km above the surface C. 100 km above the surface D. 56 km above the surface
838	How much fresh water is used for domestic purpose	A. 8% B. 23% C. 69% D. 100%
839	On earth polar ice caps and glacier contains H2O	A. 1% B. 2% C. 3% D. 10%
840	Which one of the followings is not a pollutant.	A. CO2 B. NO2 C. CO D. SO2
841	The residence time of NO is	A. Few hours B. 1 day C. 3 days D. 4 days
842	Which field produces significant amount of methane in the atmosphere	A. paddy field B. Cotton field C. Can sugar field D. Wheat field
843	The mean residence time of methane in atmosphere in	A. 1-2 years B. 3-4 years C. 3-5 years D. 3-7 years
844	The smog which have high contents of SO2 in it, is called.	A. Reducing smog B. Oxidizing smog C. Natural smog D. Neutral smog

845	The yellowish colour in photochemical smog is due to the presence of.	A. NO B. NO2 C. SO2 D. CO2
846	Which one of the following is secondary pollutant of atmosphere.	A. CO B. NO2 C. SO2 D. H2SO4
847	Acid rain fist of all was observed by	A. Augus Smith B. Robert Hook C. Mosley D. Watson
848	Photochemical smog mainly consist of	A. Higher hydrocarbons B. Oxidising agnts C. Reducing agent D. All of these
849	In which layer of atmosphere, ozone is present.	A. Thermosphere B. Mesosphere C. Stratosphere D. Troposphere
850	The thickness of ozone layer is.	A. 25 to 50 km B. 25 to 28 km C. 3 km only D. 1 km only
851	Thickness of the atmosphere is	A. 100 km B. 500 km C. 1000 km D. 1500 km
852	The normal amount of overhead ozone is.	A. 300 DU B. 350 DU C. 400 DU D. 450 DU
853	Ozone is mostly produced in	A. Tropical region B. North polar region C. South polar region D. Thermosphere zone of atmosphere
854	Ozone depletion in stratosphere region is mainly due to the reaction of O3 with	A. O2 B. SO2 C. CFCs D. All of these
855	Whcih water will be considered as polluted water.	A. High value of COD B. Low value of COD C. High value of DO D. Low value of BOD
856	A quality of raw water is improved by	A. Reduction B. Aeration C. Dehydration D. Incineration
857	The colloidal particles in raw water can be removed by	A. Coagulation B. Aeration C. Chlorination D. Hydration